

BARNABAS

OR

THE GREAT RENUNCIATION

BY

G. BUCHANAN RYLEY

London

THE RELIGIOUS TRACT SOCIETY

56 PATERNOSTER ROW AND 65 ST PAUL'S CHURCHYARD

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PREFACE.

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DURING the last few years several series of excellent monographs on Bible characters have been published, the effect of which on the minds and hearts of Bible readers can be only very good. From all these series the name of Barnabas has been conspicuously absent. Most likely this has been due to an under-estimate of the man; and, perhaps, also to a deficient valuation of the crisis in which his name is made prominent. This book is an attempt to re-estimate the character of Barnabas, and the services that he so willingly and brightly gave to the infant Church of Christ. In the opinion of some I may have erred in the direction of faint praise. But the book may help the reader to a better valuation; and to any one who wishes, or is quick to see the signs of Christ's overwatch of His Church, it may serve to show how amply Christ has fulfilled, in the period under review, His promise to His covenanted people of a Presence that would be their guide.

One or two things may be commended as deserving fuller discussion and expansion than I

have been able to give them in the compass of this volume. The Christian life and the range of special Christian experience that are indicated by the phrase 'full of the Holy Ghost,' might be suggested to devout minds and Christ-like hearts as worthy a thoughtfulness that they have not yet received. The words are used in the New Testament in such varied circumstances, and of such different types of Christian character, that they must embody one of those large thoughts of God that are often passed by because of their greatness: as the glory of the mountains and 'the strength of the hills' seldom quicken a greater brightness in the eye of the Swiss peasant.

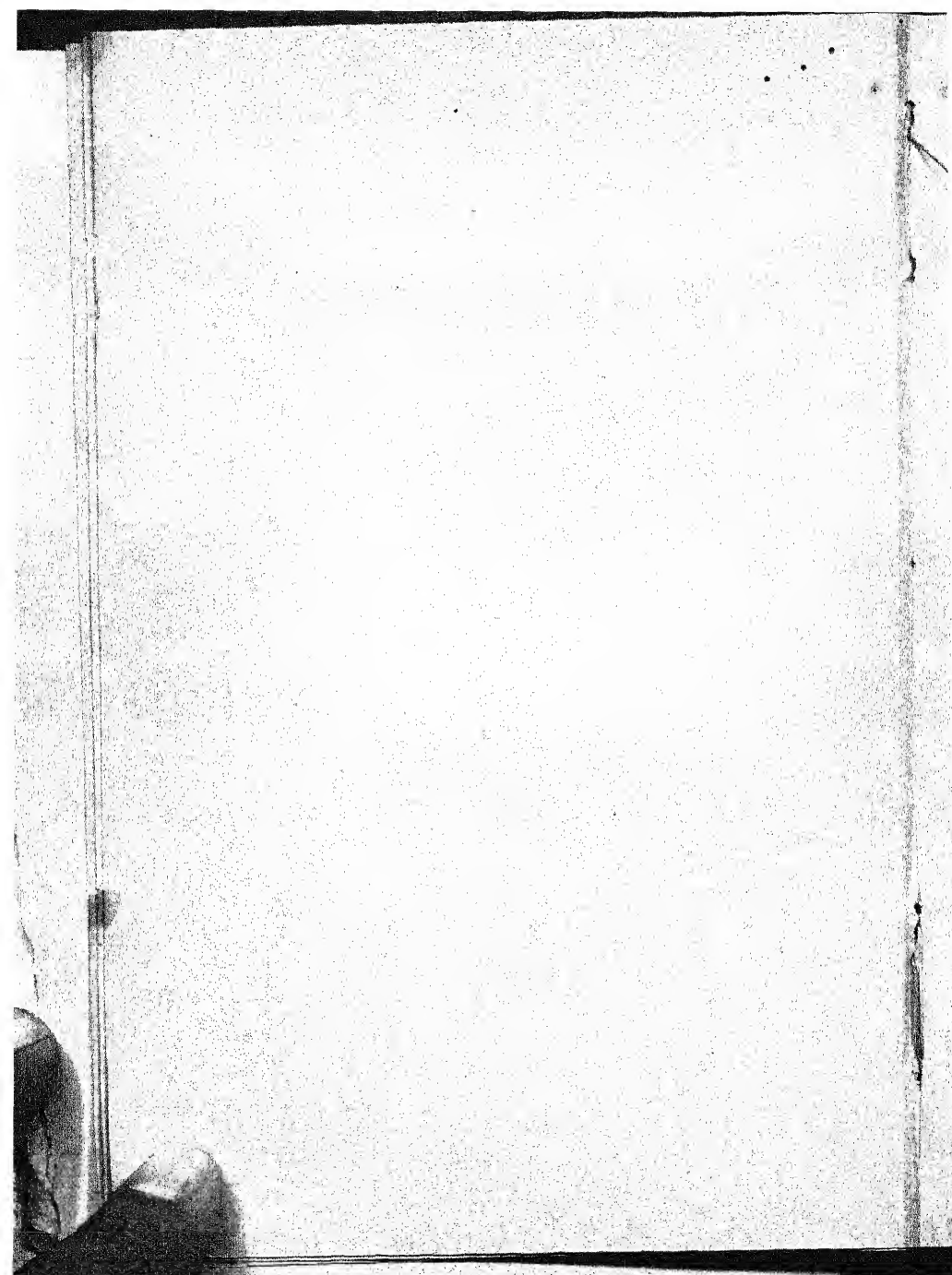
I may also suggest that the socialistic experiment in which Barnabas' name shines with so Christ-like a radiance as that of one who gave up all for Christ and the Brotherhood, is deserving of reconsideration by many who look on it as a good-natured freak of impulsive, new-born love. It may be safely asked whether all the hopefulness of this age is not turning towards the desire for a similar spirit to theirs who tried honestly, because of the need of their day, to follow Him who, 'though He was rich, for our sakes became poor, that we, through His poverty, might become rich.' To His approval and blessing I commend the book.

G. BUCHANAN RYLEY.

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BARNABAS.

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CHAPTER I.

THE STORY OF BARNABAS.

THE nearness of the men and women of the New Testament to the great central and radiant figure of Christ is, for purposes of biography, both an advantage and a disadvantage. On the one hand it is natural to expect from their proximity to the Sun of Righteousness that their spiritual movement, their progress along the track of light assigned to them, would have been rapid and ardent, with few intervals of shadow or eclipse; while their reflection of His radiance would have been clear and glorious. On the other hand, their closeness to the fresh spiritual influences of our blessed Lord; their opportunity of drinking at the very fountain head of the water of life freely; and, in many cases, their

having known Christ in the flesh, make comparison with Jesus Himself almost unavoidable.

And none can stand that test. In all such expectation and comparison the words of Christ's regret to one of His apostles are remembered almost involuntarily:—'Have I been so long time with you, and yet hast thou not known Me, Philip?' And these find almost an echo in the words of the most illustrious of the apostles when he, after many years of knowing and serving the Lord, almost with self-accusation, longs to 'know Christ and the power of His resurrection.' The men that were near the Lord saw their own defects in the great light of His presence, quite as readily as any who now try to tell their story or criticise their lives.

Happily the New Testament cannot be brought down to a number of biographies or life sketches. There is no nobly solitary individuality but that of Christ. The greatest men in the first Churches have their greatness and the fine development of their high qualities, from the flow through them of the life and influence of Christ, whose power

in them all, and variously in each one, makes the Church of the New Testament what it is, and what it has of nobleness and dignity in the story of the New Testament. Great and worthy as many of the leaders were in their natural qualities, they would have been comparatively inert, if not dead to any possibility of wide influence such as they now have, if it had not been for the vitalizing, the regenerating energy of Christ in them. They would have been but as the cold iron of the galvanic coil unconnected with the battery; while, touched by Christ, they became as the same coil when wired to the battery and filled with the potency of the electric current. Paul's intense words were true of such as he: to them, to live was Christ.

Hence it is evidently with purpose that the writer of the *Acts of the Apostles* tells his friend, that his 'former treatise was made concerning all that Jesus *began* both to do and to teach, until the day in which He was received up;' and then goes on, without break of continuity, to tell something of the outworking of the life of the glorified

Saviour in the saints of the first generation of the Church. Primitive Church history was thus declared to be the continued revealing and out-working of the Saviour's will. It was not dependent on the life story of a few great men, without whose particular individuality nothing worth doing would have been done. But the noble men who do appear in the front of the story, were great and influential for their high uses and destiny, because the energy of the life of the glorified Saviour had caught them in its tremendous sweep, and had touched them to the finest issues of which they were capable on earth. This power of the glorified Jesus also brought them forward and gave them scope at the right time, and in the right place; with such fitness of person to circumstances, and of gifts and aptitude to imperative needs, as to enable even dull souls to see the working of the Divine life, and to impel them to glorify God and the grace of Christ in His chosen servants.

Perhaps none of the New Testament saints shows these things more brightly and beautifully

than Barnabas. His fine spirit seems to have caught the light of the glory of God in the face of Jesus Christ in a way, and to a degree, that show the hallowing and glorifying of naturally good qualities; and, also, the addition to a noble nature of special graces from the Holy Spirit, of which the Church of Christ is the only province and sphere. These things make the story of Barnabas not only interesting, but spiritually instructive in a very special way.

Yet no one can read the story of Barnabas without seeing that he is one of the saints who have been pressed out of their proper place. Unintentionally they have been dwarfed in history by the greater prominence and wider reach, in visible influence, of some of their comrades. The mental intensity and greatness of Paul, together with the consequently large and permanent publicity of his work, have unfortunately lessened both the recognition and estimate of the services of Barnabas. It ought not to have been so. The life and work of Paul ought to have been as added light on the lustre of Barnabas.

For a true judgment of Paul cannot but aid in giving to Barnabas his proper value ; while, at the same time, the right estimate of 'the Son of Exhortation' will assign to 'the Apostle of the Gentiles' only additional and more evidently deserved honour.

The two are inseparable in some of the most important events in the history of the Church of Christ. They were, in their union and comradeship, involved in such momentous crises of the brotherhood in Christ, and together had so to determine the direction of the Church's action at important turning-points, as to make it more than desirable to judge neither of them apart from the other ; but each in his dependence on his brother, for a right valuation of their work in the Lord. They have been almost regarded as moving in separate circles, and as coming only for a little time into influential relationship with each other ; Barnabas being but the satellite of Paul. They were, rather, for the most important part of their lives, as twin stars having concentric revolution ; of different colour and size, certainly, but therefore

all the more interesting in their nearness and companionship, and all the more suggestive to devout thought because of their action and reaction on one another.

Paul's life and work have been well and deservedly studied; and will yet be even more thoroughly and adequately judged. Barnabas' labour and influence have not been duly valued. But if Paul could speak for his old friend and colleague, and if the relative status of a worker were of any moment in the Church of Christ, then Paul would rank his partner's work and value even higher far than our warmest estimate could reach. And that would be because he knew Barnabas as no one else could; and was indebted to Barnabas as to no other soul on earth. Paul's load of obligation to Barnabas was less only than his indebtedness to the Lord Jesus. In one of the chief experiences of his life Saul owed almost his soul to Barnabas. He had, also, through his friend, his great opportunity of proclaiming the universal gospel of Christ, as will be seen. For that opening of life, not only was Paul, but the

Church and the world are still the unredeemable debtors of the Son of Consolation.

Nor should it be forgotten that in his own person and by his special character and distinctive individuality, Barnabas shows title to the reverent regard of Christians that few can equal, and fewer still surpass. To show this in some detail will be the purpose of what follows. It may be well, however, first to give a simple outline of all that is known of this great saint of God, though the meagreness of the information will always be provocative of regret. All our reliable knowledge of Barnabas, or Joses, as he was first named, will be found in the *Acts of the Apostles*, and from one or two allusions in the Epistles of Paul. Legendary stories about his life and work both preface and follow the credible and authentic information of the Scriptures. And, for the most part, the legends are such as that all might safely and reasonably wish that they were beyond doubt. The mystical odour of sanctity that was reported to linger round a saint's body, might fitly symbolize

the happy fact, that very little that has gathered round the name and memory of Barnabas is other than sweetly and graciously attractive, truly saintly, and like the memory of the just who are at rest in God.

And this is the more to be valued because much of it has accumulated among the fungoid influences of ecclesiastical tradition, and in the distorting light of a love of the marvellous. Who could but long, for instance, that it were a true tale that says that Joses of Cyprus studied with Saul of Tarsus, under Gamaliel, in Jerusalem, and that thus they first knew one another? All may harmlessly wish that the legend were true that tells us that Joses knew the Lord Jesus in His earthly life, and was converted to Christ through witnessing the miracle at the pool of Bethesda. All would gladly believe the statement, if there were real warrant for it, that Joses was one of the seventy whom the Lord sent out on evangelistic work. All these things have been written in primitive Church history about Joses of Cyprus; and, going before

our Bible knowledge of him, would, if they had been true, have added greater interest to his most alluring story and tender personality.

All would gladly believe, if it were possible to warrant the belief of those legends that take up the story of Barnabas after our scanty authentic information ends. These give him the high honour of being the first preacher of Christ's gospel in Rome; and add thereto the distinction of being the means of converting Clemens Romanus to Christianity, whose name is still used in the Roman Catholic service of the mass. By these *post-mortem* legends we are also told that Barnabas founded the honourable Church of Milan, and began Church history in the city of Brescia. These legends almost fascinate the reader of Church history. For it would add a fitting nobility of origin to the great story of Christian Milan, and her august Church life, if we were able to believe that the discerning mind and warm heart of Barnabas, were the channel through which the life of the glorified Lord first flowed to her, and gave her that high

impulse of noble Christian life that her history has so frequently honoured.

So, too, Rome itself and the Church of Rome would have an honour greater than any that can be asserted by civic pride, or assumed by priestly arrogance, if it could be believed that the most generous heart of the earliest believers, was the first who bore to the cruel metropolis of the world, the message of the universal love of God, and the story of the omnipotence of the cross of Calvary ; and that one of her earliest bishops was the convert and disciple of the Son of Consolation. While of Brescia it might be well said, that if Barnabas had founded the Church of Christ here, it would have been a fitting and honourable thing that from her midst should have come one, whose protest against debased Christianity in Rome, and the consequent degradation of her civic life, shook the infamy of Papal pride and made the name of Arnold of Brescia to be a word of lofty impulse and inspiration, and a call to noble public deeds in Italy even to this day.

But these traditions have little warrant, though no one will get harm if he accept them as veracious. Inclination would lead to belief that Barnabas was martyred. It is difficult to think that such a man could escape the martyr's crown. If he were martyred, the crown never rested on a saintlier head. Legend says that he sealed his faith with his blood in his own native island of Cyprus, after bringing many to the faith of Christ. Having come to Salamis about the year 64, he is said to have preached the gospel in the synagogue with great results. But some Jews who had followed him from Syria stirred up the people against him, one Barjesus being their leader. Barnabas is said to have foreboded his approaching death: and one day, having taken the communion with his brethren of the Church, he bade them farewell. Having given his nephew instruction as to his burial, and commanded him to go to Paul after his death, he went once more to the synagogue to preach Christ. He had scarcely begun before the Jews laid hands on him; and keeping

him till night, they stoned him to death. The mangled body, that they tried to burn, could not be destroyed—so says one account. And Mark, getting possession of it, secretly buried it in a cave at some distance from the city, afterwards joining Paul at Ephesus, and going with him to Rome. The persecution that followed the death of Barnabas destroyed all trace of his place of burial.

The account of the discovery of his body is quite in the vein of ecclesiastical myth; and is, perhaps, the only thing to be resented in all the traditions of his life and death. The 'invention' of the saint's body, and the production of it at a suitable time as the means of determining an ecclesiastical quarrel between Antioch, supported by Constantinople, on one side, and the Cyprian Church on the other, even though the recorded result was the victory of the right side, is too much in the line of debased ecclesiasticism, to warrant other than a smile, restrained by regret. The legend is to the effect that Peter the Fuller tried to compel

the Cyprian Church to give up their independence, and to submit to his patriarchate. The Bishop of Salamis, who, though good, was a poor reasoner, was summoned to defend the rights, if such they were, of the Church founded by Barnabas, publicly, at Constantinople. And being sorely troubled at the summons, he was visited three times by Barnabas in the night-time, who gave him promise of success; and as a guarantee, told him where he could find the long-lost body of the martyr, with a copy of the Gospel by St. Matthew lying with it, or on it. The next morning the bishop called his presbyters together, and they found the body, according to direction. This secured their triumph at Constantinople. Peter the Fuller, of whom it has been said that he 'was formed to promote discord and controversy,' was degraded: the manuscript of the gospel found in the grave was placed in the palace at Constantinople, the Easter lessons being publicly read from it; and the emperor built a church on the place where the corpse of Barnabas had

been discovered. This legend is the most pitiful thing in connection with the memory of Barnabas.

Of the so-called *Epistle of Barnabas* little need here be said. The evidence for and against its authenticity is delicately poised; and honourable names support each side. Barring some things that seem to make the supposition of Barnabas being the author quite untenable, the spirit of the epistle is not unworthy such a man as the *Acts of the Apostles* pictures in Barnabas of Cyprus. The work is distinctly marked by the quality for which there is no other word than—spirituality. The interpretation of Judaism and the institutions of Moses moves on a high but perilous level of thought. The condemnation of the Jews that the document contains, is for their unspiritual reception of the Divine oracles, and their carnal estimate of God's dealing with their nation.

This general feature of the epistle is quite in harmony with the New Testament presentation of Barnabas' life and spirit. For, in the too brief

story of the *Acts of the Apostles*, it is he who, before any of the original apostles, looked over and beyond the Jewish pale, for the true promise of the conquests of Christ, and an adequate field for the work and battle of the Church. At the same time, a comparison of the epistle with any of the canonical letters, shows at once the profound difference between them, and makes it impossible to give to it the honour and position that were assigned to it by the early Churches.

The historic account of Barnabas is not lengthy in detail ; but the events of his life are of great importance. In brief, the whole story as we have it can be thus given. Barnabas is first mentioned as Joses or Joseph, a Cypriote Levite, and owner of land. This property he voluntarily gave to the Church in Jerusalem, thus literally obeying the Master's command to the young ruler, 'to sell all and give to the poor.' The writer of the *Acts of the Apostles* says without any comment, as if such things might be expected,— 'and Joseph, who by the apostles was surnamed

Barnabas (which is, being interpreted, Son of Exhortation), a Levite, a man of Cyprus by race, having a field, sold it, and brought the money, and laid it at the apostles' feet.'

He next appears as the brotherly and Christ-like champion and guarantor of Saul of Tarsus, when the Jerusalem Church feared to receive the notorious convert into fellowship.

Subsequently Barnabas was commissioned by the Church in Jerusalem to go to Antioch, and there investigate the great development of the Church's work, that some of the persecuted fugitives from Judea had started, on lines never before contemplated by the apostles and the Jerusalem Church. In confirming this hitherto untried work, he joined to himself Saul of Tarsus once more; and remained with Saul in Antioch a year, earnestly consolidating and building up that most eager-hearted and enterprising Church. The latter part of the eleventh chapter of the *Acts of the Apostles* gives the story in most winning and gracious simplicity. From Antioch, Barnabas and Saul went to Jerusalem, with the contributions that the

piety of the newly formed Church had made to the needs of the poor and famine-stricken brethren in Judea. When they returned to Antioch, they were commissioned once more by the Church, at the inspiration of the Holy Spirit, to go on an evangelistic journey to Cyprus and Asia Minor. The story of this mission is given in the thirteenth and fourteenth chapters of the Acts; and is notable for having in it Paul's (for he was no longer Saul) emphatic committing of himself to the world-wide work of the gospel, in a way that must have fully identified him with the large-heartedness of the thoughts of Barnabas. At the end of their mission journey they returned 'to Antioch, from whence they had been committed to the grace of God for the work which they had fulfilled. And when they were come, and had gathered the Church together, they rehearsed all things that God had done with them, and how that He had opened a door of faith unto the Gentiles. And they tarried no little time with the disciples.'

In the account of this journey it is that we have the only hint given of Barnabas' personal appear-

ance ; and it bears out what any one might have supposed he would look like from the story of his character. At Lystra, the cure of the crippled man caused the people to cry out that 'the gods are come down to us in the likeness of men. And they called Barnabas, Jupiter ; and Paul, Mercury, because he was the chief speaker.' There must have been something of nobleness in appearance, and stateliness of demeanour in Barnabas, for him to have been identified with Zeus. Of course, this ought not to be made too much of ; but remembering the impulsive cry of the Lystra people, it is at least pleasing to find, in all the story of Barnabas, qualities that would be expected to go with a noble and dignified personal presence.

On their return to Antioch, the trouble with the Judaizers and limiters of the gospel and grace of God, led the Church to send Barnabas and Paul again to Jerusalem, there to lay the cause of dissension before the apostles and elders, whose judgment they bore back to their much-loved Church. After this, Barnabas agreed with Paul to go on a second missionary journey. But a quarrel,

about what may have been the cowardice of John Mark, cousin of Barnabas, parted the two great friends. Barnabas took Mark with him to his native island, Cyprus.

This is all that is told in the New Testament about Barnabas, save that there are one or two incidental allusions to him in Paul's letters to the Corinthian Church and the Churches of Galatia. These references add little, if anything, to our real knowledge of the facts of his life, though they throw light on one or two matters that can be best dealt with in their historic place.

CHAPTER II.

THE GREAT RENUNCIATION.

THERE is no preface to the mention of the first most notable deed of Barnabas in the *Acts of the Apostles*. Yet that first record implies much; so much, in truth, as to warrant considerable liberty of conjecture. For his proper name was Joses or Joseph; and he had been 'christened,' in the true sense, by the apostles; receiving the better and most significant name of Barnabas, or Son of Exhortation. He was a Levite, and, therefore, was not likely, at the outset, to be free from very hampering prejudices and warping misconceptions; or to have very charitable feelings toward any of the name and following of Jesus of Nazareth.

Barnabas was of Cyprus, in which island the Jewish colony was very numerous and very bold.

There the glamour of the myth of Venus Aphrodite had left its witchery and curse; a remnant of which lasts to this day in the worship of the Mother of Jesus by the Cypriotes, under the name of Aphroditissa. No one would expect from a man who had been born and bred in Cyprus, the upward look of a soul that was consecrated to spiritual things; or the saintly look round on earthly affairs of one who sought the spiritual and Divine in the earthly and human. Belief in the influence of environment, however, may very easily be carried too far. It is loaded with too much now-a-days by many for whom a new fate seems to have risen, a fresh and cold-blooded supremacy from which the labouring heart of man can make no appeal. Still, experience would not have led any to expect a Barnabas to be the outcome of a Cypriote Levite. For even if it be supposed that the vileness of Venus worship wrought only antagonism and bitter resentment in a few, that very hate and scorn of the way of the Gentiles would be the more likely to make it impossible for a Hebrew to have the wide and generous feeling and thought for all.

men that were the distinctive marks of Barnabas when he became a Christian.

But be this as it may, there must have been almost from the first very fervent character and peculiarly specialized faith in Christ and love of the brotherhood in Jesus ; and these things must have made their mark quickly and distinctly in the short life of the Church, for Joses to have won so soon a distinctive name, and such honourable regard as the name betokens. It is not too much to suppose that from the first he may have been a refuge for the poor, 'a strength to the needy in his distress,' an inspiration to the downcast, a quickener of the dispirited, a consoler of the sad. Certainly, it would not be reading more than would be right into his name, to say such things of 'the Son of Consolation.'

Be this so or not, yet for some good and blessed reason the apostles called Joseph of Cyprus the Son of Consolation, or of Exhortation. Notwithstanding the presence of miraculous gifts in the Church, and the ministration of apostles, the need of gentle and tenderly helpful hearts must have

been very great in the young community, if the manifold grace of God were to be adequately shown. And very deep and hallowed must Joseph's service of Christ and the Church have been for him to be known as the first believer, after the ascension of our Blessed Lord, who is recorded to have had a new Christian name. This renaming with a religious name had been begun by Christ Himself; and Peter, who was the notable instance of this, must have felt less worthy his sacred and almost august name than Joseph was of his more tender and gracious nomination.

Barnabas was not at first a poor man, though he became poor for the sake of Christ and the Church. He owned what every one naturally wants to have—land; the freehold of a bit of God's earth. And there were others in the Church of Jerusalem who were like Barnabas, both in ownership and sacrifice. By them and their renunciation of their property it was, that the Holy Spirit showed a feeling to be natural to the Church of Jesus, that has never ceased to be an inspiration to the world. Some-

times, it is true, this spirit has led to extravagance of judgment and action, to wrong views of life and unwise dealing with life's trusts ; but it has always leaned towards self-forgetting, and to sacrifice for others. And that is more than a compensation for the mistakes that have been made. Earth will never be so spiritually and emotionally rich as that it can afford to dispense with, or lightly regard, this willingness to self-spoiling for the sake of others. Less than ever can it afford such a thing now. The hitherto unequalled wealth of the Church, and the world, rather, with necessity, calls for imitation of the *spirit*, if not the precise form of action, that to this day glorifies the name of Barnabas and the memory of his unknown brethren.

But it is too much to hope that the world will admire to emulation the spirit that puts self-sacrifice before it as right and true and needed. On the contrary, it follows with imitation, if not approval, the extravagance of those who aim at getting much for themselves, and love to glitter in selfish show and waste. All the more,

therefore, are believers in Jesus to honour the great event that showed so deep and tender a revolution of judgment as is told in the story of the 'communism' of the first Church of Christ. For, no matter whether they were few or many who felt the holy impulse, it was a wondrous uplifting of the whole level of life that made it possible for such fair fruits to grow from the human heart, and such a noble record to be made of the Christian Church in Jerusalem, that 'as many as were possessors of lands or houses sold them, and brought the prices of the things that were sold, and laid them at the apostles' feet: and distribution was made unto each, according as any one had need.' Evidently they could do without owning lands and houses; but they could not bear to see a brother in need. They could stand empty-handed where they had been full-handed. They could dare to be poor in order that others, through their poverty, might have enough. Even an age of luxury, that can look back on the movement with a pitying smile, and wonder where political economy would be if such

things were common, must acknowledge that the motive was laudable, and that the act itself was something like, nay, very like, the doings of Jesus of Nazareth. And it is even probable that to have acted like Jesus would be deemed quite honour enough in the estimation of those primitive believers, who 'counted all things but loss for the excellency of the knowledge of Christ Jesus their Lord.'

Only one name is honourably associated with this communistic movement, and that is Barnabas. He sold his field, brought the money, and laid it at the apostles' feet. If nothing else, he showed that he was not only an exhorter, a consoler by word, a speech-maker, but also a willing sacrifice for Christ. He gave up his all to his Lord and the poor of the Church. He was not 'a speaker that forgetteth, but a doer that worketh'; and he was blest in his doing. Maybe James thought of Barnabas when he wrote such words.

Is it too much to suggest that Barnabas may have been the first in the great sacrificing movement of the Church in Jerusalem to give up all

for the poor of Christ? and that, because he was the leader in this, his name is so honourably placed in the story of the faithful in Christ Jesus? His act must have been, in some way, prominent in the self-stripping and full consecration of the Church members. Else, why name him at all? Or, why put him and his deed in awful contrast with the living lie to the Holy Ghost of Ananias and Sapphira? Without doubt, when Barnabas made his great renunciation of property, it was either so notable in itself as to be worthy a record that would make him stand out before the Church in admirable outline of character and nobility of deed; or his was the case that best served to put into the lurid light of a deserved and awful condemnation, the blot on the movement that Ananias made in company and conspiracy with his wife. As Abraham was to Lot and his wife, so was Barnabas to Ananias and Sapphira. Perhaps the horror of the lie to the Holy Ghost, whose fulness of indwelling made the true sacrifice possible, and its terrible doom, have diminished the appearance of Barnabas' excellence in the

great surrender, by fascinating the reader at this point in the story. But the writer must have had distinct purpose in putting the characters so closely together, and thus making each the contrast and foil of the other. Barnabas shows, in holy deed, what Ananias might have been in the same Church life and condition, and was not. Barnabas is the soul ascending, by means of Christlike sacrifice. Ananias is a soul that turned the public work of the Holy Spirit, and the love of the brotherhood in Christ, into an occasion of mortal sin. Equally great privileges, and equally grave perils, attend the communion of saints now; though the final issues of conduct are not so immediately visible. Where the Holy Ghost can be most powerfully, and yet most tenderly met (and that is in the Church of Christ), there the soul's spiritual opportunities and perils are of the greatest and most intense nature. Men need walk warily among the powers of the world to come.

The impulse that wrought in this 'socialistic' movement (as it may be called for lack of a better

word) in the early Church, must not be forgotten. It was, what is later on noted pre-eminently as a mark of Barnabas—'fulness of the Holy Ghost.' If it had not been sufficiently mentioned in connection with the movement, and with all the Church life of which the donation of property to the brotherhood is one expression, yet a devout reader would naturally associate so great a sacrifice and the impulse to it with a special endowment of the Holy Spirit. But the assertion that 'all were filled with the Holy Ghost, and spake the word of God with boldness,' in Acts iv. 31, is too close to the story of the upgiving of property that begins in the thirty-second verse to be separated from it.

Nor would anything less be adequate to produce so great a surrender, than this large reception of the Spirit of God. It is important that this practical self-devotion of some to poverty should be recognised as a fruit of the Spirit's fulness, in order that the movement may be kept distinctive from theories and schemes of social reform that are sometimes classed with it. More-

over, it was the impious and insulting burlesque of this self-devotion in Ananias and Sapphira that brought upon them more than the grief of the Holy Ghost. Peter's awestruck words are— 'Ananias, why hath *Satan filled* thy heart to lie to the Holy Ghost, and to keep back part of the price of the land?' Where Barnabas and the other saints had been full of the Holy Spirit in their voluntary poverty, Ananias had let his heart be filled with Satan. And to Sapphira Peter said, with equal awe of reverence, as became one who was speaking for the Holy Spirit, 'How is it that ye have agreed together to tempt the Spirit of the Lord?' 'Fulness of the Holy Ghost,' then, was the distinctive mark and cause of that first attempt at what may be called 'Christian socialism.'

It is in having such an origin and character, that the generous movement which is chiefly associated with the name of Barnabas contrasts so vividly, to Christian minds, with principles, schemes and movements that bear similar names now-a-days, but are born of a sadder parentage and in a less wholesome atmosphere.

What may be called 'socialism' or 'communism,' is one and a very specific thing when the full indwelling of the Holy Spirit of the God of all the earth, and the Father of all men, makes a man's money or land seem nothing to him compared with the bond of Divine and brotherly love that links him with the poor of this world, in whom, through the gift of Christ, the same Holy Spirit is equally dwelling. In such case, they that know Him 'who for their sakes became poor, that they through His poverty might be made rich,' will not, cannot look on the gifts or entire renunciation of land and money and all earthly gains, for the sake of others, with contempt or criticism, or even a half-sympathetic pity; but, rather, with admiration and thankfulness. They will think that it was good for the Church and the world to see what the fresh impulse of the Holy Spirit, working in hearts that were consecrated by the love of Christ, could do for the poverty and weakness of the Church and the world. And if they think that it is not possible to repeat the gracious and glorious experiment in this age, it

will not seem to them that the impossibility is because there is now more faith and zeal and love, and a fuller reception of the Holy Ghost than there was among the followers of the poor Christ of old, but rather because there is less.

Also, it will be probably felt that this has something to do with the difference between apostolic Church life and the modern condition of Christian fellowship, that we do not estimate union in Christ that binds rich and poor in Church life equally, so highly as we ought to do, nor hold it so dearly as the great martyr spirits in Jerusalem, who became poor for the sake of those who were already poor. They that thus suffered loss were as truly martyrs as if they had suffered the spoiling of their goods, instead of gladly surrendering them to others. For it should not be forgotten that it may betoken an even greater consecration to Christ for a man or woman to give up all that this world holds or has brought for poorer brethren (the Lord Himself being invisible and inaudible), than to leave all at the wondrous call and authoritative summons of a visible and audible Saviour. Matthew, leaving

the receipt of custom, may not be placed higher for so doing than the unknown of the Church in Jerusalem, who sold all and gave it to the apostles for their poorer brethren. Their names are only in the honour roll of the Lamb's Book of Life. Such a 'socialism' or 'communism,' call it what men please, is most admirable, and God is to be thanked for showing the spirit of it so early in the history of the Church.

But 'socialism' or 'communism' is very different from this when it only formulates and gives voice to the angry or heart-sick demands of the empty-handed of the world, who, looking at others that are more than full-handed in this world, and have more than heart could wish, without being more worthy of the same than others, and yet will not really spare for the needy, protest against the earth being any longer made, as they judge, a paradise for the rich and socially powerful at the cost, as they deem, of being also and at the same time an inferno for the weak and poor. It makes all the difference in the world whether 'socialism' express the life of the Holy Spirit in a man, mak-

ing him glad to work and share, and more anxious to do this than to heap up wealth for himself; or whether it only give tongue to the envy and heart-soreness of a man whose every breath is, and ought to be, a protest against the needless inequalities and poverty of the world. The former is the rich gift of God, and ought to be greatly valued; the latter may be at best only the premonitory warning of the Judge of all the earth, and should be reverently feared.

In the presence of the social disquiet of the time, and among the imaginings and speculations, theories and controversies that express, and to a certain degree ease off, the dangerous pressure from below—all of which gather round what is called 'socialism' or 'communism'—no Christian man who would do his part in the day in which he lives can afford to pass by the action of Barnabas and his comrades and the distinction that it so clearly marks out. For, sympathize as much as any one may with the unrest of the straining and groping after better things, in social and international life that are among the best signs of the

times, the Christian man can have little, if any, hope from the plans of the mere social theorist, and still less from the mean contention of the political partizan. The only Christian hope for man is in brotherhood, and the only way to that comes from the recognised equality of the position of all in Christ Jesus to God the Father, who has given the members of His human family charge of one another in days of need and darkness. From Jesus Christ alone, and only by the Holy Spirit given through Him, has there ever been seen in any hopeful way and degree an adequate impulse to the communion of a love from which brotherhood would naturally come.

Away from Christ the struggles to which the world seems committed must take the form of a wretched conflict between the 'haves' and 'have-nots,' the 'have-nots' trying to get from the 'haves,' and the 'haves' striving to keep fast hold on what they have, each abusing the other meanwhile. The course and issue of this need no prophet to forecast. But where the Spirit of Christ is there will be at the least a hope that the future will see

effort on the part of those who have this world's good to arrange how they can most wisely and helpfully share with the 'have-nots;' and in the sharing, raise the level of the destitute, and exorcise the pauper spirit in the sunlight of Christian brotherhood.

Meanwhile, it will be, and is open to those who live in the refuse of the nations and among the dregs of society, to ask, Has the professed Church of Christ sought her alliances in the form of brotherhood with the lowly, and her joy in ministration to the poor by a sympathetic and equalizing self-sacrifice? That she is doing, or showing inclination to do this now more than ever, is a half confession of neglect and sin in the past. But this and subsequent increase of earnestness in such social and national work ought not to be taken up by Christians in order to meet a crisis and stave off a worse state of things, because of the cowardly cries of those who fear to lose their state and possessions, and who know not God in Christ. It ought to be, and can be, truly blessed of the Father of all, only when it is the outcome of a larger re-

ception of the Holy Ghost, and the consequently more kind recognition of the brotherhood of man and the family of God. Let them throw the child to the wolves who are only seeking to save their own skins.

A good act does not carry with it the power of being always an equally good example. Men are, as a rule, poor imitators of a good deed. They are, too frequently, like children writing a copy. The first line written under the copper-plate heading may be fair and good, but the next will be, most likely, an imitation of the first written quite as much as of the accurate headline. And so it will go on down the copy, each line farther from the top being worse than the next before, through unconscious imitation of it; until the bottom line will have very little resemblance to the perfect model at the head. In like manner, while it is possible to believe that the act of Barnabas and his friends may be looked at as the first and noblest of all the donatives to the Church and ministers of Christ, the story of which may be found on almost every page of Church history;

and while it would be right to think that the great example in Jerusalem has acted as a copy for imitation and an impulse in such later gifts, yet it is very difficult, nay, impossible, to trace in most of the ecclesiastical donatives any real resemblance to the beautiful marks of the first renunciation of property to the Church and its leaders.

In truth, as the act of Barnabas fills up one of the brightest and cleanest pages of the Church's story, so many of the later donations on which Church possession and power are in some communities most boldly, yet precariously built, only draw the eye to some of the foulest and most blotted pages of the records of the Church on earth. When, as with vast numbers of them, they were called 'the price of sin' for the living, or being given on death-beds, can only be understood in the nature of a safe and presumably profitable investment; and when, as frequently was the case, they were extorted by terror, while the poor soul was stretched on the rack of remorseless superstition; or, as with the donations of Constantine the Great, are most intelligently understood, not so much as

the pure outcome of piety, but at least equally as the result of civil and political prudence : in all such cases, and many others, it is impossible to associate them with the deed of Barnabas, while yet it is conceivable that his action may have been the far-off impulse to unspeakably poor and mean imitations.

Though, however, the noble sacrifice of the Son of Consolation may have had very imperfect following, yet the poor copies have not been without good in many instances. Anything that will make a man lose himself to any degree, or give up anything from a motive that has elements of goodness in it, is thankworthy, and carries with it some blessing to others. Thus, the sight, by a poor peasant, of a king's crown hung over the altar of a church with a lamp or lamps pendent to it, must ever have had a power to teach the infinite superiority of spiritual things and Divine mysteries to temporal power, social rank and its gaudy signs. And it might not be destitute of help now-a-days, for some to be reminded that in the very form of church lighting by circlets with lights attached to

them, called *coronæ*, there is in all likelihood the modern representative of the act by which a king gave his crown for the honour of Christ's Church, and the Church received it, turning it into the decorative and useful furniture of the house of the Lord.¹

The donation of Barnabas is high above all such things. It was no atonement for sin. It was not extorted by terror trembling under ecclesiastical threats. It was no investment for the days to come, with probability of good interest. It was not an appropriation, by an ecclesiastical caste, of means by which power could be handed on with increasing strength, to each generation of an exclusive order, and by which an earthly-minded Church could fight the coarse world with its own base weapons. It was not a gift that was to be so kept through generations, as to become, because of the meanness of men, a premium by endowment on the niggardliness of other ages. It was an upgift of property by a living man, to be used for the then needs of the poor; and to be exhausted, as to its

¹ The ancient *coronæ lucis*.

money value, in so doing. And as such, and as being the immediate fruitage of the fulness of the Holy Ghost, it is worthy the imitation of all times. For God calls every generation to meet its own necessities ; not by the inheritance of dead men's money, but by the living consecration of living men and women. And God summons each Christian age to leave a legacy, not of unfulfilled obligations, but of well done duties, that have met, as far as man can, the needs that God has brought to the surface of time. For the other, and it may be, larger needs of following generations, it is to be held as true that 'the Lord will provide.'

It was in such a spirit that Barnabas and his co-workers made the great surrender. And it cannot be too much insisted on, that whether the act were prudent or not, the impulse was born of the Holy Ghost. In order to make such action possible, every faculty and capacity must have been, for the time, under the influence of God. Every heart chord must have vibrated responsively to the gentle, yet strong, flow of the wind-like

Spirit of our God. As a consequence, Barnabas held, valued, and at last entirely gave up his property for the poor. From which great act more commonplace Christians might learn, if nothing else, that they are to hold and value all things by the Holy Spirit; and that it will depend entirely upon the extent to which any one is receiving the Holy Ghost, as to how he is able to hold for God, and with due remembrance of the needs of men, the wealth of this world, the perishable but needful riches of the earth.

Nor must it be put into the background of forgetfulness that these primitive believers offered up their all in the spirit and by the same power that wrought the greater renunciation of Calvary. For it is with a spiritual and sensitive amplitude of meaning that takes in the true motive power of all Christian sympathy, and the impulse to the simplest of Christian sacrifices, and binds them all to the great sacrifice of the Cross, that the Scripture declares that He to whom the Holy Spirit was given without measure, 'through the Eternal Spirit offered Himself without spot to God.' In

the same power the lowliest believer in Christ Jesus still offers his sacrifices to God.

Later ages of the Church have seen other admirable renunciations of saintly and self-forgetting men, like Francis of Assisi, whose choice of poverty has had thrown round it all the halo of romance, and the perilous adornment of marvels that require credulity unspeakable for their acceptance. But, without lessening by one atom the real glory of such (and very great their honour is), it may be well to note how different is the setting, and how much simpler and more innocent is the story of the great surrender of Barnabas and his brethren in the Church of Jerusalem. The simplicity and ingenuousness of the story in the *Acts of the Apostles* make a deeper and more abiding impression on a devout heart than any tale of marvels could produce. The effort to understand the meaning, the working, and the likely effects of so strange a movement compels the mind to the conclusion that the holy heart almost naturally reaches, that the scriptural is the only explanation ; and that it was the outcome of the full

reception of the Holy Spirit, disposing believers to such a state of heart and mind as enabled God to work in them 'to will and to do of His good pleasure.' And surely that great event of sacrifice was permitted, nay, brought about, in the very early days of the Church's life, in order that for all time Christ's people might see how the Holy Spirit helps the faithful in Christ Jesus to know, as their Master knew and said, that 'it is more blessed to give than to receive;' and how 'the Lord loveth a cheerful giver.'

CHAPTER III.

THE GREAT ACT OF CHAMPIONSHIP.

THERE is nothing more told of Barnabas till Saul of Tarsus comes on the scene of Church history. Saul's coming was brought about by strange ways. It would seem that a system of daily ministration for the help of the poor and widowed had been formed in the Church of Jerusalem, with the apostles as responsible almoners. Evidently communism, as ordinarily understood, was not practised in the Church. As Peter said to Ananias, the example of Barnabas was not obligatory on any of the faithful. But the care of the poor was inseparable from Church life. In this work it was that serious difficulty rose. Christ's grace was then, as ever, working transformation in very poor material. And it ought not to occasion the least surprise that jealousies and recriminations rose about the administration of the Church's

alms. They who have had similar work to do in modern Churches and communities will often have found the like kind of difficulty and danger rising, and will have felt that it needed the great grace of God to keep a gentle and loving spirit in the work, and to prevent the soul sinking to the level of the mechanical duties of a relieving officer.

The apostles evidently felt that an emergency had come ; for their call to the Church in the difficulty is full of instruction. ' The twelve called the multitude of the disciples unto them, and said, It is not fit that we should forsake the word of God, and serve tables. Look ye out therefore, brethren, from among you seven men of good report, full of the Spirit and of wisdom, whom we may appoint over this business. But we will continue steadfastly in prayer, and in the ministry of the word. And the saying pleased the whole multitude : and they chose seven whom they set before the apostles : and when they had prayed, they laid their hands on them.'

The name of Barnabas is not in the list of the

seven. If he was in Jerusalem at the time, it is almost singular, and nearly warrants the supposition that he was counted among the apostolic men who were to continue steadfastly in prayer and in the ministry of the word. To reserve such a man for such work was wise. For while the apostolic request was for men full of the Spirit and of wisdom to administer the Church's alms and freewill offerings (the lowliest Church service being thus declared worthy, and needing highest qualities for its Christ-like discharge), yet the economy of energy and time, and the right use of varied spiritual gifts, required that the apostolate and eminent teachers should be left free for the special duties that were attached to their gifts. The inclusion of Barnabas in such a supposition would go with the meaning of his Christian name, and would additionally support the assured belief of the high esteem in which the Church of Jerusalem held the Son of Exhortation.

The election of Stephen among the seven, and the apparent flaming out of his ardour and devotion to Christ on being appointed to the respon-

sible service of the Church, practically changed the fortunes of the Christian community. He being in all likelihood a Hellenist, the many synagogues of foreign Jews in Jerusalem would give him special opportunities for work in proclaiming Christ. If he spoke in those synagogues, as is likely, it will be very significant to remember that among those who disputed with Stephen were some of the synagogue of Cilicia and Asia. The probability being that these had a separate synagogue, it is almost certain that Saul of Tarsus frequented it. If so, it is likely that he heard, to his horror and indignation, Stephen's bold declaration that Jesus of Nazareth, who was crucified at Calvary, was the Messiah of God. If such were the case, what he, later in life, calls his 'exceeding madness' against Christians, would have been born and stimulated into fierce growth by the word of the Hellenist Stephen, whose arguments neither he nor his bigoted friends were able to withstand. And certainly it would not have surprised the early Christians in Jerusalem to find that, when the plot against Stephen was concocted by men of

Cilicia, among others, and was carried out in the stoning of Stephen, 'the witnesses laid down their garments at the feet of a young man named Saul,' of Tarsus, in Cilicia, who 'was consenting unto his death.' In this grievous way, and with all the mournful memories that were attached to the death of the first martyr, the name of Saul comes into the Christian record : and comes to stay.

It was a certainty that with the advent of such a man as Saul of Tarsus into the area of the Church's life, something of distinctive and clearly outlined character would appear. And this was so almost immediately ; for it was to his animosity the Church owed the scattering of believers from Jerusalem, that led, as will be seen later, to the most fruitful sowing of the seed of the kingdom that the Acts of the Apostles narrates. He was one of the men who bring events with them. Saul himself raised the storm that drove him to the shelter of the Rock of Ages.

And when the omnipotent sweep of Christ's arm had gathered him safely in, the new man who appeared out of the revelation of Christ to the

persecutor, was as fervid as of old, and the same distinctness of effect accompanied his life. For when at Damascus his conversion to Christ was completed, Saul's proclamation of the Lord Jesus, and the confusion of Christ's enemies before his attack, crystallized the fluent hate of the Jews into murderous purpose and effort. Day and night the city gates were watched by the bloodthirsty Hebrew plotters, in case the arch apostate tried to get out. Death was at every portal waiting for his step. And Saul escaped assassination only by the bold and anxious plan of the men who are called so early—'his disciples,' who let him down through the wall in a basket. Their designation and action show how quickly the strong personal influence of Saul had marked the believers in Damascus. The heat of his enemies' hate, and the depth of his disciples' love, both witness to his impelling power, and the intensity of his devotion to Jesus Christ.

Once clear of the city, the fugitive must have looked up for guidance to his new Lord and Leader, as never before even to the God of his

fathers. The morning must have found him far on the way to Jerusalem. That was the place where his enemies would not think his daring could lead him ; and where he and his friends might hope that, for a while at least, he might find a safe asylum with the Church of the metropolis. Thither Saul went, and seems to have found a lodging with Peter. But when 'he assayed to join himself with the disciples, they were afraid of him, not believing that he was a disciple.' It was of a day later than this that Paul afterwards wrote of the Churches of Judea,—'They only heard say, He that once persecuted us now preacheth the faith of which he once made havoc ; and they glorified God in me.'

On that first visit to Jerusalem, after his conversion, Saul could scarcely look for a warm welcome ; and it is certain he did not receive one. The tired, hunted man must, at first, have been the loneliest spirit in the city that, only a few years before, was full of light, social happiness and warranted dreams of ambition for him. Nor ought his solitariness to be wondered at. A few years before Saul had left Jerusalem with what he thought an honourable

commission from the leaders of his nation. Nor did it trouble him that the blood of Stephen was on his garments, and that his name was cursed by the sorrows of those whom he had haled to prison, or whom he had bereaved of their dear ones. He had left the metropolis and shrine of his ardent faith, thinking of Calvary as the place where the chief deceiver of his day had been rightly hung between heaven and earth on the accursed cross, as a warning to blasphemers and deceivers. Returning to Jerusalem, his heart was full of the love that made him at a later day, when writing to the Galatians an epistle that refers to this very time, use words that have been the stimulus of the faithful in Christ Jesus ever since :—‘ Far be it from me to glory, save in the cross of our Lord Jesus Christ, through which the world hath been crucified unto me, and I unto the world.’

When he stole into Jerusalem, as secretly as he had left Damascus, Saul could not but have wondered whether he had done wisely in coming there. Old Salem friends were now his bitter enemies. Naturally his tender and tried heart

looked to the disciples of Jesus, his new brethren, for the welcome, the fostering love, and the gentle care that would soothe a heart already burdened with a heavy cross. But Saul must also have remembered how some had hidden in terror from him a few years before. Long after he remembered it, and told Agrippa : 'In Jerusalem I shut up many of the saints in prisons, and when they were put to death, I gave my vote against them.' It was no wonder, then, that the believers in Jerusalem were afraid of him. Everywhere cold dread, or fear, or active hate would meet him. Jews were mad against him, and Christians, at the best, suspected him. In such case Saul might have recoiled in almost despair, and have been, in all seeming, lost to the Church and the world. If so, he would not have been the last, by many, of those who have been lost, so far as we can judge, to the high uses of God in the Church of Christ, because of not receiving the warm and brotherly welcome they had a right to expect from the community of the lovers of Jesus.

It is at this point that Barnabas reappears in the

history ; and his coming into prominence again is precisely in the character that might have been expected from his previous conduct. It is with the same gentle thought of others' needs that already has been seen to mark him, and with a keen intelligence of the quality and worth of men that will be found in him in the later events in Antioch. For, in the emergency caused by Saul's sudden appearance as a Christian fugitive in Jerusalem, Barnabas, humanly speaking, saved the new and unrecognised apostle for God and man. He 'took him and brought him to the apostles, and declared unto them how he had seen the Lord in the way, and how that He had spoken to him, and how at Damascus he had preached boldly in the name of Jesus.'

Whether they had known each other before or not, it seems certain that Saul had opened his heart to Barnabas. And it may have been more solace than can be imagined, for the wearied fugitive to find in the Son of Consolation one in whom he could put trust, and in whose heart he could find a sympathy and tender brotherliness which, at that

time, must have been sorely needed. But Barnabas was always more than a comforter by word and sympathy, good though that was and is. Barnabas ever put his feeling and judgment into action ; as at the surrender of his property. While the company of the disciples were afraid, and had infected the apostles with their weakness, Barnabas dealt with the difficulty at head-quarters. It is not too much to say that he championed Saul before the evidently doubting and hesitant apostles. He saw in the new convert a man fit for the great service of God.

The readiness to help others and forget himself, that even a superficial reading of the story of Barnabas shows, shone out here brightly and steadily. Not even good men are often ready to find in those who need their help the presence of high and noble qualities. Rather the reverse is frequently the case. The need of help of any kind is often allowed to veil excellencies and obscure powers that therefore droop and languish till they seem to be no longer alive. It is a fine spirit that can quickly recognise another soul's need ; and, at

the same time have a no less quick perception of that needy soul's powers, and its capacity for being touched to fine issues. And it betokens most gracious magnanimity, even in a Christian man, for him to lay himself out to use or make his opportunities of advancing other lives, that are for the while in lowly case, in any high and honourable service. We are so accustomed to think of Paul as the champion of the cross, and as the wonderful apostle in whom Christ was pre-eminently glorified, as that it is almost difficult to realize that once, after good service for Jesus, he stood at the door of the Church, wistfully and with hunger of heart longing to enter into its warmth and light, and met with no kind look or loving outstretching of hand, but rather with what was almost the shutting of the door in his face.

From the reality of that he was saved by Barnabas. It may not seem to have been much to do. But if the risk of Saul is remembered, and the possible consequences of that risk in Saul's life and Church history are borne in mind, then the act of Barnabas in championing Saul of Tarsus,

before the fear of the Church, and against the reluctance of the apostles, will stand out as one of the most momentous and fruitful events in the history of Christianity. It is, of course, inconceivable that Saul would have forsaken the work of the Lord if the apostles and Jerusalem Church had passed him by; for he was Christ's 'chosen vessel.' But it is at least supposable that; if Saul had been once driven from Jerusalem in the early days of his Christian faith, by the fear, coldness and suspicion of the Church and apostles, as well as by the hate of his Hebrew enemies, his heart and faith would have been strained by a trial that would have marked all his after history, even though he had mastered the disappointment and sadness of such an experience.

Be this as it may, his nobleness of action, and hearty assertion of Saul's worth and service are to the great honour of Barnabas, and worthy of all imitation. He saw the power and purpose of God, where others, even apostles, had chiefly the fear of man. He recognised a nature and a life that were full of promise for the service of man

and the great glory of God. And in this warm recognition of Saul's worth, Barnabas not only contrasts with the whole apostolate, but shows a beautiful kinship to the Master who, looking at the high qualities of one who was turning away from the life of sacrifice that He had opened to him, is said to have 'loved him ;' even though, for the time, they parted. The Church, in a sense, owes Paul the Apostle to the magnanimous judgment with which Barnabas welcomed Saul of Tarsus ; and to the earnestness of his interposition against the closing of the door on him by the timid believers in Jerusalem.

CHAPTER IV.

BARNABAS' GREAT LIBERALITY AND WISDOM.

TENDER feeling and wisdom are not always joined together. Often the presence of either seems to involve the absence of the other. The practical man, in many cases, seems rough and ungentle. Piety and prudence are not always in alliance. Holiness and common sense, impulsive sensitiveness and wise forethought, are not necessary adjuncts. Yet such associations and combinations they were that Christ wanted, and still wants, in His people and workers ; and such He counselled them to cultivate. When He sent out the first evangelistic band of disciples, He said to them, ' Be ye wise as serpents, and harmless as doves.' They were to be practical and gentle. Their love and zeal for God and man were not to be of the invertebrate type. Nor was their practical wisdom and strength to be coarse, rude and exacting. They were to be strong sons of God,

yet tender as only the Holy Spirit can enable a man to be. They were to be meek and lowly, yet of a fibre most tenacious, and of a most foreseeing wisdom. And, certainly, never were men and women banded together in enterprises that called for more of both classes of qualities than those in which the leaders of the Church were involved in the first years of the Christian brotherhood.

The part of Church history to which we now come shows this high endowment of the Holy Spirit in Barnabas to a degree unsurpassed in any of the Church leaders. Zeal and wisdom, gentleness and forethought, go hand in hand with him. To see his wise, broad-hearted and tender judgment and influence, it is necessary to review the story of the Church up to the time when the growth of the Christian community required a larger area than Palestine, and a broader field than that in which the life and sympathy of the Church in Jerusalem moved. That time and demand for catholicity, as against localism and nationalism, came from what, at first, must

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have appeared to be an awful and unqualified sorrow for the Church, viz., the martyrdom of Stephen. But, like many of the great mercies of God, the depth of the darkness was significant of the great light that was coming. This will be seen in the course of the present chapter.

No attentive reader of the life of our blessed Lord can be other than impressed with what, looked at any way, or from any point of view, is a strange and admirable thing: the calm, all-assured confidence with which Jesus looked forward to the whole race of man listening to the story of His earthly life and death, as the result of His people's persistent effort to tell that history to all. This was to be the Church's greatest duty and highest honour.

Two special instances of this will come to the forefront of memory with all Bible readers. One is the account of the woman who anointed the Lord in the house of Simon the leper. Of that act the Saviour lovingly and confidently said, in answer to the mean economy of the apostles—
'Whosoever the gospel shall be preached in the

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whole world, that also which this woman hath done, shall be spoken of for a memorial of her.'

This saying, quietly and confidently uttered, in the near prospect of an ignoble death and burial, and in circumstances that are altogether alien from any warrant of such a hope, tells much of the inner thought of Jesus. It shows Him to be One, the horizon of whose hopes stretched far beyond the near foreground of the life of a generation or two, and, in truth, was not co-terminous with the earthly history of the human family.

The other instance was in surroundings that were as little likely to justify an expectation of universal dominion, though it was very different from the scene in the leper's house. It was in what St. Matthew gives as his last recorded words of Jesus. It was that claim of supreme authority in heaven and earth on which the Lord grounded His people's loving and confident allegiance, and by which He enforced the duty of His Church to disciple all nations in the doctrine of the Father, Son, and Holy Ghost. This He strengthened still more by the oversweeping

promise or assertion—'And lo, I am with you always, even unto the end of the world.' From the elevation of that august moment it is evident that Christ's vision overlooked the ages, and saw the nations everywhere learning of Him; being taught by His Church, as the height of all human duty on earth, to 'observe all things whatsoever I have commanded you.' Many other words of our Lord bear similar interpretation; but these two incidents are known of all.

With such things green in their memories, it would be natural to expect that the apostles, and the body of believers under their care in Jerusalem, would, at the very least, have felt their hearts ever more widely opening before the pressure from within of a vast hope and desire for all men; and would have strained the full energy of their ability towards revealing unto all men, without distinction, the unsearchable riches and world-wide sweep of the grace of Christ. Surely, it might be thought, John's heart would have followed his Master's, and his eyes have been with his heart, in a love-lighted and prophetic vision of all men listening

and yielding to the love of God the Father in Jesus Christ. Without doubt it might be imagined that Peter's impetuous nature would have caught fire from his Master's glowing hope ; and in that flame all barriers between him and others would have been burned away. Who could have imagined anything but that he would have seen only the universal expectation of Jesus waiting its prophesied fulfilment, and would have hastened to bring it towards accomplishment ? To Peter, the middle walls of partition between men, it would be natural to think, would have been broken down for ever. And so might a sanguine faith have expected of the whole apostolate and Church in Jerusalem.

What, however, are the plain facts of the apostolic history ? They are such as show the spotless honesty of the writer in telling the story of the Church. They betoken a daring faith, born of integrity, that could put down for the perusal of after ages everything needed for true instruction and edification ; even though it might seem to be at the risk of a somewhat sullied

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repute for zeal in the apostolic body, and a diminished estimate of the intelligence and charity of the Jerusalem Church. For it is clear, from the first eleven chapters of the *Acts of the Apostles*, that neither the Church nor the apostles, at that time, thought of going beyond the narrow circle of Judaism with the gospel of salvation.

But even in this there may have been a temporary advantage. For is it not quite possible that if, from the first, the apostles had preached to Jew and Gentile, and held intercourse with all of the uncircumcision quite as readily as with Hebrews, they would have roused against them at the outset a madness of antagonism and spite that would have made impossible the time of comparatively quiet growth and Church consolidation, that passed between the day of Pentecost and the martyrdom of St. Stephen? It would be to get a gain out of loss, so to believe. But even that would hardly compensate the loss of catholicity in the earliest feelings of the primitive leaders of the Church. The wide sweep of Barnabas' feeling, and of Paul's later thought,

their fuller vision and more faithful apprehension of the purpose of the Father and the meaning of the death of Jesus, do not seem to have drawn the apostles' eyes or lured their hearts to the real longings of God for all. The horizon of Jesus was, at that time, too wide for the apostles. The probability is that in the Church at Jerusalem, and in the whole apostolate, there could not have been found one man, unless it were Barnabas, to whom the unrestricted preaching of the gospel to all nations would have seemed other than presumptuous, unauthorized and worthy of sharp reproof.

No wonder the Judaizers had such baneful power in the early Church. The marvel is that they had not more. The story of the way in which the Church was emancipated from its bondage and narrowness is one of the most important parts of apostolic history. Opening the first great crisis of the Church, and leading to a most blessed issue, it witnesses to the presence of Christ with His people even in unworthiness; and to that gentle leading of His most tolerant love which

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shone before the apostles during His earthly life, and the spirit of which is shown in the memorable words of John, that 'having loved His own which were in the world, He loved them to the end,' notwithstanding their petty thoughts of His kingdom at that time, and in spite of their rivalries for place and power in His dominion.

The crisis, and the emancipation that Barnabas confirmed, began with Peter, and the account of it opens with his mission to Cornelius the centurion. The sword of the Spirit was pointed, for swift thrust, at the very heart of Judaic narrowness and limiting of the Holy One of Israel, when Peter the apostle was sent to Cornelius, an officer of the Roman army of occupation. But before this it is significantly enough recorded that Philip the deacon, flying from the persecution that raged against the Church upon the death of Stephen, went down to Samaria, there working many miracles, and bringing many of the people to Christ. On receiving news of this in Jerusalem, the apostles sent Peter and John to examine what had been done. The two apostles confirmed

the work, and on their way back to Jerusalem they preached in many villages of the Samaritans. To the preaching of the gospel in Samaria there seems to have been no objection. Perhaps the semi-Judaism of the Samaritans, their being circumcised, and the remembrance that Jesus had visited and won converts in Sychar, had force in influencing the apostolic judgment. And, also, they may have remembered their Lord's words immediately before the ascension:—'Ye shall be My witnesses, both in Jerusalem, and in all Judea and Samaria, and unto the uttermost part of the earth;' though this would involve their limiting the application of 'the earth,' naturally enough, to their native land. Again, when Philip was sent to the Ethiopian eunuch, no censure was passed upon his reception of such a man; the truth probably being that he was a proselyte of Judaism. The Jewish circle, the Hebrew area was, even in these instances, still the contracted limit of Church work.

The same thing is more evidently and sadly the case in the events linked with the mission to

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Cornelius. The piety of Cornelius, the true piety of loving communion with God and good works among men, is emphatically attested. Its acceptance with God is notably affirmed. Yet it is no less clear that this man was regarded by the apostles and the Jerusalem Church as being outside the pale. The message of the love of God in Christ Jesus was not for him so long as he was uncircumcised. Remembering such things, it ceases to be a matter of wonder that the Gentiles thought that Christianity was only another form of Judaism. This misconception would easily explain much of the suspicion with which the Christian name was soon received. The preparatory vision given to Peter, in which the true Jew speaks, presumes an objection on Peter's part to have any intercourse with the uncircumcised. The universal cleansing, the Divine abolition of the scornful judgment between things common and sacred, between men common and divinely favoured, clean and unclean, was not yet recognised by the apostle. But the vision cleared his sight, gave character to his faith, widened his

thought of God, and ennobled his actions in the Church, for a time, at least. Peter's perplexity about the vision gave place to recognition of the Divine love as being broader than the narrow measure of a Jew's mind. The fog of national bigotry and sectarianism was cleared from his thought by the clearer influence of the Sun of Righteousness. In the fresh light he even saw farther than he understood. Though he never seems to have grasped the fact of the universality and impartiality of the love of God in Christ as Paul did, yet in the incident of the vision by which God prepared him for going to Cornelius, Peter learned more of the truth than ever he knew before; and to some degree the truth made him free from the fetters of the Jew.

How influential Jewish narrowness was, or had been, in his spirit, even in thoughts of the work of Jesus, is evident from Peter's first words to the friends and kinsfolk of Cornelius: 'Ye yourselves (Romans) know how that it is an unlawful thing for a man that is a Jew to join himself or come unto one of another nation; and yet unto

me hath God showed that I should not call any man common or unclean.'¹ Up to that time the restriction of the Jew had more influence than the liberty of the Christian. He was still, and, as he almost said, only a Jew. And how the old curse of a base nationalism and sectarianism still lurked in his mind, and how he fought against it, can be seen in this—that when the Holy Ghost fell on the whole company of Italians, and the six Jew companions of Peter were 'amazed, because that on the Gentiles also was poured out the gift of the Holy Ghost,' without affiliation to Judaism, Peter turned to his comrades, as if fearing that they might oppose what he was about to do, and said, 'Can any man forbid the water, that these should not be baptized, which have received the Holy Ghost as well as we?' The silence of the six gave consent. Thus the Church, in the person of Peter and six other Hebrews, was committed to stepping over or breaking down the middle walls of partition between Jew and Gentile, and was started on its true path of faith and love—the

¹ Acts x. 28.

faith and love that recognise, assert, and enforce among men the great truth of the family of God, that no earthly, religious, social, political, or national distinction is to stand for a moment athwart the coming together of all, as brethren in Him, with whom 'there is neither Greek nor Jew, circumcision nor uncircumcision, barbarian, Scythian, bondman, freeman.' The Church was now to be even compelled to learn that man, because he is man, is dear to the eternal Father, and is within the compass of all the efforts and hopes of our Lord Jesus Christ.

Such a great forward step, such a startling innovation, not to say outrage, on all that had been dear to the Jew, and had given him abounding self-confidence, could not happen without difficulties rising, and those most serious. Light is slowly received into the mind by the majority of men and women; and still more slowly used, when it shows an unaccustomed line of duty, or a perilous, but unavoidable course of action. Even religious persons, perhaps they most of all, if there be anything sectarian about them, cast the old

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skin of partial truths, or errors, or limiting beliefs, most slowly and painfully, and in the process are most sensitive and irritable. It is quite natural that it should be so. It is not easy to bury dead gods. A good lesson it is to teach and learn, that says—

‘Heartily know,
When half-gods go,
The gods arrive.’

But the half-gods go very slowly, and often leave behind hearts too weak and strained to give a great welcome to the real gods. The altar fire may be quenched, and the altar itself broken and desecrated, but memory will circle round it still. Peter could not have had bright anticipations at having to announce to the Jerusalem Church the dethronement of Hebrew pride. He must have foreboded trouble when he got back to the metropolis; and it certainly was ready for him. Before he reached the capital, the tale of what he had done had been told, as a scandal. With evident heartburning and nursing of what they thought was righteous indignation, ‘they of the circum-

cision' waited for Peter's return. Their charge was ready. They were prepared to arraign even the leader of the apostles in a matter of Hebrew privilege and prejudice. If he were not put upon trial, yet Peter had in a fashion to stand on his defence, as before a session of the Church, for a breach of sacred custom and for unauthorized actions.

Very wisely Peter limited his answer to a simple statement of what had taken place at Joppa and Cæsarea; and with a comprehension of his six colleagues in the honourable indictment, he finally, warming with his defence, and more like the true Peter than he showed himself to be when, later in life, he was involved in a similar controversy, said: 'Who was I, that I could withstand God?' Then, as men utterly amazed, 'they held their peace, and glorified God, saying, Then to the Gentiles also hath God granted repentance unto life.' Their thoughts of God, their judgment of the meaning and scope of the work of Christ, and their conception of the Church's duty, may be judged from their own words. It ought, all round,

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to beget great gentleness of judgment and feeling concerning doctrinal differences and even errors, to remember that so highly endowed a Church as that in Jerusalem, was in one direction so contracted and mean as to deny the possibility of repentance unto life to any but Jews, and to do this in the belief that God thought the same.

Meanwhile, a movement in the same direction of liberty and universality for the Church, as the custodian of the faith and work of Christ, had been going on, even more notably, elsewhere. Away from what seems ever to have been the benumbing shadow of the Jerusalem Church, the Spirit of the love of Jesus showed itself in a great longing for 'all to be saved and come to the knowledge of the truth.' Flying from the persecution in which Stephen was offered up to God, some of the lowly, unknown believers who had been in Jerusalem at that time, 'laymen,' as they would now probably be called, 'travelled as far as Phœnicia and Cyprus and Antioch, speaking the word to none save only to Jews.' The ecclesiastical ritual and national prejudices of their race clung

to them, as to others, like a curse. But some of the wanderers were of a greater spirit, a nobler mould, a diviner heart; and these 'men of Cyprus (like Barnabas) and Cyrene, when they were come to Antioch, spake to the Greeks also, preaching the Lord Jesus. And the hand of the Lord was with them: and a great number that believed turned to the Lord.'

This is altogether more notable, more influential and more noble than the act of Peter. The exiles had no authority given to them by heavenly vision. No appeal came from God-acknowledged piety to stimulate their zeal, or to suggest that even among the Gentiles there might be some who feared God, and who ought to be gathered into the fold of Christ. They were simple-hearted, earnest, private members of the Church, who spontaneously, yet deliberately, did that which Peter would not have done, and which it needs no imagination to picture the rest of the apostles questioning. Whether this preaching to the Gentiles took place before or after Peter's visit to Cornelius cannot be positively determined, but it seems likely that it

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was before. If so, it should not be forgotten that it is to the unnamed, unknown brethren, to the wandering martyrs for the faith in Antioch, that the high honour of first openly showing the broad spirit of the love of Christ to all men must be given. The first real act of true mission work for the Lord; the earliest recorded act of obedience to the world-wide command of Jesus; the first true comprehension of the Spirit of the Lord wrought out in holy, loving act, is to be associated, not with the apostolate, not with prominent believers whose names are in the apostolic records, but with the few bright and noble-hearted fugitives in Antioch, of whom it can only be said—'their names are in the Lamb's Book of Life.' For the time, Antioch was of more interest and importance to the world and Christianity than Jerusalem: and the doings of the exiles near the grove of Daphne were of more consequence than the acts of the apostles, under the shadow of the temple of Jehovah. Like many, very many of the most useful and blessed events in the history of the Church and of nations, this great act of self-

emancipation, and consequent enlargement of the Church and enlightenment of the world, came from the lowly, and unrecorded by name, the almost forgotten—save by God.

The broad and generous mission work in Antioch, joined with the act of Peter at Cæsarea, could not be other than a serious business to what may readily and rightly called the mother Church in Jerusalem. They who had been quick to judge and censure Peter the apostle, would not be eager to endorse the record of the doings at Antioch by private members of the Church. Peter, with a vision from God and evident warrant from heaven, would be, to those of the circumcision in Jerusalem, a different person from the non-commissioned believers who had crossed all the prejudices of the Jew, without waiting either for a sign from heaven or an injunction from earth. It was a crisis of the Church's history in which some of the most momentous questions waited answer and decision. Unwise action at that time might have committed Christianity to a long course of ignobility and narrowness, the marks of which would have lasted to this day.

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Nor, if there had been un wisdom shown, could it rightly be regarded with surprise. Seeing how dim the vision of many Christians is even now, as to the sweep of 'the love that passeth knowledge,' it would be easy to expect too much from Hebrew Christians eighteen centuries ago, when 'the report concerning the things done in Antioch came to the ears of the Church which was in Jerusalem.' And nothing could more truly witness to the spirit of wisdom being given to that Church, in fulfilment of the promise of our Lord, than the action taken at that time. The story at this point is suspiciously brief, as if the writer would pass over, almost in silence, events that it would not be of use to any for him to write out at length. The commission of Barnabas to Antioch (for that was the resolution of the Church) was too serious a thing to have been lightly decided. And when it is found, in the fifteenth chapter of the *Acts of the Apostles*, that the Church in Antioch a little while after was troubled by some from Judea, who would have destroyed the good broad foundation laid by the martyr exiles, and who

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taught the brethren, saying, 'Except ye be circumcised after the custom of Moses, ye cannot be saved ;' and when it is seen that the Judaizers caused such misery by their narrowness and ignobility, that 'Paul and Barnabas had no small dissension and questioning with them,' and had at length to go up to Jerusalem to consult with the apostles and mother Church on this very matter; and when it is read that in that Council some of the Church at Jerusalem 'rose up and said, It is needful to circumcise them and charge them to keep the law of Moses,' it will not seem likely that the commission of Barnabas, the most charitable of all at Jerusalem, was decided easily, quietly and unanimously.

But the sending out of Barnabas is the bright spot in the whole. He was the one man who was fit, by largeness of heart and self-sacrificing love, and also by devout intelligence, to judge the believers in Antioch, and wisely estimate their work. His greater fitness, say than Peter's, is clear from what Paul tells us of Peter's doings when he came to Antioch: 'When Cephas came to Antioch, I

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resisted him to the face.' And so great was the antagonism of Jewish narrowness at that time, that Paul says, 'before certain came from James, Cephas did eat with the Gentiles: but when they came, he drew back and separated himself, fearing them that were of the circumcision. And the rest of the Jews dissembled likewise with him; inso-much that even Barnabas was carried away with their dissimulation.' Barnabas was stronger away from Peter than with him.

When Barnabas arrived at Antioch, the double joy waited him of finding there some of his old Cypriote friends, and also of seeing that it was their action, in conjunction with their Cyrenian brethren, that was giving a catholic character to the Church, and so an adequate opportunity to the grace of Christ. Barnabas' large heart took in at once the blessedness of the work that he saw. Such a widening of the possibilities of the faith of Christ, and such an onward step by the Church's energy were congenial to all his desires and cravings. 'When he had seen the grace of God, he was glad; and he exhorted them all, that with

purpose of heart they would cleave unto the Lord.'¹ Or, probably, and better, as some ancient authorities put it—'that they would cleave unto the purpose of their heart in the Lord.' If this latter be the true word, it surely can refer only to their intention and act in preaching the gospel to all, without the mean limitation of Hebraism. So understood, it might suggest that Barnabas felt that such boldness of faith would certainly bring great opposition from some of the brethren, if not actual alienation. That was the almost quick result. But they must have been the better prepared for it in Antioch by the exhortations of 'the Son of Exhortation.'

It is beautiful to note the same spirit of consolation in Barnabas at Antioch, that had been so attractive to the brotherhood long before in Jerusalem. And as in the city of David the apostles had renamed him for his loving and tender service there, so the comment made on his influence among those who were doing a truly apostolic and inspired work in the capital of

¹ Acts xi. 23.

Syria is equally characteristic and significant: 'for he was a good man, and full of the Holy Ghost and of faith.'¹ Such notes and comments are very rare in the New Testament, and ought to carry the greater meaning, wherever they occur, because of their rarity.

By the approval of Barnabas as the representative of the Jerusalem Church, the catholic-spirited believers of Antioch were thus confirmed in their new departure. The mission of the Church to the Gentiles was made to appear of no less moment than its work among 'the lost sheep of the house of Israel.' Another man might have spoiled the larger venture of faith and hope. But the critical time was as good as safely passed, when Barnabas gave and the Antioch Church received his loving and altogether approving counsel. Henceforth, whatever Jerusalem might think or be, the light of the true world-wide faith, and the radiance of a full obedience to the command of Christ in its universality, shone brightly and attractively at Antioch. And Barnabas secured the benefits of having safely

¹ Acts xi. 24.

passed the critical moment, and made a relapse almost impossible, by the singular wisdom of his decision concerning the guidance of the Church. Apparently he thought that he himself was not the most fitted to lead the Antioch Church in its new departure, or to steer it through the broken water that must be ahead. With that lowly estimate of himself which is one of his chief marks, he stepped into the background, in order that a better and abler man, as he thought, might come to the front. Barnabas had the greater grace of Christ that enables a man to take a second place, of freewill, and prefer another to the foremost rank. Ten men would think themselves fit or worthy of the first place; where one would see the greater qualification of another for the honourable eminence of leadership. It needs more grace of God to enable a man to step back in order to bring forward another, than to contend for, or to aim at pre-eminence. Conceit will make a Diotrophes,¹ but only the love of Christ can make a Barnabas.

¹ 3 John 9.

Was there such a man as Barnabas might well make way for? Where was there one whose native vigour and ability marked him as a leader of men? and whose culture, wisdom and sagacious charity would find a genial air in the large-hearted Church of Antioch? Was there any one in the Church at that time, if not Barnabas, who could take the leader's place on that broader road, and in that wider march which had only just been opened? Barnabas never for a moment seemed to have thought of any one in the Jerusalem Church. But there was a man, hiding, it almost seems, in the city of his birth, to whom Barnabas turned as by the inspiration of the Lord. Soon the Son of Exhortation was on his way to Tarsus, to seek Saul and exhort him to a duty for which no one else was fitted. To all appearance Saul was forgotten by the Church, and might be losing touch of the brotherhood. In reality he was in Tarsus a fugitive from the persecution of the Jews. There, or in unrecorded places, he had been for several years. But certainly there had been no waste of time with him. He was being tried in the refiner's

fire. He was being slowly tempered for the wars of the Lord.

The visit of Barnabas to Tarsus, and his exhortation to Saul that he ought to go to Antioch, and throw his vigour and zeal into the new life and wide enterprise there beginning, was like Farel's arrest of Calvin in Geneva, pointing him to an open door that none could shut, but that he, at peril, might refuse to enter. Calvin's purpose was to pass through Geneva, and so on to Basel, in hope of there studying in quiet. No sooner did Farel know that Calvin was in Geneva, than he begged him earnestly and often to stay at Geneva and help on the work of the reformation in that city. Calvin for a time refused, until Farel adjured him by the name of the Lord to stay, denouncing that 'if he refuse to employ himself in that Church in the work of the Lord, God would curse the repose that he sought for the convenience of his studies.' Calvin stayed.

There is neither need nor warrant to suppose any want of much persuasion on Barnabas' part to bring Saul to Antioch. Happily, as with Calvin,

Saul heard God's voice in his brother's, and returned with Barnabas to Antioch, thus being brought forward a second time for the Church and the world by the Son of Exhortation. For a year Saul taught in Antioch, as only he could teach, the fuller word of the life and love of God. The significance of his work, and the special character that Saul impressed on the Antioch Church, are evident in this, that 'the disciples were called Christians first in Antioch.'

This name, cast, in all likelihood, at the believers in Jesus, on the top of a sarcasm or sneer, by the vicious and pleasure-loving people of Antioch, tells very distinctly of the character of the Church in that city. Under the influence of Barnabas and Saul, it compelled attention. It made its mark where Apollo and Daphne had lured the vile and the pleasure-lovers of all the world, to the seduction of the groves and streams that were shamefully consecrated to their names. The shallow wit of the heathen labelled the new community by one of its own most frequently used words. No Jew could have given the name. Probably no Christian of

himself, would have assumed it. The names of the brotherhood of believers among themselves were very distinct and very beautiful. 'Disciples,' 'saints,' 'brethren,' etc., are the titles by which they designated one another in the Lord. The nickname of 'Christian' shows that the office of Jesus as Messiah, His redeeming work for the world, to which He was sent by the Father of all, was, as might be expected, the prominent feature in the teaching of Barnabas and Saul, and in the belief of the Church that they taught. As the effect of this, even the shallow Antiochian heathen soon perceived a very marked line of cleavage between the Church of Jesus and Judaism. They laid their hand, in quite a clumsy fashion, on the one chiefly prominent thing that they could hear of and see in connection with the faithful in the Lord ; and calling them 'Christ's men and women,' Christians, gave a name that has stuck ever since, with increasing honour, to the Church of the living God. Followers of Christ, of *The Christ*, as believers used the phrase, was the new name of the Church. But it was not so much the Christ of

the Jew, but the Messiah of humanity, whom Barnabas and Saul preached, and whom they, with the Church, adored.

In Saul's ministry at Antioch there was a coincidence that has had the most important meaning for Christianity. Can it be doubted that, as certainly as he printed his character on that Church, so surely the special circumstances of the Church's origin, its peculiarly broad life and catholic zeal, must have had a special influence on him, and must have exercised a great formative power on his comparatively early Christian life and work? Saul's life and ministry among the believers in Antioch must have largely qualified him for the mission work among the Gentiles, to which, with Barnabas, he was ordained by that Church. For it is very noteworthy that it was not by the apostles, nor by the mother Church in Jerusalem, but by the Church that was first opened to the Gentiles by the handful of earnest believers in Antioch, that Barnabas and Saul were commissioned to the true work of the Church.

Nothing could have been more opportune,

stimulating, and educating for Saul, than to be the leader with Barnabas of the broad-hearted believers of Antioch. And when, by evident signs from heaven, and the inspiration of the Holy Spirit, the first formal mission to the Gentiles was undertaken by Saul and Barnabas, they being ordained thereto by the presbyters and teachers of Antioch, to whom, 'as they ministered to the Lord, and fasted, the Holy Ghost said, Separate me Barnabas and Saul for the work whereunto I have called them,'—then was finally made clear to all who would see, that the narrow ring of Jewish misconception of God was broken, and that the love of God was to all men, without exception, as that of a father to his children. How slowly the Hebrew element in the churches came to receive this in its fulness, is evident from Paul's life-long conflict with Judaizers and limiters of the love of God.

But through and over all that weary strife, Saul's clear trumpet-like cry can ever be heard, like the voice of a fighter sure of victory: 'Is He the God of the Jews only? is He not also of the

Gentile?' 'There is one God and Father of all.' It may be that the truth is not fully learned by some, even yet. Still the faith of some, if it be a true faith, lives in a ring narrow as that of the contracted hearts of the Jews, who would have wrecked the Church as it began its voyage, and even before it had cleared the harbour. That they did not, was due to the great wisdom of Barnabas in summoning Saul to the new work of the Lord.

Let any one try to imagine what the consequences to religion and to the world at large would have been if the universal command of Jesus had been kept in abeyance by the confirmed or unremoved scruples of Peter and others, or by the dominance in the Church of 'those of the circumcision,' especially at Jerusalem, or, to any extent, by the continuance of even the appearance of an authoritative and absolute metropolitan character in the Church that met under the shadow of the Hebrew temple. As Paul said, when writing of this very matter, Christ would have 'died in vain.' God would still have been mis-

conceived. The brotherhood of man would have been less than a dream. The love of the Father of all, revealed in the life and death of the only begotten Son, as the atonement for the sins of all the earth-born prodigal children of God, would have been clouded or distorted into a degrading and repellent partiality. The world would have been left still seeking a God who might be worthy a sinner's love. That men now know that the love of God is wider than the best thoughts of the best men, and that His ways are not the Jews' ways, nor His thoughts the Jews' thoughts, but wider, deeper, higher, more blessed ; and that this was wrought into the faith of Christendom, though not yet adequately believed, is due to the lowly men, the unknown faithful ones, who at Antioch stepped out from the bounds of Jewish ignorance, and to Barnabas, who was so wise and liberal and loving in his judgment of their innovation and the needs it created ; and to Paul, who, commissioned of God through them, took his proper place as the great leader of the larger hearts of the Church, and with his own heart constrained by the love

of Christ, ever maintained that 'in Jesus Christ neither circumcision availeth anything nor uncircumcision, but faith which worketh by love.'

It was a great crisis through which Christ led His Church by the hands of Barnabas and Saul, to truer thoughts of the Father and Himself, and of the purpose for which the Holy Spirit had been given to the Church. And when Jerusalem became only a memory, and the last symbols of the temple were carried to Rome as part of the plunder of a Roman victory—the Church of Jerusalem being at the same time scattered—then the thoughts and feelings of the faithful in Christ Jesus could be, and were, more distinctly than ever turned from local and limited judgment of God's purposes and ways, to Christ's own all-comprehending word, as the sum of all needed belief: 'God so loved the world, that He gave His only begotten Son, that whosoever believeth on Him should not perish, but have everlasting life'; and also to Christ's parting command, as the marching orders of the ransomed of the Lord: 'Go ye, and make disciples of all the nations.'

CHAPTER V.

THE GREAT CHARITY OF BARNABAS.

IT has been seen in the previous chapter how Barnabas met a great emergency of the Church of Christ by what was a practical self-effacement, when he brought Saul of Tarsus to the front at Antioch. Barnabas served Christ incomparably by taking a second place. He gave larger opportunity to the Holy Spirit, by bringing Saul from his unsuitable, if not wasteful, retirement. Few see the evidence of the Spirit of God working by them in their having assigned to them an inferior position. Very few would discover a prime condition of the growth of the Church, and the progress of the truth as it is in Jesus, in their own voluntary abandonment of a prominent position, and their equally voluntary effort to bring forward another man, who could more effectively meet the need of the time and place. This it was that Barnabas

did. To it Paul owed his great opportunity of making the Church catholic and all-comprehensive, instead of remaining sectarian and local.

From that time Saul became the leader; the two comrades marking out lines of Church activity and evangelistic extension on which no subsequent workers have improved. They toiled with the aim of proclaiming the gospel universally and without restriction. Probably unmarried, and untrammelled by worldly cares (for both had counted all things but loss, for the excellency of the knowledge of Christ Jesus their Lord), the once Jewish aristocrat and the Hebrew landowner went forth, earning their scanty earthly living as they journeyed. They left behind them for ever a world of restriction of holy purposes and limitation of desire in God; and everywhere they revealed that new life in which 'there cannot be Greek or Jew, circumcision or uncircumcision, barbarian, Scythian, bondman, freeman, but Christ is all and in all.'

What enfranchisement of spirit theirs must have been! Their only limit was their knowledge of

the love of God ; and that grew vaster and more glorious every day. The farther they went, and the more they saw of man's sore need, the wider and the more comprehensive the grace of God in Christ appeared to their thought and feeling, and the more imperative became the call for them to reveal the heart of God by preaching the cross of Christ. Often the two beautifully balanced men, who were so suited to one another, must have felt the word of the Lord and the work of God's Spirit to be dearer to them because of the sight in each other of the varied effects and graces of the Holy Ghost. Christ's preciousness as the Redeemer of men, and the importance of the trust that was committed to them, as to no others, of declaring the gospel to the Gentiles, must have been continually intensified by the sight of the heathen and corrupt towns and villages through which they journeyed. The more they travelled, the more the world must have stirred them, as it stirred the Master, when the people appeared to Him as sheep not having a shepherd. And the Lord must often have blessed the companionship

and co-operation of two such men, to the deeper consecration of both. Each must have felt what has been specially written of Paul,—

‘Oft when the word is on me to deliver,
Open the heavens, and the Lord is there;
Desert or throng, the city or the river,
Melts in a lucid paradise of air.

Only like souls I see the folk thereunder,
Bound who should conquer, slaves that should be kings,
Hearing their one hope with an empty wonder,
Sadly contented in a show of things.

Then with a rush the intolerable craving
Shivers throughout me like a trumpet call;
Oh, to save these! to perish for their saving,
Die for their life, be offered for them all!’

Quite in harmony with this sure and sweet sympathy of souls covenanted together for Christ and man, is the scene that shows Paul and Barnabas once more together in Antioch, after the trouble in the Council of Jerusalem. There, notwithstanding the congenial atmosphere of that most Christlike Church, the longing grew in Paul to know how the Churches that they had founded thrived, and how their newly-found brethren fared

'And Paul said to Barnabas, Let us return now, and visit the brethren in every city wherein we proclaimed the word of the Lord, and see how they fare.' Who would have thought that at this point, and in consequence of Paul's wise and loving proposal, the long union of the two saints would be broken, and the two great hearts parted in work, never again, on earth, so far as we know, to go on the business of the Lord together? But such a sadness was brought about. And, strangely enough, it came through the good qualities of the two apostles. Each had the defects of his virtues. Among the best of men living, they were but men, at the best. The loving-kindness of Barnabas prompted him to take with them his kinsman, John Mark. The whole-heartedness of Paul resented, as an affront, the company of one who had tired of the work, if he had not cowardly deserted his great leaders at Perga.

We can only conjecture the real reason why Mark proved recreant; and many conjectures there are. Whether he dreaded the journey through the Taurus mountains to Antioch in Pisidia, and

the undoubted perils of the way; or, because of his connection with Peter and the Jerusalem Church, could not thoroughly go with the larger-hearted purposes of Paul and Barnabas, that the Judaizers had condemned; or whether, as has even been surmised, Mark was envious for Barnabas' sake, because of the greater prominence that Saul had at Paphos, after which he assumed the name of Paul, cannot be decided. But from the fact that John Mark went on at once from Paphos to Jerusalem, back to his mother's house, it may be fairly enough conjectured that the impulse which led him to go with Barnabas and Saul from Jerusalem had spent itself, and that his young life had, for the time, exhausted its power of consecration to so active, dangerous and criticized a work; and that, in consequence, he longed for the rest and refreshing of home and loving surroundings. Paul's finer soul evidently was strongly roused by Mark's defection. The reasons for it were assuredly not such as Paul could honour, and he saw Mark part from them and the work with fervid indignation. To Paul, John

Mark was one more who made 'the great refusal'; another of those sad ones who put their hand to the plough and look back, thus showing themselves unfit for the kingdom of God. To refuse to go on in the work of winning souls from sin and for Christ, was an act that Paul could not understand. For any man to draw back from the privilege of opening the kingdom of heaven to believers must have seemed to Paul a miserable weakness at best.

It need not, then, excite surprise that when Paul wished to go over the old ground, and consolidate the work that had been done, he should refuse to take with them one who had preferred sunning himself in Jerusalem to carrying the cross over the Taurus mountains. 'Paul thought not good to take with them him who withdrew from them from Pamphylia, and went not with them to the work.' That was the sting of the falling away. The work had been left undone, so far as Mark was concerned. Christ's call to service, and the sight of an open door for duty, had found no response in one whose ardour ought to have been

as a flame. But may it not be safely judged that Mark's being again in Antioch (which is a quite natural and warrantable supposition) betokens repentance, and a longing to redeem the past? He must have known that Antioch was the place where, sooner or later, he would meet again the indignant eyes of Paul, and the gentler reproaches of Barnabas. It may be rightly enough believed that Mark's return to Antioch meant a renewed longing for the company and work of his kinsman and their great comrade. In that case, it would be like the gentle goodness of Barnabas, to wish to take the penitent back again, with full restoration to brotherhood and service. And it would be like Paul, at that time, to refuse.

The result was the most piteous thing in the *Acts of the Apostles*. 'There arose a sharp contention, so that they parted asunder one from the other, and Barnabas took Mark with him, and sailed away unto Cyprus; but Paul chose Silas, and went forth, being commended by the brethren to the grace of the Lord.' Evidently the Church in Antioch was involved in the sad dissension; and

it is not meaning too much to read the commendation of Paul and Silas to the grace of the Lord as implying that the earnest feeling of the brethren in Antioch went with him whose zeal for the honour and work of the Lord consumed him. There is no record that Paul and Barnabas ever met again till they stood together, each glorified with the martyr's crown, before the throne of God and the Lamb.

It is not easy, nor is it pleasant, to apportion credit or blame in this pitiful story. Both of the apostles may be honoured for something admirable in the conflict of judgment. Deep tenderness in one, and intense earnestness in the other show themselves attractively. But without wronging Paul, it may be fairly enough wished that he had remembered how once he needed a champion and friend, and had found both comforter and advocate in Barnabas. Nor would it have been to Paul's dishonour if he had understood that when Barnabas as good as said, 'Give John Mark another chance,' he was doing what was, in spirit, like taking Saul of Tarsus by the hand, and

justifying him before the apostles and elders in Jerusalem. Barnabas, in this contention, is consistent with all that we know of his goodness and kindness to weak and tried brethren. It may be also that Paul is in harmony with all that we know of his consuming surrender to the service and honour of his glorified Lord. But zeal for the Lord, even in an apostle, does not necessarily carry perfection with it. Nay, the very reality and intensity of a noble spirit's surrender to God and Christ may liberate and give play to feelings, the activity of which may be as scorching fire to any who do not fully agree with the devotee, or who seem to him to come short of the high standard of his own jealous consecration. It was this spirit, sometimes in coarser forms, that refused re-communion and almost hope in the early Church to those weaker ones who had lapsed from open faith in times of persecution, and that barred such from ordination in the Western Church, and restored a lapsed but penitent presbyter only to lay communion.

Barnabas, however, was true to the lines of his

gentle and more gracious character. Under the shelter of his great charity and tenderness John Mark found a new prospect of consecration to Christ, that gave him the opportunity of struggling manfully against the weakness of the past, and proving himself to be one who could 'endure hardness as a good soldier of Jesus Christ.' The charity that hopeth and believeth all things, of which Paul became the great exponent and example, showed its living and life-giving power in Barnabas. It made him, in this instance, the protector and helper of Mark, as once he had been of Saul. Perhaps Barnabas could sympathise with John Mark the more easily because of his own weakness. For if it is accurate to think that the influence of Jerusalem and the Judaizers was strong on Mark, and made him shrink from the wider work that so distinctly and irreversibly ended all the exclusive Hebrew claims before God, Barnabas could feel tenderly to such a sensitiveness. For he had known a similar weakness in yielding to 'them of the circumcision.' And while Peter was largely responsible for it, yet

there was less excuse for the surrender of Barnabas than for the half-heartedness or cowardice of Mark.

This ought to be remembered in judging the antagonism of Paul and Barnabas about the reinstatement of Mark. Paul tells, in his letter to the Galatian Churches, of the sore trouble he had in Antioch with the Judaizers. And he writes with an evidently special feeling about Barnabas, 'When Cephas came to Antioch, I resisted him to the face, because he stood condemned. For before that certain came from James, he did eat with the Gentiles: but when they came, he drew back and separated himself, fearing them that were of the circumcision. And the rest of the Jews dissembled likewise with him; insomuch that even Barnabas was carried away with their dissimulation.' Paul was incapable of surrendering the catholicity of the gospel and the Church for James or any party; and it may be that this matter was at the bottom of the difficulty about John Mark. But because of this—his very passion of glorious conviction, for which the Church ought

ever to be thankful—Paul may have been unable at that time to judge dispassionately his younger and unmatured brother. No one is fit to estimate another's sins, or sit as in judgment, from which there is to be no appeal, on a brother's weaknesses.

This, too, Paul learned later in life. For where one is strong, another is weak. The power, or pressure, or allurements that seems irresistible to one soul may be as an idle wind to another. One will stand where his brother will fall, or collapse where his comrade will be as granite. A great pity is becoming in every one concerning others' sins and weaknesses. Even of the worst it can be said, and ought to be remembered, that no one knows what is resisted. The return of a soul to light and duty should ever be hailed and helped; and the spirit's return and vigour of upspringing, rather than its previous fall, should be the data on which the character is estimated. So judged, John Mark will not lose honour in the long run, for it will appear that, in all likelihood, the gentle and warm nurture of Barnabas developed in him

a stronger fibre than that which excited Paul's vigorous condemnation.

It is certain that later in life and work Paul's judgment of Mark entirely altered; and it is equally certain that the changed estimate had warrant. The once weak comrade won his way to the very heart of Paul. Nothing could more tenderly evidence this than the fact that when he was a prisoner in Rome, Paul wrote most tenderly of John Mark, as one of the 'men who have been a comfort to me.' And later still, he longed for Mark's company—wanting him even when he was anticipating the coming of his 'true child in faith'—Timothy. To Timothy it was that Paul sent the message, 'Take Mark, and bring him with thee: for he is useful to me for ministering.' This need not be understood as meaning that Paul thought himself wrong in his earlier judgment, but that he magnanimously let the past be past, and as lovingly welcomed the brother who had redeemed past mistakes, as he had earnestly condemned the wrong when it was done.

It is with positive relief to feeling that we look

across the breach between Paul and Barnabas and Mark to such a reunion towards the end. We only wonder, where was Barnabas when Paul wrote those words? Was he alive or dead? Had he been alive on earth at that time, surely we should have learned something of him in the life and letters of Paul. It seems very probable that the legend that gives the Son of Consolation early martyrdom in his native Cyprus is true. If so, Paul must have found, in reunion with Mark, some healing for the deep wound his heart must have felt when he was separated from the loving man who had been more than a brother to him, and who now wore the martyr's crown. To receive Mark to his heart would be, for Paul, like a fragrant offering to Barnabas' memory.

But at the time of the contention they parted, it seems, for ever on earth. Barnabas, in his great charity, took the young and mortified life of his kinsman with him, and went to Cyprus. There, abruptly, for him, authentic history ends. Sudden as it seems, the end is fitting. The record leaves Barnabas, as later it leaves Paul, working for

Christ and His Church. Who can doubt that the Lord's work was bravely and wisely carried on in Cyprus by the gentle and thoughtful Son of Consolation? And if the tradition of martyrdom there be true, we may also confidently believe that the spirit of Barnabas in death crowned and glorified a life that was peerless in the primitive Church for rare qualities of consecration, and that has been a noble example to all ages of the Church of a soul that was possessed, in some things, more amply than any of the apostles, of the mind that was in Christ Jesus.

CHAPTER VI.

THE EPITAPH OF BARNABAS.

THE high Christian excellence of Barnabas, that must be evident to even a casual reader of the fragmentary notice that we have of him, is very beautifully accounted for by the writer of the *Acts of the Apostles* in a passage that has been noted in its place. 'He was a good man, and full of the Holy Ghost and of faith.'¹ There is a tone about these words that sounds almost like the reverent and tender mention of one who is already dead. It reads like an epitaph. While it is the endorsement of Barnabas in the momentous events at Antioch, it might also stand for the explanation of all the good of his Christian life. And though the meaning of the words 'a good man' were to be watered down till it came to little more than the negative amiability that passes for goodness,

¹ Acts xi. 24.

and that may be found in many who are out of Christ, yet, from the circumstances of its application to Barnabas, there would be a need of giving it a distinction and positiveness of character that would redeem it from triviality. For the 'goodness' of Barnabas was of that fine and sensitive quality that made him sympathetic with any advance of others' goodness and increase of their spiritual life. It answers fully to what is meant by saying that 'only the good discerns the good.' It was an inner delicacy of fibre that made it possible for his nature to respond quickly, spontaneously, to any sign of worth and virtue in the world or the Church. Like as a soft note drawn under a great bell of similar pitch will set all the mighty mass of sonorous metal vibrating and responding with the same note more grandly uttered, so the note of goodness near Barnabas' large nature brought out at once the answering but nobler tone of sacred joy.

Souls of which such tender goodness as this can be truly declared are among God's most gracious and refining gifts to the world or the

Church. They are souls made truly after 'the pattern in the mount.' They multiply, by their quick recognition and sympathy, the force of feeble virtues, and witness to the vital presence of the active and mighty grace of God in the Church, where others see and hear only the signs of ordinary human intercourse. They see God and the movement of the Divine life, where others only know of the grasp of human hands, or only see the friction and fire that come from the interchange of human thoughts. In the case of such, and when, as with Barnabas, self seems as good as dead, and gentle helpfulness of others abounds, it will seem natural to look for more elevated sources of life and excellence than any merely earthly beauty of character can make known.

Such a higher level of life, such a nobler fount of thought and action, is found in that fulness of the Holy Ghost with which the writer of the Acts hallows the name and memory of Barnabas and his goodness. This was the hill whence came his help. A personal God with him; a

personal Holy Spirit dwelling in him and having him dwelling in His Divinity; a Holy Spirit filling him and vitalizing him, as the sea fills and at the same time bears the life that floats in it,—this made Barnabas' true being; and this made his goodness what it was. It made it to be Divine goodness, the goodness of God visible in a man. For this, what willingness in the day of God's power Barnabas must have shown! What sensitiveness to the movement of God's Spirit! What openness to the waiting and ever-rising tide of the Divine energy, that at length filled every creek and nook of nature till grace was all in all! For him, in the quiet ways of every-day life, it was a Pentecost continued in the gentleness of the dove-like Spirit of God.

And good would it be for all Christians to keep such a disposition to God, such a fulness of the Holy Ghost, ever before them, as the vital contrast of so much that passes for adequate and satisfying Christian life, but is in reality, that sad restraining, or resisting, or

grieving, or going perilously near quenching the Holy Spirit, of which God so anxiously speaks to His people. Fulness of the Holy Ghost in common things is Christian life in truth. The very body of humiliation is to be the temple of the Divine Spirit. The very thorniest paths of temptation and sorrow, as well as the bright ways of joy and peace, may be the chosen courses of the Holy Spirit; as it was with Christ Himself, of whom it is written that, 'being full of the Holy Ghost, He was led by the Spirit in the wilderness during forty days, being tempted of the devil.' But whether tempted or triumphant, rejoicing or in grief of heart, the Christian life may be, true and mature Christian life will be, full of the Holy Spirit of our God.

This, however, is attendant, if not dependent on a quality and power of heart and spirit of which also Barnabas is one of the most notable possessors in the history of religion. That is the 'fulness' of faith with which he is honoured in the apostolic story. Faith may be called the

filial recognition of God, in consequence of which the soul can stay itself and be at peace, where, but for the assurance of God the Father being with it, there would be only disquiet and unrest. With the confidence of God's love and grace working for man in all ways, it can see God in Christ reconciling the world to Himself, and can leave the ordering of all things to the love and wisdom that it recognises are over all. It can rest in the Lord and wait patiently for Him. Fulness of faith takes everything to be in its province; it allows no exceptions, nor is it perplexed by any emergency. Everything, sin excepted, is among the 'all things that work together for good to them that love the Lord.' There is no such thing as blind faith. A full faith always sees God. Christ dwells in the heart by means of such faith.

So, too, faith is the soul's trustful answer to all that God has revealed Himself to be, and to have done for it. It commits itself to the Lord alone, and in consequence the soul has sight of the higher life of man in God, that only a believer

can see or discern at all. It also disposes the heart to give itself up to that exalted life, no matter what it may involve on earth, or lead to when death is left behind. Having fulness of faith, the spirit is attracted to the object of its confidence and desire more and more, and becomes increasingly like Him in whom it believes. It possesses Christ, and with Him the righteousness of God which is by faith. It identifies itself with its Lord. It keeps ever an open channel of communication with God through Christ, in whom the fulness of faith rests. Thence and by means of the faith that grows to fulness, the whole man can be transformed. The righteousness of God in Christ; the perfect work of redemption; the complete sanctification of the Holy Spirit and preparation for eternal glory—these and all other gracious mysteries of the love of God come through the channel of fulness of faith. A less faith may and will be a limitation of blessing. That is why our Lord said, 'According to your faith be it unto you!' The shallow and blocked channel

will limit the power of receiving the blessing, and will also diminish the entry of God the Spirit. Therefore the fulness of faith is to be desired, sought and rejoiced over.

Such a fulness Barnabas had. And though his record is but scanty on earth, yet it is enough to show what tender fruits of holiness and self-sacrifice can be brought to maturity by any heart that will fully receive the Lord Jesus by fulness of faith, and the fulness of the Holy Ghost by the same spirit of filial trust. 'Barnabas was a good man, full of the Holy Ghost and of faith.'

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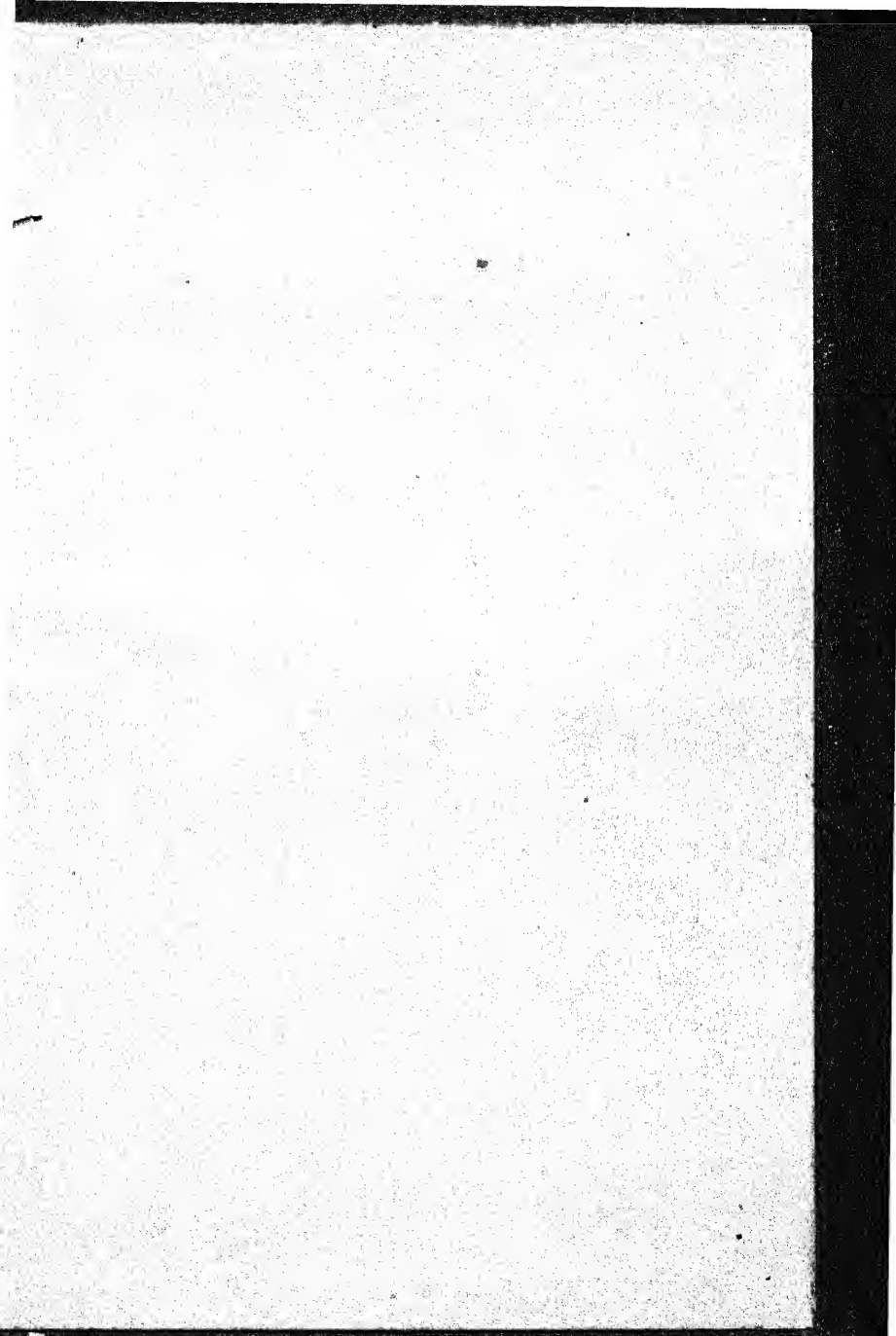
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THIN-LAYER CHROMATOGRAPHY

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Foreword

Thin-layer chromatography (as it is known today) was developed because of a specific need for a rapid method which would separate small amounts of compounds. Soon after its development it was apparent that the method was more than a micromethod for separating compounds. It was a means of (1) investigating absorbents and solvents from column work, (2) following the course of elution chromatography of colorless compounds, (3) checking the course of reactions and (4) carrying out certain reactions such as oxidations, reductions, dehydrations and so forth, directly on the strip or plate. By these means and other reactions applied directly to the unknowns, an insight into the type of compound could be gained. All these functions of thin-layer chromatography are enhanced by the speed of the method and the fact that only minute amounts of material are needed.

In general, a new technique that is developed for the separation and identification of compounds is not a method which, per se, displaces a previous technique. Rather it is an addition to the numerous tools available to the scientist for solving his problems.

A man that adopts the idea that a given technique is the solution to all his problems is missing an opportunity. Many times combinations of various techniques give results which surpass anything that can be attained with any of the individual techniques. Keeping this in mind it is to be hoped that this book will be an introduction to a method to those who have not used it and an aid to those who wish to further their knowledge of the subject.

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Preface

Chromatography on thin layers of adsorbent was conceived as early as 1938 and developed largely during the early 1950's. However, it has been used more extensively since about 1958 and now appears to be in an era of rapid development and almost universal adoption. At such a time, a book setting forth the basic technical details and surveying the field is in order. The author first observed the technique in the laboratory of Professor Hans Schmid in Zürich, Switzerland and was much impressed. This impression and the successful application of the method in his own laboratory have prompted the writing of this book.

The text and the bibliography cover the majority of the literature through December 1962 and a few particularly pertinent papers in 1963.

It is the intention of the author to write a condensed and non-theoretical description of this simple laboratory technique. Thus, the experimental aspects and "tricks" will be taken up, catalogued and discussed in considerable detail. In the last Chapter, an attempt will be made to tabulate, in a concise and meaningful way, as many specific applications of the method as possible. Any theoretical treatment of the subject must consist of two portions. The first is a discussion of general chromatographic theory and the second is a consideration of the deviations from the theory which arise in thin layers. Such a discussion is beyond the scope of this book.

The author would like to express his appreciation to the following persons who have made his career—and, thereby, this book—possible: Professors James L. Hall, C. L. Lazzell, L. A. Pursglove, Katherine Wilson and Elizabeth Frost Reed who are or were on the Staff of

West Virginia University; Professor Carl Djerassi of Stanford University; Professor Hans Schmid of the University of Zürich; Dr. Ulrich Weiss of the National Institutes of Health, and, finally, his parents, Mr. and Mrs. J. S. Bobbitt of Bluefield, West Virginia. Beyond doubt, the author owes his greatest debt to Professor Melville Lawrence Wolfrom of the Ohio State University and to him this book is respectfully dedicated.

The actual writing of the text would have been impossible without the assistance of the author's wife, Jane Ann Hickman Bobbitt, Drs. Roy J. Gritter, Ana Rother and R. G. Jensen of the University of Connecticut, Dr. H. K. Mangold of the Hormel Institute, University of Minnesota, and Mr. Herbert R. Simonds of the Reinhold Publishing Corporation who read portions of the manuscript at various times. The following co-workers actually carried out most of the thin-layer chromatography and assisted in the final checking and proofreading of the text: Mr. Christopher Allen, Dr. Robert Doolittle, Mrs. Judith Douville, Dr. Robert Ebermann, Dr. K. L. Khanna, Mr. Donald Kiely, Miss Judith McNew, Dr. Sardar Mahboob, Dr. Ana Rother and Mr. David W. Spiggle. The author is further indebted to his secretary Mrs. Jean Jemielity for typing the manuscript, to Mrs. Eileen Stock who made all of the drawings and to Mr. Ernest F. Spitzer and the Chas. Pfizer Co. of Groton, Connecticut for the use of their fine library.

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CHAPTER 1

Introduction, History and General Applications

INTRODUCTION

Since the very beginning of the science of chemistry, its practitioners have been plagued by two technical problems. The first is to learn the purity of a given preparation or, conversely, how many components are in a given system. This information is necessary regardless of whether the system is a mineral mixture, an extract of some plant or animal, or the result of a chemical reaction. The second problem is to resolve the system into its pure components so that they can be characterized and studied. The earlier chemists performed phenomenal feats by fractional crystallization and distillation, but it has only been in the last fifty years or so that the various techniques of chromatography have promised true solutions to these problems.

It follows from the stated problems that chromatography can be considered from two viewpoints. One is *diagnostic** or qualitative and the second is *preparative*.* The object of the former is to determine the number of components in a system and to learn, if possible, what they are without actually isolating them. The latter viewpoint involves the separation of a mixture into its components in such a way that reasonable amounts can be isolated and studied. Still a third aspect is *quantitative*, or concern about how much of each component is present.

* These concepts and others in this section have been well reviewed by Keulemans²⁸ and will be only briefly discussed in this volume.

All of the techniques of chromatography are based upon the same simple principle. They involve a moving system of some type (liquid or gas) which is in equilibrium with a stationary phase. These phases are so designed that the mixture to be separated will be distributed between the two. When the stationary phase is a solid and the forces acting between it and the mixture are adsorptive in nature, the technique is called *adsorption chromatography*. When the stationary phase is a simple liquid or a liquid held on some type of support, the chromatography is considered to be *partition chromatography*.

In general, adsorption chromatography involves a relatively non-polar moving phase and works best when the substances to be separated are not very polar. The major advantages over partition chromatography are that larger quantities can be separated in comparable systems and that a controlled temperature is not necessary. Partition chromatography, on the other hand, generally involves polar solvents and mixtures of very polar compounds such as carbohydrates or amino acids. Since it is basically dependent upon the distribution coefficients of the substances in question, which are, in turn, highly sensitive to temperature and other conditions, a carefully controlled atmosphere is required. The line of demarcation between adsorption and partition chromatography is, however, often blurred.

At this point, it would be appropriate to define some of the terms as they will be used throughout this book. The *adsorbent* will be the finely divided powder which makes up the stationary phase in adsorption chromatography, or holds the stationary liquid in partition chromatography. The *layer* will refer to a thin layer of adsorbent, bound or unbound, deposited on a glass plate. The term *spotting* will refer to the application of the substance to be separated to the thin layer. The passing of a liquid through the layer to affect a separation will be referred to as *development* and the liquid itself will be consistently called the *developer*. The term *eluent* will be used only when the complete removal of a substance from the adsorbent is intended. The term *visualization* will refer to the rendering visible of the results of a developed chromatogram. The *R_f* will be the distance traveled by a given substance divided by the distance traveled by the solvent front. Both are measured from the origin.

Column chromatography, as originated by Tswett^{409,410} and since

practiced by legions of chemists, was primarily an adsorption technique used for preparative work. Column chromatography using an adsorbent such as silica gel or cellulose as a support for a liquid was conceived by Martin and Synge²⁴¹ and is partition chromatography for a preparative purpose. The countercurrent distribution of Craig and his co-workers^{59,60} may or may not be defined as chromatography but at least involves similar phenomena. It might also be considered a partition technique with a preparative goal. None of the above techniques have been particularly successful as diagnostic methods due to the extensive labor involved in the preparation of a single chromatogram.

The first major development in diagnostic chromatography was the paper chromatography pioneered by Consden, Gordon and Martin.^{57a,57b} The method was fantastically successful and was rapidly adopted all over the world and in every type of laboratory. The major advantages of paper chromatography are its extreme simplicity and the fact that relatively inexpensive equipment is needed. However, since it is a partition technique with the stationary liquid phase held on a piece of paper, it works best with polar developers and small amounts of polar substances. While techniques of preparative paper chromatography have been worked out and used successfully, they are laborious.

The second, primarily diagnostic technique of chromatography was gas chromatography. This was either an adsorption chromatography, as pioneered by Turner,⁴¹² Claesson⁵⁸ and Cremer,^{61,62} or a partition chromatography, as suggested by Martin and Synge²⁴¹ and introduced by James and Martin.¹⁷⁰ The advantages of gas chromatography are its speed of operation, its almost unbelievable degree of resolution and the fact that the results can be interpreted quantitatively. The disadvantages are that relatively expensive and complex equipment is needed and that the substances to be separated must have at least some vapor pressure at workable temperatures. The use of gas chromatography as a preparative technique has been quite successful although hampered by technical difficulties.

Thus, it would appear that a diagnostic technique which is adsorptive in nature and which combines the technical simplicity of paper chromatography and the speed of gas chromatography is lack-

ing. Ideally, such a method should lend itself to quantitative interpretation and should be useful also as a preparative method. Most of the conditions are fulfilled by *thin-layer chromatography*, or, as it is sometimes called, "thin-film," "open-column," "chromatostrip" or "chromatoplate" chromatography. In essence, this is a type of adsorption chromatography where the adsorbent is a thin layer of some solid deposited on a glass plate support. In operation, it is analogous to paper chromatography; that is, the substance to be separated is placed a short distance from one end of the layer and is resolved by a solvent passing through the layer by capillary action. The development is carried out in a simple closed system as in paper chromatography, but is much more rapid. When the proper solvent mixtures are used, the method can become a partition technique. Thus far, the quantitative interpretation has involved errors of 3 to 5 per cent, but the method shows promise as a preparative means for quantities of one gram or less.

HISTORY OF THIN-LAYER CHROMATOGRAPHY

It is always difficult to sort out the various originators of some concept or technique and to assign credit, because any such development is the result of contributions and ideas from many people and laboratories. Such contributions vary from the initial flash of genius to the more practical aspects such as the introduction of new or more convenient equipment or materials. The historical development of thin-layer chromatography is best divided into two such phases. The first phase is the conception and initial development of equipment and techniques followed by their slow but steady adoption. The second phase is an extensive and rapid development following the invention of new equipment, the standardization of the method and, particularly, the commercial availability of apparatus and adsorbents. A third phase will surely follow in which the method becomes a routine technique applied universally and without particular thought and attention.

In 1938, Izmailov and Shraiber¹⁶⁹ described the use of thin layers of adsorbent on glass plates for the separation of galenicals. The English summary of this paper is as follows:

"A method for chromatographic adsorption analysis is elaborated, based on the observation of the division of substances into zones on a thin layer of adsorbent, using one drop of the substance. The results obtained by the method proposed are qualitatively the same as those obtained by the usual chromatographic adsorption method of analysis. The method enables one to obtain satisfactory results using one drop of the substance under test, very small quantities of the adsorbent and minimal time. The method may be used for the evaluation of galenical preparations and their identification, as well as for a preliminary test of the adsorbent and the kind of developer."

In 1941, Crowe⁶³ described the use of thin layers of unbound adsorbent to help predict the best solvents for column chromatography. In a book entitled "Introduction to Chromatography"⁴⁴³ which was published in 1947, Williams reported the use of thin layers of adsorbent held between horizontal glass plates. The top plate had a small hole for application of the sample and the developer, and the chromatogram was made in a circular fashion.

The use of a binding agent (starch) to hold the layers in place was introduced by Meinhard and Hall²⁴⁶ in 1949. The resulting layers of alumina—"Celite" prepared on microscope slides, were used for the separation, again in a circular fashion, of inorganic salts.

The next and, by far, the most extensive and comprehensive contributions to the development of thin-layer chromatography were made by J. G. Kirchner, J. M. Miller and their co-workers at the U. S. Department of Agriculture Laboratories in Pasadena. In an extensive series of papers (1951–1957) this group investigated various adsorbents and binding agents,¹⁹³ designed equipment for the preparation of "chromatostrips"²⁵³ and used the technique for the investigation of the terpenoids.^{190-194,251-253} Narrow glass plates called "chromatostrips" (0.5 by 5.25 in.) with a silicic acid-starch layer were found to be most useful for their work although they also introduced the use of larger plates.¹⁹¹ The larger plates called "chromatoplates" by Reitsem³²³ were used by him in extensive work on essential oils.³²³⁻³²⁷

Starting a little later, but roughly paralleling the work of Kirchner and Miller, Mottier and Potterat²⁶⁸ worked out a method for analyzing food dyes on layers of nonbonded alumina and even separated amino acids by this technique.²⁸⁷ Some of the other workers in the

earlier phase (prior to 1956) were Ito, Wakamatsu and Kawahara,^{167,168} Labat and Montes,²⁰⁹ Grüner and Spaich,¹³¹ Coveney, Matthews and Pickering,⁵⁸ G. Wagner,⁴²⁶ Rigby and Bethune,³²⁸ Bryant,⁵¹ Lagoni and Wortmann,^{210,211} Demole,⁷⁶ Fukushi and Obata,¹⁰⁸ Onoe²⁸⁴ and Maruyama.²⁴² The earlier phases are well reviewed by Demole in French⁷⁷ and English.⁷⁸

The major reason for the slow development and adoption of the technique during the time from 1951 to 1958 was the fact that equipment and chemicals were not commercially available. More inertia must be overcome in the construction of equipment and the preparation of adsorbents than in signing a check or placing an order.

This last obstacle was removed by Egon Stahl in Germany. Stahl, working with C. Desaga, G. m. b. H., in Heidelberg, devised a system involving standard-size glass plates (5 by 20 cm and 20 by 20 cm), a new apparatus for the preparation of layers, and a standard adsorbent (silica gel-plaster of Paris). Subsequently, the apparatus became commercially available from Desaga and the adsorbent became available from E. Merck, A. G., in Darmstadt. Stahl and his co-workers studied the variables of the technique in a thorough manner,³⁸¹ applied it to many new types of organic compounds³⁶⁶⁻³⁸² and pioneered the use of new adsorbents. These new adsorbents immediately became commercially available. Stahl called the method "Dünnschicht-Chromatographie" or "thin-layer chromatography" ("TLC") and this name seems to be widely accepted.

The applications of the thin-layer technique mushroomed after 1958 so that it is impossible to mention all of the people who have contributed. However, some of the more prominent workers in the field have been Stahl, of course, and Kaufmann, Demole, Gänshirt, Mangold, Stanley and Seiler. The contributions of these people and many others will be discussed later.

The use of thin-layer chromatography is now almost routine in natural product, pharmaceutical and lipid laboratories, but has not been so well accepted by the organic chemists as it might. The applications to various aspects of nonlipid biochemistry have been hampered by the very polar molecules involved, but these problems are rapidly being overcome by the introduction of new adsorbents.

The field has recently been reviewed again by Demole in French⁷⁹

and English;⁸⁰ by Stahl,³⁷² Schorn³⁴⁶ and Wagner⁴²⁸ in German; by Mangold²³² and Wollish, Schmall and Hawrylyshyn⁴⁵² in English; by Michalec²⁴⁸ and Procházka³¹² in Czech; by Vioque⁴¹⁸ in Spanish; by Kokoti-Kotakis;²⁰⁵ and by Jensen.¹⁷⁶

At present, three books have been published on the subject. Two of these, one edited by Stahl³⁷⁵ with contributions from Bolliger, Brenner, Gänshirt, Mangold, Seiler and Waldi and one written by Randerath,^{320a} are in German. The third, by Truter⁴⁰⁴ is in English.

GENERAL APPLICATIONS

The motives for the application of thin-layer chromatography are many and varied. Some of these will be considered in the following discussion. The prejudices of an organic chemist with inclinations toward natural product work will be apparent. The various points will all be fairly obvious, but may, on occasion, suggest a solution to a research problem. Examples will be cited to illustrate the discussion, although it should be noted that they are chosen at random and no attempt has been made to cite all of the possible references.

Preliminary Study of a System or Situation

The small amounts of material involved and the speed with which results are obtained make thin-layer chromatography a valuable tool for preliminary explorations of many kinds. The pharmacognosist or plant chemist can learn something about the complexity and, by the judicious choice of spray reagents, even something of the components of a plant extract. He can then use the technique to follow his gross separation procedures and to learn where the products of interest are. The same reasoning applies, of course, to various aspects of plant and animal biochemistry.

The organic chemist, or indeed, any chemist, can explore a large number of reaction conditions, using small amounts of material, in a short time. This applies to both normal synthetic reactions and attempts to degrade a natural product or unknown substance. The operator can readily learn which reactions produce new products and which yield only starting material; which reactions produce complex product mixtures and which do not. If a given material is sought and

is available for comparison, conditions can readily be found which produce it in the best yield and the cleanest form. The author and his co-workers have used this last general idea in the synthesis of the alkaloids, pilocereine³⁸ and anaferine.³³²

Intensive Study of a Single Reaction

After some idea of the best reaction conditions or best degradation scheme has been obtained for a given problem, thin-layer chromatography can be employed to learn considerably more about the reaction. A crude kinetic study of the reaction can be made by removing samples at timed intervals and chromatographing them simultaneously with starting material and the expected product, if available. This may tell whether stable intermediates are involved, whether the desired product is an end product or an intermediate and something of the course of reaction. At least, such a study will tell when all of the starting material is gone and will aid in establishing an optimum reaction time.

The author and his co-workers have used this approach extensively in the structure elucidation of the glucoside, catalposide.⁴⁰ An example from this work is shown in Figure 1.1. Others have used it in the study of the sequential methylation of phenolic acids,²²⁵ the preparation of *p*-bromophenylosazones of xylose derivatives,⁷ the glycerolysis of linseed oil³³⁵ and the biogenetic production of terpenes in mint plants.²⁰ The presence of an unstable intermediate was shown in the synthesis of pseudoionone.²³⁷

The Isolation of Reaction Products

Thin-layer techniques can be applied to the isolation of reaction products in several ways. First of all, if acid-base partition, distillation or crystallization methods are to be used, one can easily learn something of the precision and completeness of the separations and the purity of the various fractions. Furthermore, the short times involved in learning where the product is will satisfy even the most impatient chemist.

If simple separation methods fail and the products must be separated by chromatography, thin layers offer several possibilities. Quantities less than one gram can be separated by preparative thin-

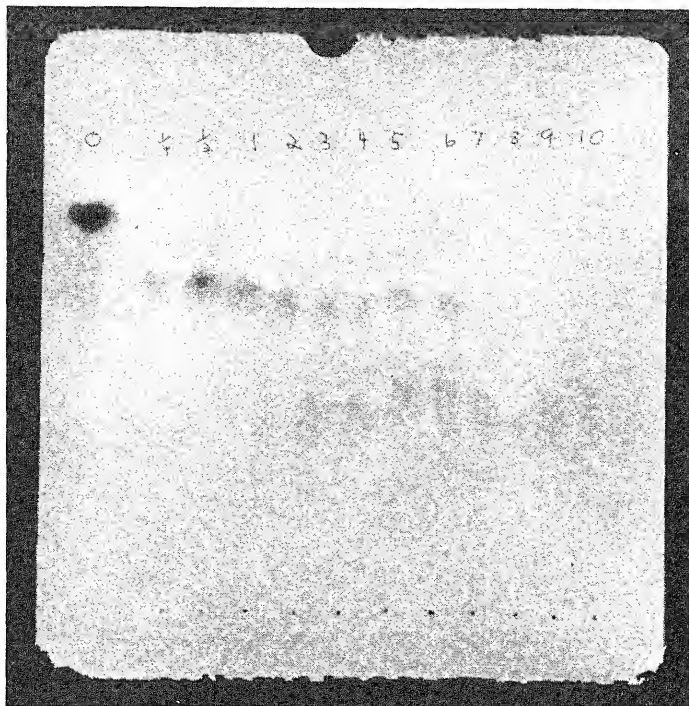


Figure 1.1. Kinetic study of the saponification and epoxide opening of the glucoside catalposide.⁴⁰ The starting material shown on the left at zero time is converted after one hour to a deshydroxybenzoyl derivative (middle spots) which is, in time, converted to a glycol (lower spots). The example clearly shows the presence of an intermediate and indicates the optimum reaction time for the preparation of either product. (Photograph by S. E. Wollman)

layer chromatography. This is, beyond doubt, easier, quicker and more precise than column chromatography and is discussed extensively in Chapter 9. In the author's laboratory, this has proved to be an ideal method for purifying samples prior to a final crystallization for microanalysis. Specific examples of this are noted in the tables in Chapter 11.

Applications to Column Chromatography

The simplest type of column chromatography involves the removal of small amounts of contaminants or tars by passing the sam-

ple, in a suitable solvent, over a short column of adsorbent. Such a process is adaptable to relatively large quantities and gives high yields. Thin-layer techniques can predict, in a short time, the ideal solvent and adsorbent. One needs only to find a system in which the product moves and the contaminant stays at the origin and to transpose this to a short column of the *identical* adsorbent. Normally, the adsorbents used for thin layers are of such small particle size that a pressure system or a vacuum system must be used when they are placed in columns. Such a scheme has been used by Kirchner and Miller¹⁹⁰ for the preparation of terpeneless essential oils. The reverse of this notion, that is, the isolation of small amounts of product from mixtures, has been used to isolate insecticide residues from plant material.

If the quantities involved and the degree of separation required are such that normal column chromatography cannot be avoided, thin-layer techniques can be used to predict the best solvent system and, above all, to analyze the effluent from the column. Several authors have stated or implied^{90,251,365} that the solvents used to give good separations on thin layers can be used directly on a column of the same adsorbent. Duncan⁹⁰ has published the following formula for predicting whether two substances will separate on a column.

$$r = \frac{a}{b + 0.1a}$$

a = R_f of faster moving substance on a thin layer

b = R_f of slower moving substance on a thin layer

Separation can only be expected under the conditions prescribed by Duncan (very high adsorbent-material ratios) when r is greater than unity.

It should be pointed out, however, that all of these situations depend upon the actual magnitude of the thin-layer R_f values. This is particularly true in respect to the Duncan expression which would certainly not be valid when the substances run close to the solvent front. The separations on a column will depend primarily on the retention times of the substances involved and these retention times, as stated for thin-layer chromatography by Zoellner and Wolfram,⁴⁵⁸ are inversely proportional to the R_f values of the substances on the same

adsorbent. Thus, a small difference when R_f values are small can be expected to produce an acceptable difference in retention times and a separation, whereas a large difference when R_f values are close to unity will not give a separation. The author has had some success using the following scheme. A pair of solvents is found which will produce an acceptable separation on a thin layer. The polar component of the solvent system is then reduced until the substances to be separated show R_f 's of about 0.2. This modified system is then used directly on a column of adsorbent identical in every way with that in the thin layer.

The monitoring of the effluent of a column is normally done by analyzing fractions obtained using a fraction cutting device. On a typical large layer (20 by 20 cm), as many as eighteen samples can be analyzed simultaneously. The fractions are then combined according to their components. Such applications are also noted in the tables in Chapter 11, and an example taken from the author's work on catalposide is given in Figure 1.2.

Routine Quantitative Assay

A number of procedures have been published for quantitative assay by thin-layer chromatography. These are discussed in Chapter 10 and documented in Chapter 11. Most of the work has been done in clinical chemistry where speed is an important factor. The major limitation of the technique is the error of about 3 to 5%.

Thin-Layer Chromatography in Conjunction with other Methods

The use of thin-layer chromatography in conjunction with gas chromatography has been especially useful in the lipid field.^{178,232,233,236} Normally, the class separations are made preparatively on thin layers and the classes are resolved, generally after chemical modification, by gas chromatography. In this manner, it is possible to take advantage of the ease of quantification and the high degree of resolution of the latter method. It should be noted, however, that the classes are also often resolved by "reversed phase" thin-layer chromatography.^{181,182,184,230,232 and others}

A method has been suggested¹⁴⁷ for the combined use of thin-layer chromatography and the mass spectrograph. The silica gel taken

extracts and biochemical preparations. The comparison of these with patterns of adulterated or diseased samples can provide interesting results. Such techniques have been used to study "chemical races" in plants,⁴⁵³ drug contaminants,¹⁶¹ mint oils³²³ and lipid components of multiple sclerosis patients.¹⁵⁷

Classification Systems

In the alkaloids,⁴³⁶ the estrogens²²¹ and mint oil constituents,³²³ thin-layer chromatographic properties have been correlated to some extent with structural types, and classification procedures have been developed. Such systems make it easier to identify known compounds and to obtain information about unknown ones.

CHAPTER 2

Adsorbents

INTRODUCTION

The adsorbents most commonly used in thin-layer chromatography are, in the following order, silicic acid or silica gel,* aluminum oxide or alumina, kieselguhr or diatomaceous earth and powdered cellulose. Those which have been used to a lesser extent are polyamide powders, ion-exchange powders (modified cellulose), "Florisil," calcium sulfate, polyethylene, "Magnesol," hydroxyl-apatite, "Sephadex," zinc carbonate and various mixtures of these. Kirchner, Miller and Keller¹⁹³ explored several adsorbents using starch and plaster of Paris as binding agents. These data are shown in Table 2.1. Although they pertain only to the separation of essential oils by adsorption processes, they show the large number of substances which can be formed into usable layers. Only calcium hydroxide layers are really poor and, as will be discussed later, the addition of silica gel produces usable layers of this substance. Essentially any substance which has desirable adsorptive and chemical characteristics can be successfully used in thin-layer chromatography.

The four most widely used adsorbents, *silica gel*, *alumina*, *kieselguhr* and *cellulose*, however, offer a wide range of properties and can be used in most cases. Waldi, Schnackerz and Munter⁴³⁶ made a comparative study of Silica Gel G, Aluminum Oxide G and Kieselguhr G with the following conclusions. Silica gel and alumina are, respectively, essentially acidic and basic in character while kieselguhr is

* The major difference between silicic acid and silica gel appears to be solely in the method of preparation. They are chemically identical, but the gel is prepared in such a manner as to enhance its adsorptive power.

TABLE 2.1. CHARACTERISTICS OF CHROMATOSTRIPS MADE WITH VARIOUS ADSORBENTS^{193*}

Adsorbent Coating	Physical Characteristics of Strip	Resolution of Essential Oils
Magnesium oxide	soft	none
Alumina	excellent	good
Alumina + silicic acid	excellent	good
Calcium hydroxide	soft, crumbly	none
Starch	good	none
Dicalcium phosphate	fair	some resolution
Bentonite	good	oils decomposed
Calcium carbonate	good	slight resolution
Magnesium carbonate	fair	slight resolution
"Filtrol"	good	separation, but oils decomposed
"Filtrol" X202	good	separation, but oils decomposed
"Filtrol," neutral E	good	separation, but oils decomposed
"Florisil"	good	fair separation
Talc	good	slight separation
Silicic acid	excellent	excellent

* Reproduced from Kirchner, Miller and Keller, *Anal. Chem.*, 23, 420 (1951) through the courtesy of the authors and The American Chemical Society.

neutral. Furthermore, silica gel has the highest capacity (the ability to separate the largest quantity of a mixture) followed by alumina and kieselguhr in that order.

Alumina chromatography is essentially adsorption chromatography; silica gel can function in both adsorption and partition chromatography depending upon the solvent system and kieselguhr is best characterized as a support for partition phenomena. Thus, alumina would be the adsorbent of choice for the separation of non-polar basic or neutral mixtures. Silica gel can be used for nonpolar acidic mixtures and, because of its higher capacity, would even be preferred for the nonpolar neutral materials. Silica gel, kieselguhr and cellulose are all suitable for the separation, through partition processes, of polar molecules. Of course, the properties of these adsorbents can be extensively altered by addition of various acids, bases and buffers so that many purposes can be served which are not covered by the above generalizations.

Although artifact formation on thin layers is rare, it is not unknown. Steroid 16- β -esters have been shown to undergo reactions on alumina layers³⁴⁷ and ethylene ketals have been hydrolyzed on silica gel layers.³⁶³ Alumina in columns is known to catalyze ester hydrolysis, isomerization of double bonds and other reactions.²³² The chromatography of sugars on silica gel layers with ammonia solutions has been shown⁴³⁹ to produce amination reactions.

The various adsorbents can be and are sometimes used in a pure form, but more frequently, they are used in combination with small amounts of added components. These additives can be blended with the solid adsorbent before it is made into layers, or they can be dissolved in the water portion of the slurry which, in most cases, is the vehicle for application to plates. In general the following purposes are served by the additives.

(1) They may serve as binders or substances to make the adsorbent more cohesive and to hold it to the plate. These make the layers much easier to handle and use and make it possible for them to be sprayed with visualizing reagents. The most common binder is plaster of Paris (calcined calcium sulfate, $\text{CaSO}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$) which is added to the dry adsorbent in amounts up to 20 per cent. The first binder used by Meinhard and Hall²⁴⁶ and by Kirchner and his group¹⁹³ was starch, which, although it gives the layers an added stability over plaster of Paris, limits the spray reagents which can be used to visualize the finished chromatograms. Other binders which have been used are polyvinyl alcohol,²⁸⁴ collodion³¹⁷ and an alcohol-soluble polyamide, "Zytel 61."^{*153} Two binders which have been added at later stages are decalin,²⁵⁴ which is added to the developing system and dichlorodimethylsilane,¹⁸⁵ which is applied to the developed plates so that they can be washed like paper chromatograms. The relative advantages and disadvantages of these binders will be discussed in more detail in the section of this chapter devoted to silica gel.

(2) The additives may serve to alter the properties, both physical and chemical, of the adsorbents. Thus, acids, bases and buffers are commonly incorporated into the layers. Reagents which will complex or chelate to a greater or lesser extent with the substances to be sep-

* Du Pont de Nemours, Geneva, Switzerland

arated, are often added. The latter concept is a relatively recent one and will, in the author's opinion, become important.

The impregnation of layers with hydrophobic liquids so that "reversed phase" chromatography can be carried out will be discussed in Chapter 3.

(3) On occasion, materials are added to adsorbents as "fillers." These reduce the activity of the adsorbent and make it somewhat more porous, resulting in shorter developing times. This concept was originally used by Meinhard and Hall²⁴⁶ who added "Celite" (see p. 24) to starch-bound alumina. More recently, Bennett and Heftmann^{26,27} have added Kieselguhr G to Silica Gel G for a similar purpose.

(4) Finally, reagents may be added which will aid in the visualization of the completed chromatogram. The incorporation of phosphors or fluorescent compounds in the adsorbent so that unsaturated compounds can be seen or easily rendered visible in ultraviolet light is a common practice. On occasion, sulfuric acid is added to the adsorbent. This alters the character of the adsorbent and allows the visualization of the completed chromatogram by heating (to char the organic compounds).

Procedures have been developed by Stahl³⁷³ and by Heřmánek, Schwarz and Čekan¹⁴⁴ for relating the activity of alumina layers to the Brockmann⁵⁰ activity scale.

A number of these adsorbents, both modified and pure, are commercially available. These sources will be listed in Table 2.4 at the end of this chapter, although, at the appropriate places, procedures will be given so that an investigator can prepare his own.

Silica Gel or Silicic Acid

Silica gel, used with a plaster of Paris binder is, by far, the most extensively used adsorbent in thin-layer chromatography. It forms a versatile layer which can serve as the solid stationary phase for adsorption chromatography, as a support for the polar liquid phase in normal partition chromatography and as a support for the nonpolar liquid phase in "reversed phase" chromatography. The distinction between the adsorption and normal partition chromatography in this case lies in the solvent system chosen for the chromatogram. If

the system consists of solvents such as hexane, benzene, ether, methanol or similar substances, adsorption phenomena will result. If a system is chosen which contains both a nonpolar liquid and a polar liquid such as water, the result will probably be partition chromatography. In "reversed phase" chromatography, the silica gel layer is impregnated with a nonpolar liquid such as a Silicone oil or a hydrocarbon before use. This will be discussed in Chapter 3.

Two binders, starch and plaster of Paris, are commonly used with silica gel. Both originated with Kirchner, Miller and Keller.¹⁹³ Starch gives a more mechanically stable layer^{8,120,193} which can actually be written on with a blunt lead pencil, but the inclusion of starch in the layer forbids the use of corrosive spray reagents for visualizing the finished chromatogram. (Cold concentrated sulfuric acid may be sprayed on the layers without difficulty, but the sprayed layers cannot be heated to char the spots.) The use of these corrosive spray reagents (sulfuric acid, chromic acid, nitric acid, etc.) as essentially universal reagents for organic substances, is one of the greatest advantages of thin-layer chromatography. On the other hand, the presence of calcium ion (from the plaster of Paris) in the layers has caused some difficulty in the chromatography of inorganic ions³⁵⁵ and nucleotides.³¹⁷

Silica Gel with No Binder. Silica gel has been used only rarely^{1,51,312,322} without some sort of binder. This situation will surely change, however, due to the recent introduction of Woelm "Silica Gel for Thin-layer Chromatography," which contains no binder and depends, for the stability of its layers, on a small particle size. The commercial preparation "Anisil" has no binder as such, but does contain 10 to 15 per cent of magnesium oxide which enhances its layer stability and makes it a basic adsorbent.

A method for the preparation of a silica gel suitable for thin-layer chromatography has been given by Adamec, Matis and Galváněk.¹ The starting material was water glass (mixed, hydrated sodium silicates) and the adsorbent was used without a binder.

Silica Gel with a Plaster of Paris Binder. The commercial preparation Silica Gel G (G for *gips* or gypsum) is of this type. It is manufactured by Merck in Germany according to the specifications of

Stahl and, without doubt, is the most widely used adsorbent in thin-layer chromatography. Mangold²³² has stated that a reasonable approximation of the commercial product can be prepared by thoroughly mixing Mallinckrodt Silicic Acid, 200 mesh, with 10 to 15 per cent of newly calcinated calcium sulfate (prepared by heating reagent grade $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ at 180° for 24 to 48 hours²⁹⁶) of the same grain size. A 200-mesh screen produces particles smaller than 0.12 mm.

Silica Gel G contains appreciable amounts of ferric ion. This does not interfere with the separations of organic compounds, but must be removed for successful work with inorganic ions.³⁵⁵⁻³⁶⁰ Seiler and Rothweiler³⁵⁸ give the following instructions (freely translated) for doing this.

"Five hundred grams of Silica Gel G is treated with 1000 ml of 6N hydrochloric acid (500 ml of concentrated hydrochloric acid and 500 ml of distilled water), stirred and allowed to stand. The supernatant liquid, colored yellow by the iron, is decanted and the silica gel is washed twice more with acid and then with three successive 1000-ml portions of distilled water. The adsorbent is then filtered and washed with distilled water until the filtrate is only slightly acidic. It is finally washed with 250 ml of ethanol and 250 ml of benzene and dried in an oven at 120° ."

The material from the above treatment *now contains insufficient binder*. Two grams of plaster of Paris³⁵⁹ or one gram of starch³⁵⁸ are added to 28 g of purified silica gel to obtain a usable adsorbent.

When materials are to be recovered from thin-layer adsorbents by elution, it is suggested^{244,394} that the adsorbent be prewashed. This can be done for Silica Gel G by extracting it three times with boiling methanol²⁴⁴ or by allowing it to stand overnight with the same solvent.³⁹⁴ Honegger found¹⁵⁸ that 50 g of Silica Gel G yielded the following amounts of impurity on elution: with chloroform, 6.8 mg; with benzene, 5.5 mg and with acetone, 10.5 mg. Or, the layers can be prewashed with the developing solvent.

Silica Gel With a Starch Binder. To the author's knowledge, no commercial preparation of this type is available. Kirchner, Miller and Keller¹⁹³ combine 19 g of Merck reagent grade silicic acid, 100

mesh, one gram of "Clinco"* starch,** 0.15 g of zinc silicate and 0.15 g of zinc cadmium sulfide. The latter two reagents are added as phosphors, but the same proportions would surely serve for the preparation of layers without them. The preparation of starch bound layers is inherently different from the preparation of plaster of Paris bound layers in that the slurry of adsorbent, starch and water is heated before it is applied to plates. This will be discussed in Chapter 3.

Silica Gel with Fluorescing Agents. Several reagents are commonly incorporated into thin layers so that the resulting spots on the developed chromatogram can be seen without using a spray reagent. This is particularly helpful in preparative work. The fluorescing agents, with one exception, are usually blended with the dry adsorbent before the layers are prepared. The exception is sodium fluorescein which is normally dissolved in the liquid used to make the slurry. When the developed plates are placed in short wave ultraviolet light, compounds containing conjugated double bonds show up as dark spots. The various fluorescent materials, the amounts used and their commercial sources are given in Table 2.2.

Acidic Silica Gel. Acids have been incorporated into thin layers for two reasons. The first is to modify the properties of the adsorbent and the second is to aid in visualization. Thus, Stahl³⁶⁹ used 0.5N aqueous oxalic acid rather than pure water to prepare layers. The adsorption activity of such plates is slightly increased.

In contrast to this, Reichelt and Pitra³²² deactivated silica gel by adding 25 per cent water or 43 per cent of dilute (50 per cent) acetic acid to the dry activated powder. In this case, the silica gel was then used without a binder for the separation of cardenolides.

Layers prepared from adsorbents containing small amounts of sulfuric acid can easily be visualized by heating the developed chromatograms on a hot plate or in an oven to char the organic substances. Peifer²⁹⁶ added 2.5 ml of concentrated sulfuric acid to 100 ml of chloroform-methanol, (70:30) v/v, which was subsequently slurried

* Clinton Foods, Inc., Clinton, Iowa.

** Dr. Kirchner has informed the author that still another starch binder is made from ordinary corn starch and "Superior AA Tapioca Flour" mixed in a ratio of 2 to 1. The tapioca flour is a product of Stein, Hall and Co., Inc. of New York. Furthermore, the amount of starch can conveniently be cut in half from the original proportions.

TABLE 2.2. FLUORESCENT AND PHOSPHORESCENT CHEMICALS USED IN THIN-LAYER CHROMATOGRAPHY

Name	Amount Used	Commercial Source	Ref.
1. Zinc silicate luminescent material P 1, Type 118-2-7	2% in adsorbent	General Electric Co., Cleveland, Ohio	46
2. Rhodamine 6G	0.003% in adsorbent	Matheson, Coleman and Bell, East Rutherford, N.J.	323
3. Luminescent substance ZS-Super	2% in adsorbent	Riedel-de Haen, 3016 Seelze bei Hannover, Germany	117
4. "Ultraphor"	0.005% in adsorbent	Badische Anilin and Soda Fabrik, Ludwigshafen (Rhein), Germany	293
5. Luminescent chemical No. 601	1% in adsorbent	E. I. du Pont de Nemours & Co., Photo Products Dept., Wilmington 98, Delaware	232
6. "Sodium fluorescein"	0.04% sol. in water phase		366
7. Zinc silicate-zinc cadmium sulfide	0.75% of each in adsorbent	Du Pont phosphors No. 601 and No. 1502 respectively	193

with 50 g of Silica Gel G. The developed chromatograms (in this case on microscope or lantern slides) were visualized by heating.

Basic Silica Gel. Stahl³⁶⁹ described the preparation of basic layers using 0.5N aqueous potassium hydroxide rather than water to make the slurry. The layers were found to be somewhat less active than normal. Such plates can be used to separate acids, which stay at the origin, from neutral compounds or bases which move with the solvent. Adsorbents consisting of 80 per cent calcium hydroxide and 20 per cent Silica Gel G⁴⁴⁹ will be more properly considered as a separate adsorbent system. The commercial material "Anisil," as noted above, is a basic adsorbent.

Buffered Silica Gel. In general, buffered silica gel layers have been used for the separation of polar or ionic compounds by partition chromatography. The addition of the buffer is normally made in the same manner as with acids and bases, that is, by using the desired buffer solution in place of water to make the slurry. Such slurries (at least those made with Silica Gel G) must be spread quickly since they thicken more rapidly than usual.²³² The layers should be allowed to stand several hours at room temperature before they are oven-dried.²³²

TABLE 2.3. THE USE OF BUFFERED LAYERS IN THIN-LAYER CHROMATOGRAPHY

Buffer Sol.	Concentration	Compounds Separated	Ref.
1. Sodium acetate	0.3 <i>m</i>	flavones	379, 380
	0.02 <i>m</i>	sugars	307, 377
2. Monopotassium phosphate and disodium phosphate	equal amts. of 0.2 <i>M</i>	amino acids	271
3. Citric acid-phosphate buffers	pH 7	sugars	307
4. Boric acid	0.1 <i>N</i>	sugars	307
5. Ammonium sulfate	10%	acidic lipids and amphoteric phospholipids	234

This helps to prevent flaking which is also more pronounced with buffered layers. Some of the uses of buffered layers will be summarized in Table 2.3.

The use of boric acid in the separations of sugars is interesting because of the formation of borate complexes between adjacent *cis*-hydroxyl groups. Such complexes are probably responsible for the effectiveness of the buffer.

The use of buffered layers for thin-layer ionophoresis by Honegger¹⁵⁶ is a rather special case since such solutions are normally used to carry current and to form complex ions which can be caused to migrate. Honegger used 0.1 *M* sodium citrate and acetic acid-formic acid solutions.

Silica Gel Containing Complexing Agents. The addition of complexing agents to thin layers to increase the power of resolution is an intriguing notion and will surely be further developed. At present, the only clear case of this is the use of silver nitrate on silica gel by Barrett, Dallas and Padley¹⁶ and by Morris²⁶⁰ to separate certain closely related lipids. The basis of this separation is the tendency of silver ion to complex with the pi electrons of double bonds to a greater or lesser degree. The former authors used 12.5 per cent silver nitrate to make a slurry with Silica Gel G. A similar result was produced by Morris²⁶⁰ who sprayed layers with a saturated silver nitrate solution or a saturated methanolic boric acid solution (see p. 48) or both.

As previously stated, the use of boric acid layers for sugar separation may well be an example of complex formation. The use of 2,5-hexanedione in the moving phase for the separation of inorganic ions³⁵⁹ is an example of the application of this notion in a reverse manner.

Alumina

Historically, alumina was the first adsorbent used for thin-layer work. Izmailov and Shraiber,¹⁶⁹ Meinhard and Hall²⁴⁶ and the Kirchner group¹⁹³ all used or explored this adsorbent. It is strange, then, that alumina has not been more extensively used than it has. One possible reason for this is that alumina is a more reactive adsorbent than silica gel and is known to catalyze ester hydrolysis, isomerization of double bonds, and other reactions.²³² Also, alumina is a basic substance⁴³⁶ in contrast to the acidic silica gel and has a somewhat lower capacity.⁴³⁶ Thus, one would tend to choose it for the separation of basic or neutral substances only when silica gel separations were poor.

Alumina with No Binder. Again, in contrast with silica gel, alumina has been used to a large extent with no binder. The reasons for this are obscure but probably can be traced to the ready availability of chromatographic grades of alumina and the large body of knowledge which has accumulated on alumina chromatography. Furthermore, layers prepared by simply leveling dry alumina powder on plates give quite good separations. Thus, no slurries or elaborate application apparatus are needed. These layers are mechanically unstable and must be sprayed with caution.⁷⁰

The alumina which is used in this manner is normal, commercially available alumina of any desired activity grade having a particle size of 0.1 mm or 200 to 300 mesh. Recently, Woelm has announced a special "Alumina for Thin-layer Chromatography" which is available in acid, basic and neutral grades and which is more cohesive than usual. This enhanced cohesiveness is produced by smaller particle size.

Huneck¹⁶⁵ has published directions for the preparation of a fibrous alumina which is quite cohesive and which forms stable layers with no binder.

Alumina with a Plaster of Paris Binder. This adsorbent is com-

mercially available as Aluminum Oxide G and as Fluka Aluminum Oxide D5 but, strangely enough, few reports have been made on its application. It has, however, been used to separate alkaloids.⁴³⁶ A reasonable approximation of the commercial product can be prepared²³² by blending Alcoa Activated Alumina,* 200 mesh, with 5 per cent of its weight of plaster of Paris (see p. 19 for preparation of this substance).

Alumina with a Starch Binder. This adsorbent was used by Meinhard and Hall²⁴⁶ in the first recorded use of a bound adsorbent. However, it does not appear to have been used since and would seem to offer no particular advantage.

Kieselguhr or Diatomaceous Earth

Kieselguhr with a Plaster of Paris Binder. Silica gel is acidic and applicable in both adsorption and partition chromatography while alumina is basic and useful in adsorption chromatography. Kieselguhr completes this picture in that it is neutral⁴³⁶ and is primarily a support for the stationary phase in partition chromatography. As such, it has a low capacity and is used for the separation of very polar molecules such as carbohydrates and amino acids. The commercial product, Kieselguhr G, has been used in a buffered form (with sodium acetate) for the separation of the simple sugars³⁷⁷ with reasonable success. Furthermore, it has been impregnated with paraffin^{6,183} and undecane^{181,185,186} for "reversed phase" chromatography.

Kieselguhr Containing Complexing Agents. Halmekoski made slurries of Kieselguhr G with 0.01 mole of sodium molybdate, sodium tungstate or borax (per 30 g of adsorbent and 65 ml of water) and used the resulting layers for the separation of phenolic carboxylic acids.¹³⁶

Cellulose

The use of powdered cellulose in thin layers was essentially an attempt to carry out "paper chromatography" in a short time. This has been successful. The cellulose is, of course, only a support for the stationary liquid phase and has been used in both normal (impregnated with formamide,^{399,400}) and reversed phase chromatography. The capacity of the adsorbent is low.

* Aluminum Company of America, 230 Park Avenue, New York 17, New York

Cellulose without a Binder. The fibrous nature of powdered cellulose permits it to be shaped into stable layers without a binder.^{321,450}

Cellulose with a Plaster of Paris Binder. The commercial product, Cellulose Powder MN 300 G, is of this type. Randerath³¹⁹ has systematically compared this material with paper for the separation of various nucleic acid derivatives. He found that the cellulose layers gave smaller, more compact spots and better separations in a shorter time (90 min) than paper (6 to 8 hr). Wollenweber⁴⁵¹ found similar results with food dyes.

Cellulose Ion-Exchange Powders. Two cellulose powders, modified to produce ion-exchange resins, have been used for the separation of nucleotides.^{316-318,320} These are DEAE-cellulose (diethylaminoethyl cellulose) and ECTEOLA-cellulose (epichlorohydrin linking triethanolamine with cellulose). They have been used with a collodion binder^{316,317} or without a binder.³¹⁸

A large number of cellulose powders, plain and modified, bound (G series) and unbound, are commercially available (Table 2.4). In fact, several are available which have not yet been mentioned in the literature.

Ion-Exchange Resins

Berger, Meyniel and Petit²⁹ have reported the use of "Dowex" 1 and "Dowex" 50 (both anion and cation exchangers, 200 to 400 mesh) for the separation of organic and inorganic compounds. They used a 1:1 mixture of the resin and commercial plaster of Paris-bound cellulose, Cellulose Powder MN 300 G.

Polyamide Powders

Polyamide powders are normally used without binders and have been developed by two groups. Davidek and his co-workers^{69,71,74} have spread commercially available polyamide powder on plates and used the resulting layers for the separation of such polar molecules as flavonoids and gallic acid esters. Wang and his co-workers^{219a,436a} allowed 20 per cent solutions of ϵ -polycaprolactam resin in 80 per cent formic acid to evaporate on glass plates and used the resulting thin layers for the separation of phenols.

A commercial product is now available from Woelm (see Table 2.4).

Hydroxyl-Apatite

Hofmann has reported the use of hydroxyl-apatite (a complex calcium phosphate hydroxide*) for the separation of 1- and 2-mono-glycerides¹⁵⁵ and of proteins.¹⁵³ In the former case, he used plaster of Paris as a binder and in the latter case, "Zytel 61" (see p. 16).

"Celite"

"Celite" No. 545 with a plaster of Paris binder has been used⁴¹⁴ for the separation, by partition processes, of steroids. The layers were impregnated with formamide. Such layers have extremely short developing times (3 to 7 minutes).

"Sephadex"

"Sephadex" G 25, a cross-linked dextran which differentiates according to the molecular weight of the molecules being separated, has been used for the chromatography of proteins.⁸² A modified adsorbent, "DEAE-Sephadex" has been used for the separation, by gradient elution, of proteins and nucleotides.⁸²

Polyacrylonitrile "Perlon" Mixtures

Mixtures of polyacrylonitrile and "Perlon" (a polycaprolactam powder) have been used, in a buffered state, to separate anthocyanins and such very polar molecules as sugars.³⁶

Zinc Carbonate

Zinc carbonate with 5 per cent starch as binder has been used for the separation of 2,4-dinitrophenylhydrazones.⁹

Polyethylene Powder

Polyethylene powder, "Hostalen S," has been used²³² for the separation of fatty acids and their methyl esters.

Calcium Sulfate as Adsorbent

Layers prepared from pure calcium sulfate have been used for the separation of steroids²⁴³ and lipids.¹⁸⁰ The major advantage of such

* Prepared according to Anacker and Stoy^{4b}

plates is that they can be mechanically washed like paper chromatograms. Thus, even more spray reagent systems can be used.

"Magnesol" (Adsorptive Magnesium Silicate)

This adsorbent with a starch binder was preferred by Bryant⁵¹ and by Pryor and Bryant³¹³ for the separation of the essential oils. A commercial product containing no binder but especially prepared for thin-layer work is available.

Calcium Hydroxide with Silica Gel G

Layers prepared from a mixture of calcium hydroxide and Silica Gel G (6:1) and (4:1) have been found to produce good separations in the carotenoid field.^{168,449}

"Florasil" (Mixed Silica Gel and Magnesia)

Peifer²⁹⁶ has described a procedure for the preparation of micro-chromatoplates from "Florasil"-plaster of Paris (see p. 46).

CONCLUSION

Obviously, any finely divided material having the desired adsorptive, structural or chemical properties can be used as an adsorbent for thin-layer chromatography. Probably, the area which will be developed most fruitfully is the use of mixed adsorbents. Several of these are given above, and others, such as Silica Gel G-Kieselguhr G mixtures, have been reported.^{26,27}

It is rather interesting that the one factor which delayed the development of this technique, that is, commercial availability of materials, has now progressed so far that several adsorbents are available (the various types of cellulose) for which no published uses are recorded.

The commercial sources of most of the adsorbents are given in Table 2.4 and the specific adsorbents which have been used for specific separations are detailed in Chapter 11.

TABLE 2.4. COMMERCIAL SOURCES OF THIN-LAYER CHROMATOGRAPHY ADSORBENTS

Commercial Name	Nature	Manufacturer	U. S. Source
"Anisil"			
Silica Gel Woelm	silica gel-magnesium oxide pure silica gel ¹	Analytical Engineering Co., Hamden, Conn. M. Woelm, Eschwege, Germany	Alupharm Chemicals, 616 Commercial Pl., P.O. Box 755, New Orleans, La. Instru- ments, 115 Cutter Mill Rd., Great Neck, N.Y. Brinkmann Instr. Brinkmann Instr.
Silica Gel G	silica gel with plaster of Paris	E. Merck, A. G., Darm- stadt, Germany	Brinkmann Instr.
Silica Gel H Silica Gel HF	pure silica gel pure silica gel with inorganic fluorescent material	E. Merck E. Merck	Brinkmann Instr.
Silica Gel GF	silica gel with plaster of Paris and in- organic fluorescent material	E. Merck	Brinkmann Instr.
Silica Gel D5	silica gel with 5% plaster of Paris	Fluka, A. G., Buchs, S. G., Switzerland	Gallard-Schlesinger, Chemical Mfg. Corp., 1001 Frank- lin Ave., Garden City, N.Y.
Aluminum Oxide DO Alumina Woelm Basic Alumina Woelm Neutral Alumina Woelm Acidic Aluminum Oxide G Aluminum Oxide D5 Aluminum Oxide D5F	pure aluminum oxide pure basic aluminum oxide ¹ pure neutral aluminum oxide ¹ pure acidic aluminum oxide ¹ aluminum oxide with plaster of Paris aluminum oxide with 5% plaster of Paris aluminum oxide with 5% plaster of Paris and a fluorescence indicator	Fluka Woelm Woelm Woelm E. Merck Fluka Fluka	Gallard-Schlesinger Alupharm Alupharm Alupharm Brinkmann Instr. Gallard-Schlesinger Gallard-Schlesinger
Kieselguhr G Cellulose Powder MN 300	diatomaceous earth with plaster of Paris pure cellulose powder	E. Merck Macherey, Nagel & Co., Düren, Rhld., Germany	Brinkmann Instr. Brinkmann Instr.
Cellulose Powder MN 300G	cellulose powder with plaster of Paris	Macherey, Nagel & Co.	Brinkmann Instr.
Cellulose Powder MN 300 F254	cellulose powder with an inorganic fluores- cent material	Macherey, Nagel & Co.	Brinkmann Instr.
Cellulose Powder MN 300 GF254	cellulose powder with plaster of Paris and an inorganic fluorescent material	Macherey, Nagel & Co.	Brinkmann Instr.
Cellulose Powder MN 300/AC	pure acetylated (10%) cellulose	Macherey, Nagel & Co.	Brinkmann Instr.

Cellulose Powder MN 300 G/AC	acetylated (10%) cellulose with plaster of Paris	Macherey, Nagel & Co.	Brinkmann Instr.
Cellulose Powder MN 300 CM	pure carboxymethyl cellulose	Macherey, Nagel & Co.	Brinkmann Instr.
Cellulose Powder MN 300 G/CM	carboxymethyl cellulose with plaster of Paris	Macherey, Nagel & Co.	Brinkmann Instr.
Cellulose Powder MN 300 P	pure cellulose phosphate	Macherey, Nagel & Co.	Brinkmann Instr.
Cellulose Powder MN 300 G/P	cellulose phosphate with plaster of Paris	Macherey, Nagel & Co.	Brinkmann Instr.
Cellulose Powder MN 300 DEAE	pure diethylaminoethyl cellulose	Macherey, Nagel & Co.	Brinkmann Instr.
Cellulose Powder MN 300 G/DEAE	diethylaminoethyl cellulose with plaster of Paris	Macherey, Nagel & Co.	Brinkmann Instr.
Cellulose Powder MN 300 ECTEOLA	pure anion exchange cellulose	Macherey, Nagel & Co.	Brinkmann Instr.
Cellulose Powder C-1000	cellulose powder with plaster of Paris	Excorna O. H. G., Mainz, Germany	Brinkmann Instr.
Magnesium Silicate Woelm	pure magnesium silicate ¹	Woelm	Alupharm
Polyamide MNP	pure polyamide powder	Macherey, Nagel & Co.	Brinkmann Instr.
Polyamide Woelm	pure polyamide powder ¹	Woelm	Alupharm
"Hostalen S"	polyethylene powder	Farbwerke Hoechst, A. G., Frankfurt a.M., Germany	
"Florisl"	activated magnesium silicate	Floridin Co., P.O. Box 989, Tallahassee, Fla.	
"Sephadex"	cross-linked dextrans	Deutsch Pharmacia, G.m.b.H., Frankfurt a.M., Germany	
Polyacrylonitrile		Farbenfabriken Bayer, Dormagen, Germany	
"Perlon"	polycaprolactam	Farbwerken Hoechst, Bo- bingen, Germany	
e-Polycaprolactam resin		Toyo Rayon Co., Nakano- shima, Kita-ku, Osaka, Japan	
"Dowex" 1 and 50	anion and cation exchangers	Dow Chemical Co., Mid- land, Mich.	
"Celite" No. 545	diatomaceous earth	Johns-Manville, Interna- tional Corp., New York	

¹ These adsorbents are of an especially fine particle size for thin-layer chromatography and produce mechanically stable layers without binders.

CHAPTER 3

Preparation of Thin Layers

INTRODUCTION

The thin layers of adsorbent used for thin-layer chromatography are supported and held in place by glass plates. The plates may, in theory, be any convenient size, but in fact the size is usually dictated by the apparatus used to prepare the layers. Layers are applied to the plates by spreading or spraying a slurry, by spreading a dry powder, or in the case of small plates, by dipping them in a slurry. The thin film of slurry is dried to produce the thin layer. The layer may then be treated in various ways before it is used.

THE VARIABLES IN LAYER PREPARATION

In general, the variables to be considered are the nature of the adsorbent (Chapter 2), the type of glass support plate, the thickness of the layer and the moisture content.

The Glass Plates

The glass plates used in thin-layer chromatography may be flat and smooth, flat and lightly ground^{180,254,328,334} or ridged as shown in the cross-section diagram in Figure 3.1. Flat, smooth plates have been almost universally used, but the ground surfaces are said to increase the adhesion of the layers. The ridged plates are used solely to simplify layer preparation. Thus, one simply pours the adsorbent-water slurry on a ridged plate, wipes off the excess with a spatula and dries the finished layer. The squared-ridge plates (Figure 3.1a) have been advocated by Gamp, Studer, Linde and Meyer,¹¹⁴ but the rounded ones (Figure 3.1b) are more readily available from normal glass shops,

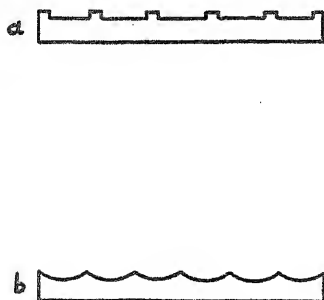


Figure 3.1. Cross-section views of two types of ridged glass plates.

where they are used to make semiopaque windows.* The ridged plates and the microchromatoplate system described later (p. 42) probably offer the most convenient systems for student use. To prevent accidents, all plates should have slightly beveled and smoothed edges.

The size and thickness of the glass plates are essentially dictated if a mechanical device such as those described below is chosen to prepare the layers. It has been found convenient to use, routinely, at least two sizes (exclusive of microplates). The smaller size should be wide enough to carry out concurrent, one-dimensional chromatography on three to four samples (about 2 in. or 5 cm minimum). Most workers allow the solvent front to rise 10 to 15 cm so that an over-all length of 20 cm or 8 in. is satisfactory. The larger size should be a square plate suitable for two-dimensional chromatography of a single sample or simultaneous, one-dimensional chromatography of eighteen or more samples. In a commercial apparatus system, the dimension of the square plate is equal to the length of the small plate (20 cm or 8 in.) and a multiple of its width. Such two-size systems are available with Stahl-type apparatus from Desaga-Brinkmann (p. 37) (5×20 cm and 20×20 cm) and Research Specialties Co. (p. 40) (2×8 in. and 8×8 in.). The Kirchner-type** apparatus manufactured by

* The author's attention was first called to the use of rounded plates by Dr. Louis Long of the Quartermaster Research and Development Laboratories in Natick, Mass.

** The author has elected to give this name to the type of apparatus in which the glass plate is passed under a stationary hopper because the Kirchner group was the first one to describe such an apparatus.^{27a}

Camag (p. 36) is normally designed for only one width, but two sizes of apparatus (10 cm and 20 cm width) are available. On the other hand, the Kirchner-type apparatus can be used with any type and thickness of glass, while the Stahl design works best with the plates supplied by the manufacturer. Although soft glass plates are routinely used, "Pyrex" plates are available from the manufacturers of the spreading apparatus.

If layers are prepared by spraying or by any of the simpler techniques discussed below, the plates can be of any size or type.

The original "chromatostrips" of Kirchner, Miller and Keller¹⁹³ were only wide enough for one sample (0.5 in.) although these workers also used wide plates. Such small strips are prepared in an apparatus designed by Miller and Kirchner²⁵³ and are still used³⁸³⁻³⁸⁹ by Stanley and his co-workers. However, they preclude a concentration study in which the sample is applied to a single plate in several concentrations, as well as a more precise comparison of two or three samples. The narrow plates are, however, more suitable for descending chromatography (p. 68).

The use of microscope slides for the preparation of "microchromatoplates" was described by Meinhard and Hall²⁴⁶ in the first paper advocating the use of bound layers. Lately, the use of such plates has again become fashionable^{152,437} and Peifer,²⁹⁶ in an exceptionally fine paper, has described a complete system based upon microscope and lantern slides (see p. 42).

A somewhat different type of layer support was worked out by Lie and Nyc,²¹⁹ who coated the inside of test tubes with thin layers of adsorbent. Rosi and Hamilton³³⁰ have prepared layers on both sides of glass plates so that they may be utilized more efficiently.

The Thickness of the Layer

Within a reasonably wide range (0.15 to 2.0 mm) the thickness of the layers is unimportant with respect to R_f 's and the degree of separation. The thickness itself should be prescribed by the type of information or results sought. For diagnostic or qualitative work, very thin layers are best because the spray reagent is much more sensitive when it does not have to "search" for tiny amounts of substance in a large amount of adsorbent. On the contrary, when the work is pre-

parative in nature, it is more efficient to use as thick a layer as possible so that the maximum amount of material can be separated in a single chromatogram.

The minimum thickness of about 0.15 mm was determined by Stahl, Schröter, Kraft and Renz³⁸¹ through a study of *Rf* vs. thickness. Only above this value were the *Rf*'s relatively constant (on silica gel layers). The maximum thickness is determined by the difficulty of drying thick layers without cracking and the degree to which there is unequal migration on the top and bottom of a layer. Honegger¹⁵⁸ investigated layers up to 5 mm thick and concluded that the best results were obtained between 1 and 3 mm. Ritter and Meyer,³²⁹ in a similar study, obtained better results with 1-mm layers than with 2-mm layers. The vast majority of diagnostic work has been done with layers 0.25 mm thick because the original Stahl-Desaga apparatus produced layers of this thickness. The Kirchner group¹⁹³ used 0.02 in. (0.51 mm).

Thin layers of unbound adsorbents which are prepared by spreading the dry powders on a plate are generally somewhat thicker, ranging from 0.5 mm²⁶⁹ to 1 mm.⁵²

The Moisture Content

The activity of adsorbents to be used for adsorption chromatography is greatly reduced by small amounts of water. Since most thin layers are prepared from aqueous slurries, this variable is controlled largely by the length of time and the temperature used to dry and activate the layers. This will be discussed in detail in a later section entitled "Drying of Layers" (see p. 46).

LAYER PREPARATION

Cleaning the Glass Plates

It is essential that the glass plates be as clean as possible and that they be particularly free of grease. Normally, washing the plates with detergent and wiping them with cotton soaked with hexane just before application of adsorbent will suffice. Sometimes, however, it is necessary to clean the plates with some type of cleaning solution. Peifer²⁹⁰ has recommended that his microchromatoplates be washed

successively with detergent solution, water and finally with 50 per cent methanol-water slightly acidified with hydrochloric acid.

The Preparation of Slurries

As previously stated, the majority of thin-layers are prepared by spreading a uniform film of a slurry of the adsorbent on glass plates and allowing it to dry. Normally, the liquid portion of such a slurry is water, which may contain acids, bases, buffers or complexing agents, and the solid portion is adsorbent and binder. The major problem associated with these slurries is the obtaining of a viscosity suitable for the manner of spreading. If the slurry is too thin, it runs through the spreader too rapidly and produces excessively thin layers. When it is too thick, it does not run through the spreader rapidly enough and may clump.

When plaster of Paris is used as a binder, this viscosity is controlled by two factors: the relative amounts of adsorbent and water, and the time elapsing between the addition of water to adsorbent and the actual spreading operation. The hydration of plaster of Paris ($\text{CaSO}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$) to gypsum ($\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$) is so rapid that the slurry becomes unmanageable in two to three minutes. When buffers or other modifying reagents are added to the slurry, this time is appreciably shortened.²³² The actual mixing of adsorbent and water can be carried out by shaking them in a stoppered flask, by triturating them in a mortar or by stirring them in some other vessel. Table 3.1 shows the amounts of water suggested for various commercially available adsorbents. When possible, the type of spreading device is also given. Commercial sources of the adsorbents have been cited in the preceding chapter.

For the preparation of very thick layers, Honegger¹⁵⁸ suggests that these ratios be slightly modified by decreasing the amount of water. Furthermore, for layers from 3 to 5 mm thick, he added an additional 2 per cent of plaster of Paris and dried the layers under an infrared lamp. His data are given in Table 3.2.

The procedure for preparing a slurry containing starch as a binder is essentially different in that it involves a heating step. Kirchner, Miller and Keller¹⁹³ give the following procedure for such a preparation containing phosphors.

"The specified amounts of material [19 g of Merck reagent grade

TABLE 3.1. RECOMMENDED AMOUNTS OF WATER TO BE ADDED TO VARIOUS COMMERCIAL ADSORBENTS

Adsorbent	Adsorbent-water, w/v, g/ml	Apparatus	Ref.
Silica Gel G	30/60	Desaga	Manuf.
Silica Gel Woelm ¹	30/45	—	Manuf.
Silica Gel D5	20/50	Camag	Manuf.
Silica Gel, "Anisil"	15/30	Desaga	Manuf.
Aluminum Oxide G	25/50	Desaga	232
Aluminum Oxide D5	20/65	Desaga	232
	20/50	Camag	Manuf.
Aluminum Oxide D5F	20/50	Camag	Manuf.
Alumina Woelm ^{1, 2}	35/40	—	Manuf.
Kieselguhr G	25/50	Desaga	377
Cellulose MN 300 G	15/100	Desaga	319
Magnesium Silicate Woelm ¹	15/45	—	Manuf.
Polyamide Woelm ¹	5 g in 45 ml of chloro- form-methanol (2:3)	—	Manuf.

¹ These adsorbents contain no binder, but are of an especially small particle size.² Same proportions for acid, basic or neutral.TABLE 3.2. HONEGGER'S DATA FOR PREPARATION OF VERY THICK LAYERS¹ *

Adsorbent	Thickness, mm	Adsorbent—H ₂ O, g/ml	Air-dry, min	Infrared dry- ing, min
Silica Gel G	1-1.5	1:1.7	5	
	2-3	1:1.6	30	30
	3-4 ²	1:1.6	30	30
	4-5 ²	1:1.57	30	30
Aluminum Oxide G	1-1.5	1:1	5	30
	4-5	1:0.9	10	30

¹ The data were obtained using the Camag apparatus.² Added 2% more plaster of Paris.* Reproduced from Honegger, *Helv. Chim. Acta*, 45, 1409 (1962) through the courtesy of the author and Verlag Helvetica Chimica Acta, Basel, Switzerland.

silicic acid passing 100 mesh sieve, 1 g of "Clinco 15" starch, 0.15 g of zinc silicate and 0.15 g of zinc cadmium sulfide] thoroughly blended while dry, were mixed with 36 ml of distilled water in a 250 ml beaker. The slurry was then heated on a water bath held at 85° with constant stirring until it thickened (1.75 minutes) and was then held at this temperature for 30 seconds longer with stirring. The beaker was removed from the bath and 2 to 7 ml of water were added immediately to form a thin paste." As stated on p. 20, a mixture of corn

starch and "Superior AA Tapioca Flour" is better than "Clinco" starch. If the amount of starch is decreased by half, it is not necessary to watch the temperature during the heating. The mixture must only be warmed thoroughly to thicken it.

Although water is the most common liquid for the preparation of slurries, several workers have preferred nonaqueous systems. In addition to the test-tube system of Lie and Nyc²¹⁹ and the microchromatoplate system of Peifer,²⁹⁶ which are quoted elsewhere in this chapter, Hörhammer, Wagner and Bittner¹⁶⁰ have used ethyl acetate, Duncan⁹⁰ has used water-methanol, 1:1, v/v, and Müller and Honerlagen²⁷⁰ have used acetone. None of these authors used the conventional spreading apparatus.

Mechanical Devices for the Preparation of Layers

At the present time, these devices fall into three major categories, the Kirchner type (see footnote on p. 31), the Stahl type and the spray type. While these "gadgets" are considerably more expensive and complicated than those used in some of the methods which will be described later, the author believes they have a distinct advantage. They produce uniform layers of any desired thickness in a very efficient manner.

The Kirchner Type. This type of spreading device is characterized by a stationary reservoir or hopper which deposits a thin film of adsorbent slurry on glass plates which are slowly passed under it. The layer-thickness is determined by the distance between the plate and the back edge of the hopper, and the width of the plates is predetermined by the size of the apparatus. The first such apparatus was designed by Miller and Kirchner in 1954²⁵³ to produce "chromatostrips" 0.5 in. wide. The apparatus has since been modified by Applewhite, Diamond and Goldblatt⁸ and shop drawings of the modification are available from the authors (Western Regional Research Laboratory, U. S. Department of Agriculture, Albany, Calif.).

A commercial model of this type of apparatus is shown in Figure 3.2 and is manufactured by Camag* in Switzerland. It is based upon the design of K. Mutter and J. F. Hofstetter at Hofmann-LaRoche A.

* Camag A. G., Homburger Str. 24, Muttentz, B. L., Switzerland. U. S. representatives: A. H. Thomas Company, Vine Street at 3rd, Philadelphia, Pa. and Microchemical Specialties Co., 1825 East Shore Highway, Berkeley 10, Cal.

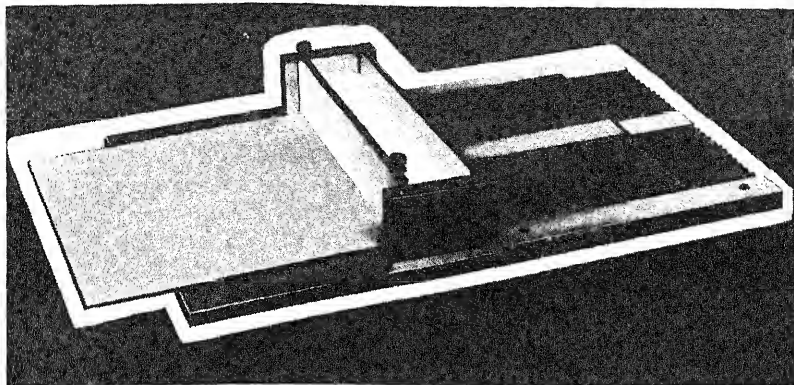


Figure 3.2. The commercial model of the Kirchner-type apparatus as designed by Mutter and Hofstetter. (Reproduced through the courtesy of Camag A. G.)

G. in Basel. Two models of this apparatus are available; one of these (Model A) coats plates 20 by 10 cm for one-dimensional chromatography and the second (Model B) coats plates either 10 cm or 20 cm in width. The Camag apparatus has been slightly modified by Wollish, Schmall and Hawrylshyn.⁴⁵² A similar apparatus is manufactured by the Kensington Scientific Corp.* in California.

It would appear to the author (who has had no experience with this apparatus) that the Kirchner type would be somewhat more tedious to use and less flexible with respect to plate width, than the Stahl type. On the other hand, it is less sensitive to plate thickness and probably yields more consistently uniform layers on small plates.

The Stahl Type. This type of spreading device is characterized by a moving slurry reservoir and stationary plates. The plates are arranged in a continuous surface on a support (the commercial models use plastic slabs) having a raised edge on two sides to hold the plates in position during the coating operation. The reservoir, machined in such a way that it produces the desired layer thickness, is then passed over the plates.

The first commercial model of this design was worked out by Stahl and is manufactured by Desaga** in Germany. The apparatus is pic-

* Kensington Scientific Corp., 1717 Fifth St., Berkeley, Cal.

** C. Desaga, G.m.b.H., Hauptstrasse 60, Heidelberg, Germany. U. S. representative: C. A. Brinkmann and Co., Inc., 115 Cutter Mill Rd., Great Neck, N.Y.



Figure 3.3. The Stahl-Desaga apparatus showing, in detail, the variable thickness applicator, a, and the standard applicator, b. (Reproduced through the courtesy of Brinkmann Instruments, Inc.)

tured in Figure 3.3. The spreader is available in two models; the Standard Model which gives layers 0.25 mm thick and the Model S 11 which gives layers of any thickness up to 2 mm. The standard plate sizes are 5×20 cm and 20×20 cm. Twenty of the small plates or five of the large ones can be coated in a single operation. While the technique works beautifully with the larger plates, the author has had difficulty obtaining uniform layers on small plates. This difficulty has also been encountered by Bennett and Heftmann²⁵ who solved it by placing a series of the large size plates (20×20 cm) under the small ones.* A glass slab or a smooth desk top can also alleviate

* Shandon Scientific Co., Ltd., 65 Pound Lane, London, N.W. 10 (U.S. representative, Consolidated Laboratories, Inc., P.O. Box 234, Chicago Heights, Ill.) manufactures an apparatus to avoid this difficulty. It contains a special leveling device and layer thicknesses are guaranteed to ± 0.010 mm.

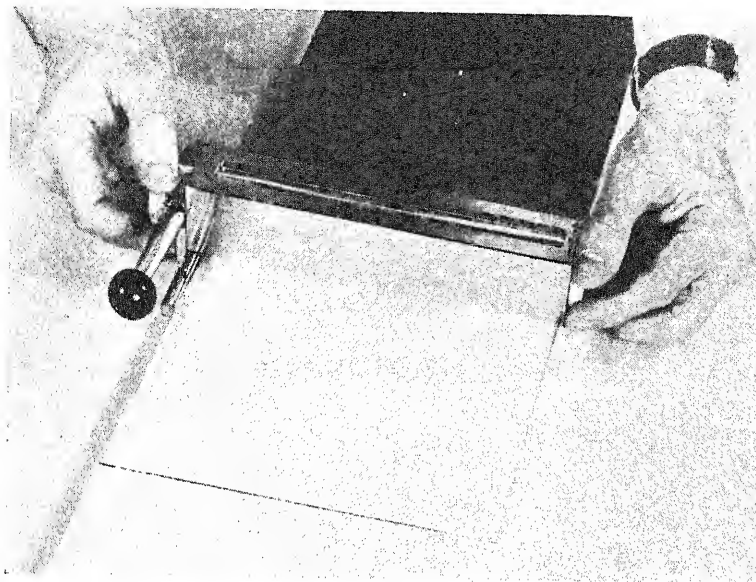


Figure 3.3a.

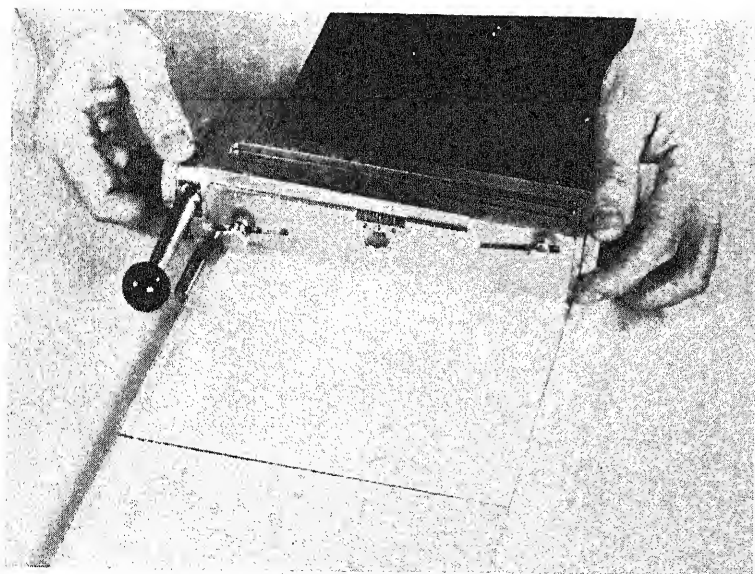


Figure 3.3b.

the problem. Sometimes wavy layers are produced by the "bump" caused by going from one plate to the next. As long as the developing solvent is passed in a direction parallel with the spreading operation, this is no problem.

A second commercial model similar to the Desaga apparatus is manufactured in the United States by Research Specialties Co.* This apparatus can also be purchased in a standard model or a variable thickness model. It is shown in Figure 3.4.

Home-made apparatus of the Stahl type has been devised by Barbier, Jäger, Tobias and Wyss,¹³ Machata,^{227**} Gamp, Studer, Linde and Meyer¹¹⁴ and Schulze and Wenzel.⁸⁴⁹

Sprayed Layers. The notion of spraying a thin slurry of adsorbent on glass plates has been suggested, notably by Reitsema,³²³ who used a small paint sprayer. Bekersky²⁴ has worked out a procedure using a normal laboratory spray bottle. Dilute slurries of Silica Gel G (15 g to 35 ml of water) or Aluminum Oxide G (20 g to 50 ml of water) are sprayed from a distance of 7 to 9 in. onto plates in a horizontal position. The method produces fairly uniform smooth layers, but the final thickness is unknown and only approximately controllable. The method has not yet been commercialized.

"Poor Man's Layers"

The development of a cheap method for the preparation of uniform, thin layers has challenged the ingenuity of chemists everywhere. This has resulted in the evolution of several acceptable methods. While the author is a firm believer in the more complex spreaders, these methods certainly offer possibilities for student use and for persons who wish to explore the system before investing in an apparatus.

The simplest way of preparing a layer of slurry is surely that of Lees and DeMuria²¹⁴ as shown in Figure 3.5. One to three layers of tape, either surgical or masking, but not "Scotch," are put on two opposite edges of a single glass plate or on the outside edges of the outside plates of a set arranged on a flat surface. The slurry is poured

* Research Specialty Co., 200 South Gerrard Blvd., Richmond, Calif.

** A plastic apparatus of this type is manufactured by Applied Science Laboratories, College Park, Penn.

on an open end and smoothed by drawing a glass rod over the plate or plates in a direction parallel with the taped edges. The tape supports the rod and determines the layer thickness. After a few minutes, the tape is removed and the layers are dried. The method is not very satisfactory for preparing thick (ca. 1–2 mm) layers. A similar method was devised by Duncan⁹⁰ who used removable metal strips on the edge of the plates to support the leveling rod. In this case, the thickness of the metal determines the thickness of the layer. Duncan used methanol-water, 1:1, v/v, to prepare his slurries. The detail of such a strip is shown in Figure 3.6. Boll⁴¹ arranged a set of plates in an apparatus with raised edges. The slurry is poured on and leveled with a straight edge. Still another method is to machine out an indentation in a large spatula to a depth corresponding to the desired thickness. Such a spatula (Figure 3.7) is then used to level a slurry on a glass plate which is large enough to support the two ends of the blade.* As in many other aspects of thin-layer chromatography, the first mention of this type of thing was made by Kirchner, Miller and Keller¹⁹³ who used an arrangement of glass plates, some placed 0.02 in. above the others and serving as support for a leveling rod. Thus, this concept is as old as the method.

Still other workers prefer to pour a slurry of adsorbent in ethyl acetate on a plate and to level it by tipping it.¹⁶⁰ Such layers are said to require no activation and to function satisfactorily in a chamber which is not saturated. The same procedure has been used with aqueous starch-silica gel slurries.⁴³ A similar scheme was used by Wang to prepare polyamide layers.^{436a} He prepared 20 per cent solutions of polyamide in 80 per cent formic acid, poured these solutions on glass plates, leveled them by tipping and allowed them to dry.

A special problem arises with the preparation of thin layers on the inside walls of test tubes.²¹⁹ Such layers were prepared and used by the following method. The test tube to be coated (13.5 × 150 mm) is filled with a suspension of silicic acid (no binder) and chloroform (1:2.5, w/v) and allowed to stand for about 5 seconds. The contents are then slowly poured out of the tube in one continuous slow tilting motion. After most of the slurry is discharged, the vessel is held in an

* This method was mentioned to the author by Mr. Robert Conkin of the Monsanto Chemical Co., St. Louis, Mo.

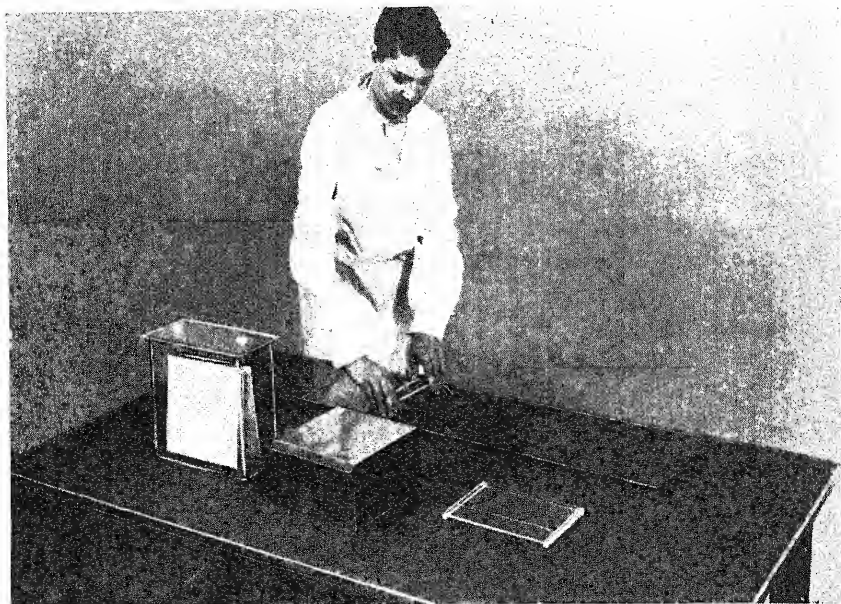


Figure 3.4. The Research Specialties apparatus showing, in detail, the variable thickness applicator, a, and the standard applicator, b. (Reproduced through the courtesy of Research Specialties Co.)

inverted position for 2 to 3 seconds and then slowly tilted, without rotation, to a horizontal position for drying. The top portion (in the horizontal position) is marked because it is the linear surface where the sample will be placed. An additional coating is placed on the open end of the test tube by immersing the tip end about 5 mm, at an angle of 45° , in the adsorbent slurry. In this position it is rotated once. The sample is placed 1 cm from the open end on the side previously marked. The open end is dipped in a suitable developing solution.

Microchromatoplates

Thin layers have been prepared on microscope slides and lantern slides by spreading a paste with a spatula²⁴⁶ with an especially designed gadget (for microscope slides only)⁴³⁷ and using the Stahl apparatus with an alternate arrangement to hold the slides during the slurry application.¹⁵² The author has prepared satisfactory layers using the complete Stahl assembly.

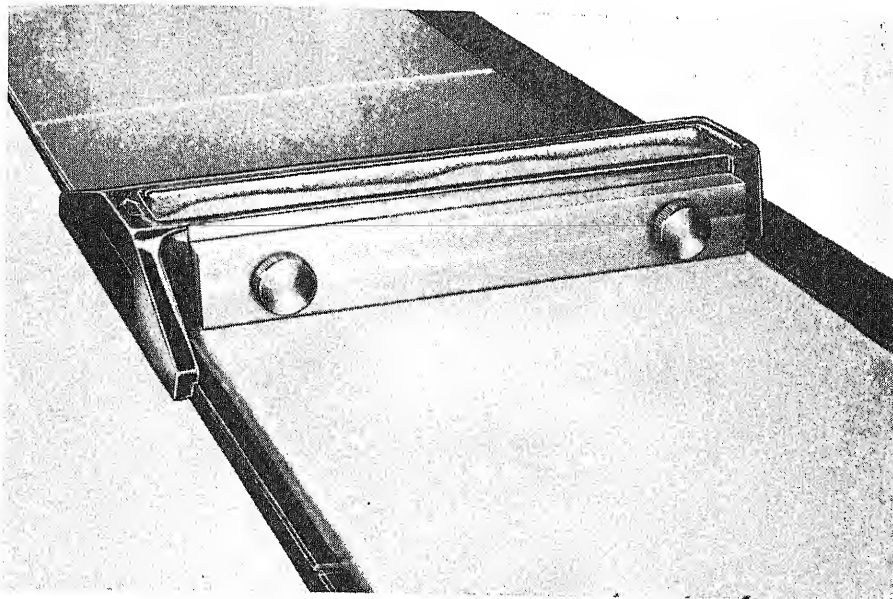


Figure 3.4a.

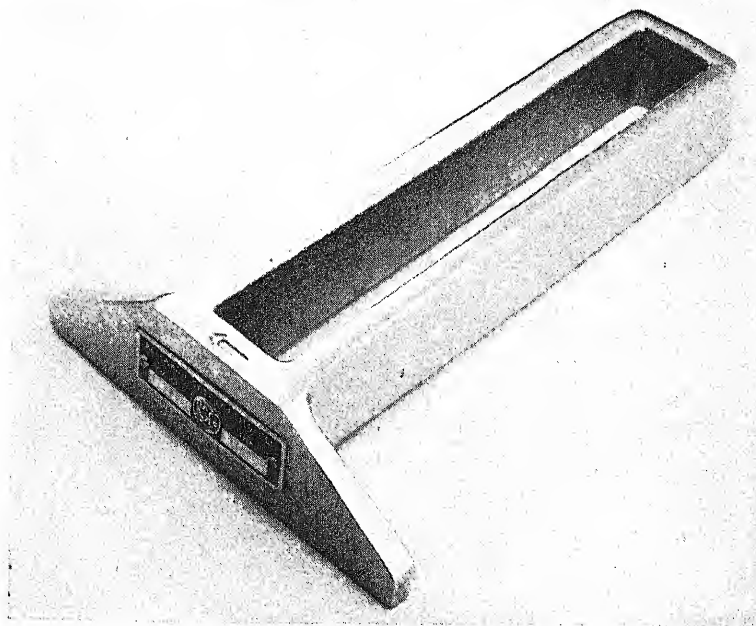


Figure 3.4b.

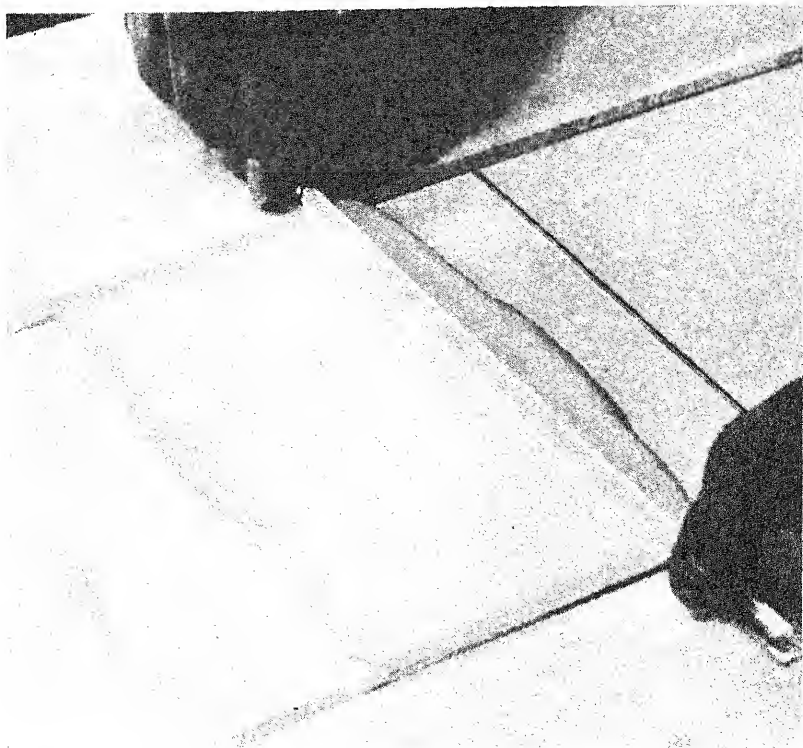


Figure 3.5. Preparation of thin layers using tape and a glass rod. (Reproduced from Lees and DeMuria, *J. Chromatog.*, 8, 108 (1962) through the courtesy of the authors and the Elsevier Publishing Co.)

Peifer²⁹⁶ has developed a complete scheme for chromatography on such small plates. He uses microscope slides for one-dimensional work and lantern slides for two-dimensional work. The slides, cleaned with detergent, water and water-methanol (50:50) acidified with hydrochloric acid, are *dipped*, two at a time, back to back, into a slurry of adsorbent in a nonaqueous medium. The slurry should be well stirred (about 1 min.) before dipping. They are then lifted slowly out of the slurry and allowed to drain on the edge of the container. They dry rapidly, are separated and the edges are wiped clean with tissue. In order to obtain a more durable layer, the plates are exposed briefly to steam to hydrate the plaster of Paris binder. The steam treatment is not actually necessary with silica gel layers. The layers are then

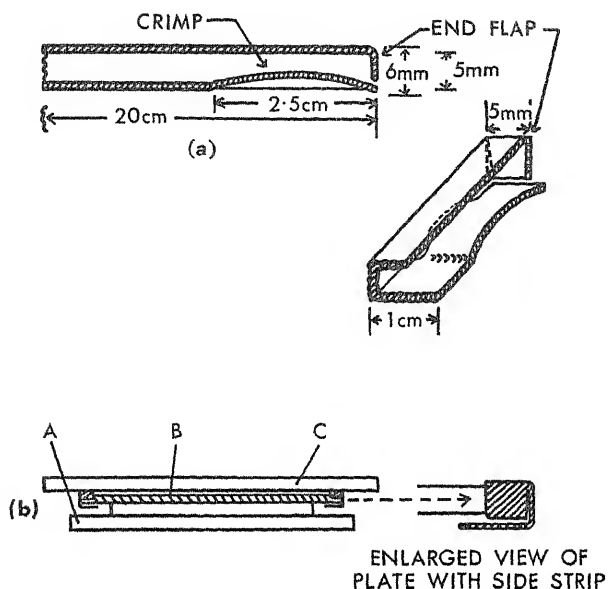


Figure 3.6. Detail of Duncan side strips, a, and the side strips in use, b. The side strips are fastened to the opposite edges of a glass plate (B) and held in place by the 2.5 cm long indentation. Slurry is poured on the plate and leveled with a glass rod (C) which rests on the side strips. The apparatus is shown on a mounting board (A). The thickness of the metal in the strips determines the layer thickness. (Reproduced from Duncan, *J. Chromatog.*, 8, 37 (1962), through the courtesy of the author and the Elsevier Publishing Co.)

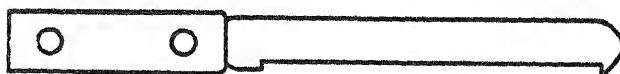


Figure 3.7. A machined spatula for the preparation of thin layers.

activated by suspending them about 6 mm over a hot plate for 1 to 3 min. As mentioned previously, Peifer sometimes added sulfuric acid to his adsorbents to facilitate visualization. The recipes for Peifer slurries are summarized in Table 3.3. The chromatograms containing sulfuric acid are visualized by warming them over a hot plate and the others are sprayed in the normal manner.

TABLE 3.3. RECIPES FOR PEIFER SLURRIES²⁶⁹

Adsorbent	Slurry Medium	Proportions, g in ml
1. Silica Gel G	chloroform or chloroform-methanol (2:1, v/v)	35 g in 100 ml
2. Silica Gel G and sulfuric acid	chloroform-methanol-sulfuric acid (70:30:2.5, v/v/v)	50 g in 102.5 ml
3. Cellulose powder ¹	chloroform-methanol (50:50, v/v)	50 g in 100 ml ¹
4. Alumina ²	chloroform-methanol (70:30, v/v)	60 g in 100 ml ²
5. "Florisl" ³ (60-100 mesh)	chloroform-methanol-acetic acid (70:30:1, v/v/v)	55 g in 101 ml ³

¹ 35 g of cellulose powder and 15 g of plaster of Paris are triturated with a minimum volume of methanol. The resulting viscous mixture is then diluted to give the above ratio.

² 45 g of activated alumina and 15 g of plaster of Paris are triturated with a minimum volume of chloroform-methanol and then diluted to the above ratio.

³ 45 g of "Florisl" and 10 g of plaster of Paris are triturated with 1 ml of glacial acetic acid and a minimum volume of chloroform. The result is diluted to the above ratio.

Layers Prepared From Dry Powders

The preparation of layers from dry powders of unbound adsorbents is a slightly different problem. The first method was that of Mottier and Potterat^{267,269} who used the apparatus shown in Figure 3.8. An even simpler affair is illustrated in Figure 3.9. A glass rod equipped with rubber or tape rings to hold it slightly above the plate and a larger ring to keep it centered is rolled over the plate containing an excess of alumina. This latter system is, of course, based upon the Mottier work, but has been more widely used.^{52,71,74,144,146}

Drying of Layers

After the slurry is spread on a plate, it should be allowed to stand for at least 30 min. so that the binder can "set." Indeed it is even recommended by several workers^{46,47,249,319} that more durable layers will result if they are allowed to stand in the air overnight. This would prevent cracking due to "casehardening." For adsorption chromatography, the layers are activated by heating them at 100 to 110° for various periods of time, generally one to several hours. There is little change in the activity when the time is lengthened beyond one hour or the temperature is raised above 110°. For layers containing plaster of Paris, the maximum drying temperature is 128° which is the point

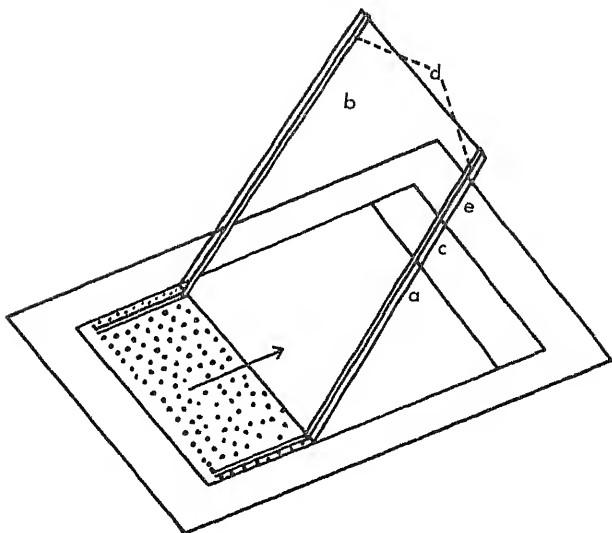


Figure 3.8. Preparation of thin layers of dry, adsorbent powders by the method of Mottier and Potterat. The dry alumina is placed on the glass plate a in front of the moving and leveling guide b. Plate c is a stopping board and d is two rubber bands which determine the layer thickness. Guide b is then moved over the plate. (Reproduced from Mottier, *Mitt. Gebiete Lebensm. Hyg.*, 49, 454 (1958) through the courtesy of the author and the Eidgenoessische Drucksachen-und Materialzentrale, Bern, Switzerland).

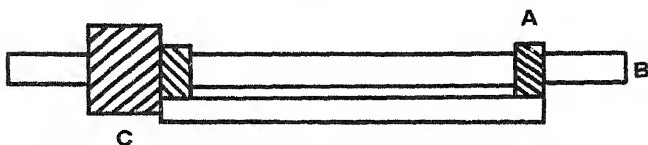


Figure 3.9. End view of a glass-rod leveling device for preparing layers of dry powders. The small spacers (A) are rubber tubing rings placed on a glass rod (B). The large ring (C) centers the gadget. Dry powder is placed on a plate and leveled by rolling the device back and forth.

at which gypsum again dehydrates. For partition chromatography, the layers are normally not activated since the residual water acts as the stationary liquid phase.

Better results with less cracking will be obtained if the layers are

dried in a vertical position as suggested by Desaga. All of the manufacturers of thin-layer equipment sell a metal drying rack which will hold several plates and which can be placed in the drying oven in either a horizontal or vertical position.

The activated plates should be stored in a desiccator or dry box of some type until they are used. The manufacturers of thin-layer equipment provide such containers. However, in the author's laboratory, the plates are carefully stacked in a normal desiccator over calcium chloride or, if large, are kept in a large box containing a beaker of sulfuric acid. After removal from the desiccator, the layers remain active for a time which is inversely proportional to the relative humidity. Normally, plates are active for ten minutes or more, but it is not wise to leave a plate out for several hours during the monitoring of a reaction or the elution of a column.

Other Post-Preparative Operations

The adsorbent sticking to the edges of the plates should be carefully wiped off before using the layers. The layers may then be prewashed^{388,389} with an appropriate solvent. A 0.5 per cent solution of "Versene" (ethylenediaminetetraacetic acid trisodium salt) was used by Schweiger³⁹⁰ to prewash unbound cellulose layers. Figure 3.10 shows a suggested³⁸⁹ apparatus for carrying out this operation.

The finished plates may, at this stage, be sprayed with a complexing agent such as silver nitrate or boric acid.²⁶⁰ This serves the same purpose as the incorporation of these substances in the slurry as discussed in Chapter 2, but spraying is more convenient in that the plates can be prepared one at a time as needed or in that only a portion of a plate need be treated. Morris²⁶⁰ sprayed with a saturated aqueous solution of silver nitrate or a saturated methanolic solution of boric acid or both. After spraying, the plates were dried at 110° for one hour.

On occasion, it is desirable to use partition chromatography in which the polar, stationary liquid is something other than water. To this end, Teichert, Mutschler and Rochelmeyer^{399,400} have impregnated cellulose layers by dipping them into a 20 per cent (v/v) solution of formamide in acetone. Vaedtke and Gajewska sprayed layers of plaster of Paris bound "Celite" with formamide, propylene glycol

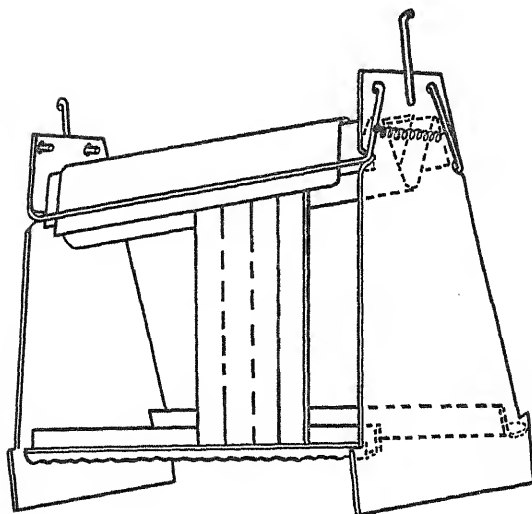


Figure 3.10. An apparatus for prewashing layers. The layers are placed in the apparatus and the washing liquid flows out of a reservoir by means of filter paper and over the layers in a descending manner. (Reproduced from Stanley, Vannier and Gentili, *J. Assoc. Off. Agr. Chemists*, 40, 282 (1957) through the courtesy of the authors and the Association of Official Agricultural Chemists.)

or paraffin oil.⁴¹⁴ The plates were weighed before and after the spraying to determine the extent of impregnation. Knappe and Peteri²⁰⁰ added polyethylene glycol (15 g to 30 g of Kieselguhr G and 45 ml of water) to the adsorbent slurry before the layers were prepared. In the latter case 0.050 g of sodium diethyldithiocarbamate was also added to prevent peroxide formation in the polyglycol. Such chromatograms are developed with nonpolar solvent systems.

The Preparation of Reversed Phase Chromatograms

The impregnation of thin layers with a high-boiling, nonpolar liquid such as Silicone fluid or a hydrocarbon, leads to a partition chromatogram in which the moving phase is more polar than the stationary phase. Such chromatography is called reversed phase chromatography and normally gives improved separations in a homolo-

gous series of nonpolar compounds. This has been particularly adapted to the fatty acid series.

Malins and Mangold^{230,232} prepared Silicone layers by slowly immersing Silica Gel G layers in a 5 per cent solution of Silicone Oil* in ethyl ether. The operation is carried out at room temperature and both the plates and the liquid must be at this same temperature. In addition, the operation must be carried out slowly. Failure to observe these precautions will produce crumbly layers.

Kaufmann, Makus and Khoe,^{182,186} and Kaufmann and Ko¹⁸¹ have used either a 10 to 15 per cent solution of undecane in petroleum ether (b.p. 40 to 60°) or a 5 per cent solution of tetradecane in the same solvent. Reversed phase layers have been made using paraffin by Winterstein, Studer and Rüegg⁴⁴⁹ (5 per cent in petroleum ether) and by Michalec, Šulc and Měšťan²⁴⁹ (0.5 per cent in ether). All of these were prepared by dipping.

Purdy and Truter³¹⁵ prepared such layers by allowing Silica Gel G layers to be "developed" with a 15 per cent solution of decane in petroleum ether. A similar process was carried out by Wagner, Hörhammer and Dengler using 5 per cent paraffin in ethyl ether.⁴³¹

After the impregnation, the layers are allowed to dry in the air for a few minutes. The development is carried out with polar solvents.

*Dow-Corning 200 fluid viscosity 10 cs. Dow-Corning Corp., Midland, Mich. This preparation is also available from Research Specialities Co. (p. 40) in a spray package for spraying layers.

CHAPTER 4

Application of the Sample

INTRODUCTION

The various factors to be considered in the application of samples to thin layers are the application solvent, the amounts to be applied and the mechanics of the application. These are largely controlled by the type of information or results desired. Thus, for diagnostic or qualitative chromatography a small amount of substance, which need not be accurately known, is applied to a single small spot. This is then developed simultaneously with spots containing possible components of the mixture (reactants, products or known constituents). For quantitative chromatography, a precise amount of the mixture to be analyzed must be applied. For preparative chromatography, large quantities of the mixture to be separated must be applied in a thin uniform band along one edge of a plate.

After the sample has been applied, it can be caused to undergo various reactions which will assist in its identification or will reveal something of its properties.

VARIABLES IN SAMPLE APPLICATION

Application Solvent

The solvent used for the application of samples to thin layers is not particularly critical, but the following factors should be considered. The ideal solvent should boil between 50° and 100° so that it does not evaporate too rapidly but can be easily removed at will. It should completely dissolve the sample to yield at least a one per cent solution, but should be as nonpolar as possible. This allows the

gous series of nonpolar compounds. This has been particularly adapted to the fatty acid series.

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Kaufmann, Makus and Khoe,^{182,186} and Kaufmann and Ko¹⁸¹ have used either a 10 to 15 per cent solution of undecane in petroleum ether (b.p. 40 to 60°) or a 5 per cent solution of tetradecane in the same solvent. Reversed phase layers have been made using paraffin by Winterstein, Studer and Rüegg⁴⁴⁹ (5 per cent in petroleum ether) and by Michalec, Šulc and Měšťan²⁴⁹ (0.5 per cent in ether). All of these were prepared by dipping.

Purdy and Truter³¹⁵ prepared such layers by allowing Silica Gel G layers to be "developed" with a 15 per cent solution of decane in petroleum ether. A similar process was carried out by Wagner, Hörhammer and Dengler using 5 per cent paraffin in ethyl ether.⁴³¹

After the impregnation, the layers are allowed to dry in the air for a few minutes. The development is carried out with polar solvents.

* Dow-Corning 200 fluid viscosity 10 cs. Dow-Corning Corp., Midland, Mich. This preparation is also available from Research Specialties Co. (p. 40) in a spray package for spraying layers.

CHAPTER 4

Application of the Sample

INTRODUCTION

The various factors to be considered in the application of samples to thin layers are the application solvent, the amounts to be applied and the mechanics of the application. These are largely controlled by the type of information or results desired. Thus, for diagnostic or qualitative chromatography a small amount of substance, which need not be accurately known, is applied to a single small spot. This is then developed simultaneously with spots containing possible components of the mixture (reactants, products or known constituents). For quantitative chromatography, a precise amount of the mixture to be analyzed must be applied. For preparative chromatography, large quantities of the mixture to be separated must be applied in a thin uniform band along one edge of a plate.

After the sample has been applied, it can be caused to undergo various reactions which will assist in its identification or will reveal something of its properties.

VARIABLES IN SAMPLE APPLICATION

Application Solvent

The solvent used for the application of samples to thin layers is not particularly critical, but the following factors should be considered. The ideal solvent should boil between 50° and 100° so that it does not evaporate too rapidly but can be easily removed at will. It should completely dissolve the sample to yield at least a one per cent solution, but should be as nonpolar as possible. This allows the

sample to be concentrated at the center of the spot and not in a ring around the edge. The application solvents should, in any case, be removed as completely as possible before the chromatogram is developed.

Amount of Sample Applied

The easiest and best way to ascertain this quantity is by experimentation with the specific compounds in question. Every compound or mixture should be investigated at several different concentrations. This can be done on a single plate and, simultaneously, information about trace impurities, relative concentrations and separabilities can be obtained. An amount is then chosen which is large enough to be effectively visualized and to show trace impurities, and small enough to give discrete spots with a minimum of tailing. This amount will depend upon the layer thickness and the sensitivity of the visualization procedure which are inversely proportional to each other. Further, it will depend upon the adsorbent and whether the chromatography is adsorption or partition. The former has a much higher capacity. When the sample is too large, tailing³⁰⁷ and high *R_f* values³⁵² will result.

For diagnostic and quantitative chromatography, samples ranging from 10 to 100 γ are normally used with layers about 0.25 mm thick. For preparative work, amounts ranging from 10 to 300 mg can be separated on a single chromatogram depending upon the size of the plate, the thickness of the layers, the ease of separation, and the type of chromatography.

Specific examples of quantitative and preparative thin-layer chromatography will be given in the chapters devoted to these subjects.

MECHANICS OF APPLICATION

The application of samples to bound layers is carried out by touching the tip of a filled capillary, micropipette, or microburette to the adsorbent layer in a manner analogous to that used in paper chromatography. Multiple applications with solvent evaporation after each are usually necessary. The sample is placed or "spotted" about 2 cm or 1 in. from the end of the plate so that the solvent level will be at least 1 cm below the center of the spot. The diameter of the

spot should not exceed 0.5 cm and should be as small as possible. The spot can be kept small by a rapid evaporation of solvent using a stream of dry nitrogen²³² or a low-temperature hot plate.^{223,244,253,329} Honegger¹⁵⁷ and Wagner, Hörhammer and Wolff⁴³² advocate the application of the sample in thin bands about 1 cm long rather than the conventional circular spots. The bands are at a 90° angle to the direction of development. Tate and Bishop³⁹⁸ applied samples with a thin wire loop (about 1 mm in diameter) rather than the usual capillary and Metz²⁴⁷ used a paper gadget to make the application point very small.

The manufacturers furnish a "spotting template" to help with the application of a large number of samples. This is a transparent sheet of plastic which fits over a thin-layer plate but does not touch the layer. It is suspended on two side pieces and is marked at 1-cm intervals. The Desaga-Brinkmann model is shown in Figure 4.1.

When a large number of samples must be applied over a long period of time, it is a good idea to cover the layer above the sites of application with a clean glass plate to protect it from moisture.⁴⁶

Diagnostic Chromatography

For diagnostic work, the sample is usually applied through a melting point capillary tube of indeterminate size or a micropipette calibrated in microliters. Such a pipette is furnished by all manufacturers of thin-layer equipment.

Quantitative Chromatography

For this purpose, precise application methods are necessary. A microsyringe* used as such or with a mechanical push-button dispenser (Brinkmann, p. 37) is shown in Figure 4.2. The Agla** micrometer syringe has also been widely used. An appropriate microburette could be used.

Preparative Chromatography

The application of samples for preparative thin-layer chromatography is quite a different problem. Large quantities (50 to 300 mg)

* The Hamilton Company, Inc., Whittier, Calif.

** Burroughs Wellcome Co., The Wellcome Bldg., Euston Rd., London N.W. 1, England.

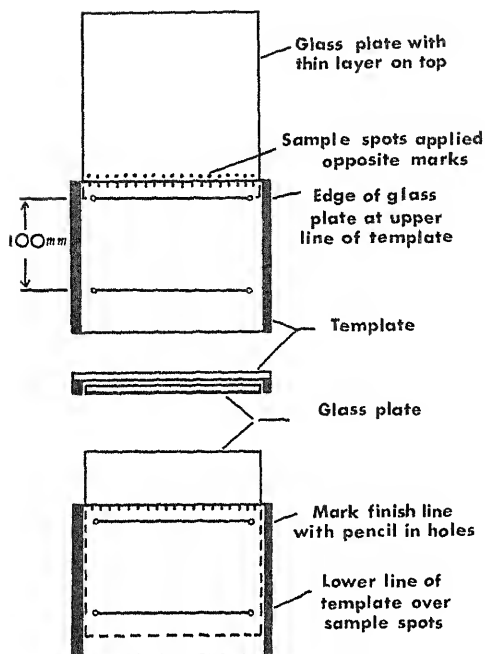


Figure 4.1. The Desaga-Brinkmann "spotting template" in position (Reproduced through the courtesy of Brinkmann Instruments, Inc.)

must be applied in a reasonable time as a uniform band without destroying the layer. Multiple spotting with a single capillary is quite unsatisfactory because it is so tedious and almost always results in a disturbed layer.

The simplest approach to this problem is to blow a fine jet of sample and solvent into the layer from a suitable pipette while moving the pipette in a uniform motion over the plate. The pipette should have a small orifice and the operator needs a deft pair of hands.* Honegger¹⁵⁸ has suggested that the sample be applied in a shallow V-shaped ditch made on the starting line. Care must be taken so that the ditch does not go through to the glass plate.

In addition to these methods, two mechanical gadgets are commer-

* The author has seen this done by Mr. P. J. DeMuria at the Chas. Pfizer Laboratory in Groton, Conn. and it has been reported elsewhere.

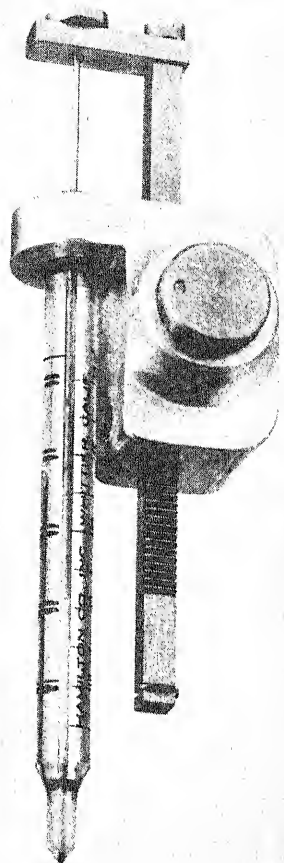


Figure 4.2. Microsyringe with push button dispenser for use in quantitative thin-layer chromatography. (Reproduced through the courtesy of Brinkmann Instruments, Inc.)

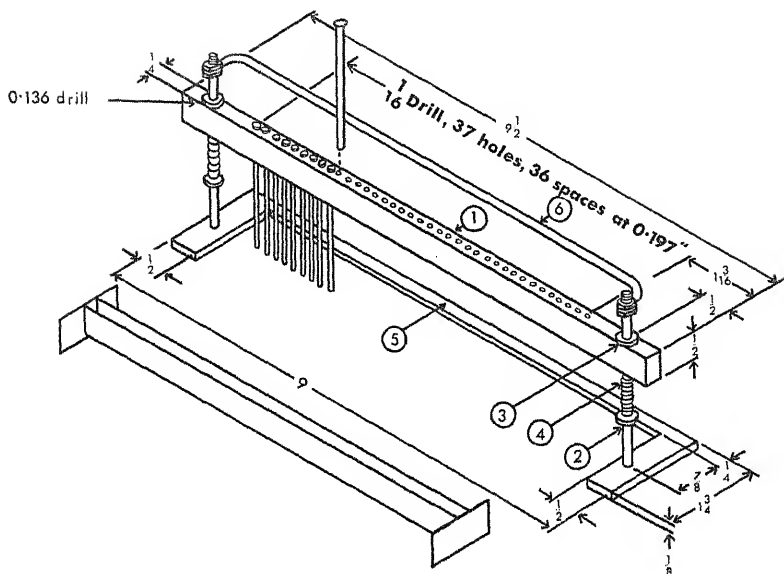


Figure 4.3. The Morgan apparatus for sample application in preparative thin-layer chromatography. The apparatus consists of an aluminum bar (1) with holes drilled near its ends which permit it to slide freely on two vertical $\frac{1}{8}$ inch brass rods (2) and is held up against adjustable stop washers (3) by partially compressed springs (4). The top and bottom stop washers are $\frac{1}{4}$ inch OD and $\frac{3}{8}$ inch ID and are enlarged slightly with a punch so that they may be forced onto the vertical rods. The bottom washers which support the springs are soldered in place while the top washers remain in place by friction and may be adjusted by forcing either up or down. The two vertical rods threaded, screwed into tapped holes, and soldered in place in the cutout base plate (5) and the upper horizontal tie rod (6) form the rigid support for the apparatus. Thirty-seven precisely drilled $\frac{1}{16}$ inch holes spaced 0.197 inch (0.5 cm) on centers in the aluminum crossbar carry the spotting pipettes. These are made from pieces of 1.5 mm OD Pyrex melting point capillary tubing (Corning #9530) selected so that they will just pass through the holes in the crossbar with no friction. A funnel stop is blown at the upper end of each tube, the tubes cut at a length such that the tips will extend to within $\frac{1}{16}$ inch of the surface of the chromatographic plate when placed in the crossbar. The delivery tips are fire polished and squared off by grinding on a flat piece of fine carborundum wetted with water. Progress in the latter operation should be observed under suitable magnification and the grinding continued only until the contact surface of the tip is smooth.

A sample trough of suitable volume and design to be used with the apparatus in spotting preparatory plates may be constructed from metal, glass, or plastic depending upon the chemical nature of the samples and solvents employed. A trough of the design shown was constructed of No. 23 gauge Monel metal and has proven to be satisfactory for use with non-corrosive solvents. (Reproduced from Morgan, *J. Chromatog.*, **9**, 379 (1962) through the courtesy of the author and the Elsevier Publishing Co.)

cially available for sample application in preparative chromatography. These are the multiple capillary arrangement of Morgan²⁵⁸ and the moving syringe system of Ritter and Meyer.³²⁹

The Morgan Apparatus. The Morgan apparatus is, simply, some thirty-seven capillary tubes arranged in a movable holder and is shown in Figure 4.3 with instructions for its construction. All of the tubes are filled by dipping them into a trough containing the sample in solution and they spot simultaneously when touched to a thin layer. It has been used with great success by the author and elsewhere at the University of Connecticut, and is commercially available from the A. H. Thomas Co.*

The Ritter and Meyer Apparatus. The Ritter and Meyer apparatus is a syringe equipped with a pressure valve that is supported by two bars on which it slides freely. The sample is placed in the syringe, a slight pressure is applied and the syringe is moved back and forth spraying a thin jet of sample. The plate holding the layer is placed on an asbestos-covered aluminum block which is heated to facilitate the removal of solvent. The apparatus is shown in Figure 4.4. It is commercially available from Desaga-Brinkmann (see p. 37).

Sample Application to Unbound Layers

The unbound layers (excluding the Woelm and Anisil special adsorbents) are not mechanically stable enough to touch so that samples are applied in small drops from a short distance away.⁵² When unbound layers are developed in a near horizontal position, the sample can be adsorbed on a small amount of adsorbent and placed on the thin layer as such.²⁶⁹ This latter technique is especially useful if one wishes to rechromatograph a sample on a new layer. The adsorbent containing the sample can simply be scraped off of one plate and applied to another without elution. Samples are applied to Woelm and Anisil layers in the same manner as with bound adsorbents.

CHEMICAL REACTIONS ON APPLIED SAMPLES

Sometimes, it is advantageous to modify or prepare a derivative of a sample after it has been spotted on a thin layer. This can be accom-

* Arthur H. Thomas Co., Vine Street at 3rd, Philadelphia, Pa.

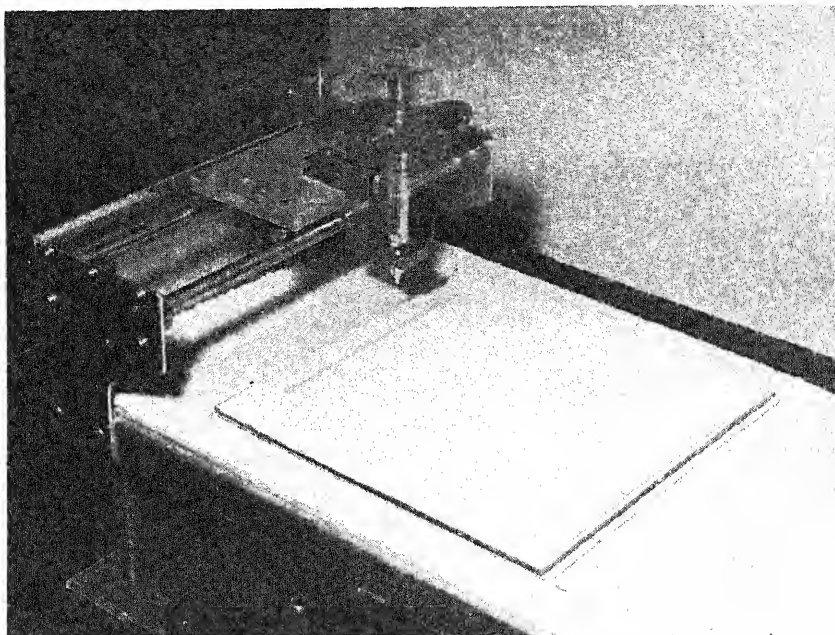


Figure 4.4. The Ritter and Meyer apparatus for sample application in preparative thin-layer chromatography. (Reproduced from Ritter and Meyer, *Nature*, 193, 941 (1962) through the courtesy of the authors and Macmillan and Co., Ltd.)

plished by adding the reagent to the developing solvent. Thus, bases can be regenerated from their salts and compounds can be brominated or oxidized by putting the appropriate reagents in the developing system. Alternately, a drop of reagent can be added to the spotted sample before chromatography. This technique has been used for oxidation, dehydration, hydrogenation and the preparation of derivatives. These reactions are carried out in the following manner.

Regeneration of Bases

The amine or alkaloid is spotted on the layer as its hydrochloride, and a base, diethylamine⁴³⁶ or ammonia,³³² is added to the developing solution to the extent of 1 or 2 per cent. A similar reaction was performed in the thin-layer separation of the halide ions³⁵⁷ by Seiler and Kaffenberger. The halides were spotted as their sodium and po-

tassium salts, but the developing solution contained 10 per cent of concentrated ammonium hydroxide and the ions migrated as their ammonium salts. Mottier²⁶⁶ neutralized dye solutions before chromatography on alumina by placing a drop of 1 *N* sodium hydroxide solution on the layer, drying it and then spotting the sample.

Bromination¹⁸⁶

The sample is spotted in the normal way and the chromatogram is developed by solvents containing 0.5 per cent (by volume) of bromine.

Oxidation

Oxidation is brought about by addition of an oxidizing agent to the developing system or by direct application over the sample spot. Thus, Mangold^{230,231,232} separated saturated and unsaturated fatty acids on reversed phase layers with a mixture of acetic acid-peracetic acid-water (15/2/3). The unsaturated compounds were oxidized and moved to the solvent front whereas the saturated compounds were separated in a normal manner.

In an alternate technique,²⁵² the sample to be oxidized is spotted in the normal way and covered with a drop of glacial acetic acid, saturated with chromic anhydride. The chromatogram is then developed with solvents which are not easily oxidized.

Dehydration²⁵²

The sample, spotted as usual, is covered with a drop of concentrated sulfuric acid and developed with solvents such as hexane which are resistant to sulfuric acid.

Phenylhydrazone Formation²⁵²

The sample, spotted as usual, is covered with a drop of phenylhydrazine and then developed with a noncarbonyl solvent.

Semicarbazone Formation²⁵²

A 10 per cent solution of semicarbazide hydrochloride is neutralized with sodium hydroxide and spotted with the sample. The chromatogram is developed with a noncarbonyl solvent.

Hydrogenation¹⁸⁰

The spotting region of a layer is sprayed with a colloidal solution containing 1 to 2 per cent palladium in an electrolyte free solution.* The plate is dried in an oven, the sample is spotted in the region containing the palladium, and the layer is exposed to a hydrogen atmosphere in a vacuum desiccator. The plate is developed in the normal manner.

Irradiation³⁷¹

The spotted sample is irradiated with a suitable energy source in the desired atmosphere and developed in the normal manner.

Salt regeneration and bromination normally go to completion but the other reaction products are contaminated with starting materials which can be used, untreated, on the chromatograms as standards. The reactions listed certainly represent only a small portion of the possibilities for this technique.

* The solution is commercially available from Firma Dr. Th. Schuchardt, Ainmillerstr. 25, Munich 13, Germany.

CHAPTER 5

Choice of a Solvent System

INTRODUCTION

The type of solvent system chosen to make a given separation will depend primarily upon whether the chromatography is adsorption or partition. Adsorption chromatography is the simpler of the two because the solvents are completely miscible in one another and their ability to move a sample follows the normal eluotropic series. On the other hand, partition chromatography involves a selection of a two-phase system which will have desirable partition coefficients for the particular sample to be separated. Furthermore, partition chromatography can be of two types depending upon whether the stationary phase is polar (normal) or nonpolar (reversed phase). An even more complex situation involving various ionic species arises with ion-exchange chromatography on thin layers.

It is sometimes advantageous to add a small amount of some special reagent to developing systems to bring about a desired chemical reaction, to improve separations or to change the layer in some manner.

ADSORPTION CHROMATOGRAPHY

The normal worker will spend more time searching the literature for solvent systems to use in adsorption chromatography than he will in determining these systems by experimentation. This is particularly true since the exact adsorptive properties of layers tend to vary from one laboratory to another. The classic eluotropic series of Trappe,⁴⁰³ Strain³⁹³ and Knight and Groennings²⁰² apply to thin-layer chroma-

tography²³² and are given in Table 5.1. The more polar solvents producing the greatest migrations are at the bottom of the columns.

A preliminary study can be made in two ways. The first, devised by Izmailov and Shraiber,¹⁶⁹ Crowe,⁶³ and further developed by Stahl,^{366,369,370} is the simpler. It involves the spotting of the sample in several places, an inch or so apart, on a single layer. Pure solvents of various polarities are deposited at the centers of these spots by means of a capillary and the outer edge of the solvent spot is noted. The layer is sprayed or visualized in some manner and the solvent which moves the sample about halfway between the origin and the outer edge is the best solvent. The second way to make a preliminary study is by running chromatograms with various solvents. In some laboratories, a series of chromatography jars containing certain key solvents like hexane, benzene, ether and methanol are kept ready for this purpose. In one of these ways, the ideal pure solvent or a pair of solvents whose properties "bracket" the material to be separated (one moving it less and one moving it more than desirable) are selected. In the latter case, various mixtures are investigated, or a gradient elution

TABLE 5.1. ELUOTROPIC SERIES OF SOLVENTS

Trappe ⁴⁰³	Strain ⁴⁰³	Knight and Groennings ⁴⁰²
Light petroleum	light petroleum 30-50°	heptane
Cyclohexane	light petroleum 50-70°	diisobutylene
Carbon tetrachloride	light petroleum 50-100°	benzene
Trichloroethylene	carbon tetrachloride	isopropyl chloride
Toluene	cyclohexane	isopropyl ether
Benzene	carbon disulfide	ethyl ether
Dichloromethane	anhydrous ethyl ether	ethyl acetate
Chloroform	anhydrous acetone	sec-butyl alcohol
Ethyl ether	benzene	ethyl alcohol
Ethyl acetate	toluene	water
Acetone	esters of organic acids	acetone
n-Propanol	1,2-dichloroethane	methanol
Ethanol	alcohols	pyruvic acid
Methanol	water	
	pyridine	
	organic acids	
	mixtures of acids or bases,	
	water, alcohols, or pyridine	

system³³⁴ can be used. The gradient elution technique will be discussed more extensively in Chapter 6.

The author and others^{144,232} strongly recommend that developing mixtures be kept as simple as possible. In most situations, a mixture of two components will produce satisfactory results. For example, the author has separated a wide range of compounds such as glycosides, aglycones and nitrogen heterocyclics with mixtures of methanol and ether.

One should not become too concerned about whether the substance to be chromatographed is soluble in the solvent system. There appears to be no direct relation between R_f and solubility. This is probably because migration involves solution from a molecular or amorphous state while normal solubility requires the breakup of a crystal structure. Thus, sucrose and glucose can be moved readily on silica gel by methanol-ether solutions in which they are apparently insoluble.

NORMAL PARTITION CHROMATOGRAPHY

In this case, the substances to be separated (generally very polar compounds like sugars or amino acids) are caused to distribute themselves between a polar stationary phase (water, acetic acid, phenol or formamide) on the layer and a moving nonpolar phase. The nonpolar phase normally consists of two organic liquids having different polarities. Since the actual developing solution must contain a miscible mixture of all three of these components, it is likely that the literature will provide easier information than random experimentation.

The two liquids constituting the nonpolar phase provide the major point of variation for partition chromatography. Within the bounds of mutual solubility, their ratio can be varied to provide a greater or lesser dissolving power for the moving phase. The literature on paper chromatography provides at least a starting point in search for a system and is directly applicable when the layer consists of powdered cellulose³¹⁹ or Silica Gel G.⁴⁴ The literature summarized in Chapter 11 cites many solvent systems which have been used in thin-layer work.

REVERSED-PHASE PARTITION CHROMATOGRAPHY

In this case, the stationary phase is nonpolar and is usually a high boiling hydrocarbon or a Silicone liquid which has been applied to the layers are described in Chapter 3, p. 49. The moving, more polar phase, might be water, acetic acid, acetonitrile, alcohol, or some combination of these solvents. The method is useful for separating compounds of a homologous series such as the fatty acids or their derivatives. Table 5.2, taken mainly from Mangold²³² suggests several possible solvent systems.

TABLE 5.2. DEVELOPING SYSTEMS FOR REVERSED PHASE PARTITION CHROMATOGRAPHY*

Solvents ¹	Ratios, v/v	Stationary Phases ²	Lipid Classes Fractionated	Ref.
Methanol	—	higher paraffins	carotenals	449
Acetic acid-water	24/1	undecane	fatty acids	182
—	3/1	Silicone oil	fatty acids and their methyl esters	230
—	17/3	Silicone oil	fatty acids and their methyl esters and aldehydic cores derived from lecithins	230
Acetic acid-acetonitrile	1/1	undecane	fatty acids	182
Acetic acid-acetonitrile-water	2/14/5	Silicone oil and squalane	methyl-esters of fatty acids	230
Methyl ethyl ketone-acetonitrile	7/3	higher paraffins	cholesteryl esters of fatty acids	184
Chloroform-methanol-water	5/15/1	undecane	diglycerides	182
Acetone-acetonitrile	7/3	undecane	triglycerides	182
Acetone-ethanol-water	6/1/3	polyethylene	methyl esters of fatty acids	232
Acetic acid-formic acid-water ³	2/2/1	Silicone oil	fatty acids saturated/unsaturated	230
Acetic acid-peracetic acid-water	15/2/3	Silicone oil	fatty acids saturated/unsaturated	230

¹ All solvent mixtures must be saturated with the stationary phase.

² Silica Gel G served as a carrier for the stationary phases. Only polyethylene was applied directly to the plates.

³ Chromatographed at 4-6°. (All others at room temperature.)

* Reproduced from Mangold, *J. Am. Oil Chemists' Soc.*, **38**, 708 (1961) through the courtesy of the author and The American Oil Chemists' Society.

ION-EXCHANGE CHROMATOGRAPHY

The modified-cellulose ion exchangers DEAE and ECTEOLA-celluloses have been used by Randerath^{316-318,320} for the separation of nucleotides. He used dilute hydrochloric acid (0.02 to 0.04 *N*) to develop his chromatograms. Determann, Wieland and Lüben⁸² used a gradient elution system, adding increasing amounts of sodium chloride to a phosphate buffer (*pH* 7.2), to separate enzymes and nucleotides on "Sephadex." Berger, Meyniel and Petit²⁹ used 1 *M* sodium nitrate to develop halogen ions on "Dowex 1" and acetic acid-methanol-acetone to separate fluorescein dyes on the same resin.

MISCELLANEOUS SPECIAL ADDITIVES

Several substances have been added to developing solutions for one reason or another. As mentioned (p. 58), bases,^{332,436} bromine^{186,295} and oxidizing agents²³⁰⁻²³² are added to bring about some reaction with the sample or some part of it. Seiler and Seiler³⁵⁹ have added small amounts (0.4 per cent) of the complexing agent, 2, 5-hexanedione, to facilitate the separation of inorganic cations. Mistryukov²⁵⁴ added a small amount of decalin to developing solutions to be used on nonbound layers. The decalin increased the mechanical stability of the layer so that it could be visualized by spraying.

Small amounts of acidic or basic solvents are often added to developing systems to facilitate the separation of acids and bases, respectively, and to prevent "tailing" or smearing. "Tailing" is due, among other things, to the presence of more than one ionic species of the substances being chromatographed.²⁷⁴ Thus, Mangold²³² and others²⁶⁴ suggest the addition of 1 to 2 per cent of acetic acid to systems for developing chromatograms of fatty acids. In a similar manner, systems containing ammonia^{232,256} and diethylamine^{332,436} are used for separating nitrogenous lipids and other bases. These additives buffer the system and keep the material being separated exclusively in the non-ionic form.

CHAPTER 6

Development of Thin-Layer Chromatograms

INTRODUCTION

The development process, that is, the actual passing of solvents through thin layers to cause separations, is a major point of variation in thin-layer chromatography. Like the preparation of layers, it challenges the imagination of the laboratory worker and many ingenious techniques have resulted.

Thin-layer chromatography has been carried out in an ascending manner on vertical or near horizontal plates, in a descending manner on vertical plates, or in a linear or circular manner on horizontal plates. The development can be performed discontinuously, one pass or multiple pass, in a stepwise fashion with different solvents, continuously or by gradient elution methods. The layers can be scored in such a way that the chromatography is carried out in wedge or modified wedge-shaped areas. The process can be carried out in a two-dimensional fashion with different solvent systems being used in each direction. The layer or the sample may also be modified between two developments. These techniques will be discussed in detail in this chapter.

Finally, the various factors which influence the development or, more precisely, the R_f values resulting therefrom will be considered. These include the nature of the adsorbent and of the developing system, the nature and amount of the sample, the chamber saturation, the use of multicomponent solvent systems and the temperature. Some

methods of reporting R_f values and the causes and alleviation of tailing will be briefly discussed.

METHODS OF DEVELOPMENT

Ascending Chromatography

Most thin-layer chromatograms are developed by placing a glass plate, containing the sample on a bound, thin layer, in a vertical position in a closed, saturated system such that the bottom of the layer dips into the developing solvent. After the solvent ascends 10 to 15 cm, the plate is removed, dried, and visualized in some manner. Two or more layers can normally be processed in the same container provided that care is taken to completely saturate the system.

Kirchner, Miller and Keller¹⁹³ found that when water is a component of the developing system, the adsorbent layer tends to slide off of the plate below the solvent level. This problem was solved by standing the plate containing the layer on a cotton wad saturated with solvent. The author has observed a similar effect with very polar solvent systems, but found that even though the adsorbent below the liquid surface slides off, enough remains to carry the solvent into the layer satisfactorily.

While the plates are normally leaned against the side of the chromatographic jar or tank, several pieces of apparatus have been devised for holding them. A stainless-steel rack which will hold two large plates (20×20 cm) is available from Desaga-Brinkmann (p. 37), but a more versatile apparatus is available from Kensington Scientific Corp. (p. 36). This apparatus, pictured in Figure 6.1, is constructed from glass rods and will hold two plates (10×20 cm). It is suspended in a chromatographic chamber such that it can be raised or lowered without opening the chamber and thus allows the equilibration of a layer with the solvent before actual development. Such an equilibration step is recommended for partition chromatography.

When unbound layers are used (p. 46) they must be kept in a near-horizontal position (10 to 20° from horizontal) so that the adsorbent does not slide off. While other schemes have been devised,²⁷⁴ the simplest method for supporting these plates is that of Mottier and Potterat.²⁶⁹ As shown in Figure 6.2, the plates are propped in a large

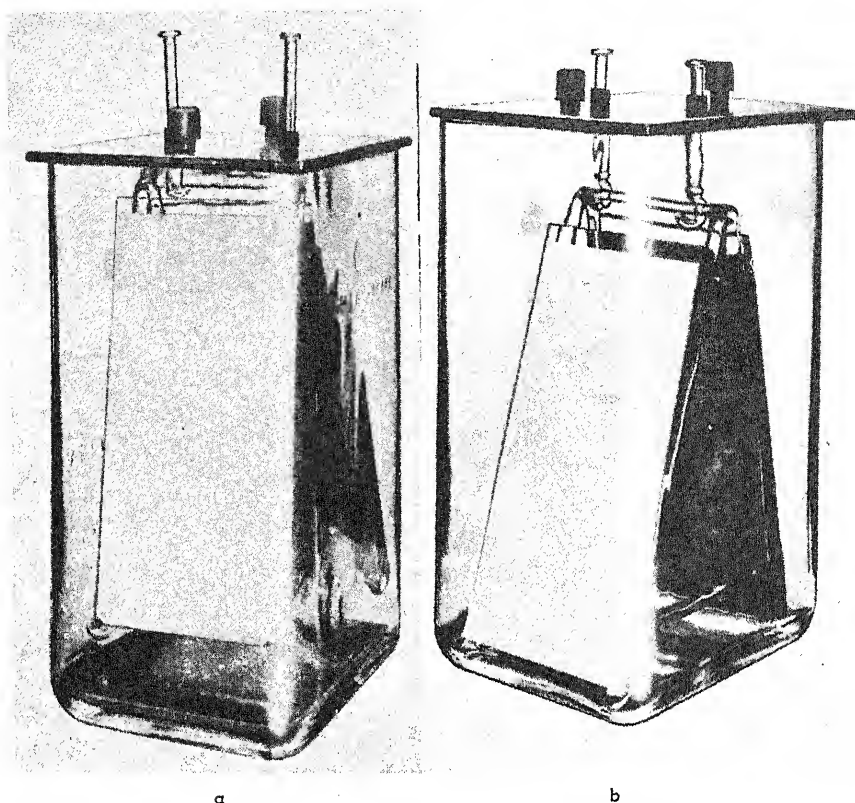


Figure 6.1. Chromatographic chamber containing a movable plate holding device which can be used to equilibrate thin-layer chromatograms prior to development. It is shown in a raised position, a, and a lowered position, b. (Reproduced through the courtesy of the Kensington Scientific Corp.)

Petri dish or crystallizing dish and the dish is tilted so that the solvent is confined to the portion touching the layer.

Descending Chromatography

The technique of descending thin-layer chromatography has essentially been worked out and applied by W. L. Stanley and his co-workers.^{383-385,388} They suspended 0.5 in. wide "chromatostrips"¹⁹³ coated with a silicic acid starch layer in an especially designed chamber as shown in Figure 6.3. The developing solvent is carried from the reservoir by a piece of coarse filter paper which is clipped to the top

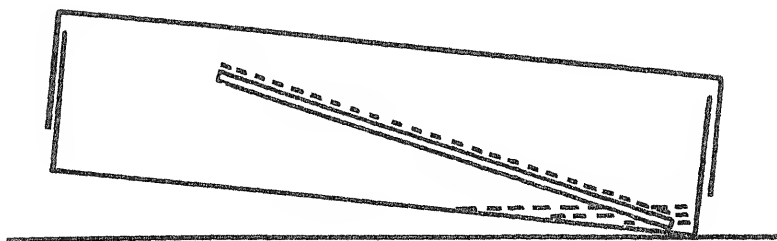


Figure 6.2. Development of a thin layer of unbound adsorbent. (Reproduced from Mottier and Potterat, *Anal. Chim. Acta*, 13, 46 (1955) through the courtesy of the authors and the Elsevier Publishing Co.)

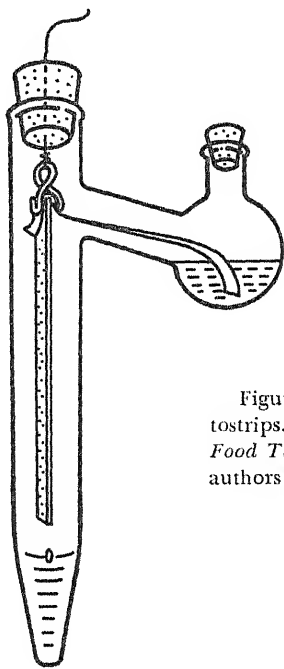


Figure 6.3. Descending chromatography on "chromatostrips." (Reproduced from Stanley, Ikeda and Cook, *Food Technol.*, 15, 381 (1961) through the courtesy of the authors and the Institute of Food Technologists.)

end of the layer. Such a system is advantageous in that it permits a continuous chromatography from which samples of eluent can be removed for analysis.³⁸⁵ Furthermore, it allows the use of relatively nonpolar solvents over a long time period and gives better separations.

Descending chromatography on wide plates containing the somewhat more fragile Silica Gel G layers has been carried out by Birkofer, Kaiser, Meyer-Stoll and Suppan³⁶ and by Zöllner and Wolfram.⁴⁵⁸ The apparatus of the latter authors is much the simpler of the two and is shown in Figure 6.4. The solvent is carried to the surface of the layer (facing down) by a piece of paper which also helps to saturate the atmosphere. The apparatus has an additional advantage in that a paper chromatography jar can easily be modified for the purpose.

Mistryukov²⁵⁵ has described an apparatus for the development, by descending chromatography, of unbound layers.

Horizontal Chromatography

Linear. Mistryukov²⁵⁴ has published a method for carrying out linear chromatography on completely horizontal unbound layers and the continuous chromatography of Brenner and Niederwieser^{45,46} as described below is on horizontal layers.

Circular. The original publication on bound layers by Meinhard and Hall²⁴⁶ involved circular chromatography of inorganic ions and since then, the technique has been used extensively.^{51,210,211,246,306,313,368} The sample is placed at the center of a layer and the solvent is fed into the center of the spot with a pipette^{210,211,246} or a cotton wick⁵¹ held in contact with the layer. The resulting chromatogram consists

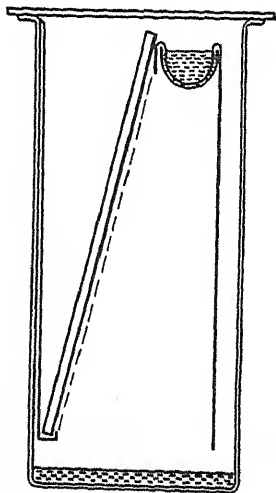


Figure 6.4. Descending chromatography on wide plates. The dotted line represents the layer. (Reproduced from Zöllner and Wolfram, *Klin. Wochschr.*, 40, 1098 (1962) through the courtesy of the authors and Springer Verlag).

of concentric rings. An alternate method which allows the comparison of several samples involves the placement of samples around and equidistant from a center point. Solvent is fed into the center point and the result consists of a series of chromatograms radiating from the center point. Figure 6.5 shows one arrangement for circular chromatography by Stahl.³⁶⁸ The apparatus requires a plate with a small hole in it.

Discontinuous Development

Single Pass. The vast majority of thin-layer chromatography has been carried out by allowing a single solvent system to pass through the layer one time. As discussed below, however, there is some reason to believe that better results could be obtained by different techniques.

Multiple Pass. Under the proper conditions, the development of a layer several times with the same solvent (with drying between developments) offers considerable advantage. The technique has been extensively used^{158,179,222,269,380,398} and has, in itself, been the subject of two important papers.

The first paper¹⁷⁵ dealt with paper chromatography, but should be equally applicable to thin layers. In it, Jeanes, Wise and Dimler state that the optimum number of passes, n_{\max} , can be predicted using Equation (6.1) where Rf_1 and Rf_2 are the values of individual com-

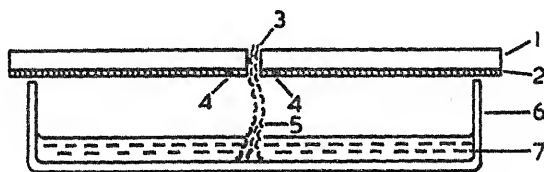


Figure 6.5. Circular thin-layer chromatography.

- | | |
|-----------------------------|----------------------|
| 1 glass plate (20 × 20 cm) | 5 cotton wick |
| 2 Silica Gel G layer (0.25) | 6 Petri dish cover |
| 3 2 mm hole | 7 developing solvent |
| 4 sample spots | |

(Reproduced from Stahl, *Parfüm. Kosmetik*, 39, 564 (1958) through the courtesy of the author and Dr. Alfred Huethig Verlag, Heidelberg, Germany).

$$n_{\max} = \frac{\log Rf_1 - \log Rf_2}{\log (1 - Rf_2) - \log (1 - Rf_1)} \quad (6.1)$$

pounds or Equation (6.2) where R_F is the average Rf value of the

$$n_{\max} = \frac{1}{R_F} - 1 \quad (6.2)$$

compounds. In the second paper⁴⁰¹ which does discuss thin-layer chromatography, Thoma gives Equation (6.3) for the calculation (using the same symbols as 6.1 and 6.2). It is apparent from these equations

$$n_{\max} = \frac{-1}{\ln_e (1 - R_F)} \quad (6.3)$$

that systems with low Rf values (0.2 or less) can be greatly assisted by repeated developments while those systems with high Rf values cannot. Thoma was able to show that better separations were, in fact, obtained by several developments with a less polar solvent system than by one development with a more polar one.

Stepwise Development. Stahl and his co-workers^{349,378} devised a technique in which the layer was developed twice to differing heights with different solvent systems. Such a system is ideal when the sample contains a group of similar polar molecules and a group of similar nonpolar molecules. The chromatogram is first developed halfway (about 6 to 7 cm) with a polar system which deposits the nonpolar compounds at the solvent front about halfway up the layer and resolves the polar mixture in the bottom half of the chromatogram. The chromatogram is removed, dried and developed with a nonpolar solvent system. This system passes over the separated polar substances and resolves the nonpolar mixture in the top half of the chromatogram. This method is illustrated in Figure 6.6³⁷⁸ by the separation of a mixture of glycosides (polar) and aglycones (nonpolar). Similar schemes have been used by Billeter and Martius³⁵ for the separation of the K vitamins and by Mangold and Kammereck²³³ for the separation of lipid derivatives.

Weicker⁴³⁸ used still a third step for the separation of complex lipid mixtures. He first developed silica gel layers to 3 cm with propanol-ammonia (2:1). This carried the fatty acids, cholesterol and cholesterol esters to the solvent front and resolved lecithins and polar

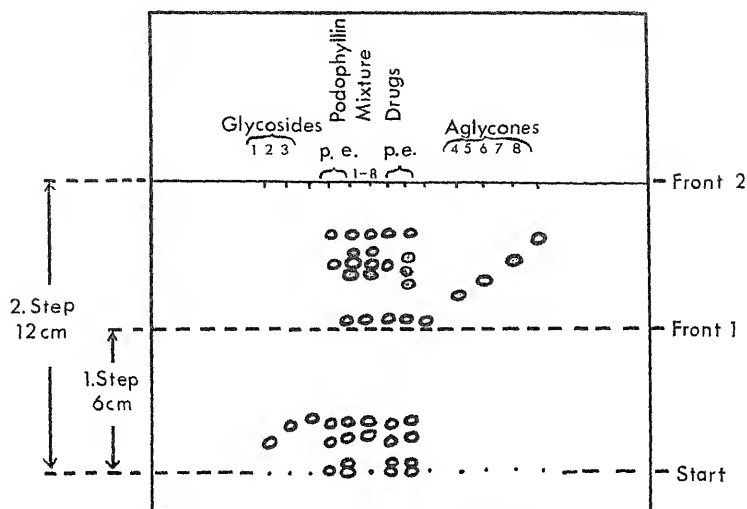


Figure 6.6. Thin-layer chromatography of *Podophyllum* components using a stepwise technique. 1. α -Peltatin- β -D-glucoside; 2. Podophyllotoxin glucoside; 3. β -Peltatin- β -D-glucoside; 4. 4'-Demethylpodophyllotoxin; 5. α -Peltatin; 6. Podophyllotoxin; 7. β -Peltatin; 8. 1-Deshydroxypodophyllotoxin. The plant extracts (marked podophyllin) and the drugs of two species of *Podophyllum*, *peltatum* L. (p) and *P. emodi* Wall (e) are shown along with a synthetic mixture of the eight compounds. The adsorbent was Silica Gel G and the solvent systems for steps 1 and 2 were 10% methanol in chloroform and 35% acetone in chloroform respectively. (Reproduced from Stahl and Kaltenbach, *J. Chromatog.*, 5, 548 (1961) through the courtesy of the authors and the Elsevier Publishing Co.)

lipids. A second development with chloroform-benzene (3:1) to 10 cm resolved the fatty acids and free cholesterol and carried the cholesterol esters to the solvent front. The layer was then rotated 180° and developed to a height of 4 cm with carbon tetrachloride to resolve cholesterol esters.

Still another variation⁴¹³ involves replacement of old adsorbent with new, and subsequent chromatography. Thus, in the separation of lipids, the crude mixture is developed first with ether-petroleum ether (5:95) on an unbound alumina layer. This moves the triglyceride fraction and the cholesterol esters. The adsorbent containing these components is removed and they are then assayed. The portion of the plate which contained these components is then cleaned

and a new layer added. Development with petroleum ether-ether-acetic acid (94.5:5:0.5), then moves and separates cholesterol, fatty acids and phospholipids on the new layer and they are subsequently removed and assayed.

Gradient Elution. Rybicka^{334,335} and Determann, Wieland, and Lüben⁸² have developed apparatus for gradient elution of thin-layer chromatograms, that is, for changing the solvent during the actual course of development. In this case a chromatography jar equipped with a magnetic stirrer is used. The chromatogram is dipped into one solvent (the less polar) in the bottom of the jar. Simultaneously, a second solvent (the more polar) is added from a burette which is connected to a glass tube extending through a hole in the top of the jar to the solvent surface. As the development and the addition continue, the properties of the solvent system slowly approach those of the added solvent. Using this technique, Rybicka separated glycerides^{334,335} on silica gel layers, Hofmann¹⁵³ separated proteins on hydroxyl-apatite and Determann, Wieland and Lüben⁸² separated enzymes and nucleotides on "DEAE-Sephadex."

Chromatography in Shaped Areas. Sometimes, layers are scored or divided with a sharp object into various shaped areas such as is shown in Figure 6.7. These shapes are such that the solvent enters through a narrow area of the layer and is forced to expand into the wider areas. The method was introduced by Mottier²⁶⁶ who used shape 6.7a in unbound alumina layers. Shapes 6.7b and 6.7c have been used by Stahl³⁰⁸ and Peereboom and Beekes²⁰⁵ and a wedge shape, 6.7d, is preferred by Prey, Berbalk and Kausz.³⁰⁷ The result of such a chromatogram is that the horizontal dimension of the spots is increased and the vertical dimension is decreased. The resulting narrow bands allow a greater number of compounds to be more completely separated than usual. An additional advantage of this technique is that small quantities of one component can be found in the presence of a large amount of another.³⁰⁷ The chromatogram 6.7c is the result of the separation of a mixture of one hundred parts of saccharose to one part of raffinose.

Two-Dimensional Chromatography. Sometimes, more components are in a mixture than will separate clearly on a one-dimensional strip. Or, perhaps, the mixture contains substances that differ from one

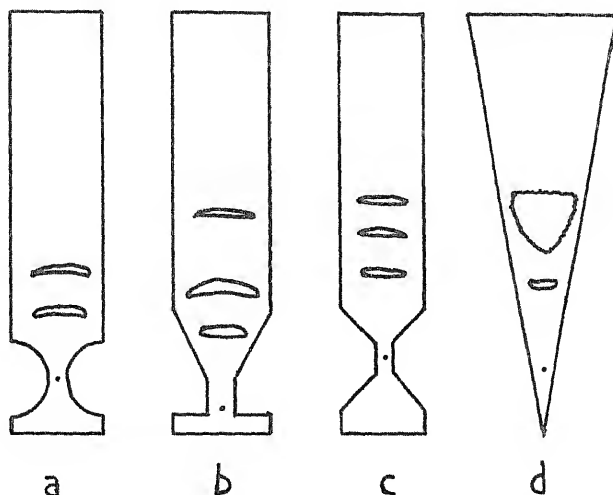


Figure 6.7. Chromatography in shaped areas according to Mottier,²⁰⁰ a, Stahl,³⁰⁸ b, Peereboom and Beekes,²⁰⁵ c, and Prey, Berbalk and Kausz,³⁰⁷ d.

another so much that two solvent systems should be used. In these situations chromatography is often carried out in two directions which are perpendicular to one another. The general procedure is as follows. The sample is spotted in one corner of a square layer and developed so that the components are resolved along one edge of the plate. The chromatogram is then removed from the developing jar, dried and placed in a second solvent so that the previously formed component spots along the bottom of the plate are again submitted to chromatography. Such a system has been used extensively in thin-layer chromatography.^{21,44,46,87,182,184,193,221,269,317,352,371} The major point of variation in this technique is what is done to the solvent system or to the layer between the two developments.

Normally, two different solvent systems are used in the two developments so that the sample is exposed to the broadest possible conditions. The layer can be modified between the developments by dipping it in a hydrocarbon or a Silicone oil solution²³² to make a reversed phase chromatogram. The latter system has been especially useful in the lipid field where adsorption chromatography is used in one direction to separate the lipid classes and reversed phase chro-

matography is used in the second direction to resolve the classes. Anet⁵ has actually changed the adsorbent between developments. He first chromatographed 2,4-dinitrophenylhydrazones on Silica Gel G in one direction with toluene-ethyl acetate (3:1); then scraped off the adsorbent above the row of spots and replaced it with Aluminum Oxide G, and finally developed with the same solvent in the new adsorbent. The layers were made in each case with the modified⁴⁵² Camag apparatus (see p. 36).

Any of the modifying reactions listed in Chapter 3 (p. 57) can be carried out after the first development. Thus, hydrogenation,¹⁸⁶ bromination¹⁸⁶ or irradiation³⁷¹ sometimes modifies a mixture in such a way that the compounds which did not separate in the first development will do so in the second. This technique, called separation-reaction-separation by Stahl,³⁷¹ is clearly shown in the chromatogram in Figure 6.8 prepared by Stahl, where an irradiation reaction was carried out between separations. The chromatogram also shows the use of standards.

Continuous Development

Continuous development, or elution through a thin layer of adsorbent, offers two advantages. First, samples can be eluted off the end of the chromatogram, thus allowing the separation of mixtures containing components of widely differing polarities. Second, it allows the use of a relatively nonpolar solvent system over a long period of time. In this respect, it should be noted that a very polar developer tends, not only to move things too fast, but to override small differences between mixture components. Of course, with increased developing times, the spots diffuse more and separations become less sharp.

One method which has been used for this purpose (p. 68) is descending chromatography.^{385,388} In this technique, the solvent drips from the bottom of the layer into the chromatography chamber and can be collected easily. A commercial apparatus is available from Research Specialties Co. (see p. 40) for this purpose. The apparatus is similar to that shown in Figure 6.3 except that there is an opening in the bottom of the chamber.

A second approach for continuous development has been described

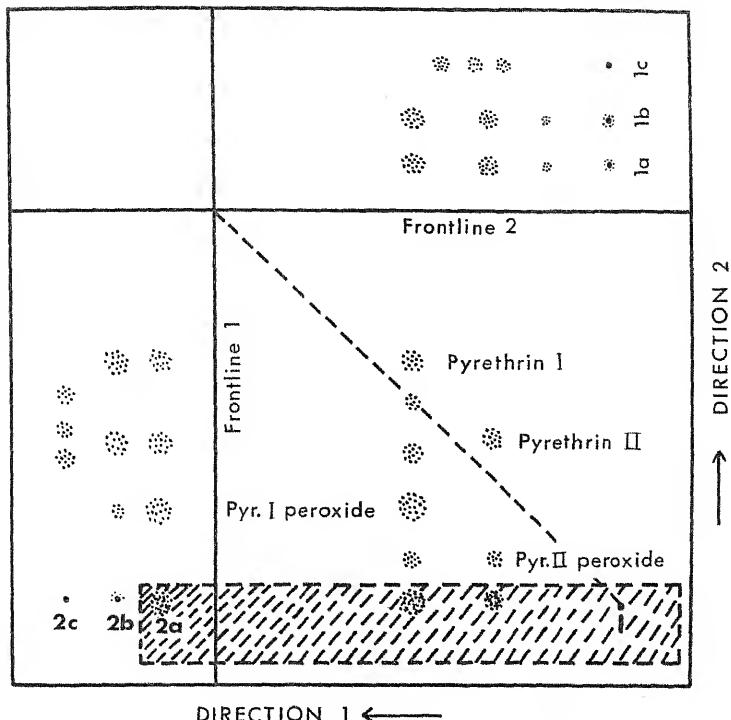


Figure 6.8. Thin-layer chromatography of pyrethrins by the separation-reaction separation-technique. The irradiation (dotted line area) was performed after the first separation. The same solvent system, 25% ethyl acetate in hexane, was used for both separations. The sample was initially placed at 1. Spots 1a, 1b and 2a and 2b are different concentrations of the starting mixture which will be developed in only one direction in order to orient the results of the two dimensional chromatogram. Spots 1c and 2c are a commercial (Desaga-Brinkmann) dye mixture of butter yellow, indophenol and Sudan Red G which is used to calibrate the layer and validate the R_f values given. The adsorbent was Silica Gel G. (Reproduced from Stahl, *Arch. Pharm.*, 293/65, 531 (1960) though the courtesy of the author and Verlag Chemie.)

by Brenner and Niederwieser.^{45,46} In this technique, the solvent is carried by a paper wick to one edge of a thin layer, passes through it and evaporates at the opposite edge. Samples cannot be taken from the end of the layer in this apparatus, which is illustrated in Figure 6.9. This has been used for the separation of amino acids⁹⁹ and for

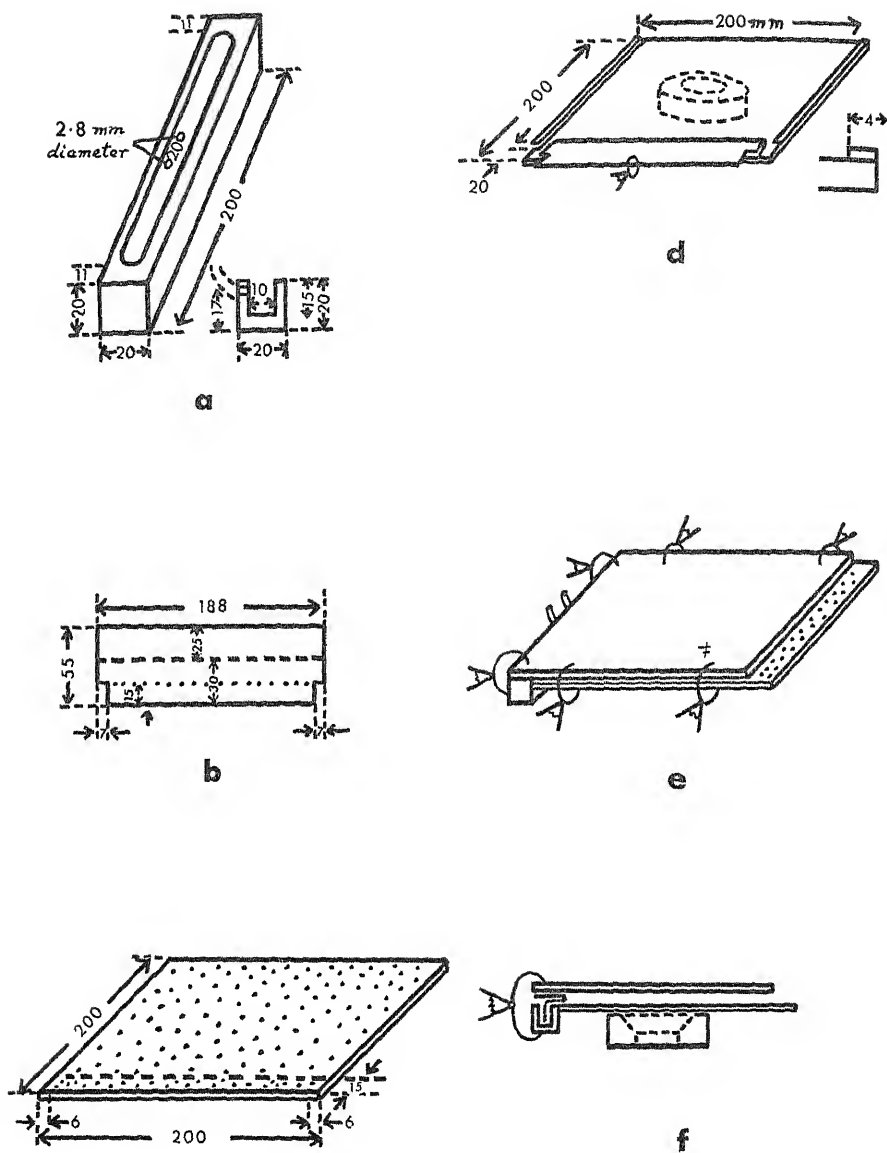


Figure 6.9. Brenner-Niederwieser apparatus for continuous, horizontal thin-layer chromatography. All measurements are in mm. Drawing a shows the steel solvent trough and the holes which feed solvent into it. Drawing b is the filter paper wick which carries the solvent to the layer. Drawing c shows a thin layer which has been spotted on a line 15 mm from the bottom and from which the adsorbent has been removed on the outermost 6 mm of the sides. Drawing d shows the layer in place on a cork ring with the wick in place and a detail of the edge strip which will support the cover plate. Drawing e shows a top view of the assembled apparatus and drawing f shows a side view (without the edge strips). In operation the sample is applied, the apparatus is assembled and the solvent is added. (Reproduced from Brenner and Niederwieser, *Experientia*, 17, 237 (1961) through the courtesy of the authors and Birkhaeuser Verlag).

preparative separations.¹⁵⁸ A commercial apparatus is available from Desaga-Brinkmann (see p. 37).

Essentially the same thing has been accomplished in a simpler fashion by developing in chambers without lids or by using a chamber shorter than the layer is long. In the latter case, a cover with a slit for the plate is used as shown in Figure 6.10a. The solvent evaporates after it emerges from the slit. The author has carried out continuous development by placing a piece of filter paper over the top of the plate as shown in Figure 6.10b. The filter paper is held in place with rubber bands or paper clips and leads the solvent over the top of the plate.

Layer Electrophoresis

Electrophoresis has been carried out on thin layers by Honegger¹⁵⁶ and by Pastuska and Trink.²⁹¹ Honegger used Silica Gel G, Kieselguhr G and Aluminum Oxide G for the separation of amines and amino acids and combined the two processes of thin-layer chromatography and electrophoresis. That is, he separated a mixture by chromatography in one dimension and subjected the results to electrophoresis in the second dimension. Pastuska and Trink investigated amino acids, phenols, amines and dyes. In both cases, the layer was

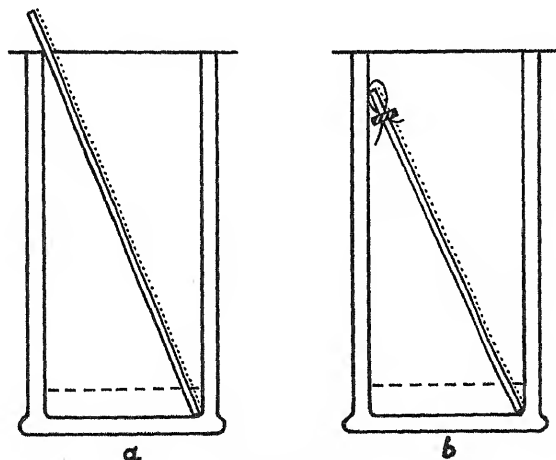


Figure 6.10. Simple arrangements for continuous thin-layer chromatography.

impregnated with an appropriate electrolyte and current was transmitted to the layer with filter paper moistened with electrolyte.

REPRODUCIBILITY OF R_f VALUES

One of the major disadvantages of thin-layer chromatography is that R_f values of substances are not exactly reproducible, at least from one laboratory to another. Furthermore, different samples of the same compound sometimes show slightly different behavior *even on the same layer*. These difficulties are more common in adsorption thin-layer chromatography and are probably traceable to the complexity of surface adsorption or to the fact that no two laboratories prepare layers in exactly the same manner. Many workers^{47,112,144,146,227,252,290} have commented extensively on this problem and their remarks will be summarized by considering the various factors which determine and affect R_f values.

Nature of the Adsorbent

The nature of the adsorbent and the thickness and moisture content of the layers are major factors in the determination of R_f values and have been discussed previously in this book. The use of standard, commercial adsorbents, layer thicknesses over 0.15 mm and consistent drying times will largely reduce the variations. It has been noted,⁴⁷ however, that there is an appreciable difference between batches of Silica Gel G. The worker should also remember that an activated layer becomes less active after exposure to a moist atmosphere for appreciable times.

Developing System

The nature of the developing system is, of course, a prime factor in any discussion of R_f values and should be controlled as closely as possible by the use of reagent grade solvents. Solvent mixtures should be made up frequently. This is especially true when very volatile solvents such as ether or pentane are being used. The mere opening of a chamber containing such a volatile solvent in a mixture can change the concentration appreciably.

Nature and Amount of Applied Sample

The nature of the sample, that is the properties of its components, is the third major factor in the determination of R_f values and is not controllable. The amount of sample does not have too much effect so long as it does not exceed the capacity of the system. Above this point, the R_f increases³⁵² or decreases and tailing or bearding results, depending upon whether the substance has a convex or a concave adsorption isotherm (see p. 82). The worker is urged to look at a given compound or mixture at several concentrations since the capacity of a layer will vary from one substance to another.⁴⁷

The R_f of a given substance may not be exactly the same in a mixture as it is in the pure state.²⁸⁸ If a given compound is suspected to be present in a mixture, this should be confirmed by chromatographing a pure sample against a mixture of the pure compound and the substance under investigation. The appropriate spot should increase in size if the compound is, in fact, a component.

Saturation of the Chamber

Care should be taken to saturate the chromatographic chamber with solvent as completely as possible before use.³⁶⁹ This is normally accomplished by partially lining the walls with filter paper which dips into the solvent and, when possible, by shaking the chamber before the introduction of the chromatogram.⁴⁷ Such saturation (called oversaturation by Stahl) will speed up development by about one-third,²³² will give lower, but more consistent R_f values and will produce straighter solvent fronts and rounder spots.

The so-called "edge effect,"^{42,369} that is, the difference in R_f values between the center and the edge of the plate, can largely be cured by complete chamber saturation if the layers have a fairly consistent thickness.

Multicomponent Developing Systems

There are two factors which should be watched when the developing system consists of several solvents. The first, as noted above, is the evaporation of the more volatile components and the second is the variation in concentration which will occur as the solvent front moves up the layer. To alleviate this latter situation, Brenner, Nieder-

wieser, Pataki and Fahmy⁴⁷ recommended that the ratio of the distance from the solvent level to the sample origin and the total distance traveled by the solvent front be kept constant when precise R_f values are being measured.

Temperature

Variations in temperature have little effect on thin-layer chromatography.⁴⁷ This is in contrast to the situation with paper chromatography where it is quite important. This is understandable where adsorption chromatography is concerned, but less so when the chromatography is partition.

REPORTING R_f VALUES

When reporting R_f values, it is a good idea to report the R_f values of some universally available substance or substances in the system used. The dye system of Stahl (butter yellow-indophenol-Sudan Red G)³⁶⁹ is readily available (from Desaga-Brinkmann) and is satisfactory for certain ranges of adsorption chromatography. As an alternate method, Dhont and de Rooy^{84,85} reported R_B values (sample migration/migration of butter yellow) for 2,4-dinitrophenylhydrazones and 3,5-dinitrobenzoates; Peereboom²⁰³ reports R_A values (sample migration/migration of dibutyl sebacate) for various esters used as plasticizers; and Lisboa and Diczfalussy²²¹ report R_{OE1} values (sample migration/migration of estrone) for steroids.

TAILING

This bothersome phenomenon can normally be attributed to two factors, both of which have been discussed previously. The first, that is, the presence of various ionic species of acids and bases, can be alleviated by the addition of small quantities of acids and bases to the developing system (p. 65). The second factor is sample overloading such that one is working in the nonlinear portions of the isotherms of the substances being separated.* If the isotherms are convex, tail-

* This is discussed in considerable detail by Hagdahl on pages 63 and 64 and by Giddings and Keller on page 96 in Heftmann's book "Chromatography."¹⁴²

ing will occur in concentrations above the linear range. If they are concave, bearding or tails which precede the sample will result. In any case, the situation can be improved by working at lower concentrations.

CHROMATOGRAPHY CHAMBERS

Any wide-mouth jar or bottle which can be completely sealed will serve as a chromatography chamber. The ideal chamber is not too much larger than the plates which are to be developed so that chamber saturation is rapid and, further should have a fairly flat bottom so that the plate will be straight. Special chambers are available from all of the manufacturers of thin-layer equipment.

CHAPTER 7

Visualization

INTRODUCTION

The visualization of developed thin-layer chromatograms can be carried out in many ways. The method used in a specific case will depend upon the type of compounds to be visualized, the nature of the adsorbent and the purpose of the chromatogram. Most visualization methods involve a spray reagent, but other techniques are available.

The inorganic character of most thin layers permits the use of corrosive spray reagents such as sulfuric acid, nitric acid, chromic acid or mixtures of these. These acids will convert almost any organic compound to carbon and thereby serve as universal reagents. The applicability of such reagents constitutes one of the major advantages of thin-layer chromatography. In addition, those specific spray reagents which have been developed for paper chromatography and which do not involve a mechanical washing step can be used.

The majority of spray reagents are destructive in that they destroy or modify the materials which they visualize. In preparative thin-layer work or in quantitative work where the sample is to be eluted before assay, such destructive methods cannot be used. Techniques for surmounting this problem will be discussed in Chapters 9 and 10.

Visualization techniques involving the presence of phosphors or fluorescent substances in thin-layer adsorbents have been developed as well as methods depending upon the relative transparencies of adsorbent and adsorbed compounds when sprayed with the proper reagent. Finally, techniques of radiochemistry and bioassay have been applied to the problem of visualization.

CORROSIVE AND UNIVERSAL SPRAY REAGENTS

With the exception of iodine, these reagents are strong dehydrating and oxidizing systems which function by converting organic compounds to carbon. The reagents are sprayed on the chromatograms which are subsequently heated to about 100°, and the organic compounds appear as black spots. The system is not without flaws, however. When layers containing silver nitrate are sprayed with sulfuric acid,²⁰⁰ a very strong oxidizing system is produced which oxidizes the organic compounds to carbon dioxide. A second difficulty arises when the compound is very unreactive or volatilizes before it chars. In the case of camphor,¹⁹³ this difficulty was solved by spraying the chromatogram with concentrated sulfuric acid containing 5 per cent nitric acid. The sprayed chromatogram was placed, face down, on a glass cloth and heated on a hot plate registering approximately 500°. After the acid fumes ceased to come off, the glass plate was removed and the layer (which had come loose from the plate) was turned over with the glass cloth. The compounds showed as black spots. Such a volatility has been utilized by Baehler,¹⁰ who visualized chromatograms by subliming the resolved compounds out of the adsorbent layer on to a cooled plate held 1 mm above the heated layer (see p. 115).

Iodine, either in a spray or as a vapor, is an exceptionally fine universal reagent and has been used extensively (see Table 7.1). It appears to concentrate itself in the region containing an organic compound,^{219,244} thus producing a brown spot. The normal technique is to spray with 1% solution of iodine in methanol or to place the layer in a closed container with a few crystals of iodine. In either case, the spots fade rapidly and must be marked quickly. Under normal conditions, the reagent does not appear to modify the adsorbed substances.^{17,230,244} If one wishes more permanent spots, the layers can be sprayed with a 0.5 per cent solution of benzidine in absolute ethanol while the iodine spots are still visible.⁶

A number of universal reagents are given in Table 7.1.

SPECIFIC SPRAY REAGENTS

In general, paper chromatographic spray systems which do not require mechanical washing are applicable to thin layers. If, however,

TABLE 7.1. UNIVERSAL SPRAY REAGENTS USED IN THIN-LAYER CHROMATOGRAPHY

Reagent	Composition and Use	Ref.
1. Conc. H_2SO_4	Spray with acid and heat to 100–110°.	193 and many others
50% H_2SO_4	Spray with reagent and heat to 200°. Observe in daylight and U.V.	25, 26, 27, 262, 264, 310, 420
2% H_2SO_4 in H_2O -EtOH (1:1)	Spray with reagent and heat to 100°.	221
H_2SO_4 in acetic anhydride	Spray with H_2SO_4 -acetic anhydride (1:3) and heat.	363, 368, 378
2. H_2SO_4 - KMnO_4	Sol. A: 0.5 g KMnO_4 in 15 ml conc. H_2SO_4 ¹ a. Spray with A. b. Heat.	19, 98
3. H_2SO_4 - $\text{Na}_2\text{Cr}_2\text{O}_7$	Sol. A: 3 g $\text{Na}_2\text{Cr}_2\text{O}_7$ in 20 ml H_2O diluted with 10 ml conc. H_2SO_4 a. Spray with A. b. Heat.	98
H_2SO_4 - $\text{K}_2\text{Cr}_2\text{O}_7$	Spray with sat. sol. of $\text{K}_2\text{Cr}_2\text{O}_7$ in conc. H_2SO_4 and heat.	95
H_2SO_4 -chromic acid	Spray with a sat. sol. of chromic acid in conc. H_2SO_4 and heat.	232
4. H_2SO_4 - HNO_3	Spray with H_2SO_4 - HNO_3 (1:1) and heat. Spray with 5% HNO_3 in H_2SO_4 and heat.	110, 190, 191, 193, 252
5. HClO_4	Spray with 2% (or 25%) sol. and heat to 150°. Allow 70% acid to diffuse into the layer at right angle to developing direction.	247, 296 70
6. I_2	Spray with 1% of I_2 in MeOH. Place in closed chamber containing I_2 crystals.	119, 130, 183, 186, 230, 232, 234, 235, 244, 249, 292, 361, 397, 420
I_2 -benzidine	Spray I_2 treated layer with 0.5% sol. of benzidine in absolute EtOH.	6

¹ Large quantities of KMnO_4 can lead to the formation of explosive manganese heptoxide.

Silica Gel G layers are treated with dichlorodimethylsilane,¹⁸⁵ they are stable enough to be washed. The procedure is to place the layer in a vacuum desiccator with the silane and evacuate it to 300 mm for fifteen minutes. The layer is removed and allowed to stand for

thirty minutes. It is then stable enough to be dipped in aqueous solutions.

A number of specific spray reagents are given in Table 7.2. Table 7.3 is a cross reference for Table 7.2.

Explanation of Table 7.2

Reagent. The reagents are listed alphabetically according to the major component of the reagent or, on occasion, the name of the reagent. The minor components as well as some additional names are given in the cross reference, Table 7.3.

Preparation and Use. When several similar preparations have been used, the author has chosen one to put in the table. However, the references apply to all documented applications of the reagents.

The following abbreviations are used to conserve space:

HOAc	acetic acid	PrOH	<i>n</i> -propanol
NaOAc	sodium acetate	BuOH	<i>n</i> -butanol
EtOAc	ethyl acetate	pet. ether	petroleum ether
EtOH	ethanol	EtOEt	ethyl ether
MeOH	methanol		

Type of Compounds. The types of compounds given cover most, but not all, of the total uses of a given reagent.

Color. The colors are generally taken from a single paper and should not be taken too seriously.

References. The references apply to the main bibliography.

METHODS INVOLVING FLUORESCENCE AND PHOSPHORESCENCE

The preparation of fluorescent and phosphorescent thin layers by the incorporation of chemicals in the adsorbent has been discussed in Chapter 2. Several fluorescent reagents such as 2',7'-dichlorofluorescein (0.2 per cent in ethanol,²³²), fluorescein (0.04 to 0.004 per cent sodium salt in water,^{144,227,366}) Rhodamine B (0.05 per cent in water,¹⁶⁵) and morin (0.05 per cent in methanol³⁴¹) can also be sprayed on thin-layer chromatograms.

In any case, the treated layers cause compounds containing conjugated double bonds to show up as colored spots when viewed in an

TABLE 7.2. SPRAY REAGENTS

	Reagent	Preparation and Use	Type of Compounds	Color	Ref.
1	Acid Violet-7B (6BN)	Sol. A: 1.5% aqueous sol. of reagent. Do not heat over 60° a. Spray with A b. Warm to 100° for 20 min.	cations of the alkali and alkaline earth metals	varies	315, 358
2	Aluminum chloride	Sol. A: 5% alcoholic sol. of reagent a. Spray with A b. Observe in daylight and U.V.	flavonoids	yellow spots	286
3	Ammonia vapor	Expose to vapor.	flavonol glycosides		126
4	Ammonium molybdate-HClO ₄ Ammonium molybdate-SnCl ₂	Sol. A: 3 g of ammonium molybdate in 25 ml H ₂ O, 30 ml of 0.1 N HCl and 15 ml 60% HClO ₄ a. Spray with A b. Heat to 110° for 15 min. Sol. A: 1% aqueous sol. of ammonium molybdate Sol. B: 1% SnCl ₂ in 10% HCl a. Spray with A b. Dry c. Spray with B d. Heat to 105°	universal for lipids phosphates (inorganic)	blue-black spots blue spots	157 355
5	Aniline phthalate	Sol. A: 0.93 g aniline and 1.66 g phthalic acid in 100 ml of 1-BuOH saturated with H ₂ O ⁽⁷²⁾ a. Spray with A b. Heat to 105° for 10 min.	reducing sugars	various colors	439
6	Anisaldehyde in H ₂ SO ₄ and HOAc	Sol. A: 0.5 ml of reagent in 0.5 ml conc. H ₂ SO ₄ , 9 ml of 95% EtOH and a few drops of HOAc a. Spray with A b. Heat to 105° for 25 min. Sol. A: 0.5 ml of reagent in 50 ml of HOAc and 1 ml conc. H ₂ SO ₄ a. Spray with A b. Heat to 120° for 6 min.	carbohydrates steroids	various blues	377, 441 247

7	<i>p</i> -Anisidine			honeybee larval foods		292
8	<i>p</i> -Anisidine phthalate		Spray with an alcoholic sol. 0.1 M in both <i>p</i> -anisidine and phthalic acid	reducing sugars	various colors	350
9	Antimony pentachloride in CCl_4		Sol. A: 20% (also 12 and 16) sol. of reagent in CCl_4 a. Spray with A b. Heat to 120° c. Observe in daylight and U.V.	terpenoids, steroids and essential oils	various colors	117, 171, 286, 299, 368, 394
10	Antimony trichloride in CHCl_3		Sol. A: sat. sol. of reagent in alcohol-free CHCl_3 a. Spray with A b. Heat to 100° for 10 min. c. Observe in daylight and U.V.	steroids, steroid glycosides, aliphatic lipids, vitamin A and others	various colors	13, 21, 67, 68, 128, 129, 141, 145, 171, 223, 247, 257, 287, 368, 371, 382, 392, 415, 416, 457, 460
	Antimony trichloride in CHCl_3 with acetic anhydride		Sol. A: sat. sol. of reagent in alcohol-free CHCl_3 -acetic anhydride (5:1) a. Spray with A b. Heat to 130° for 5-10 min. c. Observe in daylight and U.V.	steroids and steroid glycosides		322
	Antimony trichloride in HOAc		Sol. A: equal parts, by weight, of SHCl_3 and HOAc a. Spray with A b. Heat to 95° for 5 min. ³⁷⁵	steroids		348
	Antimony trichloride in CHCl_3 with SOCl_2		Spray with a sat. sol. of reagent in alcohol-free CHCl_3 - SOCl_2 (10:1)	Δ^4 -steroids		145
11	<i>Aqua regia</i> -ammonium molybdate-vanadyl chloride		Sol. A: <i>aqua regia</i> (3 parts conc. HCl and 1 part conc. HNO_3) Sol. B: 3 ml of 5% ammonium molybdate and 7 ml 5 N HCl diluted to 100 ml with acetone Sol. C: 2 g V_2O_5 and 20 ml conc. HCl boiled until blue green, then diluted to 37 ml with 6 N HCl and to 400 ml with H_2O ; 20 ml of this is then shaken with zinc powder until it becomes brown; decant and use immediately a. Spray with A, B and C consecutively	phosphorus in phosphates	blue spots	174

TABLE 7.2—Continued

	Reagent	Preparation and Use	Type of Compounds	Color	Ref.
12	Benzidine, diazotized	Sol. A: 0.1 to 0.085 <i>M</i> NaHPO ₄ Sol. B: 0.18 g benzidine in 50 ml 0.5 <i>N</i> HCl Sol. C: 2 ml of B diluted with 2 ml 1% NaNO ₂ , allowed to lose its color (3–5 min), treated with 2 ml of 5% urea and finally diluted with 14 ml H ₂ O a. Spray with A b. Spray with C ₁₀₀	phenols which couple		288, 290, 291
13	Bial's reagent	Sol. A: 40.7 ml conc. HCl mixed with 0.1 g orcinol, 1 ml of 1% FeCl ₃ and diluted with H ₂ O to 50 ml a. Place in HCl atmosphere at 80° for 1.5 hr. b. Spray with A c. Place in HCl atmosphere at 80° until reaction occurs	glycolipids	violet	127, 197, 427, 437
14	Boute reaction	a. Expose to NH ₃ vapor b. Expose to nitrogen dioxide (from Cu and HNO ₃)	estrogens with phenol groups	yellow	221
15	Bromcresol green	Spray with a 0.3% sol. of reagent in H ₂ O-MeOH (20:80) containing 8 drops of 30% NaOH per 100 ml.	carboxylic acids	yellow spots on green	42, 105, 106, 107, 193
16	Bromcresol purple	Spray with a 0.1% sol. of reagent in EtOH made just basic with NH ₄ OH or NaOH	a. halogen ions, except F ⁻ b. dicarboxylic acids	yellow spots on purple	357, 200, 201
17	Bromphenol blue-citric acid	Sol. A: 0.5 g of bromphenol blue and 0.2 g of citric acid in 100 ml H ₂ O a. Spray with A	aliphatic carboxylic acids	yellow spots on blue	302
18	Bronthymol blue	I. <i>alkaline</i> Spray with a sol. of 40 mg of reagent in 0.01 <i>N</i> NaOH (100 ml) II. <i>buffered</i> Spray with a sol. of 50 mg of reagent and 1.25 g of boric acid in 8 ml of 1 <i>N</i> NaOH and 112 ml H ₂ O	universal for lipids	yellow spots on blue	134, 173, 174, 208, 459

19	Ceric sulfate-trichloroacetic acid	Sol. A: 0.1 g $\text{Ce}(\text{SO}_4)_2$ suspended in H_2O ; add 1 g of trichloroacetic acid; boil and add, drop by drop, 1.84 ml of conc. H_2SO_4 until it becomes clear a. Spray with A b. Heat to 110°C		various colors	354
20	Chloramine T- $\text{HCl}\cdot\text{NH}_3$	Sol. A: 10% reagent in water Sol. B: 1 N HCl a. Spray with A b. Spray with B c. Heat to drive off Cl_2 d. Expose to NH_3 ³⁵⁵	caffeine	rose red	117
21	Chlorine-KI-starch Chlorine-toluidine-KI	a. Expose to Cl_2 for 0.5 hr. b. Drive off excess Cl_2 by waving in air c. Spray with sol. of KI and starch Sol. A: 160 mg toluidine in 30 ml HOAc ; add 1 g KI and make up to 500 ml. a. Expose to Cl_2 for 5-10 min. b. Allow to stand in air 3-5 min. c. Spray with A^{356}	esters of carbohenzoxy-amino acids 3-phenyl-2-thiohydantoins of amino acids	blue spots	456 46
22	Chromotropic acid in H_2SO_4	Sol. A: 10% aqueous sodium 1,8-dihydroxynaphthalene-3,6-disulfonate-23 N H_2SO_4 (1:5) a. Spray with A b. Heat to 105° for 30 min.	methylene-dioxy group	blue	32
23	Chlorosulfonic acid- HOAc	Sol. A: $\text{ClSO}_3\text{H}\cdot\text{HOAc}$ (1:2) a. Spray with A b. Heat to 130° for 5 min. c. Observe in U.V.	triterpenes and cardiac aglycones	violet-brown	405, 406
24	Copper acetate-dithiooxamide on stabilized layers	Sol. A: 20 ml of sat. aqueous copper acetate diluted to 1000 ml Sol. B: 0.1% dithiooxamide in EtOH a. Dip layers in H_2O b. Dip layers in A for 10 min. c. Wash with H_2O for 30 min. d. Dip in B for 10 min. e. Wash in H_2O	fatty acids	green spots on white	180

TABLE 7.2—Continued

	Reagent	Preparation and Use	Type of Compounds	Color	Ref.
25	Cyanogen bromide-benzidine	Sol. A: 0.05% benzidine in 2 N HOAc a. Expose to CNBr vapor b. Spray with A	nicotinic acid derivatives	red to brown	117
26	α -Cyclodextrin- I_2	Sol. A: 1.0% sol. of reagent in H_2O -EtOH (70:30) a. Spray with A b. Expose to I_2	monochain lipids	white spots on purple	230, 231, 232
27	<i>o</i> -Dianisidine	Spray with a 1% sol. of reagent in HOAc	aldehydes and ketones	yellow-brown	193, 252, 328, 368
28	Dibromo-R-fluorescein	Sol. A: 0.2% sol. of reagent in EtOH a. Spray with A b. Observe in U.V.	glycerides	yellow fluorescent	16
29	2',7'-Dichlorofluorescein	Sol. A: 0.2% sol. of reagent in 96% EtOH a. Spray with A b. Observe in U.V.	nonpolar lipids, saturated and unsaturated	green spots on purple	8, 28, 120, 230, 232, 234, 235, 236, 264, 333, 334, 364, 420, 423
30	2,6-Dichloroquinone-chlorimide- NH_3 2,6-Dichloroquinone-chlorimide-borax	Sol. A: 0.1% sol. of reagent in EtOH a. Spray with A b. Expose to NH_3 Sol. A: 2% quinone in EtOH Sol. B: 2% borax in water a. Spray with A and B consecutively	Vitamin B ₂ phosphate esters	blue spots blue spots	116, 117 293
31	N,N-Dimethyl- <i>p</i> -phenylenediamine-2HCl-NaOMe-irradiation	Sol. A: 0.5 g of diamine in 100 ml MeOH with 1 g Na a. Spray with A b. Moisten with H_2O c. Expose to U.V. without filter	chlorinated hydrocarbons	green-violet spots	23
32	<i>p</i> -Dimethylaminobenzaldehyde in HCl (van Urk reagent) <i>p</i> -Dimethylaminobenzaldehyde in H_2SO_4	Sol. A: 0.1 to 1% reagent in 50 ml of conc. HCl and 50 ml 96% EtOH a. Spray heavily with reagent b. Spray lightly with <i>aqua regia</i> (for alkaloids) Spray with a 1% sol. of reagent in 2% sol. of H_2SO_4 in EtOH	sulfonamides alkaloids (indole) a. essential oils b. morphine	blue rose spots	452 376 287, 437

33	2, 4-Dinitrophenylhydrazine (2, 4-DNPH)	Spray with a 0.5% sol. of reagent in 2 N HCl	aldehydes and ketones	yellow to red spots	2, 13, 105, 106, 107, 165, 171, 192, 287, 323, 326, 328, 368, 376
33a.	DPNH-(reduced diphosphopyridine nucleotide)		enzymes		82
34	Diphenylamine	Sol. A: 1% sol. of reagent in EtOH a. Spray with A b. Expose to U.V.	nitramine explosives	blue spots	216
	Diphenylamine in HCl and HOAc	Sol. A: 10% sol. of amine in EtOH a. Dilute 20 ml of A with 100 ml conc. HCl and 80 ml HOAc b. Spray with diluted A c. Cover with another plate and heat to 110° for 30-40 min.	glycolipids	blue	157, 174
35	Diphenylboric acid- β -amino-ethyl ester	Sol. A: 1% of reagent in MeOH a. Spray with A b. Observe in U.V.	flavanols, coumarins and their glycosides	varies	379
36	s-Diphenylcarbazone	Spray with a 0.1% sol. of reagent in 95% EtOH	acetoxymethoxymercurey adducts of unsaturated lipids	purple to blue spots on rose	177, 232, 233
36a.	Diphenylthiocarbazone	Sol. A: 0.01% of reagent in CHCl ₃ a. Spray with A	organotin derivatives	orange-red	411
37	Dragendorff's reagent	I. Sol. A: 1.7 g of basic bismuth nitrate in 100 ml of H ₂ O-HOAc (80:20) Sol. B: 40 g KI in 100 ml H ₂ O a. Spray with sol. made from 5 ml of A, 5 ml of B, 20 g of acetic acid and 70 ml H ₂ O II. Sol. A: Warm 2.6 g of bismuth carbonate and 7 g of NaI in 25 ml of HOAc. After 12 hr, filter sol. and dilute 20 ml of the filtrate with 8 ml EtOAc Sol. B: 10 ml of A diluted with 25 ml HOAc and 60 ml EtOAc a. Spray with B b. Colors can be enhanced by spraying with 0.1 N H ₂ SO ₄ 's	alkaloids and choline containing compounds	orange	88, 157, 188, 227, 273, 332, 344, 427, 430, 444

TABLE 7.2—Continued

	Reagent	Preparation and Use	Type of Compounds	Color	Ref.
38	Ehrlich reagent	Sol. A: 1% sol. of <i>p</i> -dimethylaminobenzaldehyde in 95% EtOH a. Spray with A b. Place in HCl atmosphere for 3–5 min.	a. tryptophan metabolites b. gangliosides	various colors	87, 208
39	a. Echt blue salt B (tetraazotized di- <i>o</i> -anisidine) b. Echt blue salt BB (diazotized 1-amino-4-benzoylamido-2,5-diethoxybenzene) c. Echt red salt B (diazotized 5-nitro-2-aminoanisole)	Sol. A: 5% aqueous sol. of reagent a. Spray with A b. Spray with 0.1 <i>N</i> NaOH	phenols and amines, which couple		283
40	Ferric chloride Ferric chloride in HOAc Ferric chloride-K ₃ Fe(CN) ₆	Spray with a 1% aqueous sol. of reagent Sol. A: Equal volumes of 5% aqueous FeCl ₃ and 2 <i>N</i> HOAc a. Spray with A Sol. A: 1 volume 5% aqueous K ₃ Fe(CN) ₆ , 2 volumes 10% aqueous FeCl ₃ and 8 volumes H ₂ O a. Spray with A	phenols pyrazolones phenacetin phenols chlorinated hydrocarbons peroxides	various colors blue	287, 292 227 117 98 43
41	Ferrous ammonium sulfate-potassium thiocyanate	Spray with a mixture of the two aqueous solutions			242
42	Fluorescein Fluorescein-Br ₂	Sol. A: 0.004% aqueous sol. of sodium fluorescein a. Spray with A b. Observe in U.V. Sol. A: 0.04% aqueous sol. of sodium fluorescein a. Spray with A b. Observe in U.V. for conjugated systems c. Expose to Br ₂ d. Observe in U.V. for unsaturates	conjugated double-bond systems unsaturated compounds	yellow spots on pink	82, 227 105, 106, 107, 171, 190, 191, 192, 193, 232, 252, 328, 365, 368

43	Folin-Ciocalteu reagent	Sol. A: Mix 10 g sodium tungstate, 2.5 g of sodium molybdate and 70 ml H ₂ O. Add, consecutively, 5 ml 85% H ₃ PO ₄ and 10 ml conc. HCl. Reflux for 10 hr and add 15 g Li ₂ SO ₄ , 5 ml H ₂ O and 1 dp Br ₂ . Dilute to 100 ml with H ₂ O. Reagent should not be green. Sol. B: 20% aqueous Na ₂ CO ₃ a. Spray with B b. Dilute A with 3 volumes H ₂ O and spray ^{77a}	a. phenols b. estrogens	163, 221, 380
44	Fuchsin-H ₂ SO ₄	Sol. A: 1 g parafuchsin is dissolved in 700 ml H ₂ O and 50 ml 2 N HCl, diluted to 1 l with H ₂ O and allowed to stand overnight with occasional shaking a. Add 1% HgCl ₂ to A b. Spray with reagent ^{78a}	plasmalogens	134
45	Hydrazine sulfate in HCl	Spray with a 1% sol. of reagent in 1 N HCl	aromatic aldehydes	396
46	Hydrogen chloride	Hold over conc. HCl	chalcones	384
47	Hydroxamic acid-ferric ion	Sol. A: equal volumes of 12.5% NaOH and 5% hydroxylamine, both in H ₂ O Sol. B: acetic acid Sol. C: 10% FeCl ₃ in H ₂ O a. Spray with A b. Allow to stand for 10 min (or heat to 100° for 10 min) c. Spray with B and C consecutively	esters of all types	134, 207
48	8-Hydroxyquinoline-NH ₃	Sol. A: 0.5% sol. of reagent in 60% EtOH a. Expose to NH ₃ b. Spray with A c. Observe in U.V.	inorganic cations	359, 360

TABLE 7.2—Continued

	Reagent	Preparation and Use	Type of Compounds	Color	Ref.
49	Iodine-potassium iodoplatinate	Sol. A: 0.5% sol. of I_2 in $CHCl_3$ Sol. B: equal volumes of 1.1% aqueous KI and 0.135% aqueous chloroplatinic acid a. Spray with A and B alternating	nitrogen-containing drugs		279
	Iodine in KI and HOAc	Sol. A: aqueous sol. which contains 5% I_2 and 10% KI a. Dilute 2 ml of A with 3 ml H_2O and 5 ml 2 N HOAc and spray with diluted sol.	alkaloids and steroids		347, 227
	Iodine-benzidine	Sol. A: 5% benzidine in EtOH a. Expose to I_2 vapor b. Spray with A	glycerides	blue	6
50	Iodine azide	Sol. A: 3.5 g of sodium azide in 100 ml of 0.1 N I_2 (reagent is explosive when dry ¹¹⁹) a. Spray with A	penicillins	yellow	103
51	Isatin in H_2SO_4	Sol. A: 0.4% sol. of reagent in conc. H_2SO_4 a. Spray with A b. Heat to 120°	thiophene derivatives	various colors	65
52	Isonicotinic acid hydrazide	Sol. A: 4 g isonicotinic acid hydrazide and 5 ml conc. HCl in 1 l H_2O a. Spray with A b. Allow to stand 3–16 hr. c. Observe in U.V. and daylight	steroids		414
53	Kedde reagent	Sol. A: 2% sol. of 3,5-dinitrobenzoic acid in MeOH Sol. B: 2 N KOH a. Mix equal volumes of A and B just before spraying b. Spray with mixture	cardiac glycosides		223

54	Morin (2',3,4',5,7-pentahydroxyflavone)	Sol. A: 0.05 to 0.005 % sol. of reagent in MeOH a. Spray with A b. Heat to 100° for 2 min. (optional) c. Observe in U.V.	a. N-blocked amino acids b. conjugated double-bond systems	yellow	32, 341
55	N-(1-Naphthyl)ethylene diamine-2 HCl (diazotized)	Sol. A: 1 N HCl Sol. B: 5 % NaNO ₂ Sol. C: 0.1 % reagent in EtOH a. Spray with A, B, and C (consecutively)	sulfonamides	reddish purple	196
56	Naphthoresorcinol in H ₃ PO ₄	Sol. A: 100 ml of 0.2 % sol. of reagent in EtOH and 10 ml H ₃ PO ₄ a. Spray with A b. Heat to 105° for 5-10 min.	carbohydrates	various colors depending upon type	120, 289, 307, 308
57	Ninhydrin	I. Sol. A: 95 ml of 0.2 % reagent in BuOH plus 5 ml of 10 % aqueous HOAc a. Spray with A b. Heat to 120-150° for 10-15 min. II. Sol. A: 50 ml of 0.2 % sol. of reagent in abs. EtOH mixed with 10 ml of HOAc and 2 ml of 2,4,6-collidine Sol. B: 0.1 % Cu(NO ₃) ₂ ·3H ₂ O in abs. EtOH a. Spray with A-B (50:3) mixed just before use b. Heat to 105°	a. amino acids b. aminophosphatides c. amino sugars	blue	44, 95, 99, 116, 117, 134, 156, 292, 424, 438, 450
58	Osmium tetroxide	Expose to reagent vapors (caution, poison) in closed chamber	cmpds. with double bonds	brown to black	459
59	Palladium chloride in HCl	Spray with a 0.5 % sol. of reagent in dilute HCl	thiophosphate insecticides	yellow spots on brown	23
60	Periodic acid-HClO ₄ -V ₂ O ₅	Sol. A: 10 g H ₂ O ₈ in 100 ml of 70 % HClO ₄ and a few crystals of V ₂ O ₅ a. Spray carefully dried layer with A Sol. A: 0.5 % HIO ₄ in 90 % HOAc Sol. B: 30 % sol. of sodium pyrosulfate-3 N HCl (1:1) (both kept at 0°) Sol. C: 5 ml of 10 % sodium pyrosulfite, 85 ml H ₂ O and 0.2 g fuchsin. Allow sol. to stand 12 hr, treat with carbon and filter a. Spray with A, B, and C consecutively b. Heat to 90° for 15 min.	thiophosphate esters glycerides	blue violet	304 174

TABLE 7.2—Continued

	Reagent	Preparation and Use	Type of Compounds	Color	Ref.
61	Phosphomolybdic acid	Sol. A: 1.5% (also 5 and 10%) sol. of reagent in EtOH or EtOH-EtOAc (1:1) a. Spray with A b. Heat to 150° for 5 min.	a. bile acids b. glycerides c. unsaturated fatty acids d. many others	blue	13, 115, 152, 154, 155, 157, 179, 181, 182, 183, 184, 186, 238, 249, 293, 295, 362
	Phosphomolybdic acid-SnCl ₄ on stabilized layers	Sol. A: 1% sol. of acid in CHCl ₃ -EtOH (1:1) Sol. B: 1% SnCl ₄ in 2 N HCl a. Spray with A b. Dry and dip in B	phospholipids	blue	185
62	Phosphomolybdic acid-NH ₃	Sol. A: 20% sol. of acid in EtOH a. Spray with A b. Dry for 2-3 min and expose to NH ₃	tocopherols	blue	354
63	Phosphotungstic acid	Sol. A: 10% sol. of acid in EtOH a. Spray with A b. Heat to 100° for 5-10 min.	cholesterol esters		249
64	Phosphoric acid	a. Spray with 50-70% acid b. Heat to 100° for 10-30 min. c. Observe in U.V.	pregnantriols and other steroids		247, 391, 414
65	ortho-Phosphoric acid-phosphomolybdic acid	Sol. A: 40% ortho-phosphoric acid Sol. B: 5% sol. of phosphomolybdic acid a. Spray with A b. Heat to 110° for 7 min. c. Observe in U.V. d. Spray with B e. Heat to 110° for 2 min.	steroids	blue-green	433, 434, 435

66	Potassium hydroxide	Sol. A: 5% KOH-acetone (2:1) a. Spray with A b. Heat to 100° c. Cool and spray (optional) with HNO ₃ -H ₂ SO ₄ (1:1)	aromatic nitro compounds and amines	110
	Potassium hydroxide-urea	Sol. A: 1 N KOH in EtOH Sol. B: 50% aqueous sol. of urea a. Spray with A b. Heat to 60° for 15 min. (citrates are yellow) c. Spray with B d. Observe in U.V.	citrates	293
67	Potassium iodide	Spray with saturated aqueous sol.	peroxides	242
	Potassium iodide-NH ₄ -H ₂ S	Sol. A: 2% sol. of KI in H ₂ O a. Spray with A b. Observe colors c. Expose to NH ₃ d. Expose to H ₂ S	copper series of inorganic cations	359
	Potassium iodide-HOAc-Na ₂ SO ₃ -starch	Sol. A: fresh sol. of 5 ml of 4% KI, 20 ml HOAc and a few drops of 1% Na ₂ SO ₃ Sol. B: 1% aqueous sol. of starch a. Spray with A and B consecutively	peroxides	371
68	Potassium iodoplatinate	Sol. A: mix 45 ml of 10% KI with 5 ml of 5% platinum chloride and dilute to 100 ml a. Spray with A	alkaloids and vitamins	57, 116, 117, 279, 285, 436
69	Potassium permanganate	Spray with 0.25% aqueous sol. of reagent	diterpenoids	213
	Potassium permanganate (alkaline)	Sol. A: equal portions of 1% aqueous KMnO ₄ and 2% aqueous Na ₂ CO ₃ a. Spray with A b. Note color after 1-2 hr.	phenols and plasticizers (esters)	290, 293
	Potassium permanganate-HOAc	Spray with mixture of equal volumes of 0.1 N KMnO ₄ and 2 N HOAc	oxidizable materials	227
	Potassium permanganate-Fe(NO ₃) ₃	Sol. A: 0.5% aqueous sol. of KMnO ₄ Sol. B: 1% aqueous sol. of Fe(NO ₃) ₃ a. Spray with A and B consecutively	phenols	340

TABLE 7.2—Continued

	Reagent	Preparation and Use	Type of Compounds	Color	Ref.
70	Procházká reagent	Sol. A: 10 ml of 35 % formaldehyde, 10 ml of 25 % HCl and 20 ml 96 % EtOH a. Spray with A b. Heat to 100° c. Observe in daylight and U.V.	indole derivatives	yellow	376
71	Pyridyl-azo-naphthol	Spray with a 0.25 % sol. of reagent in EtOH	uranium ion		360
72	Resorcinol-SnCl ₂ -H ₂ SO ₄ -KOH	Sol. A: 20 % resorcinol in EtOH with added SnCl ₂ Sol. B: 4 N H ₂ SO ₄ Sol. C: 40 % aqueous KOH a. Spray with A b. Heat to 150° for 10 min. c. Spray with B d. Heat to 120° for 20 min. e. Spray with C	esters and plasticizers	various colors	293
73	Rhodamine B	Spray with a 0.05 % (or 0.5 %) sol. of reagent in EtOH	lipids	purple spots on rose	19, 165, 181, 182, 183, 232, 427, 430, 448
	Rhodamine B-KOH	Sol. A: 0.05 % aqueous sol. of reagent Sol. B: 10 N KOH a. Spray with A and B consecutively	glycerides	bright spots on rose	6
74	Rhodamine 6 G	Sol. A: 8 % aqueous sol. of reagent a. Spray with A b. Observe in U.V.	lipids		459
75	Silver nitrate-NH ₄ -NaOMe	Sol. A: 0.3 % AgNO ₃ in MeOH Sol. B: MeOH saturated with NH ₃ Sol. C: 7 % Na in MeOH a. Mix 5 parts of A, 1 part of B and 2 parts of C b. Spray with mixture c. Note color after few min. as	acyl sugar	brown	75
76	Silver nitrate-NH ₄ OH-fluorescein	Sol. A: 1 % sol. of AgNO ₃ made basic with NH ₄ OH Sol. B: 0.1 % sol. fluorescein in EtOH a. Spray with A and B consecutively	halogen ions		357

77	Sodium periodate	Spray with a 1% aqueous sol. of reagent	ferrocenes	blue	345
78	Sodium periodate-KMnO ₄	Sol. A: 4 parts of 2% aqueous NaIO ₄ and 1 part of 1% KMnO ₄ in 2% aqueous Na ₂ CO ₃ a. Spray with A ₂₅ a	sugars		454
	Sodium periodate-ethylene glycol-thiobarbituric acid	Sol. A: 0.02 M NaIO ₄ Sol. B: ethylene glycol-acetone-H ₂ SO ₄ (50:50:0.3) Sol. C: 6% aqueous sol. of 2-thiobarbituric acid sodium salt a. Spray with A, B, and C consecutively. ^{asb}	neuraminic acid sialic acid derivatives	red	127
79	Stannous chloride-HCl	Spray with a 10% sol. of SnCl ₂ in conc. HCl	aromatic nitro compounds		110
80	Sudan Black B on stabilized layers	Sol. A: 0.1% sol. of reagent in EtOH-H ₂ O (1:1) a. Dip in A b. Wash with EtOH-H ₂ O (1:1).	saturated glycerides	blue	180
81	Sulfanilic acid, diazotized (Pauly's reagent)	Sol. A: 0.5 g of sulfanilic acid and 0.5 g of KNO ₂ in 100 ml of 1 N HCl I. a. Spray with A b. Spray with 1 N NaOH II. a. Spray with conc. KOH in MeOH b. Spray with A	phenols and aromatic amines which will couple	yellow-orange	11, 72, 74, 82, 396, 426
82	Tetrazolium blue or 3,3'-(3,3'-dimethoxy-4,4'-biphenylene)-bis-[2,5-diphenyl-2H-tetrazolium chloride]	I. Reducing steroids Sol. A: 1% sol. in EtOH Sol. B: Mix just before using 2 ml of A, 5 ml of H ₂ O and 3 ml of 4 N NaOH a. Spray with A II. Δ ⁴ -3-oxosteroids Sol. A: 2.7 N NaOH and 0.1 ml of sol. A above a. Spray with A ₁₅ a	corticoids		1, 243
83	2,3,5-triphenyltetrazolium chloride-NaOH	Sol. A: Mix just before use one volume of 4% sol. of reagent in MeOH and one volume of 1 N NaOH a. Spray with A b. Heat to 110° for 5-10 min. ⁷⁵	a. steroids b. reducing sugars c. glycosides	red spots	30, 247, 414

TABLE 7.2—Continued

	Reagent	Preparation and Use	Type of Compounds	Color	Ref.
84	Trichloroacetic acid	Sol. A: 4% sol. of reagent in CHCl_3 a. Spray with A b. Allow to stand	menthofuran	pink	19
	Trichloroacetic aminic-T	Sol. A: Mix just before use 8 volumes of a 25% sol. of acid in EtOH with 2 volumes of 3% chloramine-T sol. a. Spray with A b. Heat to 100° for 10 min. c. Observe in daylight and U.V.	K-strophanthin (in U.V.)	reddish-brown	223
			cardenolides		90, 223, 322
85	Trinitrobenzene		azulenes		108
86	Uranium acetate	Spray with a 1% sol. of reagent in H_2O	flavonoids	brown	286
87	Vanillin- H_2SO_4	Sol. A: 1% sol. of reagent in conc. H_2SO_4 a. Spray with A. b. Observe in daylight and U.V.	terpenoids	various colors	49, 167, 168, 368
	Vanillin- H_3PO_4	Sol. A: 1% sol. of reagent in 50% H_3PO_4 a. Spray with A b. Heat to 120° for 10-20 min.	steroids		247, 336, 338
88	Water on Silica Gel G	a. Spray thoroughly b. Observe in reflected light	alkaloids, steroids, sugar acetates and others	opaque areas on semi-transparent background	115, 128, 398
89	Zwicker's reagent	Sol. A: 0.5% sol. of cobalt acetate in MeOH Sol. B: 0.5% sol. of lithium hydroxide in MeOH a. Spray with A and B consecutively	barbiturates		227

TABLE 7.3. CROSS REFERENCES TO SPECIFIC SPRAY REAGENTS

Reagent or Name	Item in 7.2	Reagent or Name	Item in 7.2
Acetic anhydride	10	Iodine	26
Benzidine	25	Lithium hydroxide	89
Bismuth nitrate	37	Pauly's reagent	81
Bismuth carbonate	37	Perchloric acid	4
Borax	30	Potassium ferricyanide	40
Bromine	42	Potassium iodide	21
Chloramine-T	84	Potassium iodoplatinate	49
Cobalt acetate	89	Potassium permanganate	78
2,4,6-Collidine	57	Potassium thiocyanate	41
<i>p</i> -Dimethylaminobenzaldehyde	38	Pyruvic acid	33a
3,5-Dinitrobenzoic acid	53	Schiff's reagent	60
Diphospho-pyridine nucleotide (reduced)	33a	Sodium molybdate	43
Dithiooxamide	24	Sodium tungstate	43
Dithizone	36a	Stannous chloride	4, 72
Ferric chloride	47	Starch	21
Fluorescein	76	Thiobarbituric acid	78
Formaldehyde	70	Thionyl chloride	10
Hydrogen sulfide	67	Toluidine	21
Hydroxylamine-HCl	47	Trichloroacetic acid	19
		Urea	66
		van Urk's reagent	32

ultraviolet light.²³² If fluorescein sprayed layers are then treated with bromine vapor, the fluorescein is converted to eosin except where unsaturated compounds were present. The result is that the unsaturated compound shows up as a yellow spot on a pink background. It is best viewed in ultraviolet light.^{227,232,366}

TRANSPARENCY METHODS

If a developed layer is sprayed with or dipped into a liquid of the correct refractive index, the adsorbent and the compounds to be visualized will have relatively different transparencies and can be distinguished when held in a correct angle with respect to light. In this manner, spots or bands can be located without destruction and with great accuracy. The most commonly used system of this type is water on silica gel,^{115,128,398} in which the bands or spots stand out as semiopaque areas on a translucent background. The method would obviously fail if the adsorbed materials were water soluble.

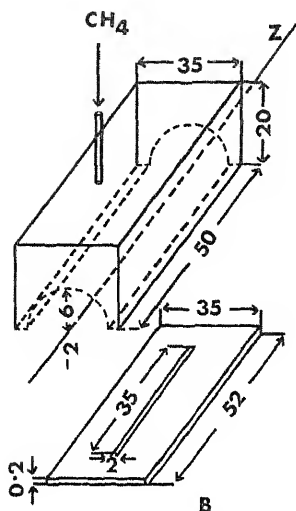


Figure 7.1. Gas flow counter with an extremely flat aperture plate (B) for measurement of weak β -emitters on thin-layer chromatograms. The aperture plate is attached to the counter tube without projecting parts. Z = counter wire ($30\ \mu$ diameter). The measurements are given in mm. (Reproduced from Schulze and Wenzel, *Angew. Chem. Intern. Ed. Engl.*, 1, 580 (1962) through the courtesy of the authors and the publishers, Verlag Chemie and Academic Press.)

In a variation of this technique, Kaufmann, Makus and Khoe¹⁸⁵ dipped Silica Gel G layers (which had been stabilized with dichlorodimethylsilane) in a mixture of ethanol-water (2:8) so that triglyceride spots became transparent. Hefendehl¹⁴¹ sprayed already visualized Silica Gel G layers with a paraffin-ether mixture (1:1). The paraffin-impregnated layers became transparent after being stored in a desiccator for a few hours and the colored spots were easier to photograph.

RADIOACTIVITY METHODS

Chromatograms of radioactive compounds can be visualized by making X-ray prints.^{20,232,234a} The layers are directly exposed to "No-Screen Medical X-Ray Safety Film"* for times varying from a few hours to a week. The films are developed with "Supermix Developer"*** for 4 to 6 min. and fixed for 30 min. with "Acid Fixer." These methods have been applied mostly to lipids, although radioactive halides have also been detected by radioactivity.²⁰

Schulze and Wenzel³⁴⁹ developed a method for continuous auto-

* Eastman Kodak Company, Rochester 3, N.Y.

** General Electric X-Ray Dept., Milwaukee 1, Wisc.

matic measurement of radioactivity distribution of weak β -emitters on thin layers. Accurate measurements of C^{14} and tritium-containing compounds were made by passing the developed layer slowly under the aperture of the especially designed gas-flow counter shown in Figure 7.1.* The gas was methane. The method was also used for locating radioactive materials in preparative thin layer chromatography.

BIOASSAY METHODS

The developed chromatogram is covered with an agar plate containing an appropriate microorganism in good health. The spots which contain antibiotic inhibit the microorganisms and can thereby be detected. Thus, Nicolaus, Coronelli and Binaghi^{275,276} located a number of penicillins, tetracyclines and rifomycins using *Sarcina lutea* and *Bacillus subtilis* in the presence of triphenyltetrazolium salts. The sensitivity was 0.01 to 0.1 γ . The macrolides have been determined³³ using *Bacillus subtilis* as the revealing organism.

* A commercial model is available from Laboratorium Prof. Dr. Berthold, Wildbad (Schwarzwald) Germany.

CHAPTER 8

Documentation of Results

INTRODUCTION

Thin-layer chromatograms are normally too fragile for preservation as such and several techniques have been devised for making a permanent record of them. In general, these are to trace them or draw them in a laboratory book, to photograph or photocopy them, or to imbed them in a paraffin layer or plastic film of some type. The latter two, more exact, documentation methods offer the additional advantage that the photograph or films can often be quantitized.

PHOTOGRAPHY

The most positive way to record thin-layer chromatograms is to photograph them.^{46,137,141,223,262,273,323,352,416} Chromatograms which are visible in normal light can be photographed with any available camera system. Hefendehl¹⁴¹ found that better pictures resulted when the chromatograms were made semitransparent after normal visualization procedures (see p. 104).

Thin-layer chromatograms can also be copied with an office letter-copying device such as the "Zerox 914,"^{*148} "Photorapid"^{**121} or the "Copease" apparatus.^{†232} A process using a direct positive blueprinting paper, Driprint HC (F speed)[‡] has been suggested.⁹⁶ In some

* Manufactured by Zerox Corp., Rochester 3, N.Y.

** Manufactured by Bürogeräte, A. G., Zürich, Switzerland, Pronto model with Geva-copy paper GS, 8½ × 11 in.

† Manufactured by Copease Corp., 425 Park Avenue, New York, N.Y.

‡ E. Dietzgen Co., 407 10th Street, N.W., Washington, D.C.

cases^{96,121} the layer is covered with cellophane film before photography to protect both the layer and the photocopying apparatus. It has also been suggested¹⁴⁸ that better copies will result if the spots are outlined with a stylus or even completely removed before duplication.

Sylvania F8T5/BLB, "Blacklite" blue tubes* and a yellow filter are recommended by Bürke and Bolliger²³² for black and white photography in ultraviolet light. Lukas has made photographs in ultraviolet light using an ultraviolet filter,²²³ and Hansbury, Langham and Ott¹³⁷ have devised a method using a "Polaroid" system for the same purpose.

The use of X-ray film for recording chromatograms of radioactive substances was discussed in Chapter 7 (p. 104).

METHODS OF LAYER PRESERVATION

Meinhard and Hall²⁴⁶ pressed a piece of "Scotch" or cellophane tape on thin layers to remove a portion of the layer. The tape was then placed directly in a laboratory book. The applicability of this technique is governed by the visualization procedure and the spray reagents used.

Mottier^{265,267} and Machata²²⁷ prepared semipermanent records of chromatograms by imbedding the layer in paraffin. In the case of unbound adsorbents, paraffin is applied to the plate as though it were another developing solvent, that is, by dipping the lower part of the layer in the melted paraffin and allowing it to rise by capillary action.^{265,267} The bound layers are impregnated by dipping them in melted paraffin.²²⁷ Semipermanent records can also be made by covering the layer with another glass plate and thus preserving it as a sandwich between two plates.^{227,232} The sandwiches are sealed by taping them around the edges.

A more permanent preparation can be made by imbedding the adsorbent layer in a plastic film which can be peeled off of the glass plates and preserved as such. This is accomplished by spraying the chromatogram with a 4 per cent collodion solution containing 7.5

* Sylvania Electric Products, Inc., 500 5th Ave., New York, N.Y.

23167

per cent glycerol,¹⁸ aqueous dispersions containing 10 to 15 per cent of polyacrylic acid esters, polyvinylidene chloride and polyvinyl propionate,²¹⁸ Fisher Label Glaze,*²³² or "Neatan" (Brinkmann, see p. 37).

* Fisher Scientific Company, 633 Greenwich Street, New York, N.Y.

CHAPTER 9

Preparative Thin-Layer Chromatography

INTRODUCTION

The preparative separation of reasonable amounts of a mixture (10 to 1000 mg) on thin layers has been a relatively neglected aspect of thin-layer chromatography. The high capacities of the adsorbents (when used in adsorption chromatography), the high degrees of resolution obtained, and the over-all simplicity of the method make it a near ideal tool for the organic chemist. The technique itself involves the mechanical steps of normal thin-layer chromatography (Chapters 2 to 8) with slight modifications.

Thus, the layers are thicker (0.5 to 2 mm compared with 0.25 mm) than normal so that the maximum amount of material can be separated at once. The sample is not applied in spots but in a narrow band and developed in such a manner that the resolved components lie in bands. These bands are located by some nondestructive visualization technique; the adsorbent containing a single component is removed from the band area and the sample is removed from the adsorbent by elution with an appropriate solvent. Quantities ranging from 10 to 250 mg can be separated on a single chromatogram by this technique and a number of chromatograms can conveniently be carried out simultaneously.

The development of this aspect of thin-layer chromatography has been severely hampered by the lack of methods for sample application and for the nondestructive visualization of the developed chro-

matogram. The sample must be applied in large quantities to a relatively thin band without undue destruction of the rather fragile layer. It now appears that satisfactory, though not ideal, solutions have been found for these problems.

At least two attempts have been made to apply the concepts of thin-layer chromatography to large-scale separations. These involve chromatography in cast blocks or "chromatobars"²⁵⁰ of plaster-of-Paris-bound adsorbent and chromatography in horizontal cellophane tubes filled with adsorbents.⁶⁶ A third similar technique is chromatography on layers of cellulose one inch thick packed between glass plates.¹³⁵ While each of these methods has certainly been successful in itself, the author believes that they are generally more laborious and inferior to the technique outlined below.

The various aspects of preparative thin-layer chromatography will be dealt with in the order and in the same manner as has been done in the preceding chapters for the general technique. Numerous examples of preparative work are cited as such in the tables of Chapter 11.

ADSORBENTS

Both bound and unbound adsorbents can be used with no special treatment other than perhaps a prepurification step (see pp. 19 and 48).

PREPARATION OF LAYERS

For the sake of convenience, most preparative thin-layer chromatography has been carried out on the usual square two-dimensional plates (20 × 20 cm or 8 × 8 in.). Larger plates could, of course, be used to separate larger quantities, but their use becomes awkward above this size and commercial apparatus is not available for preparation of the layers.

The adsorbent slurries have usually been made in the normal manner although Honegger suggests¹⁵⁸ the use of somewhat less water and a little more binder for very thick layers (see Table 3.2). Ritter and Meyer³²⁹ found the optimum layer thickness to be about 1 mm

although Honegger¹⁵⁸ gives a somewhat wider range of 1 to 3 mm. The very thick layers tend to crack and there is an appreciable difference between behavior on the open side of the layer and the side next to the glass. Layers up to 2 mm thick can be made with the adjustable Stahl applicator (see p. 37) and layers up to 5 mm thick can be made with the Camag apparatus (see p. 36).¹⁵⁸

The layers are best allowed to dry at room temperature for several hours before final activation. This prevents casehardening and subsequent cracking. It is further recommended that, as the original Desaga instructions suggest, the layers be dried in a vertical position rather than horizontal.

SAMPLE APPLICATION

Sample application is made by applying the sample, in an appropriate solvent, in a thin band about 2 cm from one edge of the layer. A smooth application can be effected using the mechanical devices of Morgan²⁵⁸ or of Ritter and Meyer³²⁹ as discussed in Chapter 4. Alternately, the sample can be applied from a syringe moved manually along the layer. The operation is best carried out on a heated metal block^{223,329} to evaporate the solvent. Honegger¹⁵⁸ applied the sample in a V-shaped trough cut in the layer.

The amount of sample applied will depend upon the specific mixture to be separated, the adsorbent involved and upon the type of chromatography. Much larger samples can be separated by adsorption than by partition chromatography. Mangold states²³² that up to 50 mg can be separated by adsorption chromatography on a layer of Morgan²⁵⁸ or of Ritter and Meyer³²⁹ as discussed in Chapter 4. Meyer³²⁹ separated 75 mg on a plate of the same size with a 1-mm-thick layer. Honegger states¹⁵⁸ that 5 to 25 mg can be separated on such a 1-mm layer and Černý, Joska and Lábler⁵² separated 20 to 100 mg of steroids on a layer of unbound alumina 20 × 10 cm and 1 mm thick. The author considers it "safe" to apply 50 mg of any mixture to a layer 20 × 20 cm and 1 mm thick and has applied as much as 300 mg of an easily separated mixture. For partition chromatography, only about one-tenth of these sample sizes should be used. Mangold²³² suggests 5 mg on a 20 × 20 cm layer 0.25 mm thick.

CHOICE OF DEVELOPMENT SOLVENTS

The solvents for preparative chromatography are the same as those used for normal thin layers.

DEVELOPMENT

The development of preparative chromatograms is carried out in the normal way with the following precautions. If several chromatograms are developed simultaneously, special care must be taken to saturate the chamber. The author prefers to develop such chromatograms with adsorbent turned toward the chamber wall and close to, but not touching, a layer of filter paper on the wall. If several chromatograms are to be developed in sequence in the same chamber, it should be remembered that thick layers remove more solvent than thin layers and the composition of the solvent may change appreciably.

All of the tricks of development discussed in Chapter 6 except those involving two-dimensional chromatography are applicable. Multiple-step and continuous development are particularly useful.¹⁵⁸

VISUALIZATION

In the ideal case, the bands on a developed chromatogram will be colored or can be seen in ultraviolet light.^{52,224} Bands containing conjugated double-bond systems can be located by using an adsorbent containing phosphors or a fluorescent additive as discussed in Chapter 2 (p. 21). Sometimes, small amounts of fluorescent or colored impurities will be present in a sample which will allow the orientation of a preparative plate with respect to a small plate (preferably with the same layer thickness) which has been visualized in a normal destructive manner. Schulze and Wenzel³⁴⁹ used their gas-flow counter to find bands of compounds containing C¹⁴ and tritium labels.

Mistryukov²⁵⁵ describes the use of "paper prints" to locate zones. Narrow paper strips (5 to 6 mm) are moistened and pressed against an unbound layer. After removal, the strips contain a layer of adsorbent and can be visualized with any spray reagent.

Iodine appears to be a good visualization reagent for most lipids²³⁰

since it does not affect them appreciably. Gritter and Albers¹²⁸ have located steroids and other types of compounds by spraying the layer with water as discussed in Chapter 7 (p. 103). When all of these methods fail, an edge spray technique can be used. The preparative layer is covered with a glass plate except for 2 to 3 cm along one vertical edge and a line of adsorbent is scraped off along the edge of the cover plate to isolate the narrow edge strip.³²⁹ The edge is then sprayed with an appropriate reagent, and the results are extrapolated across the layer.^{230,333,420, and others}

ELUTION OF COMPOUNDS

Most workers simply scrape the adsorbent containing the desired components from the glass plate with a spatula and elute it with an appropriate solvent. However, this is difficult when using unbound layers and two gadgets have been devised for the purpose. Both operate on the vacuum cleaner principle. The first, as shown in Figure 9.1, is a piece of glass tubing with a constriction and holds a cotton wad to catch the adsorbent. It was devised by Mottier^{267,269} and has the advantage that it can be used as an elution tube and no sample is lost in transfer. The Ritter and Meyer apparatus,³²⁹ as shown in Figure 9.2, is more suitable for large quantities. The adsorbent is sucked into a Soxhlet extractor thimble from which it is eluted by extraction.

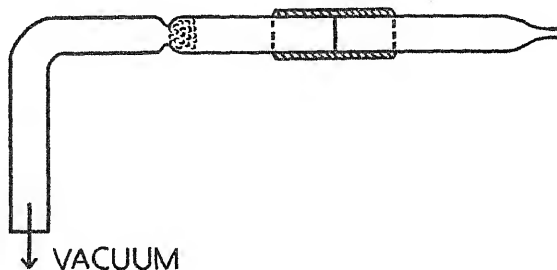


Figure 9.1. Mottier gadget for removing adsorbent from a thin layer. (Reproduced from Mottier, *Mitt. Gebiete Lebensm. Hyg.*, 49, 454 (1958) through the courtesy of the author and the Eidgenoessiche Drucksachen und Materialzentrale, Bern, Switzerland.)

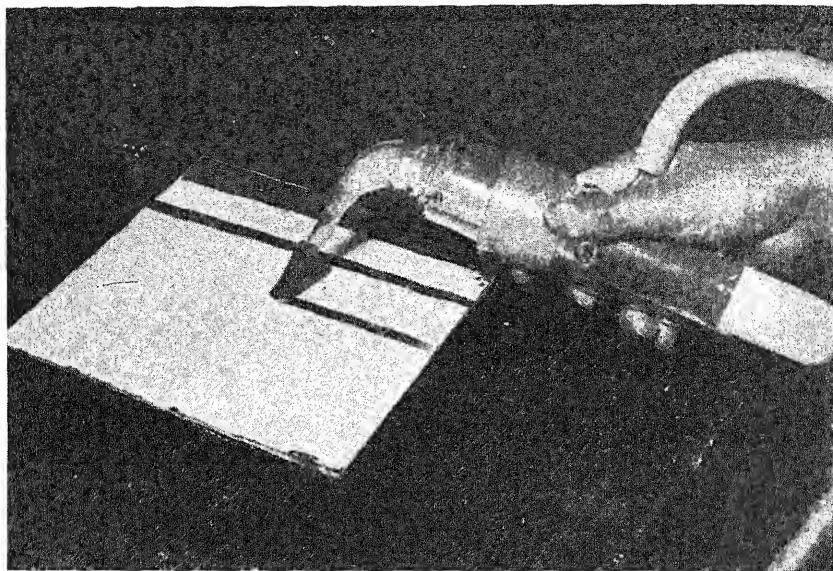


Figure 9.2. Ritter and Meyer apparatus for removing adsorbent from a thin layer. (Reproduced from Ritter and Meyer, *Nature*, 193, 941 (1962) through the courtesy of the authors and Macmillan and Co., Ltd.)

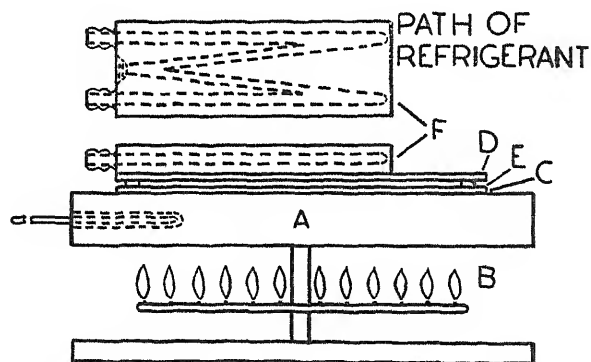


Figure 9.3. The apparatus of Baehler for direct sublimation out of thin layers. A refrigerant is passed through the metallic block (F) which rests on cover plate (D). The 1 mm spacer (E) separates the cover plate from plate (C) which bears the thin layer. Block A is a piece of metal containing a thermometer which is heated by burners (B). (Reproduced from Baehler, *Helv. Chim. Acta*, 45, 309 (1962) through the courtesy of the author and Verlag Helvetica Chimica Acta, Basel, Switzerland.)

The ideal elution solvent is one which moves the sample with the solvent front on a thin layer.²³² Aqueous eluents cannot be used with plaster-of-Paris-bound adsorbents because they will dissolve appreciable amounts of calcium sulfate. It is essentially impossible to avoid the solution of small amounts of silica gel in any eluting solvent. If the eluent is evaporated to dryness and taken up in a fresh solvent, preferably less polar than the original, most of the adsorbent will remain on the sides of the flask.

Sometimes, the substances on a chromatogram can be removed by direct sublimation to a cooled plate held about 1 mm over the adsorbent layer. Baehler¹⁰ used the apparatus shown in Figure 9.3 for the visualization of chromatograms by sublimation, but the principle is mentioned for the purpose of isolation by Ener (as reported by Mangold²³²).

CHAPTER 10

Quantitative Thin-Layer Chromatography

INTRODUCTION

Quantitative evaluation of thin-layer chromatograms can be divided into two general categories. In the first, the mixture components are assayed *on* the thin layer while in the second, they are eluted from the adsorbent before being measured. Either technique can be used with about 5 per cent error or, in ideal cases, about 3 per cent error.

The assay on the layer can be carried out by measuring the spot area, or the spot area and density on a developed and visualized chromatogram. These values are then related to the amount of substance in the spots through the use of standards or calibration curves. Spot size or density can also be measured on photocopies or photographs of visualized chromatograms and related to substance amounts. Finally, the assay can be carried out visually¹⁹⁴ by measuring the minimum visible amount of the standard on a chromatogram and making successive dilutions of an unknown solution until this minimum value is reached. The major advantages of these techniques are their simplicity and the fact that the chromatograms can be visualized with destructive spray reagents.

The elution of substances from the adsorbent introduces several sources of error which will be discussed. A spectrophotometric or colorimetric method must be available for the assay of the eluted compounds since the simple weighing of eluted samples gives poor re-

sults.⁴⁴² The elution technique has been used more extensively than the layer assay method.

ANALYSIS ON THIN LAYERS

Spot Area

This type of determination is based upon the fact that there is a mathematical relationship between the spot area for a given substance and the weight of the substance contained in the spot. The method involves only a small number of mechanical operations and is widely applicable.^{234,290,314,315,326,353,354}

There has, however, been some controversy over the type of mathematical relationship existing between these quantities. In fact, Stahl³⁷⁴ questioned the existence of any precise relation at all. Seher^{353,354} ignored the mathematics and obtained results with only 5 per cent error using calibration curves obtained by plotting spot area (from photostats) against weight. The calibration curves were not straight lines. Brenner and Niederwieser⁴⁴ and Pastuska and Petrovitz²⁹⁰ obtained a linear relationship by plotting the logarithm of the weight of sample against spot area. This agrees with Giddings and Keller,¹²³ who derived such a relationship for paper chromatography from theoretical considerations. The most thorough and successful study of this technique was made by Purdy and Truter,^{314,315} who found a linear relationship between the logarithm of the weight of a substance and the square root of the spot area. These authors noted, however, that a reasonably linear relationship resulted when the spot area or the logarithm of the spot area was used rather than its square root.

Purdy and Truter explored three general methods of analysis based upon spot size. Using quantities ranging from 1 to 100 γ , they carried out analyses for potassium and magnesium ions, hexadecanol, cholesterol, palmitic acid, cholesteryl laurate and phenylazo-2-naphthol. Both adsorption and partition chromatography were involved. In all cases, it was necessary to have a pure sample of the substance in question for use as a standard.

The Graphical Method of Purdy and Truter. This method is simi-

lar to that of the Seher^{353,354} technique, that is, it is based upon the preparation of a calibration curve by plotting spot area against weight with no regard to linearity. A standard deviation of 4.3 per cent (25 observations) was obtained and the major error was ascribed to the difficulty of duplicating conditions on the calibration chromatograms and the analysis chromatograms. In contrast, the second and third methods are algebraic and depend upon the linear relationship between $\log W$ and \sqrt{A} .

First Algebraic Method of Purdy and Truter. In this method, a standard solution of the compound to be assayed is spotted on the layer with a solution of the unknown and a solution prepared by precise dilution of the unknown. The chromatogram is developed and visualized and the spot areas are determined. The results are calculated using Equation (1) where W and A are the weight and spot area of the unknown, W_s and A_s are the weight and area of the standard, A_d is the area of the diluted unknown spot and d is the dilution fac-

$$\log W = \log W_s + \left(\frac{\sqrt{A} - \sqrt{A_s}}{\sqrt{A_d} - \sqrt{A_s}} \right) \log d \quad (1)$$

tor. In essence, the unknown and its diluted sample establish the slope of the straight line used for the analysis and the standard establishes its intercepts. This method is the most accurate of the three, and yielded standard deviations of 2.7 per cent (540 observations) for adsorption chromatography and 3.6 per cent (982 observations) for partition chromatography.

Second Algebraic Method of Purdy and Truter. This method requires a pure sample of the substance to be assayed rather than a standard solution. The unknown solution is diluted in a precise manner and the diluted solution is divided into two portions. A known amount of the pure substance is added to one of the diluted samples and all three solutions are spotted on the same chromatogram. The chromatogram is developed and visualized; the spot areas are determined and the results are calculated using Equation (2). In this case,

$$\log \left(\frac{Wd + a}{W} \right) = \left(\frac{\sqrt{A_+} - \sqrt{A}}{\sqrt{A_d} - \sqrt{A}} \right) \log d \quad (2)$$

W and A are the weight and spot area of the unknown; A_d and A_+

are the spot areas of the diluted sample and the diluted sample containing added substance, respectively; a is the amount of pure substance added to the diluted unknown and d is the dilution factor. This method is somewhat less accurate and yielded standard deviations of 3.9 per cent (374 observations) for adsorption chromatography and 3.9 per cent (982 observations) for partition chromatography.

Purdy and Truter found that it was necessary to spot the same volume of each solution on a single chromatogram. They used an Agla micrometer syringe suspended 2 mm above the thin layer for this purpose. The appropriate sample was squeezed out of the syringe and transferred to the layer by raising the layer until it contacted the solvent drop on the syringe. Chloroform was found to be an unsatisfactory spotting solvent. The chromatograms were developed in the usual manner and those containing colorless compounds were visualized by spraying with conc. sulfuric acid and charring at 160° for 10 minutes. The areas were determined by laying a sheet of transparent paper on the chromatograms, tracing the outline of the spots and measuring the areas of the figures by superimposing the tracing on millimeter graph paper.

In a related method, Eble and Brooker⁹³ measured the length of the streaks (actually the distance between the origin and the zone front) formed when tryptamine was chromatographed on Kieselguhr G with acetone-water (99:1). The method gave results that were reproducible to about 5 per cent.

Spot Area and Density

The size and intensity of spots formed by the acid spray and charring of a chromatogram provide a satisfactory basis for quantitative evaluation.^{296,310,311} The method has an advantage in that the densities of the spots can be measured on a densitometer.

Additional variants are added, however, with respect to the spot-area methods discussed above. These are the carbon density of a given compound, its volatility and whether it chars in a uniform manner. These disadvantages can largely be nullified by the judicious use of internal standards and calibration curves³¹⁰ and results can be obtained with about 5 per cent error.

Before spotting, the thin layer, either the standard size of 5×20 cm^{310,311} or microscope slide size,²⁹⁶ is scanned in the densitometer to determine the degree and the uniformity of the background density. The chromatograms are then spotted, using a micropipette or microsyringe, with the sample containing the substance to be assayed. The chromatogram is developed, sprayed with 50 per cent sulfuric acid or 25 per cent perchloric acid and charred in an oven. It is then scanned by passing it through a densitometer and the amount of unknown substance is calculated using an appropriate calibration curve. The densitometers which have been used are the Photovolt* models 52C and 521A with a stage attached for semiautomatic plotting of curves for the standard size layers^{310,311} and the Photovolt model 501A with transmission density unit no. 52 for the microchromatoplates.²⁹⁶ The slit sizes were 1×5 mm for standard size and 1×3 mm for microsize. Glycerides^{310,311} and cholesterol²⁹⁶ have been analyzed by this technique.

Spot Area and Density on Photostats or Photographs

The spot area has been determined from photostats of thin-layer chromatograms and used as a quantitative measurement by Kammerreck (as reported in Mangold²³²) and by Seher.^{353,354} The error was about 5 per cent.

Neubauer and Mothes²⁷³ photographed chromatograms of opium alkaloid mixtures and cut the photographs into strips which were passed through a densitometer. Hefendehl¹⁴¹ made chromatograms transparent (see p. 104), photographed them and measured the density of the photographs against a calibration curve. He obtained results with about 5 per cent error in the analysis of menthofuran in an essential oil. Rybicka³³⁵ followed the glycerolysis of linseed oil by this method and was able to determine the optimum reaction conditions by finding the concentrations of the glycerides at timed intervals.

Minimum Spot Visibility

In this technique,¹⁹⁴ a standard solution of the substance to be assayed is successively diluted and each dilution is chromatographed un-

* Photovolt Corp., 95 Madison Avenue, New York, N.Y. A commercial densitometer, especially designed for thin-layer work, is now available from Photovolt.

til the minimum visible amount is determined. This will, of course, depend upon the visualization method and the eye of the analyst and will vary from one compound to another. The unknown solution is then successively diluted and chromatographed until the minimum visible amount is determined. Knowing the dilution factor of the unknown and the minimum value, the amount of unknown in the original sample can be calculated. The average error for nineteen samples was 11.3 per cent.

Using a similar method, Waldi has carried out analyses for pregnandiol in urine as an early test for pregnancy^{433,435} and for adrenaline (as its triacetyl derivative⁴³⁴) in adrenal extracts. He chromatographed the solution containing the material to be assayed on the same layer with a series of standard known concentrations and estimated the concentration of the unknown by comparison.

ANALYSIS AFTER ELUTION

Analysis by Spectrophotometry

Those compounds which possess a visible or ultraviolet spectrum are conveniently assayed in the following manner. After development of the chromatogram, the spot of adsorbent containing the substance to be analyzed is quantitatively removed; the substance is eluted with an appropriate solvent and assayed in a spectrophotometer. Thus, the spectroscopic properties of such compounds provide both a means of locating the spots on a developed chromatogram and a means of quantification. Substances which contain ultraviolet chromophores are more easily observed on phosphor-containing adsorbents (p. 21).

The methods involving elution contain several more sources of error than those in which compounds are assayed on the thin layers. First, the spots on the chromatogram must be located. This, however, is a minor problem when the substances have visible or ultraviolet chromophores. Second, the adsorbent containing the substance to be assayed must be completely removed from the glass support plate. Finally, the substance must be completely removed from the adsorbent before it is measured.

It has been suggested^{194,383,389} that the thin layers be prewashed

before being used for this type of analysis. This is carried out by allowing solvent to ascend (or descend) through the layers, perhaps in an apparatus similar to that shown in Chapter 3 (p. 49). The chromatogram is then spotted with a precise amount of substance to be separated and developed in a normal manner. The spots are located by viewing them in an appropriate light and the adsorbent containing the compound to be assayed is removed, normally with a razor blade. The compound is eluted from the adsorbent with a sufficiently polar solvent (one which will move the sample to the solvent front of a chromatogram) and measured in a spectrophotometer. The gadget described by Mottier²⁶⁹ (Figure 9.1) could well be used in this technique since it provides both a method for the removal of adsorbent from the glass plate and an elution tube, thus minimizing transfer losses. The method requires the preparation of a calibration curve and compounds which follow Beer's law. It is also necessary to make "blank" readings on the adsorbent, preferably on adsorbent which has been developed in a thin layer and is removed from a point on the layer comparable to the one containing the sample to be assayed.

Normally, the method gives results with an error of 3 to 5 per cent. It was first used for the assay of biphenyl in citrus fruits (2.8 per cent error^{194,389}) and has since been used for analysis of reserpine (about 5 per cent error³⁴⁴), coumestrol (5 per cent error²²²), menthyl salicylate,³⁸³ estrogens (standard deviation, 10 per cent³⁹⁴), alkaloids (5 per cent error⁴⁰⁰), bile acids (standard deviation, 3 per cent¹¹⁵), and *p*-hydroxybenzoic acid esters (3 per cent error¹¹⁸). Quinine alkaloids²⁷⁰ and 7-geranoxycoumarin⁴¹⁷ have been eluted and determined by fluorescence methods.

Analysis by Colorimetry

This technique is almost identical with the spectroscopic method described above except that a color-producing reagent is added to the sample before it is measured in the spectrophotometer. It is subject to the same disadvantages, although the location of the substances on the developed chromatogram is now a major problem. In general, the same nondestructive visualization techniques that are discussed in

Chapter 9 (p. 112) for preparative chromatography are applicable for the location of the spots to be eluted.

The method generally gives results with about 5 per cent error. It has been used for the analysis of esters (about 5 per cent error⁴²⁰), pregnanetriols (about 10 per cent error³⁹¹), citral (2 per cent error³⁸⁶), coumarin compounds³⁸⁸ and plasma phospholipids (about 5 per cent error¹³⁴). Further, the method has been successfully applied to the assay of carbohydrates,^{120,289} steroids,^{244,278,391,460} lipids^{413,430,458} and meprobamate.¹⁰²

RADIOACTIVITY METHODS

Analysis techniques which depend upon radioactivity require prior labeling of the compounds to be analyzed. The chromatograms are spotted and developed in the usual way and have been evaluated by three methods. The first method is to make autoradiographs^{231,234a} as described in Chapter 8 (p. 104) and to evaluate them using a densitometer and calibration curve. The second method^{234a} is to locate the spots by such an autoradiograph or by nondestructive spray reagents and to elute the sample and measure it in a conventional counter. Scintillator solutions have been used^{35,349,364} to aid the counting. Finally, Schulze and Wenzel³⁴⁹ used a gas-flow counter (Figure 7.1) for the assay of C¹⁴ and tritium labeled compounds with about 3 per cent error.

The application of labeling techniques to lipids has been thoroughly discussed by Mangold.^{232,234a}

CHAPTER 11

Specific Applications of Thin-Layer Chromatography

INTRODUCTION

This chapter will contain, in a tabular form, the majority of the specific applications of thin-layer chromatography which were recorded prior to January 1, 1963. Specific R_f values will not be given because of their notorious inconstancy on thin layers and because they are too numerous. It is hoped that this chapter will serve at least two purposes.

First, it will provide a general idea of the types of compounds which have been chromatographed on thin layers and, to some extent, the conditions of the separation. It would be hazardous, however, to assume that this compilation is complete because chromatographic conditions are usually incidental to a piece of research and are frequently buried in the experimental description.

Second, it will provide some idea of the adsorbents, solvents and visualization techniques which are commonly used in a given area or with a given class of compounds. In many cases, the same group of compounds has been treated in more than one fashion and the reader can choose the one which best suits his facilities and inclinations.

The numerous blank spaces in the tables are due to the failure of the original author to give specific details or to the present author's inability to obtain or translate the original article. In either case, they are given to make the compilation as complete as possible.

In many cases, a single reference will give details of several separate experiments such as the use of two adsorbents, the use of two techniques or the separation of two distinct groups of compounds. In other cases, two papers will report different aspects of the same experiments. This has been shown, as well as possible, by a system of partial lines. On occasion, a dotted line is used. This means that the regions on the ends of the dotted line are related and divided as shown but that the category crossed by the dotted line is the same for both experiments.

(EXPLANATION OF TABLES 11.1 TO 11.23)

Compounds

The various applications have, first of all, been arranged according to the class of compounds involved. This has been somewhat arbitrary at times and considerable overlapping could not be avoided. This is particularly serious between the sections devoted to amines, alkaloids and drugs and pharmaceuticals and between the essential oils and the terpenoids. In most cases, the applications are cross-referenced and the same example is reported in two places when it involves compounds falling in two classes. When the number of compounds in a specific entry is small, they are given by name. Otherwise, collective terms are used.

Within a given class, an attempt has been made to group the applications as much as possible according to the compounds involved. In most cases, the large comprehensive papers are given first and the others follow. The papers which came to the author's attention after the tables were first prepared, however, are given at the end of each section. The 2,4-DNPH's referred to in several cases mean the 2,4-dinitrophenylhydrazones.

Adsorbent

The adsorbent is given as a specific industrial product when it is reported as such in the original paper and when it is *especially prepared for thin-layer chromatography*. When a specific product is not named, it is implied that the original authors prepared their own ad-

sorbent and the reader is referred to Chapter 2. In most cases where partition chromatography is involved, the adsorbent is equilibrated with the developing solution or some of its components. This is generally noted.

The various materials which have been incorporated into the adsorbent for one reason or another are given in the tables, but without details. The reader is referred to Chapter 2.

Developer

The solvents used for development are given as such except when the article gives a very large number. In these cases, a selected few are given, and the number not listed is mentioned in the last item. Unless noted, the different solutions represent separate experiments and are not successive.

The proportions of the various solvent mixtures, unless noted, are volume-volume. The abbreviations listed on p. 87 have been used to conserve space. Otherwise, the names or, sometimes, the formulas (CCl_4 , CHCl_3 , H_2O , etc.) are used.

Visualization

For the sake of brevity, only the names of the spray or visualization reagents are given. The reader is referred to the alphabetical listing in Chapter 7 for exact details. Two abbreviations are used in this section; 2,4-DNPH for 2,4-dinitrophenylhydrazine and U.V. for ultraviolet visualization.

Comments

The author has tried to give the basic purpose of the chromatography in each case and has, on occasion, added notes about especially interesting applications.

TABLES 11.1-11.23 FOLLOW

TABLE 11.1. ALDEHYDES, KETONES AND THEIR DERIVATIVES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Unsaturated aliphatic aldehydes and acyl- ins	Silica Gel G	a. pet. ether-EtOEt (90:10) (95:5) b. pet. ether-EtOEt-HOAc (70:30:1)	I ₂ vapor	a. diagnostic b. monitored a syn- thesis	119
2 Long-chain (over eight carbons) aliphatic ketones	Silica Gel G	a. benzene-EtOEt mixtures b. toluene-EtOEt mixtures c. pet. ether-EtOEt mixtures	phosphomolybdic acid	diagnostic	238
3 Biologically active al- dehydes					447
4 β -Dicarbonyl com- pounds	starch-bound silica gel	benzene-EtOAc (7:3)		diagnostic	14
5 Cyclohexanediones	see "Miscellaneous" No. 12				254
6 Aliphatic and aromatic aldehydes and ke- tones	unbound aluminum ox- ide	a. benzene b. benzene-EtOH (98:2)(95:5) (90:10) c. CHCl ₃ d. EtOEt e. pet. ether-benzene (1:1)	reagents used for paper chromatography	diagnostic	144
7 Aromatic aldehydes and ketones	plaster-of-Paris-bound silica gel	hexane-EtOAc (4:1)(3:2)		diagnostic	113
8 Phenolic aldehydes	see "Phenols, etc." No. 12				290
9 Vanillin and other substituted benzal- dehydes	Silica Gel G	a. pet. ether-EtOAc (2:1) b. hexane-EtOAc (5:2) c. CHCl ₃ -EtOAc (98:2) d. decalin-CH ₂ Cl ₂ -MeOH (5:4:1)	a. hydrazine sulfate in HCl b. methanolic KOH c. methanolic KOH fol- lowed by diazotized sulfanilic acid	diagnostic	396

10	Cholestane-1,3	see "Steroids" No. 45					397
11	Aldehydes and ketones in peppermint oil	see "Essential Oils" No. 6					323 324
12	Acetals and their determination in thin layers						206
13	2,4-Dinitrophenylhydrazones of aldehydes	Silica Gel G	a. benzene-pet. ether (3:1) for aliphatic b. benzene-EtOAc (95:5) for aromatic	colored	diagnostic		84
14	2,4-DNPH's of ketones	starch-bound SiO_2 -Ben-tonite (4:1)	a. benzene b. benzene-pet. ether-EtOEt mixtures	colored	diagnostic		209
15	2,4-DNPH's of aliphatic aldehydes and ketones	Silica Gel G	benzene-EtOAc (95:5)	a. colored b. alcoholic KOH	a. diagnostic b. preparative		258
16	2,4-DNPH's of aliphatic aldehydes	polyvinyl alcohol-bound silica gel	benzene- H_2O (91.8:8.2)	colored	diagnostic		284
17	2,4-DNPH's of aldehydes and ketones	unbound aluminum oxide	a. benzene-hexane (1:1) b. EtOEt	colored	diagnostic		331
18	2,4-DNPH's of aliphatic aldehydes	starch-bound ZnCO_3	small amounts of pyridine in pet. ether-benzene mixtures	colored	a. diagnostic b. monitored column separation		9
19	2,4-DNPH's of hydroxy aldehydes and ketones	a. Aluminum Oxide G b. Silica Gel G	toluene-EtOAc (3:1)(1:1)	colored	diagnostic		5

TABLE 11.2. ALKALOIDS

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Fifty-four misc. alkaloids	Silica Gel G	a. CHCl_3 -acetone-diethylamine (5:4:1) b. CHCl_3 -diethylamine (9:1) c. cyclohexane- CHCl_3 -diethylamine (5:4:1) d. cyclohexane-diethylamine (9:1) e. benzene-EtOAc-diethylamine (7:2:1)	potassium iodoplatinate	a. diagnostic b. devised an identification system	436
	Aluminum Oxide G	f. CHCl_3 g. cyclohexane- CHCl_3 (3:7) plus 0.05% diethylamine			
	Silica Gel G-NaOH	h. MeOH			
	unbound aluminum oxide	a. benzene-EtOH mixtures b. CHCl_3 c. EtOH			
2 Thirty-seven misc. alkaloids	Silica Gel G	a. MeOAc-2-propanol- NH_4OH (Ca. 25% NH_3) (45:35:20) b. CHCl_3 -HOAc (95:5)	a. Van Urk reagent b. Procházka reagent c. 2,4-dinitrophenylhydrazine	a. diagnostic b. two dimension	376
3 Twenty-one indole derivatives	Silica Gel G	EtOAc- CHCl_3 -MeOH (20:20:10)	Dragendorff's reagent	diagnostic	83
4 <i>Amaryllidaceae</i> , <i>Papaveraceae</i> and Bisbenzylisoquinoline alkaloids (total 57)	Silica Gel G Silica Gel G-NaOH	a. MeOH-HOAc-EtOEt-benzene (1:18:60:120) b. EtOH-HOAc-benzene (12:5:80) c. EtOEt- CHCl_3 -MeOH (50:50:1) d. benzene-EtOH (9:1) e. BuOH-HOAc (10:1)	a. FeCl_3 - $\text{K}_2\text{Fe}(\text{CN})_6$ b. chloramine T- HCl-NH_3 c. potassium iodoplatinate d. 2,6-dichloroquinone chlorimide e. ninhydrin f. SbCl_5 in CCl_4 g. CNBr and benzidine	diagnostic	117
5 Alkaloid-containing drug mixtures	Silica Gel G with fluorescent material				

6	Alkaloids and barbiturates in toxicology	Silica Gel G	a. MeOH b. CHCl_3 -EtOEt (85:15)	a. Dragendorff's reagent b. I_2 and KI in HOAc c. sodium fluorescein d. FeCl_3 in HOAc e. Zwikker's reagent f. KMnO_4 in HOAc	a. diagnostic b. clinical	226 227
7	Sixteen alkaloids and nitrogen containing drugs	Silica Gel G	a. EtOH-pyridine-dioxane- H_2O (50:20:25:5) b. EtOH-HOAc- H_2O (60:30:10) c. EtOH-dioxane-benzene- NH_4OH (5:40:50:5) d. MeOH-BuOH-benzene- H_2O (60:15:10:15) e. BuOH-butyl ether-HOAc (40:50:10) f. BuOH-butyl ether- NH_4OH (25:70:5)	potassium iodoplatinate	a. diagnostic b. urine assay	57
8	Opium alkaloids	Aluminum Oxide G	a. H_2O -EtOH-BuOH (1:1:9)	U.V.	diagnostic	239
9	Opium alkaloids	Silica Gel G	benzene-MeOH (8:2)	Dragendorff's reagent	a. diagnostic b. quantitative c. preparative	273
10	Ergot alkaloids	a. cellulose powder impregnated with formamide b. Silica Gel G	a. benzene-heptane- CHCl_3 (6:5:3)	a. Dragendorff's reagent b. With formamide layers the excess formamide was first destroyed by spraying with 0.25% NaNO_2 in 0.5% HCl and Dragendorff's reagent was used. c. U.V. d. $\text{K}_3\text{Fe}(\text{CN})_6$	a. diagnostic b. quantitative	399 400
	Morphine alkaloids	c. Silica Gel G-NaOH d. Cellulose powder impregnated with formamide	b. CHCl_3 -EtOH (9:1) c. benzene-heptane- CHCl_3 -diethylamine (6:5:3:0.03)			
11	Alkaloids of <i>Adhillea</i> species	Silica Gel G	CHCl_3 -MeOH (1:2)	a. potassium iodoplatinate b. H_2SO_4	diagnostic	285

TABLE 11.2—Cont.

	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
12	Lysergic acid alkaloids				diagnostic	151
13	<i>Aspidosperma</i> alkaloids	Silica Gel G (not activated)	Methyl cellosolve	U.V.	diagnostic	215
14	Quinine alkaloids	Silica Gel G (prepared in acetone)	"Shellsol" T-diethylamine-acetone (23:9:9)	exposed to formic acid vapor and U.V.	a. diagnostic b. quantitative	270
15	An isomer of rhocadin from <i>Papaver rhoeas</i>	Aluminum Oxide G	a. benzene-CHCl ₃ (1:2) b. benzene-tetrahydrofuran (95:5) c. CHCl ₃ -MeOH (9:1)	Dragendorff's reagent	diagnostic	444
16	Morphine	Silica Gel G	BuOH-HOAc (9:1) saturated with H ₂ O	<i>p</i> -dimethylaminobenzaldehyde in H ₂ SO ₄	a. diagnostic b. on microscope slides	437
17	Tropine and <i>pseudo</i> tropine	Silica Gel G	EtOH-NH ₄ OH (28% NH ₃) (8:2)	Dragendorff's reagent	diagnostic	188
18	Anaferine (a piperidine alkaloid)	Silica Gel G	a. MeOH-CHCl ₃ -diethylamine (1:9:0.5) b. EtOH-NH ₄ OH (conc.) (8:2)	Dragendorff's reagent	diagnostic	332
19	Reserpine	Silica Gel G	MeOH-methyl ethyl ketone-heptane (3.4:33.6:58)	U.V.	a. diagnostic b. quantitative c. developed in dark	344

TABLE 11.3. AMINES AND THEIR DERIVATIVES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Simple aliphatic amines	Silica Gel G	a. EtOH (95%)-NH ₄ OH (25% NH ₃) (4:1) b. phenol-H ₂ O (8:3) c. butanol-HOAc-H ₂ O (4:1:5) d. EtOH-H ₂ O (70:30)	ninhydrin	diagnostic	399
2 Strong-base amines	Silica Gel G buffered to pH 6.8 unbound aluminum oxide	a. acetone-heptane (1:1) b. CHCl ₃ /NH ₃ (sat. at 22°)-EtOH (96%) (30:1) c. acetone-MeOH-H ₂ O (8:2:1) d. methyl ethyl ketone-H ₂ O (15:1)	I ₂ vapor-U.V.	a. diagnostic b. describes descending development with unbound layers	255 256
3 Diphenylamine derivatives	Silica Gel G	a. CHCl ₃ b. toluene c. benzene	NaNO ₂ and H ₂ SO ₄	a. diagnostic b. analysis of powder	138
4 Tryptamine	Kieselguhr G	acetone-H ₂ O (99:1)	FeCl ₃ -K ₃ Fe(CN) ₆	quantitative	93
5 Aromatic amines and nitro compounds	plaster-of-Paris-bound silica gel		a. KOH in acetone b. SnCl ₄ in conc. HCl c. HNO ₃ -H ₂ SO ₄	diagnostic	110
6 Nitramine explosives	see "Miscellaneous", No. 10				216
7 Piperidones and quinolols	see "Miscellaneous", No. 12				254
8 Twelve pteridines	Silica Gel G	a. dimethylformamide-H ₂ O (190:10) (20:180) b. 5% citric acid c. NH ₄ OH (25% NH ₃) d. twenty others	U.V.	diagnostic	274
9 Simple amines and diamines by electrophoresis	Silica Gel G buffered with borax	electrolyte: EtOH-H ₂ O (80 ml. plus 30 ml.) with 2 g. NaOAc	ninhydrin	diagnostic	291

TABLE 11.4. CARBOHYDRATES, GLYCOSIDES AND THEIR DERIVATIVES

Compounds	Adsorbent	Developer	Visualization	Comments	Ref.
1 Thirteen misc. sugars	Kieselguhr G buffered with NaOAc	EtOAc-2-propanol-H ₂ O (65:23.4:11.6)	anisaldehyde and H ₂ SO ₄	diagnostic	377
2 Simple sugars	Polyacrylonitrile-Kieselguhr G-Aluminum Oxide G (5:6:1) buffered with NaHPO ₄	n-amyl alcohol-2-propanol-HOAc-H ₂ O (55:15:110:30)	aniline phthalate	diagnostic	36
3 Simple sugars	Silica Gel G buffered with boric acid	a. benzene-HOAc-MeOH (1:1:3) b. methyl ethyl ketone-HOAc-MeOH (3:1:1)	naphthoresorcinol and H ₂ SO ₄	a. diagnostic b. quantitative	289
4 Simple sugars	unbound Cellulose Powder MN-300	a. EtOAc-pyridine-H ₂ O (2:1:2) both phases b. phenol, equilibrated with 1% NH ₄ OH c. 2-propanol-pyridine-HOAc-H ₂ O (80:80:10:40)	anisidine phthalate	diagnostic	350
5 Simple sugars	Silica Gel G	n-PROH-conc. NH ₄ OH-H ₂ O (6:2:1)	a. ninhydrin b. aniline hydrogen phthalate	a. diagnostic b. preparative c. noted sugar aminations during development	439
6 Simple sugars	a. Silica Gel G b. Silica Gel G buffered with boric acid c. Kieselguhr G buffered with 0.02 N NaOAc	a. PROH-EtOAc-H ₂ O (3:2:1) b. BuOH-acetone-H ₂ O (4:5:1) c. PROH-EtOAc-H ₂ O (7:2:1) d. methyl ethyl ketone-HOAc-H ₂ O (3:1:1) e. Sol b with 5% NH ₄ OH f. EtOAc-2-PROH-H ₂ O (65:24:12) (5:2:0.5)	naphthoresorcinol and H ₂ PO ₄	a. diagnostic b. used wedge-shaped chromatograms	307

7	Glucose and trehalose	Silica Gel G	sec-BuOH saturated with H ₂ O	NaIO ₄ -K ₂ MnO ₄	a. diagnostic b. could not separate more complex mixtures	454
8	Ten sugar acetates and eight inositol acetates	Silica Gel G	benzene with 2-10% MeOH	a. H ₂ O b. ferric hydroxamate reagent	a. diagnostic b. preparative c. multiple development	398
9	Sugar acetates	starch-bound silicic acid	a. EtOAc-benzene (3:7) b. MeOH-benzene (2:98) c. EtOAc-benzene (3:97)(4:6)	silver nitrate-NH ₂ -NaOMe	diagnostic	75
10	Partially methylated sugars, isopropylidene sugars and other derivatives	a. Silica Gel G b. Silica Gel G with boric acid	a. BuOH-acetone-H ₂ O (4:5:1) b. methyl ethyl ketone-HOAc-H ₂ O (3:1:1) c. pet. ether	naphthorescinol-H ₃ PO ₄	diagnostic	308
11	Sucrose and raffinose	Starch-bound silica gel	a. 2-PrOH-toluene-EtOH-H ₂ O (10:2:5:2.5) b. toluene-EtOAc-EtOH (95%) (10:5:5)	naphthorescinol-H ₃ PO ₄ 2',7'-dichlorofluorescein	a. diagnostic b. quantitative	120
12	Maltodigosaccharides up to ten glucose units	Kieselguhr G	a. BuOH-pyridine-H ₂ O (65:20:15) and six other combinations b. BuOH-EtOH-H ₂ O (30:30:20) c. BuOH-2,6-lutidine-H ₂ O (60:30:10)	anisaldehyde-H ₂ SO ₄ -HOAc	diagnostic	441
13	Aldose 2,4-DNPH's	a. Aluminum Oxide G b. Silica Gel G	a. toluene-EtOAc (1:1) b. toluene-EtOAc (3:1) (1:1)	a. colored compounds b. NaOH	a. diagnostic b. two-dimensional development on two different absorbents	5
14	Sugar derivatives	alumina				203
15	p-Bromophenylosazones of xylose derivatives	plaster-of-Paris-bound silica gel	benzene-MeOH (9:1)	colored compounds	diagnostic	7

TABLE 11.4—Cont.

Compound	Absorbent	Developer	Visualization	Comments	Ref.
16 Cardiac glycosides	unbound, deactivated (H ₂ O or HOAc) silica gel	benzene-EtOH (3:1) saturated with H ₂ O	a. trichloroacetic acid-chloramine-T b. SbCl ₅ and acetic anhydride in CHCl ₃	diagnostic	322
17 Cardiac glycosides	Silica Gel G	EtOAc-pyridine-H ₂ O (5:1:4), upper phase	SbCl ₅ in CHCl ₃	diagnostic	392
18 Cardiac glycosides	Silica Gel G	a. CHCl ₃ -MeOH (9:1) b. CHCl ₃ -acetone (63:35)	H ₂ SO ₄ and acetic anhydride	diagnostic	378
19 Glycosides of <i>atractylis gummifera</i>	Silica Gel G	a. BuOH-MeOH-H ₂ O (8:1:1) b. PrOH-xylene-H ₂ O (7:2:1)	conc. H ₂ SO ₄	diagnostic	100
20 k-Strophanthin	Silica Gel G	BuOH-MeOH-formamide (17:2:1)	a. SbCl ₅ in CHCl ₃ b. trichloroacetic acid-chloramine-T c. Kedde reagent d. conc. H ₂ SO ₄	diagnostic	223
21 Flavone and coumarin glycosides	Silica Gel G	EtOAc-methyl ethyl ketone-formic acid-H ₂ O (5:3:1:1)	diphenylboric acid, β -aminoethyl ester	diagnostic	379
22 Flavone glycosides	starch-bound silica gel	a. EtOAc-CHCl ₃ (95:5) b. EtOAc-EtOH (95:5) c. 3-methyl-1-butanol-hexane-HOAc-H ₂ O (3:1:3:3)	a. AlCl ₃ b. uranium acetate	diagnostic	286
23 Flavanol glycosides	Silica Gel G	a. HOAc-formic acid-H ₂ O-McOH (10:2:2:1)	NH ₃ vapor	diagnostic and quantitative	126
Cardiac glycosides	—	b. methyl ethyl ketone-toluene-H ₂ O-HOAc-McOH (40:5:3:1:2:5)	SbCl ₅ in CHCl ₃		

24	Rutin	a. unbound polyamide powder b. Silica Gel G	a. MeOH b. MeOH-H ₂ O (8:2)(6:4) c. BuOH-HOAc (9:1) saturated with H ₂ O PrOH-1 N NH ₄ OH-H ₂ O (6:2:1) a. MeOH b. MeOH-EtOEt mixtures 1-pentanol-1-propanol-HOAc-H ₂ O (3:2:2:1) plus 2 parts 1-heptanol	a. diazotized sulfanilic acid b. U.V. <i>p</i> -dimethylamino-benzaldehyde in H ₂ SO ₄ conc. H ₂ SO ₄	diagnostic a. diagnostic b. on microscope slides a. diagnostic b. monitored column diagnostic diagnostic	74 437 440 39 36
25	Brain gangliosidic substances	Silica Gel G				189
26	<i>Catalpa</i> glucosides	Silica Gel G				
27	Anthocyanins and their aglucones	polyacrylonitrile-"Perlon" (7:1) buffered with KH ₂ PO ₄				
28	Misc. glycosides					

TABLE 11.5. CARBOXYLIC ACIDS AND THEIR DERIVATIVES (EXCLUSIVE OF FATTY ACIDS)

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 2,6-Dimethyl-2,6-octadienedioic acid (Hildebrandt's acid)	Silica Gel G	CHCl ₃ :HOAc (97:3)	KMnO ₄	diagnostic	428
2 Dicarboxylic acids	Silica Gel G	a. benzene-MeOH-HOAc (45:8:4) b. benzene-dioxane-HOAc (90:25:4)	bromophenol blue acidified with citric acid	diagnostic	302
3 Carboxylic acids (as ammonium salts)	Silica Gel G	a. EtOH(95%)-H ₂ O-NH ₄ OH(25% NH ₃)(100:12:16)	bromocresol green	diagnostic	42
4 Methoxy-piperonylic acids	Silica Gel G	EtOAc-hexane-HOAc (50:50:0.5)	chromotropic acid-H ₂ SO ₄	diagnostic	32
5 Phenolic acids	see "Phenols, Alcohols, etc.", No. 9 and No. 12				225 288 290
6 Aromatic acids	a. gelatin-bound Silica Gel G-"Perlon" (0.5:7:4) b. polyacrylonitrile-"Perlon" (8:4)	a. <i>n</i> -amyl alcohol- <i>n</i> -BuOH-EtOAc-HOAc-H ₂ O (60:20:30:10:5) b. <i>n</i> -octane-EtOAc- <i>n</i> -amyl alcohol-EtOEt-formic acid (40:50:15:15:10)	U.V.	diagnostic	36
7 <i>p</i> -Hydroxybenzoic acid esters	Silica Gel G with fluorescent material	pentane-HOAc (88:12)	U.V.	quantitative	118
8 Phthalates, phosphates and other esters as plasticizers	see "Miscellaneous," No. 14				293
9 Gallic acid and its esters	unbound polyamide powder	a. MeOH b. EtOH c. EtOEt d. CCl ₄ -MeOH (7:3) e. six others		a. diagnostic b. antioxidants	69

10	Lactones, lactams and thiolactones	Silica Gel G	a. isopropyl ether b. isopropyl ether-EtOAc (80:20) c. isopropyl ether- <i>is</i> -octane (20:80) (60:40)	a. NaOH-hydroxyl- amine-HOAc-FeCl ₃ consecutively b. Dragendorff's reagent	diagnostic	207
11	Dicarboxylic acids (from oxalic to sebacic) and carbocyclic acids	Kieselguhr G-Polyethyl- ene glycol (M-1000) (2:1)	isopropyl ether-formic acid-H ₂ O (90:7:3)	bromocresol purple	diagnostic	200 201

TABLE 11.6. DRUGS AND PHARMACEUTICALS

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Caffeine, theobromine and theophylline	Silica Gel G	a. EtOAc-MeOH-12 N HCl (18:2:0.05) b. EtOAc-MeOH-HOAc (8:1:1) c. CHCl ₃ -MeOH (19:1)	sublimation from layer	a. diagnostic b. preparative	10
2 Meprobamate	starch-bound silicic acid	EtOH-cyclohexane (15:85)	conc. H ₂ SO ₄	a. diagnostic b. quantitative	102 240
3 Barbiturates	see "Alkaloids," No. 6				226 227
4 Nitrogen-containing drugs	Silica Gel G	a. organic solutions b. buffer {depending upon drug solutions}	a. I ₂ in CHCl ₃ b. potassium iodoplatinate	a. diagnostic b. semiquantitative	279
5 Narcotics in forensic chemistry	see "Miscellaneous," No. 19				92
6 Alkaloidal drugs in urine	see "Alkaloids," No. 7				57
7 Urinary metabolites of psychotropic drugs	Silica Gel G	BuOH-HOAc-H ₂ O (88:5:7) (65:15:20)	conc. H ₂ SO ₄	diagnostic	101
8 Chloramphenicol	aluminum oxide			monitored synthesis	217
9 Penicillin preparations	Silica Gel G	a. acetone-MeOH (1:1) b. 2-PrOH-MeOH (30:70)	sodium azide-I ₂ solution	diagnostic	103
10 Tetracyclines, penicillins and rifomycins	a. Silica Gel G b. Silica Gel G buffered to pH 4.5 or 4.2	a. aqueous citric acid solutions b. pure alcohols c. many others	a. HCl and others b. microbiological assay	diagnostic	275 276
11 Antibiotics (macrolides)	Silica Gel G	acetone-H ₂ O (98:2)	microbiological assay	diagnostic	33

12	Sulfonamides	a. Silica Gel G b. Silica Gel G containing fluorescein	CHCl ₃ -EtOH-heptane (1:1:1) plus 1.0 to 1.8% H ₂ O	a. HCl-NaNO ₂ -N-(1-naphthyl) ethylenediamine b. U.V.	diagnostic	196
13	Sulfonamides	Silica Gel G	CHCl ₃ -EtOH-heptane (1:1:1)	<i>p</i> -dimethylaminobenzaldehyde, acidified	diagnostic	452
14	Sulfonamide, radioactive	Silica Gel G	BuOH sat. with conc. NH ₃	radioactive method	a. diagnostic b. quantitative c. preparative	349
15	Misc. constituents of plant drugs					379
16	Pharmaceutical applications				review	122 340

TABLE 11.7. DYES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Food dyes	unbound aluminum oxide	a. BuOH-EtOH-H ₂ O (90:10:10) (80:20:10)(70:30:30) (60:40:40)(50:50:50)	colored compounds	a. diagnostic b. spot made alkaline before chromatography c. circular	266 269 306
2 Food dyes	Cellulose Powder MN 300 G and MN 300	a. aqueous sodium citrate (2.5%)-NH ₄ OH(25%) (4:1) b. PrOH-EtOAc-H ₂ O (6:1:3) c. <i>t</i> -BuOH-propanoic acid-H ₂ O (50:12:36) with 0.4% KCl	colored compounds	diagnostic	451
3 Food dyes	unbound aluminum oxide	a. MeOH b. EtOH c. CHCl ₃ d. CCl ₄ e. EtOEt f. seven others	colored compounds	a. diagnostic b. quantitative	73
4 Food dyes	Silica Gel G	a. CHCl ₃ -acetic anhydride (75:2) b. benzene c. methyl ethyl ketone-HOAc-MeOH (40:5:5)	SbCl ₅ in CHCl ₃	diagnostic	257
5 Carotenoid dyes	see "Vitamins," No. 2				210 211
6 Carotenoid dyes, synthetic	calcium hydroxide-Silica Gel G (6:1)	pct. ether-benzene (2:3)	colored compounds	a. diagnostic b. monitored synthesis	166
7 Dyes in gasoline	Silica Gel G	benzene	colored compounds	a. diagnostic b. to catch gasoline thieves	139
8 Test mixture dyes: butter yellow, Sudan red III and indophenol	see "Miscellaneous," No. 16				381

TABLE 11.8. ESSENTIAL OILS

Compound	Adsorbent	Developer	Visualization	Comments	Ref
1 Essential oils of Argentina	starch-bound silicic acid containing fluorescent material	a. hexane b. pet. ether-EtOAc mixtures	a. U.V. b. fluorescein-Br ₂ c. bromocresol green d. 2,4-DNPH e. conc. H ₂ SO ₄	diagnostic	105 106 107
2 Thirteen essences	starch-bound silicic acid	a. hexane b. benzene-ethanol mixtures	a. SbCl ₅ in CHCl ₃ b. 2,4-DNPH c. KMnO ₄ d. <i>p</i> -dimethylaminobenzaldehyde-H ₂ SO ₄	diagnostic	287
3 Misc. essential oils	Silica Gel G	benzene-CHCl ₃ (1:1)	SbCl ₅ in CHCl ₃	a. diagnostic b. used with gas chromatography	382
4 <i>Perfume industry</i> a. Essential oils b. Single essences c. Balsams and resins d. Tar and tar products e. Phenols and phenol ethers f. Wool wax and derivatives g. Azulenes	Silica Gel G	a. hexane b. benzene c. CH ₂ Cl ₂ d. CHCl ₃ e. benzene f. benzene-McOH mixtures	a. SbCl ₅ in CHCl ₃ b. SbCl ₅ in CCl ₄ c. conc. H ₂ SO ₄ d. vanillin and other aldehydes in H ₂ SO ₄ e. fluorescein-Br ₂ f. 2,4-DNPH g. <i>o</i> -dianisidine	a. diagnostic b. circular c. in shaped areas	368
5 Misc. essential oils	sec "Terpenoids," No. 1	g. CHCl ₃ -McOH mixtures h. hexane i. benzene	h. SbCl ₅ in CCl ₄ i. acetic anhydride-H ₂ SO ₄ j. chloranil in CHCl ₃ k. Br ₂ in CHCl ₃		190

TABLE 11.8—Cont.

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
6	a. Misc. essential oils	a. hexane b. hexane-EtOAc mixtures	a. U.V. b. 2,4-DNPH	diagnostic	323
	b. Oils of <i>Mentha</i> genus				324
	c. Mint oils			quantitative	326, 327
					325
7	Oils of Japanese mint	hexane-EtOAc (85:15)	vanillin in H ₂ SO ₄	diagnostic	167 168
8	Oils of <i>Mentha piperita</i>	hexane	a. SbCl ₅ in CHCl ₃ b. transparency method	quant. assay of menthofuran	141
9	Oils of peppermint				297
10	Oils of <i>Oleum menthae</i>	a. benzene-EtOAc (95:5) b. hexane	a. vanillin in H ₂ SO ₄ b. SbCl ₅ in CHCl ₃ c. SbCl ₅ in CCl ₄ d. fluorescein-Bri e. 2,4-DNPH f. KI in HOAc with Na ₂ SO ₃ and followed by starch g. HNO ₃ -HOAc (1:300)	diagnostic	171
11	Oils of peppermint	"Skellysolve" B-EtOAc mixtures, up to 10% EtOAc	a. transmitted light b. U.V. c. trichloroacetic acid in CHCl ₃ d. KMnO ₄ in H ₂ SO ₄ e. Rhodamine B	a. diagnostic b. study site and mode of terpene formation in plant c. preparative	19 20
12	Menthol in peppermint oil	alumina-silicate		quantitative	133

13	Oils of Eucalyptus	starch-bound silica gel	hexane-EtOAc (85:15)	conc. H ₂ SO ₄	a. diagnostic b. circular	313
14	Oils of Eucalyptus	starch-bound magnesol			a. diagnostic b. circular	51
15	Oils of <i>Strobilanthopsis</i> <i>Limifolia</i>	a. silica gel-"Celite"- starch (16:4:1) b. plaster-of-Paris- bound silica gel	hexane-EtOAc mixtures		a. diagnostic b. monitored column	58
16	Oil of jasmine	see "Terpenoids," No. 13				76 81
17	Azulene of camphor blue oil	see "Terpenoids," No. 14				108
18	Oil of <i>Pogostemon patch-</i> <i>ouli</i> Pell	starch-bound silicic acid	a. hexane-EtOAc (85:15) b. solution a plus 5% CHCl ₃	conc. H ₂ SO ₄	a. diagnostic b. used "chromato- bars"	124
19	Oils of <i>Arnica montana</i> L.	starch-bound silicic acid and <i>terra silicata</i>	CHCl ₃	conc. H ₂ SO ₄	diagnostic	131
20	Oils of <i>grapefruit</i> a. Hydrocarbons b. Oxygenated com- pounds	a. starch-bound silicic acid b. plaster-of-Paris- bound silicic acid ----- ----- -----	a. hexane b. 2,2-dimethylbutane c. cyclohexane d. methylcyclohexane e. isopentane f. hexane-EtOAc mixtures g. EtOAc-CHCl ₃ (alcohol free) mixtures h. EtOAc-benzene (15:85) i. hexane-nitropropane (50:50) j. isopropyl ether-hexane (50:50)	a. fluorescein-Br ₂ b. HNO ₃ -H ₂ SO ₄ (5:95)	a. diagnostic b. predicted solvents for column work	191
21	Oils of orange juice	starch-bound silicic acid	hexane-EtOAc (85:15) plus a variety of others	a. fluorescein-Br ₂ b. 2,4-DNPH c. <i>o</i> -dianisidine d. conc. sulfuric-nitric acid	diagnostic	192

TABLE 11.8—Cont.

	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
22	<i>Oils of lemon</i>	starch-bound silicic acid with phosphors	a. EtOEt-pet. ether mixtures	a. U.V.	a. quantitative b. descending tech.	383
	a. Menthyl salicylates					
	b. Coumarins		b. hexane-EtOAc (85:15)	b. 2,4-DNPH c. α -dianisidine in glacial HOAc d. U.V.	c. monitored column d. diagnostic	387 388 417
	c. Hydrocarbons		c. heptane to develop and elute hydrocarbons from end of layer	e. Fluorescein-Br ₂	e. quant. assay of hydrocarbons by gas chromatography	385
	d. Citral		d. hexane-EtOAc (85:15)	f. fluorescein-Br ₂ , viewed under U.V.	f. diagnostic	386
25	e. Chalcones		e. hexane-EtOAc (75:25) f. EtOEt-pet. ether (65:35)	g. HCl gas h. U.V.	g. diagnostic	384
	23 Oils of sage	Silica Gel G	benzene	vanillin in H ₂ SO ₄	diagnostic	49
25	Volatile oil of grafted umbelliferous plants				diagnostic	198 199
26	Oils of <i>Curcuma</i>	Silica Gel G	benzene-CHCl ₃ -MeOH (200:100:3)	SbCl ₅ in CHCl ₃	diagnostic	445
27	Oils of <i>Acorus calamus</i>	Silica Gel G	benzene		a. diagnostic b. to relate plants	453
28	Oils of <i>coriander lavender</i> and muscat sage					298

29	Oils of <i>Oleum rosmarini</i> , <i>O. Lavandulae</i> and <i>O. roseae</i>	Silica Gel G	benzene-EtOAc (95:5)	a. SbCl_3 in CHCl_3 b. SbCl_5 in CCl_4 c. conc. H_2SO_4	diagnostic	172
30	2,4-DNPH's of aldehydes and ketones from essential oils	see "Aldehydes, etc." No. 14				209
31	Removal of hydrocarbons from essential oils	see "Terpenoids," No. 1				190

TABLE 11.9. FLAVONOIDS AND COUMARINS

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Isoflavones	starch-bound silicic acid with phosphors	a. EtOAc-Skellysolve B (3:1) b. acetone-EtOAc-Skellysolve B (4:3:3) c. EtOH-CHCl ₃ (1:3)(1:1)	U.V.	a. diagnostic b. preparative	132
2 Flavonoids	unbound polyamide powder (Silon)	a. MeOH-H ₂ O (80:20)(60:40) b. MeOH	U.V.	diagnostic	71 74
3 Flavonoids and coumarins	Silica Gel G-NaOAc	toluene-ethyl formate-formic acid (5:4:1)	diphenylboric acid, β -aminoethyl ester	diagnostic	379
4 Flavonoids and coumarins of lemon	starch-bound silicic acid with phosphors	CHCl ₃ -acetone mixtures	U.V.	a. diagnostic b. monitored col- umn	162
5 Coumarins	Silica Gel G	a. pet. ether-EtOAc (2:1) b. hexane-EtOAc (5:2)	KOH and diazotized sulfanilic acid	diagnostic	396
6 Coumarins in lemon juice	starch-bound silicic acid with phosphors	hexane-EtOAc (75:25)	U.V.	preparative	31
7 Coumestrol in plants	starch-bound silicic acid with phosphors	a. CHCl ₃ -EtOEt (1:1) b. EtOAc-"Skellysolve" B (3:1) c. acetone-"Skellysolve" B (1:3) d. EtOEt-"Skellysolve" B (7:3)	U.V.	a. diagnostic b. quantitative	222 224 34

TABLE 11.10. INORGANIC AND METALORGANIC COMPOUNDS

	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1	Fe ⁺² , Zn ⁺² , Be ⁺² , Ni ⁺² , and others	starch-bound alumina- ⁴⁶ Celite ³⁹	a. H ₂ O b. dilute HCl c. organic liquids	a. I ₂ b. U.V.	a. diagnostic b. circular on microscope slides	246
2	a. Copper family-Cu ⁺² , Cd ⁺² , Bi ⁺³ , Pb ⁺² , Hg ⁺² b. Sulfide group-Zn ⁺² , Co ⁺² , Fe ⁺² , Ni ⁺² , Mn ⁺² , Cr ⁺³ , Al ⁺³	purified Silica Gel G	a. BuOH-1.5 N HCl-2,5-hexanediol (100:20:0.5) b. acetone-conc. HCl-2,5-hexanediol (100:1:0.5)	a. KI b. NH ₃ and H ₂ S c. NH ₃ d. 8-hydroxyquinoline e. U.V.	diagnostic theory	359 356
3	U ⁺⁶ , Ga ⁺³ and Al ⁺³	purified Silica Gel G	a. EtOAc (sat. with H ₂ O)-tributyl phosphate (50:2) b. acetone-conc. HCl (100:0.5)	a. pyridyl-azo-naphthol b. 8-hydroxyquinoline c. U.V.	diagnostic	360
4	Alkali metals-Na ⁺¹ , Li ⁺¹ , K ⁺¹ , Mg ⁺²	purified Silica Gel G	EtOH-HOAc (100:1)	acid violet	diagnostic	358
5	Halogens-F ⁻¹ , Cl ⁻¹ , Br ⁻¹ , I ⁻¹	purified Silica Gel G	acetone-BuOH-conc. NH ₄ OH-H ₂ O (65:20:10:5)	a. bromocresol purple and NH ₃ b. ammonical AgNO ₃ and fluorescein	diagnostic	357
6	Various phosphates	purified Silica Gel G	MeOH-conc. NH ₄ OH-10% tri-chloroacetic acid and H ₂ O (50:15:5:30)	ammonium molybdate followed by SnCl ₄ and HCl	diagnostic	355
7	K ⁺¹ and Mg ⁺²	Silica Gel G	EtOH-MeOH (1:1) plus 1% HOAc	acid violet 6BN	quantitative	315
8	Radioactive halogens	⁴⁶ Dowex 1 ²⁹ -Cellulose Powder MN 300 G (1:1)	1 M aqueous NaNO ₃	radioactivity	diagnostic	29

TABLE 11.9. FLAVONOIDS AND COUMARINS

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Isoflavones	starch-bound silicic acid with phosphors	a. EtOAc-Skellysolve B (3:1) b. acetone-EtOAc-Skellysolve B (4:3:3) c. EtOH-CHCl ₃ (1:3)(1:1)	U.V.	a. diagnostic b. preparative	132
2 Flavonoids	unbound polyamide powder (Silon)	a. MeOH-H ₂ O (80:20)(60:40) b. MeOH	U.V.	diagnostic	71 74
3 Flavonoids and coumarins	Silica Gel G-NaOAc	toluene-ethyl formate-formic acid (5:4:1)	diphenylboric acid, β -aminoethyl ester	diagnostic	379
4 Flavonoids and coumarins of lemon	starch-bound silicic acid with phosphors	CHCl ₃ -acetone mixtures	U.V.	a. diagnostic b. monitored col- umn	162
5 Coumarins	Silica Gel G	a. pct. ether-EtOAc (2:1) b. hexane-EtOAc (5:2)	KOH and diazotized sulfanilic acid	diagnostic	396
6 Coumarins in lemon juice	starch-bound silicic acid with phosphors	hexane-EtOAc (75:25)	U.V.	preparative	31
7 Coumestrol in plants	starch-bound silicic acid with phosphors	a. CHCl ₃ -EtOEt (1:1) b. EtOAc-"Skellysolve" B (3:1) c. acetone-"Skellysolve" B (1:3) d. EtOEt-"Skellysolve" B (7:3)	U.V.	a. diagnostic b. quantitative	222 224 34

TABLE 11.10. INORGANIC AND METALORGANIC COMPOUNDS

	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1	Fe ²⁺ , Zn ²⁺ , Be ²⁺ , Ni ²⁺ , and others	starch-bound alumina, ^a Celite ^b	a. H ₂ O b. dilute HCl c. organic liquids	a. I ₂ b. U.V.	a. diagnostic b. circular on microscope slides	246
2	a. Copper family-Cu ²⁺ , Cd ²⁺ , Bi ³⁺ , Pb ²⁺ , Hg ²⁺ b. Sulfide group-Zn ²⁺ , Co ²⁺ , Fe ²⁺ , Ni ²⁺ , Mn ²⁺ , Cr ³⁺ , Al ³⁺	purified Silica Gel G	a. BuOH-1.5 N HCl-2,5-hexanediol (100:20:0.5) b. acetone-conc. HCl-2,5-hexanediol (100:1:0.5)	a. KI b. NH ₃ and H ₂ S c. NH ₃ d. 8-hydroxyquinoline e. U.V.	diagnostic theory	359 356
3	U ²³⁸ , Ga ⁷³ and Al ²⁷	purified Silica Gel G	a. EtOAc (sat. with H ₂ O)-tributyl phosphate (50:2) b. acetone-conc. HCl (100:0.5)	a. pyridyl-azo-naphthol b. 8-hydroxyquinoline c. U.V.	diagnostic	360
4	Alkali metals-Na ⁺ , Li ⁺ , K ⁺ , Mg ²⁺	purified Silica Gel G	EtOH-HOAc (100:1)	acid violet	diagnostic	358
5	Halogens-F ⁻ , Cl ⁻ , Br ⁻ , I ⁻	purified Silica Gel G	acetone-BuOH-conc. NH ₄ OH-H ₂ O (65:20:10:5)	a. bromocresol purple and NH ₃ b. ammonical AgNO ₃ and fluorescein	diagnostic	357
6	Various phosphates	purified Silica Gel G	MeOH-conc. NH ₄ OH-10% trichloroacetic acid and H ₂ O (50:15:5:30)	ammonium molybdate followed by SnCl ₂ and HCl	diagnostic	355
7	K ⁺ and Mg ²⁺	Silica Gel G	EtOH-MeOH (1:1) plus 1% HOAc	acid violet 6BN	quantitative	315
8	Radioactive halogens	"Dowex 1"-Cellulose Powder MN 300 G (1:1)	1 M aqueous NaNO ₃	radioactivity	diagnostic	29

TABLE 11.10—Cont.

	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
9	Twenty-four ferrocene derivatives	Silica Gel G	a. benzene b. benzene-EtOH (30:1)(15:1) c. propylene glycol-MeOH (1:1) d. propylene glycol-chlorobenzene-MeOH (1:1:1)	a. Br ₂ b. NaIO ₄	diagnostic	345
10	Chromium and cobalt complexes of <i>o</i> -hydroxy- <i>o'</i> -carboxyazo dyes	Aluminum Oxide Fluka	MeOH	colored compounds	diagnostic	342
11	Organo-tin stabilizers in polyvinyl chloride	Silica Gel G containing "Komplexon III"	H ₂ O-BuOH-EtOH-HOAc (10:5:5:0.15)	Diphenylthiocarbazone	diagnostic	411

TABLE 11.11. INSECTICIDES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1	a. Thiophosphate esters b. Chlorinated compounds	hexane-acetone (4:1) hexane	PdCl ₂ in dilute HCl N,N-dimethyl- <i>p</i> -phenyl- enediamine hydrochlor- ide in sodium methylate	a. diagnostic b. in forensic chemistry	23
2	Dichlorophene and hexa- chlorophene	heptane saturated with HOAc	FeCl ₃ -K ₃ Fe(CN) ₆	a. diagnostic b. quantitative	43
3	Thiophosphate esters	starch-bound aluminum oxide	H ₂ IO ₆ -HClO ₄ -V ₂ O ₅	diagnostic	304
4	Pyrethrins	plaster-of-Paris-bound silica gel	fluorescein-Br ₂	diagnostic	365
5	Pyrethrins and their peroxides	Silica Gel G a. benzene-methyl ethyl ke- tone (90:10) b. benzene-EtOAc (85:15) c. CCl ₄ -EtOAc (80:20) d. hexane-methyl ethyl ketone (80:20) e. hexane-EtOAc (75:25)	SbCl ₃ in CHCl ₃	a. diagnostic b. reaction on layer	371
6	Pentachlorophenol	Silica Gel G-oxalic acid	U.V.	a. diagnostic b. quantitative	83
7	Chlorinated organic pesticides				455
8	Chlorinated hydrocarbons with insecticidal properties				300

TABLE 11.12. LIPIDS—CLASS SEPARATION

1	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1	Misc. lipids	Silica Gel G	a. EtOEt b. isopropyl ether c. isopropyl ether-HOAc (98.5:1.5)	Phosphomolybdic acid	Class separation prior to analysis by reversed phase technique	182
2	Misc. lipids	Silica Gel G	pet. ether-EtOEt-HOAc (90:10:1)	2',7'-dichlorofluorescein	Class separation prior to analysis by reversed phase technique	230
3	Misc. lipids				monitored column	150
4	Radioactive lipids	silicic acid	pet. ether-EtOEt (92:8)	a. radioactivity b. α -cyclodextrin-I ₂ c. I ₂	class separation	231
5	Industrial aliphatic lipids	a. Silica Gel G b. Silica Gel G plus 10% (NH ₄) ₂ SO ₄	a. pet. ether-benzene (95:5) b. benzene sat. at 20° with 1 N NH ₄ OH c. CHCl ₃ (sat. with 1 N NH ₄ OH)-MeOH (97:3) d. acetone-14 N NH ₄ OH (90:10) e. CHCl ₃ -(MeOH-0.1 N H ₂ SO ₄ , 95:5)(97:3)(80:20)	a. I ₂ b. 2',7'-dichlorofluorescein c. K ₂ Cr ₂ O ₇ in H ₂ SO ₄	diagnostic	234
6	Misc. fats, oils and waxes	Silica Gel G	a. pet. ether-EtOEt-HOAc (90:10:1)(70:30:2)(70:30:1) b. hexane-EtOEt (95:5)	a. I ₂ b. 2',7'-dichlorofluorescein	a. diagnostic b. class separation	235
7	Tissue lipids	Silica Gel G	hexane-EtOEt-HOAc (90:10:1)	2',7'-dichlorofluorescein	class separation prior to assay by gas chromatography	236
8	Serum lipids	unbound aluminum oxide	a. pet. ether-EtOEt (95:5) b. pet. ether-EtOH (98:2) c. pet. ether-EtOEt-HOAc (94.5:5:0.5) d. heptane-HOAc (98:2)		a. class separation b. quantitative	413

9	Serum lipids and phospholipids	plaster-of-Paris-bound silicic acid	a. $\text{CHCl}_3\text{-MeOH-H}_2\text{O}$ (80:25:3) b. pet. ether-EtOAc-HOAc (90:10:1)(60:40:1)	2', 7'-dichlorofluorescein	diagnostic	423
10	Serum lipids	Silica Gel G	a. $\text{PrOH-NH}_4\text{OH}$ (2:1)-1st step b. $\text{CHCl}_3\text{-benzene}$ (3:2) 2nd step c. $\text{CCl}_4\text{-reverse}$	SbCl_5 in CHCl_3	a. stepwise technique b. two dimensional c. quantitative	21 164 438
11	Serum, liver and adipose tissue lipids	Silica Gel G				86
12	Plasma lipids	Silica Gel G	$\text{CHCl}_3\text{-MeOH-H}_2\text{O}$ (65:25:4)	a. bromthymol blue-NiH ₂ b. ninhydrin c. fuchsin- H_2SO_4 d. hydroxamate reaction	a. diagnostic b. quantitative for phosphates	134
13	Plasma lipids	Silica Gel G	a. pet. ether-diisobutylketone-HOAc (87:13:0.7) b. pet. ether-methyl ethyl ketone-HOAc (95:4:1)	a. 50% H_2SO_4 b. Rhodamine 6 G c. bromthymol blue d. osmium tetroxide	class separation prior to separation by reversed phase technique	459
14	Butter-fat lipid					245
15	Brain tissue lipids	Silica Gel G	a. $\text{CHCl}_3\text{-MeOH-H}_2\text{O}$ (24:7:1) (60:20:3) b. CHCl_3 c. CCl_4	a. ammonium molybdate- HClO_4 b. phosphomolybdic acid c. ninhydrin d. Dragendorff's reagent e. diphenylamine	diagnostic on multiple sclerosis patients	157 159
16	Brain lipid constituents (pure test substances)	Silica Gel G	a. heptane b. CCl_4 c. 1,2-dichloroethane d. CHCl_3 e. 1,2-dichloroethane-McOH (98:2) f. pet. ether-EtOAc (70:30) g. four others	a. bromthymol blue, buffered b. bromthymol blue c. conc. $\text{H}_2\text{SO}_4\text{-HOAc}$ (1:1) d. phosphomolybdic acid e. $\text{HIO}_4\text{-Schiff}$ reagent f. <i>aqua regia</i> -ammonium molybdate-vanadyl chloride g. diphenylamine	a. diagnostic b. system for exploring unknown lipids	174

TABLE 11.12—*Cont.*

	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
17	Dogfish liver lipids	Silica Gel G	pet. ether-EtOEt-HOAc (90:10:1)		separated the alkoxy-glycerides class for assay	229
18	Lipids of feces and fecal-liths	Silica Gel G	pet. ether-EtOEt-HOAc (80:20:1)	50% H ₂ SO ₄	a. class separation b. semiquantitative assay	442
19	Lipids				review review	104 232

TABLE 11.13. LIPIDS—FATTY ACIDS AND THEIR DERIVATIVES

	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1	Fatty acids	plaster of Paris impregnated with undecane	a. HOAc-CH ₃ CN (1:1) b. solution a with 0.5% Br ₂	copper acetate-dithio-oxamide	separated "critical" pairs after hydrogenation and bromination on layer	180
2	Fatty acids	Kieselguhr G impregnated with undecane	HOAc-CH ₃ CN (1:1)	copper acetate-dithio-oxamide on layers "smoked" with dichlorodimethylsilane	diagnostic	185
3	Fatty acids	Kieselguhr G impregnated with undecane	a. HOAc-CH ₃ CN (6:4), 80% sat. with undecane b. HOAc-CH ₃ CN (30:70), 80% sat. with undecane and containing 0.5% Br ₂	Rhodamine B	separated "critical" pairs after hydrogenation and bromination on layer	186
4	Fatty acids	Silica Gel G impregnated with undecane	a. 96% HOAc b. HOAc-CH ₃ CN (1:1)	phosphomolybdic acid	diagnostic	182
5	Fatty acids of Swiss Pharmacopoeia	sec "Lipids-Glycerides", No. 16				6
6	Keto- and hydroxy-fatty acids and their lactones	Kieselguhr G impregnated with tetradecane or paraffin	a. HOAc-H ₂ O (80:20), 80% sat. with impregnating solution (5% of agent in pet. ether) b. HOAc-H ₂ O (90:10), 75% sat. with impregnating solution	a. Rhodamine B b. phosphomolybdic acid	diagnostic	181
7	Fatty acids of castor oil and oiticica oil	Silica Gel G	pet. ether-EtOEt-HOAc (70:30:1)(70:30:2)	2',7'-dichlorofluorescein	diagnostic	235
8	Lignoceric, cerbronic and nervonic acids	Silica Gel G			diagnostic	109
9	Fatty acid methyl esters	starch-bound silicic acid with phosphors	a. "Skellysolve" F-EtOEt (90:10)(70:30)(50:50) b. benzene-EtOEt (75:25) (50:50)	a. 2',7'-dichlorofluorescein b. U. V. c. fluorescein-Br ₂	a. diagnostic b. monitored column	8

TABLE 11.13—Cont.

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
10 Fatty acid methyl esters	Silica Gel G impregnated with AgNO_3	EtOEt-hexane (10:90) (40:60)	2',7'-dichloro-fluorescein	diagnostic	260
11 Fatty acid methyl esters	Silica Gel G	hexane-EtOEt-HOAc (85:15:1)	2',7'-dichloro-fluorescein	preparative separation prior to analysis by gas chromatog.	333
12 Fatty acid methyl esters	Silica Gel G	hexane-EtOEt mixtures (up to 30% EtOEt)	a. I_2 b. 2',7'-dichloro-fluorescein c. 50% H_2SO_4	a. quantitative b. preparative	420
13 Fatty acid methyl esters	Silica Gel G impregnated with Silicone oil	a. CH_3CN -HOAc- H_2O (70:10:25) b. peracetic acid-HOAc- H_2O (10:75:15)	a. I_2 b. α -cyclodextrin and I_2	a. diagnostic b. separated sat. esters by using an oxidizing developer	230
14 Fatty acid methyl esters	Silica Gel G	a. pet. ether-EtOEt (80:20) b. solution a plus 1% HOAc	50% H_2SO_4	diagnostic	262
15 Epoxy fatty acid methyl esters	Silica Gel G	hexane-EtOEt (95:5) (90:10)	50% H_2SO_4	monitored various separation methods	261 264
16 Fatty acids and esters	silicic acid on the inside of test tubes	a. hexane b. hexane-EtOEt (97:3) (85:15) c. CHCl_3	I_2	diagnostic	219
17 Fatty acid esters of orujo oil	Silica Gel G	hexane-EtOEt (90:10)	50% H_2SO_4	a. diagnostic b. monitored column	422
18 Fatty acid esters of cholesterol	a. Silica Gel G b. Silica Gel G impregnated with paraffin	a. tetralin-hexane (25:75) (1:1) b. CCl_4 c. methyl ethyl ketone- CH_2CN (7:3)	phosphomolybdic acid	a. diagnostic b. two-dimension	184

19	Fatty acids and their cholesterol esters	Silica Gel G	a. PrOH-NH ₄ OH (2:1) b. CHCl ₃ -benzene (3:2) c. CCl ₄	SbCl ₅ in CHCl ₃	a. diagnostic b. stepwise technique c. two-dimension	438
20	Cholesterol and its esters	see "Steroids," No 28 and No. 27				228 249
21	Acetoxymercuromethoxy derivatives of unsaturated fatty acid methyl esters	Silica Gel G	a. pet. ether-EtOEt (4:1) b. PrOH-HOAc (100:1)	<i>s</i> -diphenylcarbazone	isolated unsaturated esters prior to assay by gas chromatography	177 233
22	<i>p</i> -Phenylazophenacyl esters of fatty acids	Silica Gel G	pet. ether-EtOEt (85:15)	colored compounds	diagnostic	419

TABLE 11.14. LIPIDS—GLYCERIDES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Glycerides	Silica Gel G impregnated with AgNO_3	CHCl_3 -HOAc (99.5:0.5)	dibromo-R-fluorescein	separation according to degree of unsaturation	16
2 Glycerides	Silica Gel G	CHCl_3 -benzene (70:30)	conc. H_2SO_4	separation of the glyceride fraction from other fractions	64
3 Glycerides	Silica Gel G impregnated with undecane	a. CHCl_3 -MeOH- H_2O (5:15:1) (for diglycerides) b. acetone- CH_3CN (7:3) (for triglycerides)	Rhodamine B	diagnostic	182
4 Glycerides	Silica Gel G	a. "Skellysolve" F-EtOEt (70:30) b. (10:90) for monoglycerides c. (70:30) for diglycerides d. (90:10) for triglycerides e. (60:40)(35:65)(85:15)	50% H_2SO_4	a. quantitative	311
5 Glycerides	plaster-of-Paris-bound silicic acid	gradient elution system of pet. ether-EtOEt varied between (9:1) and (4:6). All contained 0.1% HOAc.		b. quantitative based upon ozonolysis and periodate oxidation	310
6 Glycerides	Silica Gel G	a. 1,2-dichloroethane b. CHCl_3 c. 1,2-dichloroethane-MeOH (98:2) d. pet. ether-EtOEt (70:30) e. CHCl_3 -HOAc (96:4)	2',7'-dichloro-fluorescein	a. diagnostic b. gradient elution	334
7 Monoglycerides	hydrosyl-apatite	methyl isobutyl ketone at 10°	a. HIO_4 -Schiff reagent b. bromthymol blue	diagnostic	174
8 Triglycerides	Kieselguhr G impregnated with petroleum	acetone- CH_3CN (3:2)(7:4)	phosphomolybdic acid a. phosphomolybdic acid b. I_2 c. α -cyclodextrin- I_2	diagnostic a. diagnostic b. multiple development	155 179

9	Triglycerides	plaster of Paris impregnated with paraffin (b.p. 240-250°)	a. acetone-CH ₃ CN (8:2) b. propanoic acid-CH ₃ CN (6:4) plus 0.5% Br ₂ by vol.	Sudan Black B	a. diagnostic b. bromination on layers	180
10	Triglycerides	Kieselguhr G impregnated with tetracane	a. acetone-CH ₃ CN (8:2) 80% sat. with tetracane b. propanoic acid-CH ₃ CN (6:4) 80% sat. with tetracane plus 0.5% Br ₂ by vol.	a. I ₂ b. phosphomolybdic acid	a. diagnostic b. bromination on layer	186
		Kieselguhr G impregnated with paraffin (b.p. 240-250°) or Silicone oil	c. acetone-CH ₃ CN (8:2) d. MeOH-CH ₃ CN (5:4) e. MeOH-CH ₃ CN-CH ₃ CH ₂ CN (5:4:1.5) f. solution e saturated with silicone oil	c. Rhodamine B d. I ₂ e. phosphomolybdic acid	diagnostic	183
11	Glycerides of milk fat (lipolyzed)	Silica Gel G	pet. ether-EtOH-HOAc (90:10:1)		preparative separation prior to assay by gas chromatography	178
12	Glycerides of linseed oil	plaster-of-Paris-bound silicic acid	gradient elution from pet. ether to pet. ether-EtOEt (1:1)	conc. H ₂ SO ₄	a. diagnostic b. quantitative c. used to find optimum reaction conditions	335
13	Cl ¹⁴ and tritium assay of tripalmitin and fatty acids	Silica Gel G	pet. ether-EtOEt-HOAc (90:10:1)	a. I ₂ b. 2',7'-dichloro-fluorescein	a. quantitative b. preparative	364
14	Alkoxydiglycerides				diagnostic	212
15	Acetoglycerides	silicic acid	"Skellysolve" F-EtOEt (92:8)	I ₂	diagnostic	130
16	Glycerides of the Swiss Pharmacopeia	Kieselguhr G impregnated with paraffin	HOAc	a. I ₂ and benzidine b. Rhodamine B and KOH	diagnostic	6
17	Glycerides from plasma	Silica Gel G	pet. ether-methyl ethyl ketone-HOAc (84:15:1)	a. 50% H ₂ SO ₄ b. osmium tetroxide c. Rhodamine 6 G	diagnostic	459

TABLE 11.15. LIPIDS—GLYCOLIPIDS

	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1	Glycolipids	Silica Gel G	a. $\text{PrOH-NH}_4\text{OH}$ (12% NH_3)(4:1) b. 1,2-dichloroethane-MeOH (49:1) c. CHCl_3 -96% HOAc (95:5) all consecutively	diphenylamine	a. diagnostic b. stepwise technique	174
2	Glycolipids	Silica Gel G	CHCl_3 -MeOH- H_2O (60:35:8) (65:25:4)	a. Rhodamine B b. Dragendorff's reagent c. Bial's reagent d. diphenylamine	diagnostic	427 432

TABLE 11.16. LIPIDS—PHOSPHOLIPIDS AND SULFOLIPIDS

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Cephalins and lecithins	Silica Gel G	CHCl ₃ -MeOH-H ₂ O (11:5:1)	phosphomolybdic acid-SnCl ₂ on layer "smoked" with dichloro-dimethylsilane	diagnostic	185
2 Sphingosines, etc.	silicic acid on walls of test tubes	a. CHCl ₃ b. NH ₃ /MeOH-CHCl ₃ (20:80) (30:70) (NH ₃ /MeOH is MeOH-H ₂ O-NH ₄ OH (28% NH ₃) (93:5:2))	I ₂	diagnostic	219
3 Lecithin and eolamine-cephalin	Silica Gel G	CHCl ₃ -MeOH-H ₂ O (65:25:4)	a. Rhodamine B b. Dragendorff's reagent. c. ninhydrin	quantitative	430
4 Misc. phospholipids	Silica Gel G	CHCl ₃ -MeOH-H ₂ O (60:35:8) (65:25:4)	a. Rhodamine B b. Dragendorff's reagent. c. ammonium molybdate-HClO ₄ d. ninhydrin	diagnostic	427 432
5 Serum phospholipids	plaster-of-Paris-bound silicic acid	CHCl ₃ -MeOH-H ₂ O (80:25:3)	2', 7'-dichlorofluorescein	diagnostic	423
6 Lecithins, cephalins and sphingomyelins	Silica Gel G with 10% (NH ₄) ₂ SO ₄	CHCl ₃ -H ₂ SO ₄ /MeOH (80:20) (H ₂ SO ₄ /MeOH is 0.1 N aqueous H ₂ SO ₄ -MeOH (5:95))	a. I ₂ b. 2', 7'-dichlorofluorescein c. K ₂ C ₂ O ₇ -H ₂ SO ₄	diagnostic	234
7 Cerebron, kersin and their sulfuric acid esters	Silica Gel G	a. 1,2-dichloroethane b. CHCl ₃ -HOAc (96:4) c. PrOH-NH ₄ OH (12.5% NH ₃) (8:2)	bromthymol blue	diagnostic	173
8 Beef brain gangliosides	Silica Gel G	PrOH-H ₂ O (7:3)	a. bromthymol blue b. Ehrlich reagent	preparative	208
9 Cerebral gangliosides	Silica Gel G	BuOH-pyridine-H ₂ O (3:2:1)	Bial's reagent	diagnostic	197
10 Beef brain gangliosides	Silica Gel G	PrOH-1 N NH ₄ OH-H ₂ O (6:2:1)		diagnostic	440

TABLE 11.15. LIPIDS—GLYCOLIPIDS

	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1	Glycolipids	Silica Gel G	a. $\text{PrOH-NH}_4\text{OH}$ (12% NH_3)(4:1) b. 1,2-dichloroethane-McOH (49:1) c. CHCl_3 -96% HOAc (95:5) all consecutively	diphenylamine	a. diagnostic b. stepwise technique	174
2	Glycolipids	Silica Gel G	CHCl_3 -McOH- H_2O (60:35:8) (65:25:4)	a. Rhodamine B b. Dragendorff's reagent c. Bial's reagent d. diphenylamine	diagnostic	427 432

TABLE 11.16. LIPIDS—PHOSPHOLIPIDS AND SULFOLIPIDS

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Cephalins and lecithins	Silica Gel G	CHCl ₃ -MeOH-H ₂ O (11:5:1)	phosphomolybdic acid-SnCl ₂ on layer ^a smoked ^b with dichlorodimethylsilane	diagnostic	185
2 Sphingosines, etc.	silicic acid on walls of test tubes	a. CHCl ₃ b. NH ₃ /MeOH-CHCl ₃ (20:80) (30:70) (NH ₃ /MeOH is MeOH-H ₂ O-NH ₃ OH (28% NH ₃) (93:5:2))	I ₂	diagnostic	219
3 Lecithin and colamine-cephalin	Silica Gel G	CHCl ₃ -MeOH-H ₂ O (65:25:4)	a. Rhodamine B b. Dragendorff's reagent c. ninhydrin	quantitative	430
4 Misc. phospholipids	Silica Gel G	CHCl ₃ -MeOH-H ₂ O (60:35:4) (65:25:4)	a. Rhodamine B b. Dragendorff's reagent c. ammonium molybdate-HClO ₄ d. ninhydrin	diagnostic	427 432
5 Serum phospholipids	plaster-of-Paris-bound silicic acid	CHCl ₃ -MeOH-H ₂ O (80:25:3)	2',7'-dichlorofluorescein	diagnostic	423
6 Lecithins, cephalins and sphingomyelins	Silica Gel G with 10% (NH ₄) ₂ SO ₄	CHCl ₃ -H ₂ SO ₄ /MeOH (80:20) (H ₂ SO ₄ /MeOH is 0.1 N aqueous H ₂ SO ₄ -MeOH (5:95))	a. I ₂ b. 2',7'-dichlorofluorescein c. K ₂ Cr ₂ O ₇ -H ₂ SO ₄	diagnostic	234
7 Cerebron, kerafin and their sulfuric acid esters	Silica Gel G	a. 1,2-dichloroethane b. CHCl ₃ -HOAc (96:4) c. PrOH-NH ₄ OH (12.5% NH ₃) (8:2)	bromthymol blue	diagnostic	173
8 Beef brain gangliosides	Silica Gel G	PrOH-H ₂ O (7:3)	a. bromthymol blue b. Ehrlich reagent	preparative	203
9 Cerebral gangliosides	Silica Gel G	BuOH-pyridine-H ₂ O (3:2:1)	Bial's reagent	diagnostic	197
10 Beef brain gangliosides	Silica Gel G	PrOH-1 N NH ₄ OH-H ₂ O (6:2:1)		diagnostic	440

TABLE 11.17. NUCLEOTIDES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Misc. nucleotides	collodion-bound "ECT-EOLA" cellulose	a. 0.15 M NaCl b. 0.01 N HCl	U.V.	diagnostic	316 317
2 Misc. nucleotides	Cellulose Powder MN 300 G	a. H ₂ O b. sat. aqueous (NH ₄) ₂ SO ₄ -1 M NaOAc-2-PrOH (80:18:2) c. <i>t</i> -amyl alcohol-formic acid-H ₂ O (3:2:1) d. BuOH-acetone-HOAc-NH ₄ OH (5% NH ₃)-H ₂ O (3.5:2.5:1.5:1.5:1)	U.V.	a. diagnostic b. comparison with paper chromatog.	319
3 Misc. nucleotides	a. DEAE-cellulose b. cellulose c. ECTEOLA-cellulose	dilute aqueous HCl 0.02 to 0.04 N sat. aqueous (NH ₄) ₂ SO ₄ -1 M NaOAc-2-PrOH (80:18:2) 0.06 N HCl	U.V.	a. diagnostic b. preparative c. review	320 318
4 Misc. nucleotides	Cellulose Powder MN 300	a. H ₂ O for nucleobases b. BuOH-acetone-HOAc-NH ₄ OH (5% NH ₃)-H ₂ O (4.5:1.5:1:1:2)	U.V.	diagnostic	321
5 Nucleotides, AMP, ADP and ATP	"DEAE-Sephadex" A-25	gradient elution starting with 1.0 N formic acid and adding 10 N formic acid which is 2 M in ammonium formate	fluorescein	diagnostic	82
6 Amino acid esters as models for aminoacylribonucleic acid	see "Proteins, etc.," No. 18				456

TABLE 11.18. PHENOLS, ALCOHOLS AND THEIR DERIVATIVES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Acetylene alcohols and glycols					4
2 Chmyl, batyl and se-lachyl alcohols	see "Lipids-Glycerides," No. 14				212
3 Long-chain aliphatic alcohols	Silica Gel G	a. pet. ether-EtOEt-HOAc (90:10:1) b. pet. ether-EtOEt (20:80) (10:90) (70:30)	50% H ₂ SO ₄	diagnostic	395
4 Cyclohexanol derivatives	Silica Gel G	a. benzene b. benzene-EtOEt (80:20)	H ₂ O	a. diagnostic b. preparative	128
5 Cardanol and its methylol derivatives	plaster-of-Paris-bound silica gel	EtOAc-toluene (60:40)	diazotized sulfanilic acid	followed course of reaction	11
6 Phenolic acids and esters	see "Carboxylic Acids, etc.," No. 9				69
7 Phenols	a. starch-bound silicic acid b. starch-bound silicic acid kieselguhr (1:1)	a. xylene b. xylene-CHCl ₃ (3:1)(1:1)(1:3) c. CHCl ₃	diazotized sulfanilic acid followed by KOH	diagnostic	426
8 Pyrocatechol and eugenol derivatives	Silica Gel G, acidic and basic	benzene	SbCl ₅ in CCl ₄	diagnostic	369
9 Phenols and phenolic acids	starch-bound silicic acid with phosphors	a. "Skellysolve" B-EtOEt (3:7) b. "Skellysolve" B-EtOAc (1:3) c. "Skellysolve" B-acetone (3:1)	U.V.	followed course of reaction	225
10 Phloroglucinol derivatives	see "Miscellaneous," No. 15				380
11 Phenolic aldehydes	see "Aldehydes, etc.," No. 9				396

TABLE 11.17. NUCLEOTIDES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Misc. nucleotides	collodion-bound "ECT-EOLA" cellulose	a. 0.15 <i>M</i> NaCl b. 0.01 <i>N</i> HCl	U.V.	diagnostic	316 317
2 Misc. nucleotides	Cellulose Powder MN 300 G	a. H ₂ O b. sat. aqueous (NH ₄) ₂ SO ₄ -1 <i>M</i> NaOAc-2-PrOH (80:18:2) c. <i>t</i> -amyl alcohol-formic acid-H ₂ O (3:2:1) d. BuOH-acetone-HOAc-NH ₄ OH (5% NH ₃)-H ₂ O (3.5:2.5:1.5:1.5:1)	U.V.	a. diagnostic b. comparison with paper chromatog.	319
3 Misc. nucleotides	a. DEAE-cellulose b. cellulose c. ECTEOLA-cellulose	dilute aqueous HCl 0.02 to 0.04 <i>N</i> sat. aqueous (NH ₄) ₂ SO ₄ -1 <i>M</i> NaOAc-2-PrOH (80:18:2) 0.06 <i>N</i> HCl	U.V.	a. diagnostic b. preparative c. review	320 318
4 Misc. nucleotides	Cellulose Powder MN 300	a. H ₂ O for nucleobases b. BuOH-acetone-HOAc-NH ₄ OH (5% NH ₃)-H ₂ O (4.5:1.5:1:1:2)	U.V.	diagnostic	321
5 Nucleotides, AMP, ADP and ATP	"DEAE-Sephadex" A-25	gradient elution starting with 1.0 <i>N</i> formic acid and adding 10 <i>N</i> formic acid which is 2 <i>M</i> in ammonium formate	fluorescein	diagnostic	82
6 Amino acid esters as models for aminoacylribonucleic acid	see "Proteins, etc." No. 18				456

TABLE 11.18. PHENOLS, ALCOHOLS AND THEIR DERIVATIVES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Acetylene alcohols and glycols					4
2 Chiniyl, batyl and se-lachyl alcohols	see "Lipids-Glycerides," No. 14				212
3 Long-chain aliphatic alcohols	Silica Gel G	a. pct. ether-EtOEt-HOAc (90:10:1) b. pct. ether-EtOEt (20:80) (10:90) (70:30)	50% H ₂ SO ₄	diagnostic	395
4 Cyclohexanol derivatives	Silica Gel G	a. benzene b. benzene-EtOEt (80:20)	H ₂ O	a. diagnostic b. preparative	123
5 Cardanol and its methylol derivatives	plaster-of-Paris-bound silica gel	EtOAc-toluene (60:40)	diazotized sulfanilic acid	followed course of reaction	11
6 Phenolic acids and esters	see "Carboxylic Acids, etc.," No. 9				69
7 Phenols	a. starch-bound silicic acid b. starch-bound silicic acid kieselguhr (1:1)	a. xylene b. xylene-CHCl ₃ (3:1) (1:1) (1:3) c. CHCl ₃	diazotized sulfanilic acid followed by KOH	diagnostic	426
8 Pyrocatechol and eugenol derivatives	Silica Gel G, acidic and basic	benzene	SbCl ₅ in CCl ₄	diagnostic	369
9 Phenols and phenolic acids	starch-bound silicic acid with phosphors	a. "Skellysolve" B-EtOEt (3:7) b. "Skellysolve" B-EtOAc (1:3) c. "Skellysolve" B-acetone (3:1)	U.V.	followed course of reaction	225
10 Phloroglucinol derivatives	see "Miscellaneous," No. 15				380
11 Phenolic aldehydes	see "Aldehydes, etc.," No. 9				396

TABLE 11.17. NUCLEOTIDES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Misc. nucleotides	collodion-bound "ECT-EOA" cellulose	a. 0.15 <i>M</i> NaCl b. 0.01 <i>N</i> HCl	U.V.	diagnostic	316 317
2 Misc. nucleotides	Cellulose Powder MN 300 G	a. H ₂ O b. sat. aqueous (NH ₄) ₂ SO ₄ -1 <i>M</i> NaOAc-2-PrOH (80:18:2) c. <i>t</i> -amyl alcohol-formic acid-H ₂ O (3:2:1) d. BuOH-acetone-HOAc-NH ₄ OH (5% NH ₃)-H ₂ O (3.5:2.5:1.5:1.5:1)	U.V.	a. diagnostic b. comparison with paper chromatog.	319
3 Misc. nucleotides	a. DEAE-cellulose b. cellulose c. ECTEOA-cellulose	dilute aqueous HCl 0.02 to 0.04 <i>N</i> sat. aqueous (NH ₄) ₂ SO ₄ -1 <i>M</i> NaOAc-2-PrOH (80:18:2) 0.06 <i>N</i> HCl	U.V.	a. diagnostic b. preparative c. review	320 318
4 Misc. nucleotides	Cellulose Powder MN 300	a. H ₂ O for nucleobases b. BuOH-acetone-HOAc-NH ₄ OH (5% NH ₃)-H ₂ O (4.5:1.5:1.5:1:2)	U.V.	diagnostic	321
5 Nucleotides, AMP, ADP and ATP	"DEAE-Sephadex" A-25	gradient elution starting with 1.0 <i>N</i> formic acid and adding 10 <i>N</i> formic acid which is 2 <i>M</i> in ammonium formate	fluorescein	diagnostic	82
6 Amino acid esters as models for aminoacylribonucleic acid	see "Proteins, etc.," No. 18				456

TABLE 11.18. PHENOLS, ALCOHOLS AND THEIR DERIVATIVES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Acetylene alcohols and glycols					4
2 Chiniyl, butyl and se-lachyl alcohols	see "Lipids-Glycerides," No. 14				212
3 Long-chain aliphatic alcohols	Silica Gel G	a. pct. ether-EtOEt-HOAc (90:10:1) b. pct. ether-EtOEt (20:80) (10:90) (70:30)	50% H ₂ SO ₄	diagnostic	395
4 Cyclohexanol derivatives	Silica Gel G	a. benzene b. benzene-EtOEt (80:20)	H ₂ O	a. diagnostic b. preparative	128
5 Cardanol and its methylol derivatives	plaster-of-Paris-bound silica gel	EtOAc-toluene (60:40)	diazotized sulfanilic acid	followed course of reaction	11
6 Phenolic acids and esters	see "Carboxylic Acids, etc.," No. 9				69
7 Phenols	a. starch-bound silicic acid b. starch-bound silicic acid kieselguhr (1:1)	a. xylene b. xylene-CHCl ₃ (3:1) (1:1) (1:3) c. CHCl ₃	diazotized sulfanilic acid followed by KOH	diagnostic	426
8 Pyrocatechol and eugenol derivatives	Silica Gel G, acidic and basic	benzene	SbCl ₅ in CCl ₄	diagnostic	369
9 Phenols and phenolic acids	starch-bound silicic acid with phosphors	a. "Skellysolve" B-EtOEt (3:7) b. "Skellysolve" B-EtOAc (1:3) c. "Skellysolve" B-acetone (3:1)	U.V.	followed course of reaction	225
10 Phloroglucinol derivatives	see "Miscellaneous," No. 15				380
11 Phenolic aldehydes	see "Aldehydes, etc.," No. 9				396

TABLE 11.18—*Cont.*

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
12 Phenolic aldehydes, ketones and acids	Silica Gel G	a. benzene-dioxane-HOAc (90:25:4) b. benzene-McOH-HOAc (45:8:4) c. benzene-McOH (95:5)	a. diazotized benzidine b. alkaline KMnO_4 c. diazotized sulfanilic acid d. Rhodamine B	a. diagnostic b. quantitative	288 290
13 Phenols and glycols	a. unbound aluminum oxide with HOAc b. unbound aluminum oxide	a. benzene b. benzene-EtOH mixtures c. CHCl_3 d. EtOEt	paper chromatography reagents	diagnostic	144
14 Phenols	plaster-of-Paris-bound silica gel-oxalic acid	a. hexane-EtOAc (4:1)(3:2) b. benzene-ethyl ether (4:1) c. benzene		diagnostic	113
15 3,5-Dinitrobenzoates of alcohols and phenols	Silica Gel G	a. benzene-pet. ether (1:1) b. hexane-EtOAc (85:15)(75:25) c. toluene-EtOAc (90:10)	colored compounds	does not separate closely related compounds	85
16 3,5-Dinitrobenzoates of alcohols	starch-bound silica gel-tonite (4:1) with fluorescent material	hexane	U.V.	diagnostic	209
17 3,5-Dinitrobenzoates of butyl and isocetyl alcohols	silica gel		colored compounds		220
18 Phenols in pharmaceuticals	unbound acidic aluminum oxide	a. ether b. benzene-EtOH (95:5)(90:10)	a. $\text{KMnO}_4\text{-Fe}(\text{NO}_3)_3$ b. KMnO_4	diagnostic	340
19 Phenolic carboxylic acids	Kieselguhr G with Na_2WO_4 , Na_2MoO_4 or borax	a. organic layer of BuOH-HOAc- H_2O (4:1:5) b. butyl ether (sat. with H_2O)-HOAc (10:1) c. EtOAc-2-PrOH- H_2O (65:24:11)	Folin Denis Reagent		136

20	Xylenols	Silica Gel G	a. benzene b. benzene-McOH (95:5)	a. diazotized benzidine b. Echt blue salt	301
21	Alcohols				28
22	Phenols	Silica Gel G and boric acid	EtOH-H ₂ O (80:30) containing 4% boric acid and 2% NaOAc as electrolyte	Diazotized benzidine	291
23	Phenols	ϵ -polycaprolactam	a. hexane b. cyclohexane c. benzene d. CHCl ₃ e. EtOAc	diazotized sulfanilic acid	219a 436a

TABLE 11.19. PROTEINS, AMINO ACIDS AND THEIR DERIVATIVES

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Amino acids	Silica Gel G	a. EtOH (96%)-H ₂ O (63:37) b. PrOH-H ₂ O (64:36) c. BuOH-HOAc-H ₂ O (60:20:20) d. phenol-H ₂ O (75:25) e. PrOH-NH ₄ OH (34% NH ₃) (67:33) f. EtOH (96%)-NH ₄ OH (34% NH ₃) (77:23)	modified ninhydrin	a. diagnostic b. two dimensional	44 277
2 Amino acids	unbound aluminum oxide	a. H ₂ O b. pyridine-H ₂ O (1:1)(80:54) c. BuOH-EtOH-H ₂ O (60:40:40)	ninhydrin	a. diagnostic b. continuous elution c. spotted sample with NaOH	267
3 Amino acids	buffered Silica Gel G (buffered with equal portions of 0.2 M KH ₂ PO ₄ and 0.2 M NaH ₂ PO ₄)	a. EtOH-H ₂ O (70:30) b. EtOH-NH ₄ OH (25% NH ₃) (4:1) c. EtOH-NH ₄ OH-H ₂ O (7:1:2)	paper chrom. reagents	diagnostic	271
4 Amino acids	Silica Gel G	a. CHCl ₃ -MeOH-NH ₄ OH (17% NH ₃) (2:2:1) in first direction b. phenol-H ₂ O (75:25 g/g) in second dimension c. methyl ethyl ketone-pyridine-H ₂ O-HOAc (70:15:15:2) for continuous	ninhydrin	a. diagnostic b. continuous development c. two dimensional	99
5 Amino acids	Cellulose Powder MN 300	a. upper phase of BuOH-HOAc-H ₂ O (4:1:5) b. pyridine-methyl ethyl ketone-H ₂ O (15:70:15) c. MeOH-H ₂ O-pyridine (80:20:4) d. six others	ninhydrin	diagnostic	450
6 Amino acids	Silica Gel G	a. PrOH-H ₂ O (1:1) in first direction b. phenol-H ₂ O (10:4) in second direction	ninhydrin	a. diagnostic b. two dimensional	279
7 Amino acids and similar substances	a. Silica Gel G b. Kieselguhr G c. Aluminum Oxide G	a. 440V, 19.6 m. amp. in 0.1 M sodium citrate buffer, pH 3.8 b. 460V, 12.6 m. amp. in 2 N HOAc-0.6 N formic acid (1:1) at pH 2 c. chromatography using BuOH-HOAc-H ₂ O (3:1:1)	ninhydrin	a. diagnostic b. electrophoresis c. combined electrophoresis and chromatog.	156

8	N-(2,4-Dinitrophenyl)amino acids	a. Silica Gel G b. Silica Gel G with phosphor	a. toluene-pyridine-ethylencchlorohydrin-0.8 N NH ₄ OH (100:30:60:60), upper phase, in first direction b. CHCl ₃ -benzyl alcohol-HOAc (70:30:3) in second direction c. Sol. a in one direction and CHCl ₃ -MeOH-HOAc (95:5:1) in the second direction as continuous chrom. d. Sol. a in one direction and benzene-pyridine-HOAc (80:20:2) in second direction as continuous chrom. e. CHCl ₃ -4-aryl alcohol-HOAc (70:30:3)	a. colored compounds b. transmitted U.V. c. U.V.	a. Layers are all equilibrated with lower phase of Sol. a before use b. continuous development c. two dimensional	45 46
9	C ¹⁴ -tagged N-(2,4-dinitrophenyl)-amino acids	Silica Gel G	a. toluene-pyridine-ethylencchlorohydrin-0.8 N NH ₄ OH (100:30:60:60), upper phase in first direction b. CHCl ₃ -benzylalcohol-glac. HOAc (70:30:3) in second direction	radiographic	a. two dimensional b. layer equilibrated with lower phase of sol. a before use	89
10	3-Phenyl-2-thiobenzoyl amino acids	Silica Gel G	a. CHCl ₃ b. CHCl ₃ -MeOH (9:1) c. CHCl ₃ -formic acid (100:5) d. CHCl ₃ -MeOH-formic acid (70:30:2) e. Methyl ethyl ketone-pyridine-H ₂ O-conc. HOAc (70:15:15:2)	Cl ₂ -toluidine	diagnostic	46
11	Carbomethoxy derivatives of amino acids and small peptides	Silica Gel G	a. BuOH-acetone-glac. HOAc-NH ₄ OH (7% NH ₃)-H ₂ O (4.5:1.5:1:1:2) b. BuOH-glac. HOAc-NH ₄ OH (7% NH ₃) (5.5:3:1.5) c. BuOH-HOAc-NH ₄ OH (7% NH ₃)-H ₂ O (6:1:1:2) d. BuOH-glac. HOAc-H ₂ O-pyridine (15:3:12:10) e. in case of two-phase systems, both were used, but layer dipped only in upper phase	a. ninhydrin b. K ₂ Cr ₂ O ₇ in conc. H ₂ SO ₄	diagnostic	95
12	Amino acid and dipeptide derivatives	Silica Gel G	CHCl ₃ -MeOH (9:1)	Folin reagent	a. diagnostic b. monitored a synthesis	163
13	Polypeptides (up to five units) and derivatives	Silica Gel G	a. CHCl ₃ -acetone (9:1)(8:2) b. cyclohexane-EtOAc (1:1) c. CHCl ₃ -MeOH (9:1)	morin	diagnostic	341

TABLE 11.19—Cont.

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
14 Polypeptides (up to ten units) and derivatives	Silica Gel G	a. BuOH sat. with 0.1% NH ₄ OH b. other systems used in paper chrom.	ninhydrin	a. diagnostic b. monitored a synthesis	424
15 3-Phenyl-2-thiohydantoin and tripeptides	Silica Gel G	heptane-pyridine-EtOAc (5:3:2)	starch-iodine azide	diagnostic	54
16 Proteins as such	polyamide-bound hydroxyl-apatite	potassium phosphate buffers, pH 6.5 of various concentrations	ninhydrin	a. diagnostic b. gradient elution	153
17 Proteins and polypeptides	Sephadex G 25 (moist)	a. 0.05 M NH ₄ OH b. H ₂ O	a. Pauly's reagent	a. diagnostic b. determine mol. wt.	82
Enzymes	"DEAE-Sephadex" A 25	c. gradient elution involving addition of 0.02 M phosphate buffer pH 7.2 (which is also 1 M in NaCl) to the pure buffer	b. DPNH-pyruvic acid	c. gradient elution d. preparative	
18 Carbenoxyminoacyl derivatives of cyclic alcohols	Silica Gel G	CHCl ₃ -acetone (85:15)	Cl ₂ followed by KI and starch	a. monitored course of reaction b. monitored column chrom.	456
19 Metabolic products of tryptophane	Silica Gel G	a. CHCl ₃ -MeOH-HOAc (75:20:5) b. methyl acetate-2-ProH-NH ₄ OH (25% NH ₃) (45:35:20)	Ehrlich reagent	a. diagnostic b. urine analysis	87

TABLE 11.20. STEROIDS

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Seventeen misc. steroids	Silica Gel G	a. benzene b. benzene-EtOAc (9:1)(2:1) c. cyclohexane-EtOAc (9:1)(19:1) d. 1,2-dichloroethane e. seven more	SbCl ₃ in CHCl ₃	diagnostic	416
2 Seventy-nine steroids and sapogenins	unbound aluminum oxide, sometimes impregnated with morin	a. hexane-benzene mixtures b. benzene-EtOEt mixtures c. EtOEt-EtOH mixtures	a. conc. H ₂ SO ₄ b. U.V. c. morin spray	a. diagnostic b. preparative	52
3 Twenty-two misc. steroids	a. starch-bound silicic acid b. Silica Gel G	EtOAc-cyclohexane mixtures	a. SbCl ₃ in CHCl ₃ b. 2,4-DNPH c. phosphomolybdic acid	diagnostic	13
4 Twenty-seven steroid estrogens	Silica Gel G	a. cyclohexane-EtOAc-EtOH (45:45:10) b. EtOAc-hexane-EtOH (80:15:5) c. EtOAc-cyclohexane (50:50) d. CHCl ₃ -EtOH (90:10) e. EtOAc-hexane-EtOH-HOAc (72:13:5:4:5:10) f. BuOH sat. with H ₂ O	a. Folin-Ciocalteu reagent b. Boule reagent c. 2% H ₂ SO ₄ in H ₂ O-EtOH (1:1)	a. diagnostic b. classification system	221
5 Forty-nine misc. steroids	unbound aluminum oxide	a. benzene-pet. ether mixtures b. benzene-EtOH mixtures	a. SbCl ₃ in CHCl ₃ b. SbCl ₃ in CHCl ₃ with 10% SOCl ₂	diagnostic	145
6 Twenty-seven misc. steroids	a. unbound aluminum oxide	a. pet. ether-benzene mixtures b. benzene c. benzene-EtOH mixtures	reagents used in paper chromatography	diagnostic	144
Steroid acids	b. unbound aluminum oxide-HOAc	d. benzene-EtOH (90:10)			

TABLE 11.20—Cont.

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
7 Steroids from fermentation reactions	Silica Gel G	CHCl ₃ -acetone mixtures	a. vanillin in H ₃ PO ₄ b. anisaldehyde in H ₂ SO ₄ c. SbCl ₅ in CHCl ₃ d. H ₃ PO ₄ e. 2,3,5-triphenyltetrazolium chloride f. HClO ₄	diagnostic	247
	a. Silica Gel G with phosphor b. Aluminum Oxide G with phosphor	a. CHCl ₃ -EtOAc (3:2)(3:1) b. benzene-EtOAc (1:1)(9:1) c. hexane-EtOAc (1:1)	I ₂ vapor	a. quantitative b. diagnostic	244
9 Fifty-five misc. steroids	a. starch-bound silica gel b. starch-bound silica gel-NaOH	a. hexane-EtOAc (4:1)(1:1) b. EtOAc c. benzene-2-PrOH (4:1)	a. phosphomolybdic acid b. many others	diagnostic	363
	plaster-of-Paris-bound "Celite" impregnated, by spraying, with a. formamide b. propylene glycol c. paraffin oil	a. hexane b. hexane-benzene (1:1) c. benzene d. CHCl ₃ e. ligroin f. MeOH-H ₂ O (95:5)	a. H ₃ PO ₄ b. tetrazolium blue c. isonicotinic acid hydrazide	a. diagnostic b. distribution	414
11 Fifteen sterols and sterol derivatives	a. Silica Gel G b. Silica Gel G Kieselguhr G (1:1)	a. cyclohexane-EtOAc-H ₂ O (600:400:1) (1560:440:1) b. isooctane-CCl ₄ (19:1) c. cyclohexane-heptane (1:1)	50% H ₂ SO ₄	diagnostic	27

12	Sterols and sterol acetates	a. Kieselguhr G b. Kieselguhr G (impregnated with undecane)	a. cyclohexane-EtOAc (99:5:0.5) b. HOAc-H ₂ O (92:8)(90:10) c. HOAc-CH ₃ CN (1:3)	phosphomolybdic acid	diagnostic	295
13	Twenty-seven steroidal sapogenins and derivatives	a. Silica Gel G b. Kieselguhr G c. a and b (1:1) d. all equilibrated with water by placing moist filter paper in the chamber	a. CH ₂ Cl ₂ -MeOH-formamide (93:6:1) b. toluene-EtOAc-formic acid (57:40:3) c. cyclohexane-acetone (1:1) d. cyclohexane-EtOAc-H ₂ O (600:400:1) (1000:1000:3) e. CHCl ₃ -MeOH-H ₂ O (485:15:1) (188:12:1) f. three others	50% H ₂ SO ₄	a. diagnostic b. distribution chrom.	26
14	Steroidal sapogenins	Silica Gel G	a. EtOAc b. butyl acetate	chlorosulfonic acid-HOAc (1:2)	diagnostic	405
15	Synthetic steroids	Silica Gel G	a. benzene b. CHCl ₃ c. CHCl ₃ -MeOH (10:1)(19:1)(97:3)	SbCl ₅ in CHCl ₃	diagnostic	67 68
16	Thirty-eight 19-nor-steroids	Silica Gel G	EtOAc-cyclohexane mixtures	SbCl ₅ in CHCl ₃	diagnostic	125
17	Estrogens	Silica Gel G	benzene-EtOH (9:1)	SbCl ₅ in CCl ₄	a. diagnostic b. quantitative	394
18	Pregnatriols from urine	unbound aluminum oxide	benzene-EtOH (9:1)	70% H ₃ PO ₄	a. diagnostic b. quantitative	391
19	Estrone derivatives	plaster-of-Paris-bound silicic acid	pentane-EtOAc mixtures	SbCl ₅ in CHCl ₃	diagnostic	15
20	Pregnane derivatives	a. aluminum oxide b. Silica Gel Reidel de Haën	a. benzene-CH ₂ Cl ₂ (2:1) b. cyclohexane-CH ₂ Cl ₂ (1:2)	I ₂ in KI	diagnostic	347
21	Misc. derivatives of pregnanol, cholesterol and sitosterol	Silica Gel G	a. isopropyl ether b. isopropyl ether-EtOAc (5:2) c. benzene	chlorosulfonic acid-HOAc (1:2)	a. diagnostic b. preparative c. monitored course of reaction	408

TABLE 11.20—Cont.

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
22 Pregnandiol	Silica Gel G	CHCl ₃ -acetone (9:1)	<i>ortho</i> -H ₃ PO ₄ -phosphomolybdic acid	early test for pregnancy	433 435
23 Cholesterol and its laurate ester	Silica Gel G	benzene-McOH mixtures	conc. H ₂ SO ₄	quantitative	315
24 Cholesterol	Silica Gel G	hexane-EtOAc-HOAc (50:50:2)	HClO ₄	a. quantitative b. microscope slides	296
25 Blood cholesterol and cholesterol esters	Silica Gel G	a. benzene b. benzene-EtOAc (9:1) c. 1,2-dichloroethane d. CHCl ₃ e. two others	SbCl ₅ in CHCl ₃	a. diagnostic b. semiquantitative	415
26 Cholesterol esters	Silica Gel G	isopropyl ether-pet. ether (1:99)	SbCl ₅ in CHCl ₃	a. diagnostic b. quantitative	460
27 Cholesterol esters	Silica Gel G	benzene-pet. ether (40:60)	50% H ₂ SO ₄	diagnostic	228
28 Cholesterol esters	plaster-of-Paris-bound silicic acid impregnated with paraffin	HOAc	a. I ₂ vapor b. SbCl ₅ in CHCl ₃ c. phosphomolybdic acid d. phosphotungstic acid	a. diagnostic b. reversed phase	249
29 Bile acids, free and conjugated	Silica Gel G	a. HOAc-CCl ₄ -isopropyl ether-isoamyl acetate-PrOH-benzene (5:20:30:40:10:10) for free acids b. propanoic acid-isoamyl acetate-H ₂ O-PrOH (15:20:5:10) for conjugate acids	phosphomolybdic acid	diagnostic	154

30	Cholic acid derivatives	Silica Gel G	a. toluene-HOAc-H ₂ O (5:5:1) upper phase b. BuOH-HOAc-H ₂ O (10:1:1) c. toluene-HOAc-H ₂ O (10:10:1)	a. phosphomolybdic acid b. H ₂ O	115
31	Adrenocorticotrophic hormones	Silica Gel G	EtOAc-xylene-MeOH (90:5:5)	tetrazolium chloride	30
32	Corticoids from urine	unbound silica gel	CHCl ₃ -EtOH (95:5)	tetrazolium blue	1
33	Corticosteroids	Silica Gel G (not activated)	CHCl ₃ -MeOH-H ₂ O (485:15:1) (188:12:1)(90:10:1)	50% H ₂ SO ₄	25
34	Corticosteroids	plaster of Paris	a. acetone-benzene (1:4) b. EtOH-CHCl ₃ (3:97)	alkaline tetrazolium blue	243
35	Cholestane derivatives	Silica Gel G	pet. ether-benzene (9:1)	I ₂	361
36	Cholesterol derivatives	Silica Gel G	a. benzene b. benzene-EtOEt (70:30) c. pentane-EtOEt-benzene (20:20:60)	H ₂ O	125
37	Aglycones of oleander glycosides	Silica Gel G	EtOAc-CHCl ₃ (9:1)	SbCl ₃ in CHCl ₃	126
38	Cardinolides	Silica Gel G (not activated)	benzene-EtOH (3:1) sat. with H ₂ O	a. trichloroacetic acid-chloramine-T b. SbCl ₃ -acetic anhydride	322
39	Steroidal compounds involved in tomatin formation	Silica Gel G	EtOAc-cyclohexane (1:1)(3:7)	a. SbCl ₃ in chloroform b. 15% H ₃ PO ₄	337 339
40	Digitoxigenin and derivatives	Silica Gel G	CHCl ₃ -EtOAc (1:1)	trichloroacetic acid-chloramine T	90
41	Bile acids (free, conjugated) and derivatives	Silica Gel G	a. isoamyl acetate-CCl ₄ -isopropyl ether-benzene-PrOH-HOAc (4:3:2:1:0.5) b. isoamyl acetate-propanoic acid-PrOH-H ₂ O (4:3:2:1) c. cyclohexane-EtOAc (6:4)	phosphomolybdic acid	152

TABLE 11.20—Cont.

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
42 Steroids from bull testicles	a. plaster-of-Paris-bound silica gel b. plaster-of-Paris-bound aluminum oxide	a. benzene-EtOAc (8:2) b. EtOAc c. butyl acetate d. several others	conc. H_2SO_4 -U.V.	diagnostic	272
43 Radioactive androstane derivatives	Silica Gel G	cyclohexane-EtOAc (3:2)	radiographic	a. diagnostic b. quantitative c. preparative	349
44 β -Sitosterol from potato beetles	Silica Gel Riedel de Haen	pet. ether-EtOAc-benzene (85:10:5)	SDCs in HOAc	diagnostic	348
45 1,3-Cholestanedione methyl ether	Silica Gel G	benzene-EtOEt (4:1)	I_2 vapor	monitored column chrom.	397
46 Aldosterones from urine	Silica Gel G	a. cyclohexane-2-PrOH (7:3) b. CHCl_3 -glac. HOAc (8:2)	tetrazolium blue	a. two dimensional b. developed in dark c. quantitative	278
47 Saponinins from <i>Solanum</i> species	Kieselguhr G	EtOAc-cyclohexane (7:3)	vanillin in H_3PO_4	diagnostic	336

TABLE 11.21. TERPENOIDS

	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1	Forty-nine misc. terpenoids	starch-bound silicic acid	hexane	a. fluorescein-Br ₂ b. H ₂ SO ₄ -HNO ₃	a. diagnostic b. monitored column	190
2	Fourteen misc. terpenoid	starch-bound silicic acid with phosphors	hexane-EtOAc (85:15)	a. fluorescein-Br ₂ b. U.V. c. <i>o</i> -dianisidine d. conc. H ₂ SO ₄ e. bromocresol green f. H ₂ SO ₄ -HNO ₃	a. diagnostic b. predicted solvents for column chrom. c. monitored column	193
3	Twenty-eight oxygenated terpenoids	starch-bound silicic acid	a. hexane-EtOAc (85:15) (10:30) b. CHCl ₃ (alcohol-free) EtOAc (90:10) (95:5) c. 1-nitropropane-hexane (50:50) d. ethyl carbonate-CHCl ₃ (15:85) e. hexane-isopropyl formate (70:30) f. four others	a. fluorescein-Br ₂ b. <i>o</i> -dianisidine c. H ₂ SO ₄ -HNO ₃ d. conc. H ₂ SO ₄	a. diagnostic b. chemical reactions on layers	252
4	Twelve misc. terpenoids	unbound alumina	a. pet. ether-benzene mixtures b. benzene c. benzene-EtOH mixtures	reagents used for paper chromatography	diagnostic	144
5	Diterpenoids	silicic acid	hexane-EtOAc (4:1)	KMnO ₄	a. diagnostic b. monitored column c. review	213
6	Eighteen triterpenoids	synthetic fibrous alumina	a. hexane-benzene (8:1) b. cyclohexane-EtOAc (85:15) c. benzene d. EtOEt e. EtOEt-EtOH (98:2)	a. 2,4-DNPH b. SbCl ₅ in CCl ₄ c. Rhodamine B	diagnostic	165

TABLE 11.21—Cont.

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
7 Sixteen triterpenoids	Silica Gel G	a. isopropyl ether-acetone (5:2)(19:1) b. isopropyl ether c. cyclohexane d. benzene e. CH_2Cl_2	chlorosulfonic acid-HOAc	diagnostic	406 407
		f. sol. a g. sol. a plus 5% pyridine h. EtOAc-MeOH-diethylamine (14:4:3) i. chlorobenzene-HOAc (9:1) j. CH_2Cl_2 -pyridine (7:2)			
Twenty-three triterpene acids					
8 Triterpene acids	Silica Gel G	EtOEt-hexane (1:1)	50% H_2SO_4	diagnostic	421
9 Triterpene acids from <i>Commiphora glandulosa</i> Schinz, as methyl esters		CHCl_3 -EtOAc (4:1)	conc. H_2SO_4	diagnostic	402
10 Carotenoids	a. Silica Gel G- $\text{Ca}(\text{OH})_2$ (1:4)	a. benzene-pet. ether (1:1) b. sol. a with 1% MeOH	colored compounds	a. diagnostic b. preparative	446 449
	b. Silica Gel G- $\text{Ca}(\text{OH})_2$ (1:4) impregnated with paraffin	c. MeOH sat. with paraffin			
11 Isomeric menthols	Silica Gel G	a. benzene b. MeOH c. benzene-MeOH (95:5)(75:25)	SbCl_5 in CCl_4	diagnostic	299
12 Ipomeamarone	starch-bound silica gel	hexane-EtOAc (9:1)	2,4-DNPH	quantitative	2, 76

13	Isophytol and geranyl- linalol	starch-bound silica gel	hexane-EtOAc (85:15)		diagnostic	81
14	Azulene and cadalene	a. starch-bound sil- ica gel b. starch-bound alu- mina	a. pet. ether b. EtOH	trinitrobenzene	diagnostic	108
15	Terpenoids of mint oils	see "Essential Oils," No. 11				20

TABLE 11.22. VITAMINS

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Carotenes and fat-soluble vitamins A, D, E, K	unbound aluminum oxide	a. MeOH b. CCl ₄ c. xylene d. eleven others	a. 70% HClO ₄ b. conc. H ₂ SO ₄	diagnostic	70
2 Carotenes and Vitamin A	unbound aluminum oxide	pet. ether	a. SbCl ₅ in CHCl ₃ b. colored compounds	a. diagnostic b. quantitative	210 211
3 B vitamins	Silica Gel G with fluorescent material	HOAc-acetone-MeOH-benzene (5:5:20:70)	a. U.V. b. potassium iodoplatinate c. dichloroquinone-chlorimide-NH ₃ d. ninhydrin	a. diagnostic b. carried out in dark	116
4 Cyanocobalamin and hydroxycobalamin	Silica Gel G	a. MeOH b. PrOH c. BuOH-HOAc-H ₂ O (20:5:1) d. BuOH-HOAc-MeOH (20:10:25) e. sixteen others	a. microbiologically b. spectrophotometrically	a. diagnostic b. quantitative	55
5 Vitamins K ₂₍₂₀₎ , K ₂₍₂₀₎ and K ₂₍₁₀₎	Silica Gel G with sodium fluorescein	a. heptane-benzene (1:1) dried and followed by b. benzene	U.V.	a. diagnostic b. radioactive compounds c. quantitative	35
6 Tocopherol mixtures	a. Silica Gel G b. Aluminum Oxide D5	a. CHCl ₃ b. benzene	a. ceric sulfate b. phosphomolybdic acid-NH ₃	a. diagnostic b. quantitative	354

7	Retinen	a. Silica Gel G b. Silica Gel G impreg- nated with paraffin	a. pet. ether-benzene mixtures b. MeOH sat. with paraffin	rhodanin	a. diagnostic b. semiquant.	448
8	Fat-soluble vitamins	alumina		a. HClO ₄ -70 % b. H ₂ SO ₄ -98 %		37
9	Ubiquinones	a. Silica Gel G b. Silica Gel G impreg- nated with paraffin	a. benzene-CHCl ₃ (1:1) b. acetone-paraffin (sat. with H ₂ O) (9:1)	a. Rhodamine B b. SbCl ₃ in CHCl ₃	diagnostic	431
10	Vitamins					280

TABLE 11.23. MISCELLANEOUS

Compound	Adsorbent	Developer	Visualization	Comments	Ref.
1 Benzoquinones from <i>Myriophyes</i>	starch-bound silicic acid	hexane-EtOAc (85:15)		diagnostic	12
2 Antioxidants (phenolic)	unbound polyamide powder	MeOH-acetone mixtures	a. diazotized sulfanilic acid b. phosphomolybdic acid	a. monitored column b. diagnostic	72
3 Antioxidants, synthetic	Silica Gel G	a. benzene b. CHCl ₃ c. CHCl ₃ -methyl cellosolve (80:20)	a. phosphomolybdic acid b. diazotized sulfanilic acid c. 2,6-dichloroquinone-chlorimide, neutral and buffered	a. diagnostic b. two dimensional c. quantitative	352 353
4 Alkyl phosphenoxides, sulfides, sulfones, and sulfoxides	Silica Gel G	a. acetone b. CHCl ₃	a. Na ₂ Cr ₂ O ₇ -H ₂ SO ₄ b. KMnO ₄ -H ₂ SO ₄ c. FeCl ₃ and K ₃ Fe(CN) ₆	diagnostic	98
5 Thiophene derivatives	a. Silica Gel G b. Aluminum Oxide G	a. benzene-CHCl ₃ (9:1) b. MeOH c. pet. ether	a. U.V. b. isatin in H ₂ SO ₄	diagnostic	65
6 Aromatic amines and nitrocompounds	see "Amines," No. 5				110
7 Adulteration of the drug <i>Radix pimplae</i> DAB6	Silica Gel G	CHCl ₃	SnCl ₄ in CCl ₄	diagnostic	161
8 Diphenyl in citrus fruits	starch-bound silicic acid with phosphors	pet. ether	U.V.	quantitative	194 389
9 Polycyclic aromatics	acetylated cellulose powder	MeOH-EtOEt-H ₂ O (4:4:1)	fluorescein	diagnostic	82
10 Nitramine explosives	Silica Gel G	pet. ether-acetone (2:102)	diphenylamine	diagnostic	216

11	Peroxides	starch-bound silica gel	CCl ₄ -acetone (2:1)	a. KI b. ferrous ammonium sulfate and potassium thiocyanate	diagnostic	242
12	Piperidones, quinolones and cyclohexanones	unbound aluminum oxide	a. benzene-acetone (1:1) b. acetone-MeOH (3:1)	a. I ₂ b. U.V.	tested procedure	254
13	Honey bee larval foods	silicic acid	MeOH-CHCl ₃ (29:71)	a. <i>p</i> -anisidine b. ninhydrin c. FeCl ₃ d. I ₂ e. conc. H ₂ SO ₄	diagnostic	292
14	Plasticizers (phthalates, phosphates and other esters)	Silica Gel G with fluorescent material	a. isoctane-EtOAc (90:10) b. benzene-EtOAc (95:5) c. butyl ether-hexane (80:20)	a. phosphomolybdic acid b. resorcinol-EtOH-ZnCl ₂ followed by H ₂ SO ₄ and KOH c. alcoholic KOH followed by urea d. 2,6-dichloroquinone-chloroimide followed by borax e. five others	diagnostic	293
15	Phloroglucinol-butyrates from <i>Dryopteris</i>	a. Silica Gel G buffered with NaOAc b. Silica Gel G impregnated with paraffin	a. EtOAc b. MeOH-formic acid-H ₂ O (75:10:15)	Folin-Ciocalteu reagent-NH ₃	diagnostic	380
16	Chamazulene derivatives and dyes as test substances	a. silicic acid b. starch-bound silicic acid	hexane-EtOAc (9:1)		methods exploration	381
17	Usnic acid Aminopyrine	Silica Gel G	a. benzene b. MeOH	a. U.V. b. FeCl ₃	a. diagnostic b. on microscope slides	437

TABLE 11.23—Cont.

	Compound	Adsorbent	Developer	Visualization	Comments	Ref.
18	Neuraminic acid derivatives	Silica Gel G	a. $\text{PrOH}-1\text{ }N\text{ NH}_4\text{OH}-\text{H}_2\text{O}$ (6:2:1) b. $\text{PrOH}-\text{H}_2\text{O}$ (7:3)	a. Bial's reagent b. thiobarbituric acid-HIO ₄ -ethylene glycol	diagnostic	127
19	Narcotics in forensic chemistry	Silica Gel G with fluorescent material	piperidine-pet. ether (1:5)	U.V.	diagnostic in urine	92
20	Lumiflavin derivatives	Silica Gel G	a. 2 $N\text{ HClO}_4$ -dimethylsulfoxide-BuOH (2:2:6) b. $\text{HOAc}-\text{H}_2\text{O}-\text{BuOH}$ (2:2:6)	colored compounds	diagnostic	143
21	Bufadienolides from toad poisons	Silica Gel G	a. EtOAc b. EtOAc-cyclohexane (80:20) c. EtOAc-acetone (90:10) d. EtOAc sat. with H_2O	SbCl_5 in CHCl_3	diagnostic	457
22	Adrenaline and its derivatives	Silica Gel G	CHCl_3 -MeOH (9:1)	H_3PO_4 -phosphomolybdic acid	a. diagnostic b. quantitative	434
23	Soladulcidin, a nitrogen-containing aglycone from glycosides of <i>Solanum dulcamara</i> L.	Silica Gel G	EtOAc-cyclohexane (7:3)	vanillin- H_3PO_4	diagnostic	338
24	Anthocyanins	polyacrylonitrile-polyamide (7:1)	1-pentanol- $\text{PrOH}-\text{HOAc}-\text{H}_2\text{O}$ (3:2:2:1) plus 1-2% 1-heptanol	colored compounds	a. diagnostic b. descending	36
25	Olefins as mercury salts	Silica Gel G	$\text{PrOH}-\text{triethylamine}-\text{H}_2\text{O}$ (50:25:25)	diphenylcarbazide	diagnostic	309

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